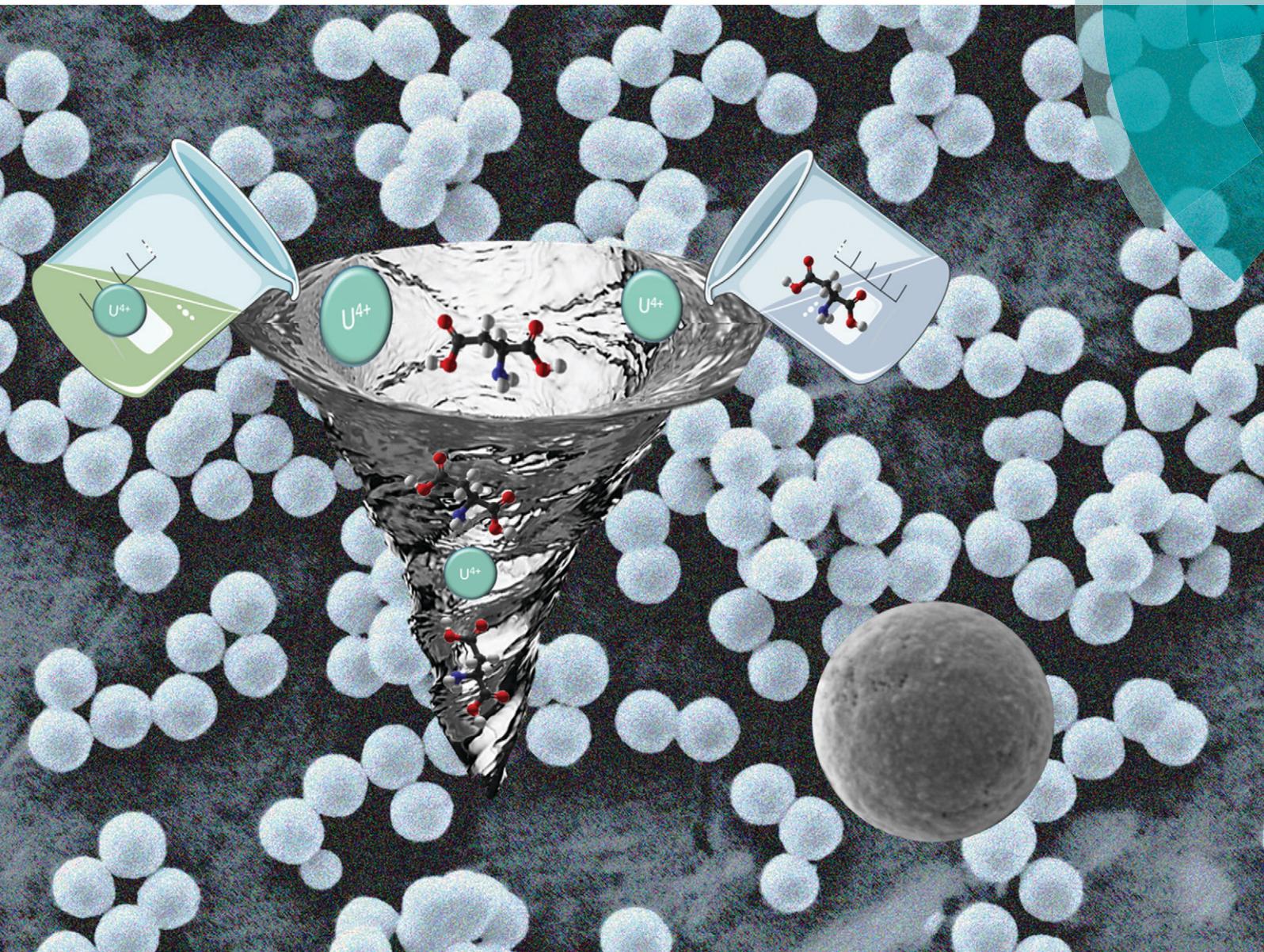


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PAPER

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Synthesis of size-controlled UO_2 microspheres from the hydrothermal conversion of U(IV) aspartate†

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A simple wet chemistry route towards micrometric spherical UO_2 particles was designed through the conversion of uranium(IV) aspartate under mild hydrothermal conditions ($T = 160$ °C). A multiparametric study examining the effects of hydrothermal treatment duration, the initial uranium/aspartic acid molar ratio and magnetic stirring allowed us to point out the role of organic species in the shaping of the particles and to specify the operating conditions leading to monodisperse and size-controlled particles. In a turbulent flow regime (*i.e.* $\text{Re}_a > 10^4$), particles with a diameter ranging from 400 nm to 2500 nm were obtained with a typical dispersion of less than $\pm 10\%$. Moreover, the protocol was found to be robust and reproducible, with only limited size variation from one batch to another (typically less than $\pm 5\%$ on the particle diameter). The effect of an additional heat treatment step was also investigated which showed that residual traces of water and organics can be removed after firing at 600 °C without altering the initial morphology. This wet chemistry route appears to be very promising for the production of spherical UO_2 particles and can be simply implemented in any nuclear chemistry lab, which paves the way to applications of such materials in various scientific areas.

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1. Introduction

Among the numerous challenges linked to the development of advanced materials for the nuclear fuel cycle, the elaboration of shape-controlled powders has recently attracted growing interest, especially for actinide-based oxides.¹ In this frame, most of the efforts concern the synthesis of spherical particles with prospective uses in various applications. First of all, as UO_2 and $(\text{U,Pu})\text{O}_2$ solid solutions are aimed at being used as fuels in future Gen III and/or Gen IV reactors,² spherical oxides with a tailored size could help to improve the flowing capability of powders used in the fabrication process^{3,4} and could enhance the green density of the pellets before the sintering step. The design of uranium oxide particles with a controlled shape and size is also required to qualify the analytical methods used for nuclear safeguards.^{5,6} Uranium oxide nanospheres have already attracted some attention in view of their catalytic properties, particularly in the framework of the destruction of volatile organic compounds.⁷ Lastly, polycrystalline spherical particles can be used as

model samples to investigate the early stages of oxide ceramics sintering.^{8,9}

All these fields of research involve the preparation of spherical grains with variable size ranges, typically going from nanoparticles to sub-millimetric objects. For example, Hudry *et al.* recently prepared thorium, uranium, and even neptunium and plutonium oxide nanospheres (typically from 1 to 5 nm) from the thermal decomposition (280 °C – 30 min) of a mixture of actinide precursor solutions in highly coordinating organic media containing benzyl ether, oleic acid and oleylamine.^{10–12} On the other side of the length scale, the production of uranium oxide spheres, typically in the 100–500 μm range, has been reported for a long time from the initial precipitation of precursors through sol-gel processes.^{13,14} Protocols based on the impregnation of resin beads using solutions containing actinides and subsequent heat treatment at high temperature were also studied to initiate dustless processes for the elaboration of AnO_2 pellets.^{4,15}

Conversely, only a few studies were dedicated to the fabrication of spherical actinide oxide particles in an intermediate size range, *i.e.* from about 50 to 2000 nm. Physical methods such as aerosol spray pyrolysis appeared to be suitable for the production of monodisperse particles, as their size is directly controlled by the diameter of the droplets forming the aerosol.⁵ Nevertheless, this kind of process requires specific equipment, which might be difficult to implement in nuclear

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facilities aiming to work with radionuclides of high specific activity. On this basis, simple wet chemistry routes based on the hydrothermal precipitation of mixtures containing either thorium, uranium(IV) or uranium(VI) in acidic solutions, a precipitating agent (urea) and a shaping agent (PEG, glycerol) were recently considered.^{16–18} However, this often yielded porous and/or very small particles (typically from 50 to 250 nm in diameter). Also, Wang *et al.* reported accurate size tuning during the synthesis of ThO₂ nanoparticles,¹⁹ but the final diameter remained generally uncontrolled when dealing with UO₂ particles.

This study reports a new and simple wet chemistry route to spherical UO₂ particles from the initial precipitation of uranium(IV) aspartate and its conversion under hydrothermal conditions. Aspartic acid was chosen as a precipitating agent owing to its good complexation properties with regard to actinides.^{20,21} Also, it readily decomposes under hydrothermal conditions, mainly through deamination.²² The aspartate precursor was first characterized, then a multiparametric study of its conversion into uranium dioxide was undertaken, with the aim of accurately controlling both the size and the shape of the particles. Finally, the thermal behaviour of the UO₂ particles was investigated to evaluate the possible use of the particles as models for the sintering of nuclear fuels or standards in nuclear safeguard analyses.

2. Experimental

Caution! Uranium is an α -emitter natural radioelement and as such is considered a health risk. Its use requires appropriate infrastructure and staff training regarding the handling of radioactive materials.

Chemicals

All the reagents used were of analytical grade and supplied by Sigma-Aldrich, except uranium sources, while deionized water (18.2 M Ω cm at 25 °C) supplied by a Merck Milli-Q apparatus was used to prepare all aqueous solutions. The uranium tetrachloride solution was prepared by dissolving uranium metal chips (kindly supplied by CETAMA) in concentrated hydrochloric acid. The metal pieces were first rinsed with dichloromethane, acetone and water, washed in 1 M HCl in order to eliminate possible traces of uranium oxide present at the surface, and finally dissolved in 6 M HCl. The high chloride concentration allowed us to maintain the tetravalent oxidation state of uranium in solution for several months.²³ The uranium concentration of the final solution was estimated to be 0.52 ± 0.02 M (4.14×10^{-4} mol g⁻¹) by ICP-AES measurements.

Preparation of the samples

The synthesis protocol for spherical UO₂ particles was based on the complexation of U⁴⁺ with aspartic acid and on the subsequent hydrothermal conversion of the obtained precursor. It was derived from our previous work undertaken on an

aspartic acid/Th⁴⁺ system.²⁴ Aspartic acid (1–1.5 mmol) was first weighed and dissolved in 10 mL H₂O. The pH of the resulting solution was raised to 2 by addition of 30% NH₄OH. The hydrochloric solution containing tetravalent uranium (0.5 mmol) was then added dropwise, while controlling the pH with ammonia. The mixture of solutions rapidly led to the formation of a greenish precipitate. The concentration of uranium remaining in the supernatant was analysed by photo-electron rejecting alpha liquid scintillation (PERALS)²⁵ (Table 1) and revealed quantitative precipitation.

Afterwards, both the solid phase and supernatant were transferred into a Teflon-lined autoclave (Parr) and treated hydrothermally at 160 °C for 1 to 30 hours. For several tests, the reacting media were agitated during hydrothermal treatment using a magnetic stirrer (2mag – MIXcontrol 20). This set-up was chosen to be close to the unbaffled stirred tank reactors envisaged for precipitation reactions in the nuclear fuel reprocessing industry.²⁶ Whatever the operating conditions, the autoclave was cooled down overnight before collecting the final precipitate. The latter was separated by centrifugation, washed twice with deionized water and twice with ethanol, and finally dried overnight at 60 °C in an oven.

SEM observations

Before the SEM study, the samples were systematically dispersed in ethanol and deposited on a mirror-grade polished aluminium plate. Scanning electron microscope (SEM) micrographs were then directly recorded from the as-deposited powders without any additional preparation such as metallization. An FEI Quanta 200 environmental scanning electron microscope, equipped with an Everhart–Thornley detector (ETD) and a back-scattered electron detector (BSED), was used to record images with an acceleration voltage of 2 kV under high vacuum conditions.

SEM images were then processed using the ImageJ software²⁷ in order to determine the average particle size as well as the size distribution. In this aim, low-magnification (5000 \times) images were recorded with a resolution of 2048 \times 1768 pixels, leading to the analysis of populations of at least 300 particles.

PXRD

Powder X-ray diffraction (PXRD) diagrams were obtained by means of a Bruker D8 diffractometer equipped with a LYNXEYE detector adopting the reflection geometry and using Cu K $\alpha_{1,2}$ radiation ($\lambda = 1.54184$ Å). PXRD patterns were recorded at room temperature in the $5^\circ \leq 2\theta \leq 100^\circ$ range, with a step size of $\Delta(2\theta) = 0.01^\circ$ and a total counting time of about 3 hours per sample. In order to avoid any radioactive contamination, the powders were placed in a special sample holder (a dome-shaped container with an anti-scattering blade) that could generate an additional peak at around 6° (2θ).

Table 1 Determination of the uranium precipitation yield from PERALS analyses for various durations of the hydrothermal process ($T = 160\text{ }^{\circ}\text{C}$)

Heating time (hours)	$n(\text{U})$ initial (mol.)	$n(\text{U})$ supernatant (mol.)	Uranium precipitation yield (%)
0	5.55×10^{-4}	$(1.1 \pm 0.1) \times 10^{-5}$	98 ± 1
1	4.77×10^{-4}	$(6.0 \pm 0.1) \times 10^{-5}$	87 ± 2
3	4.78×10^{-4}	$(3.48 \pm 0.07) \times 10^{-5}$	93 ± 2
6	4.78×10^{-4}	$(1.68 \pm 0.02) \times 10^{-5}$	96 ± 1
24	4.77×10^{-4}	$(0.23 \pm 0.01) \times 10^{-5}$	99 ± 2
30	4.78×10^{-4}	$(0.34 \pm 0.01) \times 10^{-5}$	99 ± 2

NMR

MAS NMR measurements were performed on a Bruker Avance III 400 spectrometer, equipped with a 4 mm $^1\text{H}/\text{X}$ MAS probe. Samples were packed into zirconia rotors and spun at 12 kHz at room temperature. ^{13}C spectra were recorded using TMS (liq.) as a reference. ^{13}C CPMAS (cross polarization) experiments were performed with a contact time of 2 ms and a recycle delay of 5 s.

FTIR

Attenuated total reflectance Fourier transform infrared (ATR-FTIR) spectra were recorded in the $380\text{--}4000\text{ cm}^{-1}$ range thanks to a PerkinElmer FTIR Spectrum 100 device. Powdered samples were deposited at the surface of the single bounce diamond crystal ATR accessory without any prior preparation. The spectra collected under such operating conditions exhibited a resolution lower than 2 cm^{-1} .

TG-DT analyses

Thermogravimetric analyses were undertaken thanks to a Setaram SETSYS Evolution equipped with a type S thermocouple (Pt/Pt-10% Rh). After recording a baseline using an empty crucible ($100\text{ }\mu\text{L}$), weight loss was measured during a

heat treatment up to $1000\text{ }^{\circ}\text{C}$ with a rate of $5\text{ }^{\circ}\text{C min}^{-1}$ under an Ar atmosphere. Moreover, the gaseous species emitted during the heat treatment were identified by means of a Hiden Analytical QGA analyser using mass spectrometry.

3. Results and discussion

3.1. Characterization of the initial precipitate

The precipitate rapidly obtained after the addition of uranium(IV) hydrochloric solution to aspartic acid was first characterized through XRD, SEM and ^{13}C NMR. PXRD analysis led to a noisy pattern without any detectable diffraction line, *i.e.* characteristic of an amorphous and/or nanosized sample. This finding was supported by the SEM micrographs (see Fig. S1 supplied in the ESI†) which revealed that the powder was mainly composed of small crystallites of some tenths of nanometers in length, which assembled themselves in larger aggregates. This aggregation appeared to be widely disorganized and did not result in any specific morphology.

In order to identify the chemical nature of the precipitate formed, ^{13}C NMR spectra were recorded and compared to the reference spectrum obtained for aspartic acid (Fig. 1). The latter presents two signals at about 175 ppm which were assigned to the carboxylic moieties. Compared to the

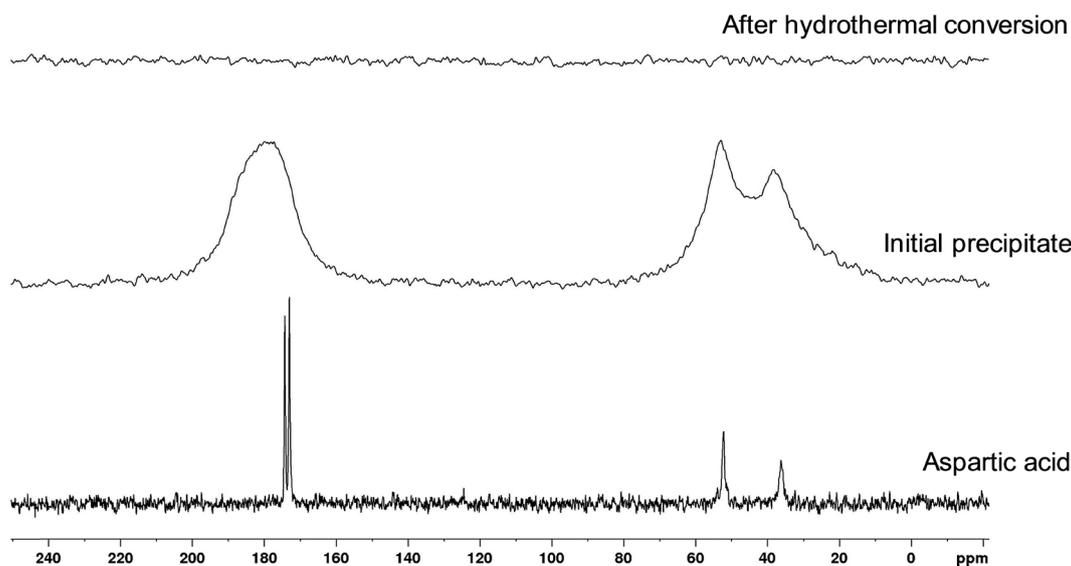


Fig. 1 ^{13}C NMR spectra of aspartic acid, of the precipitate obtained when mixed with U(IV) hydrochloric solution, and of the sample obtained after hydrothermal conversion ($T = 160\text{ }^{\circ}\text{C}$, $t = 30\text{ h}$).

reference, the spectrum collected for the unknown precipitate appeared dramatically widened, probably as a consequence of the paramagnetic character of U^{4+} and/or of the lack of sample crystallinity. Nevertheless, this large band occurred between 180 and 190 ppm, meaning that the two components shifted by about 15 ppm in comparison with the aspartic acid reference. This offset then evidences the complexation of uranium by the two carbonyl fragments, leading to the precipitation of uranium(IV) aspartate, probably as a coordination polymer. One must note that this behaviour seems to differ from that observed for thorium, for which only one carbonyl moiety was coordinated to the metal centre, resulting in a molecular species with the formula $Th(C_4NO_4H_6)_4 \cdot 4H_2O$.²⁴ On this basis, the formula of the initial precipitate should be probably written as $U(C_4NO_4H_5)_2 \cdot nH_2O$.

3.2. Multi-parametric study of hydrothermal conversion

With the aim of obtaining uranium oxide powders and to orient their morphology, the hydrothermal conversion of the amorphous U(IV) aspartate precursors was undertaken. Recent studies showed that mild hydrothermal conditions induce the decomposition of actinide carboxylates, such as uranium(IV) oxalate, leading to the direct formation of UO_2 in

solution.²⁸ As the latter results from the hydrolysis of U^{4+} after the initial decomposition of the precursor, the pH of the solution has to be carefully controlled. Preliminary tests were undertaken to determine the optimal acidity of the reacting media. For $pH \geq 3$, SEM micrographs revealed that the precipitates consist of strongly agglomerated sub-micrometric particles (Fig. 2), which probably formed through the fast nucleation of $U(OH)_4$ and further aging as partly amorphous UO_2 . Indeed, Rai *et al.* already reported that such a compound readily precipitates at $pH > 2$ at room temperature,²⁹ while the solubility of amorphous UO_2 does not sustain important variations up to 300 °C.³⁰ Under more acidic conditions, *i.e.* $pH = 1$ and 2, grains presenting a spherical habit were systematically obtained. However, they appeared to be denser and monodisperse (although significantly smaller) for $pH = 2$. This condition was then retained for the rest of our study. Similarly, we decided to maintain the temperature at 160 °C in accordance with the protocol originally designed for the thorium counterpart.²⁴ In order to optimize the operating parameters, a multi-parametric study combining the effects of heating time at 160 °C, of the initial metal/ligand ratio and of magnetic stirring was then undertaken.

3.2.1. Effect of heating time. In a first step, the hydrothermal treatment at 160 °C was extended from 1 to 30 hours. For all the heating times considered, the supernatant was

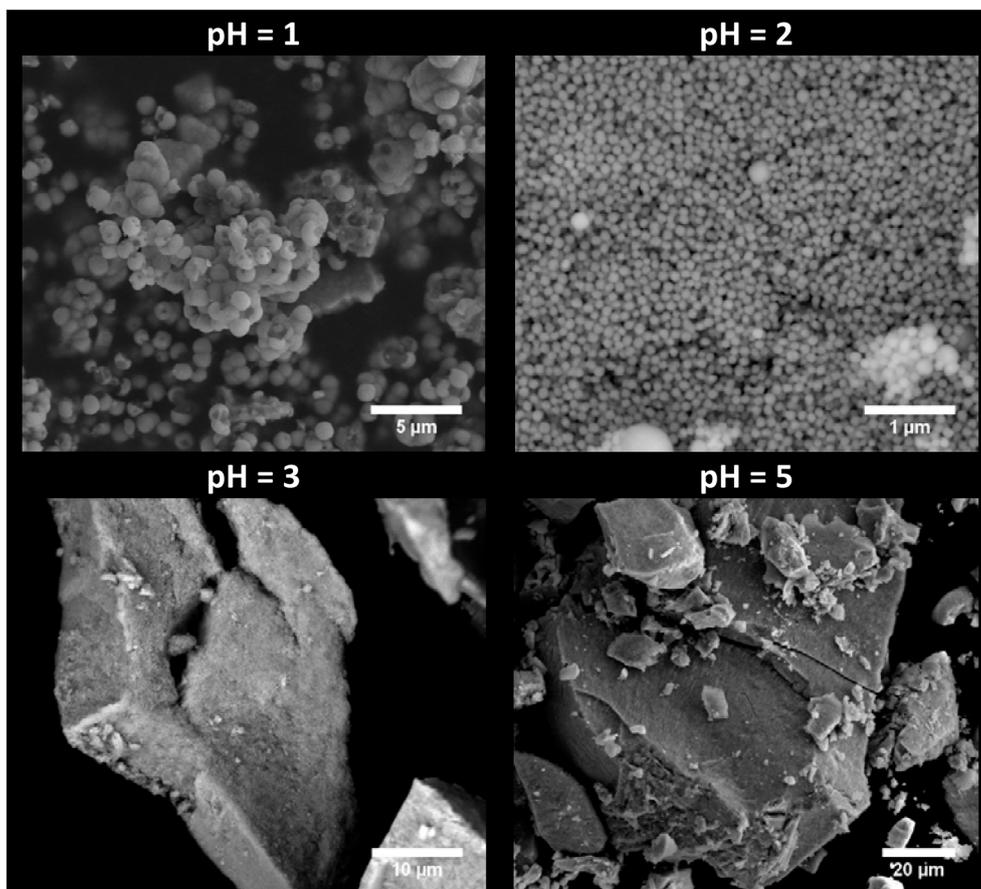
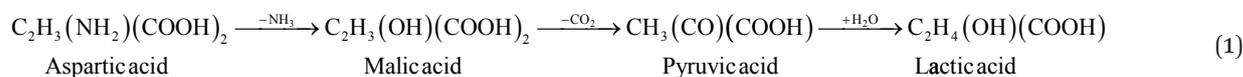


Fig. 2 SEM micrographs of the powders obtained after hydrothermal conversion of uranium(IV) aspartate at 160 °C for different pH values.

analysed by PERALS²⁵ to quantify the concentration of uranium remaining in solution (Table 1). The uranium precipitation yield was found to increase slightly during the first hours of hydrothermal treatment, uranium precipitation becoming quantitative after heating over 24 hours. On this basis, the beginning of the treatment (*i.e.* up to 6 hours) probably still corresponds to the decomposition of the initial U(IV) aspartate precursor, followed by the partial hydrolysis of the tetravalent uranium in solution. This observation is in very good agreement with the data reported by Faisal *et al.* concerning the decomposition of aspartic acid under hydrothermal conditions.³¹ These authors stated that the main reaction pathway consists in an irreversible deamination yielding malic acid, followed by a possible decarboxylation leading to a mixture of pyruvic and lactic acids:



Moreover, based on their determination of kinetic constants at 200 °C and of the activation energy associated with the decomposition reaction ($E_A = 144 \text{ kJ mol}^{-1}$), the time required to decompose 99% of the initial aspartic acid can be estimated to be about 5.5 hours at 160 °C. Shorter heating times then should be discarded if one wants to quantitatively recover U⁴⁺ from the decomposition of the aspartate precursor under these operating conditions.

Concomitantly to the increase of the uranium precipitation yield, the powder morphology appeared to be strongly modified when the duration of the hydrothermal process was increased. After 1 hour, the morphology was very close to that observed for the aspartate precursor, with a powder composed of sub-micrometric crystallites (Fig. 3). This was still found to be the major habit observed between 6 and 24 hours of hydrothermal treatment, but some scarce spherical grains tended to form. After 30 hours, the morphology of the sample drastically changed. The powder appeared to consist of spherical grains with variable sizes, typically ranging from 200 nm to 1 μm in diameter. On this basis, the operating conditions for the hydrothermal treatment were fixed to 30 hours and $T = 160 \text{ °C}$ for all the subsequent experiments.

The changes observed in the morphology of the samples were also associated with structural modifications. Whatever the conditions chosen, the XRD patterns collected present low-intensity and wide peaks characteristic of poorly crystallized and/or nanosized samples, in good agreement with the SEM observations (Fig. 4). Despite the low resolution of the XRD patterns, it was still possible to note an evolution of the structure along the hydrothermal treatment. For samples heated between 1 and 3 hours, the XRD pattern was very close to that previously reported for thorium aspartate.²⁴ Compared with the NMR data, which rather suggested the initial precipitation of an amorphous uranium(IV) aspartate coordination polymer, this result indicates that a transformation towards a molecular species probably occurred during

the first hours of heat treatment at 160 °C. Under these conditions, short hydrothermal treatment did not allow the conversion of the precursor into oxide but only resulted in the transformation and partial crystallization of the initial precipitate. Conversely, after 6 hours of heating at 160 °C and for longer hydrothermal treatments, the characteristic XRD pattern of nanosized fluorite-type UO₂ was evidenced³² while the peaks of U(IV) aspartate were no longer observed. This indicates full conversion under these operating conditions and again appears in good agreement with previous studies dedicated to aspartic acid decomposition under hydrothermal conditions.³¹ Moreover, the width of the XRD lines was not significantly modified over time, meaning that the average crystallite size remained almost unchanged. The spherical particles observed after 30 hours of hydrothermal processing

are thus formed by smaller entities, resulting in a hierarchical structure.¹⁶

Complementarily, the FTIR spectrum of the sample heated for 30 hours revealed numerous vibration bands evidencing the presence of residual organics at the surface and/or within the UO₂ particles (Fig. S2†). Also, the comparison to the reference spectra of aspartic acid and thorium aspartate, already discussed in our previous work,²⁴ showed the absence of vibration bands at around 1475 and 3120 cm⁻¹, which are characteristic of the amine functional group and were reported in the literature for uranium(IV)-aspartate complexes.³³ This feature is in good agreement with the reaction path previously described for the decomposition of aspartic acid under hydrothermal conditions,³¹ during which deamination is the predominant reaction. On this basis, the organics remaining in the sample are probably dicarboxylic acids resulting from the decomposition of aspartic acid, *i.e.* malic, pyruvic and/or lactic acid. This hypothesis is corroborated by the existence of large vibration bands at about 1400 and 1600 cm⁻¹, which can be assigned to carboxylic acids. At the same time, ¹³C NMR spectra did not show any signal (Fig. 1), possibly due to limited amounts of different organic species still surrounding the UO₂ sample after the conversion of uranium(IV) aspartate.

3.2.2. Effect of uranium/aspartic acid stoichiometry. As the organics resulting from the hydrothermal decomposition of the uranium(IV) aspartate precursor could play a role in the oxide's final morphology, the metal/ligand molar ratio (*i.e.* the uranium/aspartic acid molar ratio) in the initial reacting media was varied in order to evaluate its influence on the particle shape and size distribution. First, hydrothermal treatment of the hydrochloric solution of U(IV) was performed at pH = 2 without any addition of aspartic acid. Under such operating conditions, a small amount of precipitate was obtained after heating for 30 hours at 160 °C. It had wide XRD peaks characteristic of the UO₂ fluorite-type structure (Fig. S3†), confirming that the hydrolysis of U⁴⁺ was the key

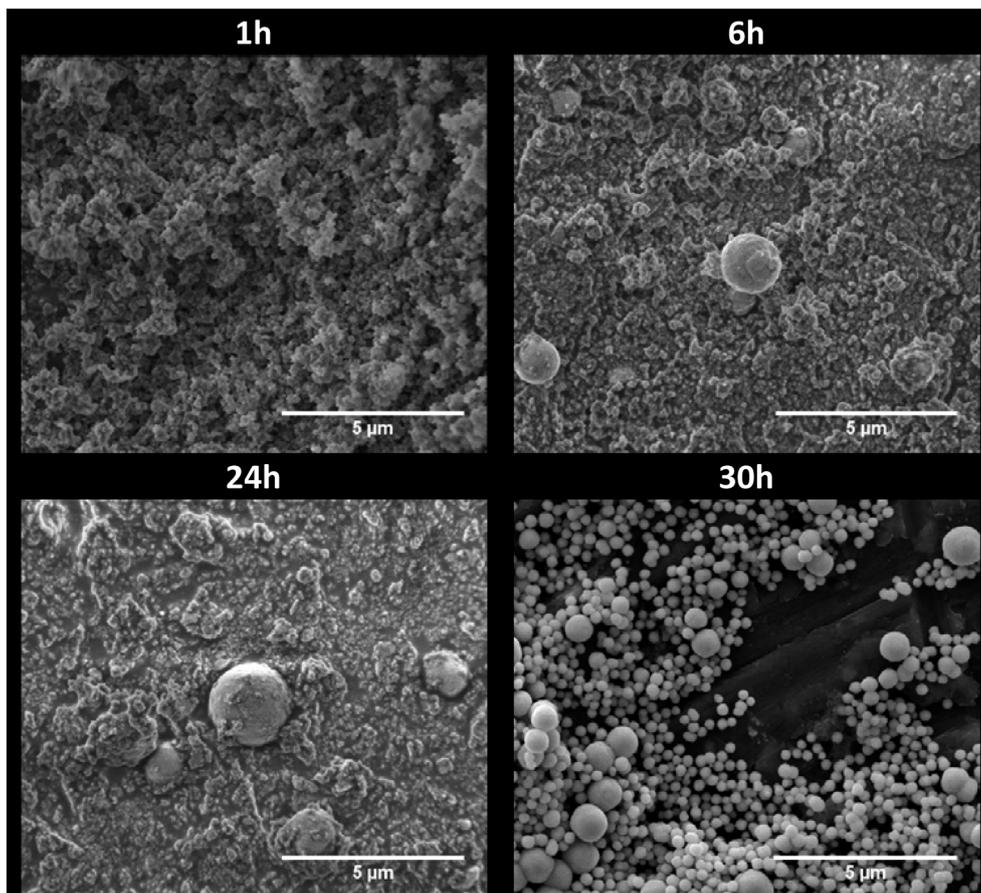


Fig. 3 SEM micrographs of the powders obtained after hydrothermal conversion of uranium(IV) aspartate at 160 °C for different durations.

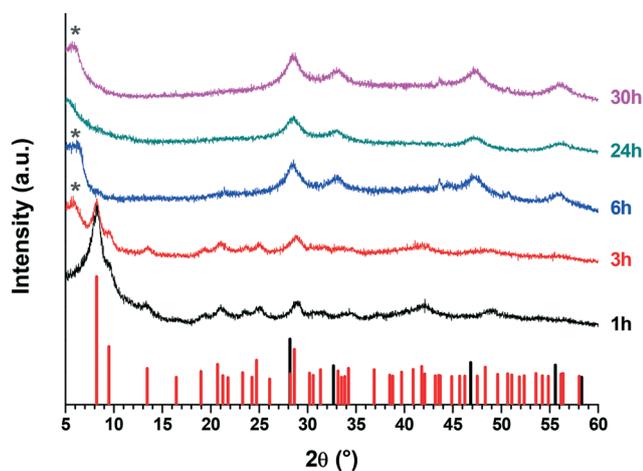


Fig. 4 XRD patterns of the samples obtained after hydrothermal conversion of uranium(IV) aspartate at 160 °C for various durations. Diffraction lines of thorium aspartate (red)²⁴ and UO₂ (black, ICDD #01-071-4823) are supplied for comparison. The * symbol indicates the signal due to the anti-contamination dome-shaped container.

step in the formation of an oxide precipitate. Nevertheless, the powder appeared to be composed of very small grains (Fig. 5), comparable to the nanoparticles obtained from synthesis *via* precipitation of hydroxides.³²

On the other hand, the introduction of aspartic acid into the reactant mixture, leading to the formation of the uranium(IV) aspartate precursor, was systematically associated with the formation of spherical UO₂ particles whatever the initial content considered. The residual organic molecules detected in the final oxide particles then probably act as stabilizing agents in the orientation of the sample's morphology towards spheres. Nevertheless, the M/L molar ratio appeared to strongly impact the size distribution of the particles. Indeed, under stoichiometric conditions (*i.e.* M/L = 1/2, based on the initial precipitation of U(C₄NO₄H₅)₂·nH₂O) and for an excess of 25% (M/L = 1/2.5), the powder was found to be poly-disperse. The powders were composed of a large population of small particles (typically between 20 and 150 nm in diameter) but included some larger spheres up to 700 nm in diameter. The latter do not account for a large number of particles but represent the main part of the volume (see Fig. S4 supplied in the ESI†). On the other hand, an excess of 50% of aspartic acid (*i.e.* M/L = 1/3) led to a very homogeneous powder. The statistical analysis of the particles showed that the size distribution basically follows the Gaussian law, associated with an average diameter of about 380 nm and a FWHM of about 60 nm. Owing to a dispersion less than ±10%, the powder can be considered as monodisperse. One can also note that increasing the initial excess in aspartic acid

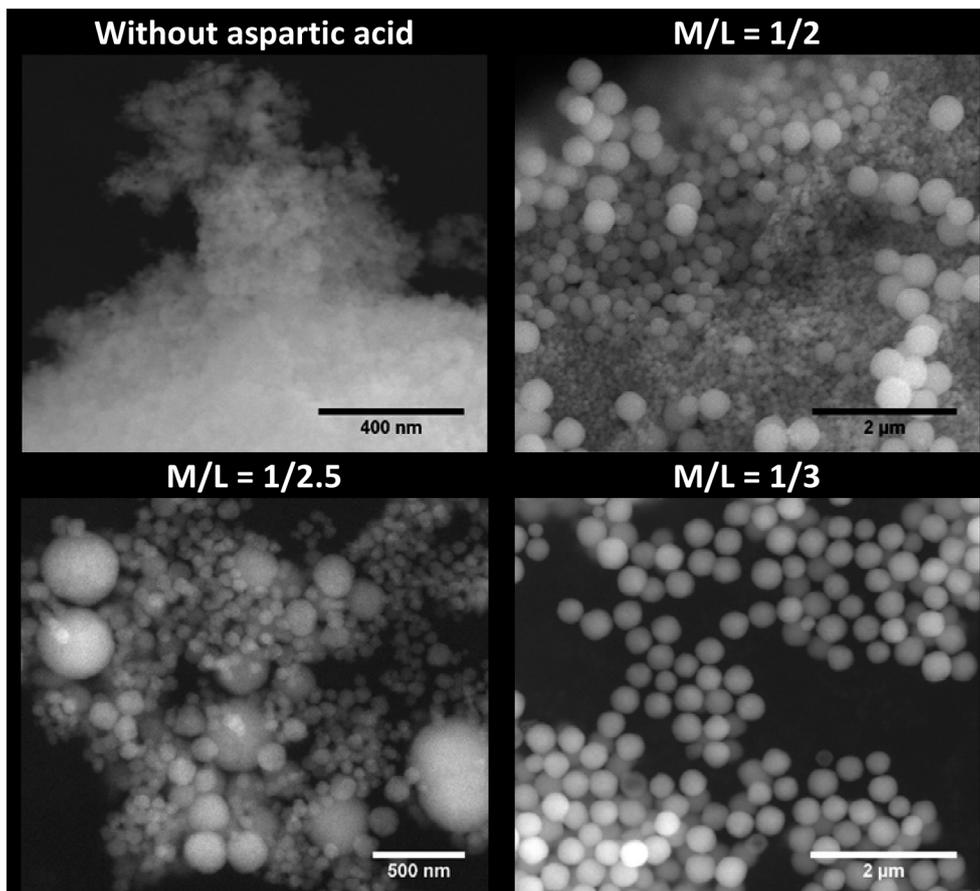


Fig. 5 SEM micrographs of the powders obtained after hydrothermal conversion of uranium(IV) aspartate ($T = 160\text{ }^{\circ}\text{C}$; $t = 30\text{ hours}$) for different metal/ligand ratios.

increases the average diameter of the main particle population, typically from around 100 nm under stoichiometric conditions to about 400 nm for 50% excess. On this basis, a molar ratio of 1/3 between uranium(IV) and aspartic acid was considered for the rest of the study in order to recover samples with a homogenous particle size.

3.2.3. Effect of magnetic stirring. In a last step, magnetic stirring was applied during the hydrothermal treatment with the aim of driving the aggregation of the crystallites towards spherical particles and to have accurate control over the final size of the particles. In order to accurately quantify the effects of stirring through a dimensionless number and characterize the flow in the reactor as laminar or turbulent, the impeller Reynolds number (derived from the conventional Reynolds number used to describe flow in pipes) was determined for each stirring speed studied in the range 200–1100 rpm as follows:

$$\text{Re}_a = \frac{N \times D^2}{\nu} \quad (2)$$

where N is the impeller (stirrer bar) rotation speed, D is the stirrer bar length and ν is the kinematic viscosity. Herein, the liquid medium was assumed to be pure water and we extracted the viscosity value from the study of Deguchi *et al.* with $\nu \approx 1.8$

$\times 10^{-7}\text{ m}^2\text{ s}^{-1}$ at $160\text{ }^{\circ}\text{C}$.³⁴ The stirrer bar length (12 mm) was chosen to be about half the diameter of the reactor (25 mm), in order to create two mixing zones, *i.e.* free and forced vortices, as described by Bertrand *et al.*³⁵

On this basis, smooth stirring up to $N \approx 200\text{ rpm}$ corresponds to $\text{Re}_a < 2000$ characteristic of a laminar flow. In this case, the particle size increased from about 400 nm (no stirring) to 1200 nm (Fig. 6a, selected SEM micrographs provided in the ESI,† Fig. S5), meaning that the establishment of a laminar flow in the reactor likely acted in favour of the agglomeration of the elementary crystallites. Between 200 and 800 rpm (*i.e.* $2000 < \text{Re}_a < 10\,000$), the average particle diameter was found to stabilize between 1.5 and 2 μm . However, the size distribution appeared to be very wide, meaning that polydisperse powders were obtained due to the establishment of a transition regime between laminar and turbulent flows. Finally, higher agitation speeds led to $\text{Re}_a > 10^4$, characteristic of a turbulent flow. In this case, a clear linear relationship was established between Re_a (stirring speed) and the average diameter of the spherical UO_2 particles (Fig. 6b). Indeed, the size of the particles was found to decrease from around 2500 nm to 400 nm when the stirring speed was increased in the 900–1100 rpm range. In this case, the interactive force assembling the crystallites as spherical particles was surpassed by

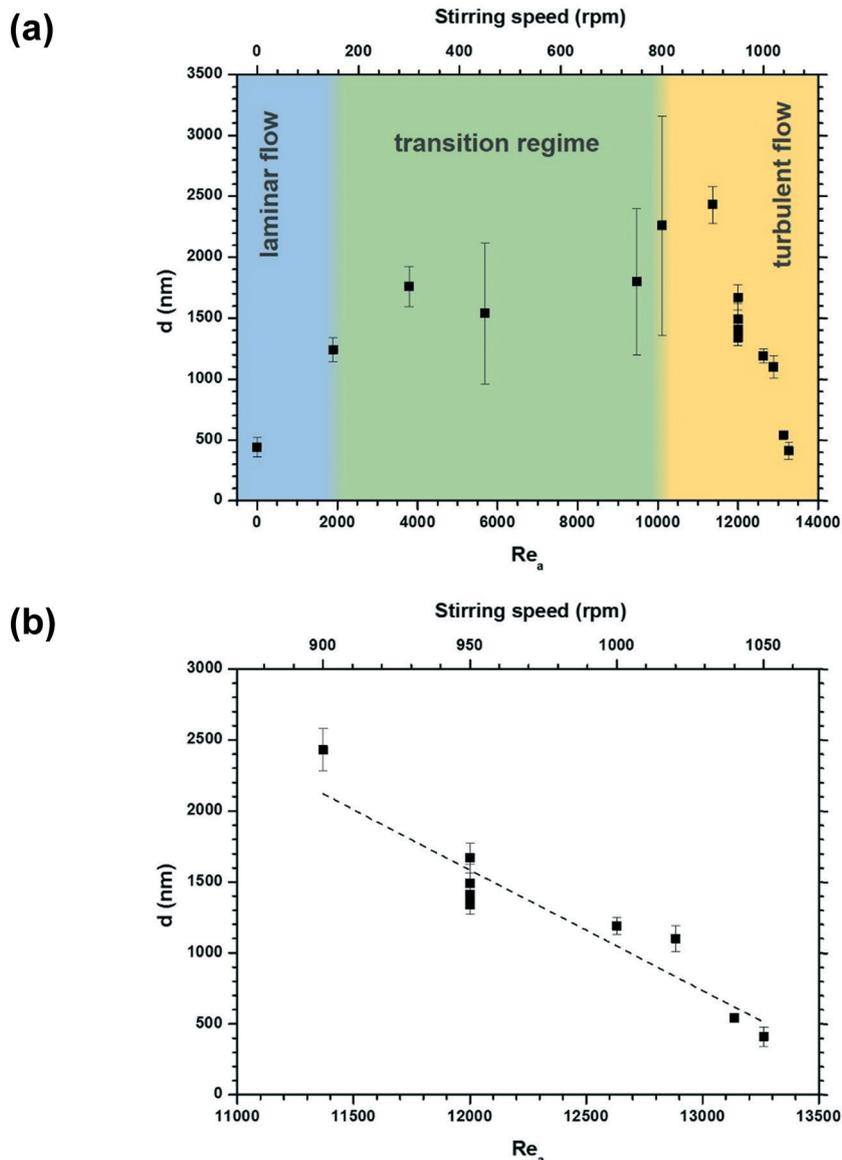


Fig. 6 Variation of the average diameter of UO_2 spherical particles versus the impeller Reynolds number and the corresponding magnetic stirrer speed (a), and focus on the turbulent flow region (b) ($T = 160\text{ }^\circ\text{C}$; $t = 30$ hours; $M/L = 1/3$).

the turbulent shear stress. Growth of the particles was then hindered even though aggregation of the crystallites is favoured by the increase in the agitation speed.³⁶ Lastly, one must note that three different batches were prepared at 950 rpm ($Re_a = 12\ 000$) to verify the reproducibility of the process developed. Under these conditions, all the powders exhibited very similar average particle diameters (*i.e.* 1340 ± 65 , 1390 ± 80 and 1410 ± 50 nm) which confirms the accuracy of the size control provided by the magnetic stirring during the hydrothermal conversion.

3.3. Thermal behaviour of UO_2 particles

As the spherical UO_2 particles prepared following the protocol detailed previously still exhibit traces of organics and/or water,

an additional heating step was envisaged to recover pure UO_2 . The thermal behaviour was then investigated under an argon atmosphere through TG analysis and SEM observations. The thermogram displayed in Fig. 7 shows that the weight loss mainly occurred through two distinct steps. Based on MS measurements, the gases eliminated were associated with the dehydration of the sample, on the one hand, and with the thermal decomposition of the residual organics, on the other. The first one, observed between 100 and 200 $^\circ\text{C}$, corresponds to the elimination of 0.2 water molecules per formula unit. This low hydration is in good agreement with the hypothesis of formation of $\text{U}(\text{OH})_4$ followed by aging into $\text{UO}_{2-n}\text{H}_2\text{O}$. Indeed, the number of water molecules reported in the literature for such compounds, which can also be written as $\text{UO}_{(2-n)}(\text{OH})_{2n}$ hydrated oxyhydroxide, is generally between 0 and 2.³⁰

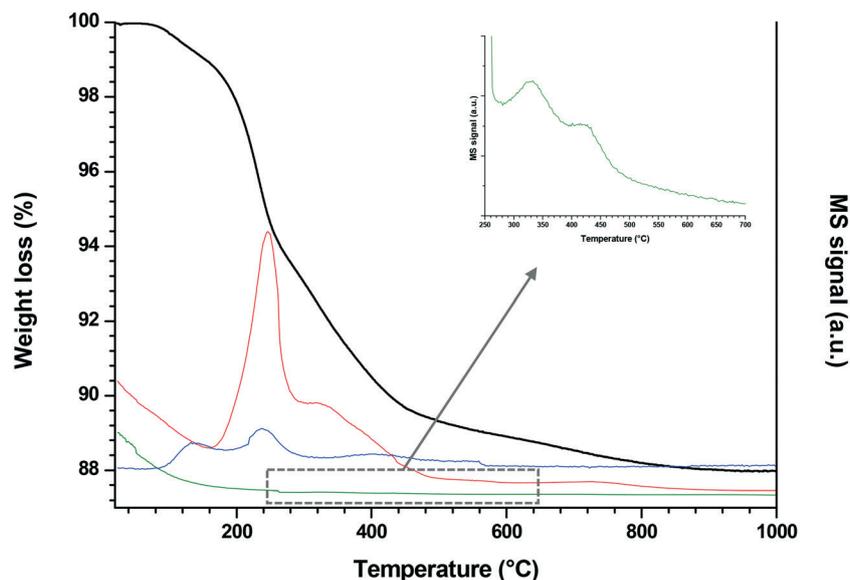


Fig. 7 TG analyses and associated MS signals of water (blue), CO + CO₂ (red) and NO + NO₂ (green) for UO₂ spherical particles prepared after hydrothermal conversion of uranium(IV) aspartate ($T = 160\text{ }^{\circ}\text{C}$; $t = 30\text{ hours}$; $M/L = 1/3$).

The second weight loss recorded between 200 and 400 °C, mainly correlated with the emission of H₂O, CO and CO₂, is by far the most important (around 10% of the initial mass) and indicates the presence of residual organics in the system. Moreover, a slight change in the slope shows that the decomposition occurred in two steps, which can be associated with the presence of different organic molecules within the sample and/or to the existence of intermediates during their thermal decomposition. For example, such behaviour was observed during the decomposition of actinide oxalates, with the formation of carbonate or oxocarbonate compounds.^{37,38} Additionally, only very limited amounts of NO and NO₂ were detected during the thermal decomposition of the organic part of the solid, which confirms that it is mainly composed of carboxylic acid resulting from the deamination of the initial aspartate groups.

Finally, it is worth noting that a slight weight loss signal was recorded between 500 and 1000 °C, associated with the detection of small amounts of CO and CO₂, which could be assigned to the elimination of residual carbon in the sample. Indeed, the thermal decomposition of MO₂ precursors, including carboxylates such as oxalates, was systematically associated with the presence of amorphous carbon within the final oxides.^{39,40} On this basis, it can be concluded that residual organics are not only present at the surface of the spherical UO₂ particles where they act as shaping agents but also within the spheres themselves, potentially adsorbed onto the elementary crystallites. Nevertheless, the amount of carbon still present in the samples remains low at this stage, and one can consider that pure UO₂ was obtained after heating above 600 °C.

As heating the samples is required to recover anhydrous and organic-free UO₂ samples, SEM observations were undertaken on powders heated between 600 and 1200 °C to study potential evolution in morphology (Fig. 8). At 600 °C, despite

the loss of water molecules and residual organics which account for about 10% of the initial weight, the spherical morphology of the particles is maintained. The more likely hypothesis is that H₂O, CO and CO₂ gaseous molecules can be evacuated through small pores remaining between the crystallites constituting the aggregates. Indeed, such mesoporosity has been previously identified in analogous systems such as CeO₂ and ThO₂ microspheres.^{9,41}

In a second step, the morphology of the aggregates slightly evolved up to 1000 °C as a consequence of crystallite growth within the particles. Once again, this two-step behaviour is in good agreement with data reported in the literature. Particularly, Claparede *et al.*^{42,43} showed that for MO₂ fluorite-type compounds (with $M = \text{Ce}$ and/or Th), the average crystallite size first slowly increases between room temperature and 700 °C through defect elimination. Conversely, the solid state diffusion processes involved in crystallite growth are triggered by higher temperature, resulting in a rapid increase in average size. In this study, such a phenomenon led to a blackberry-like morphology with crystallites of about 20–50 nm at the surface of the spheres. Finally, the spherical morphology of the particles was kept even after heating the powders up to 1200 °C. Nevertheless, such a high temperature led to the initiation of sintering phenomena, as attested by the observation of numerous necks between the particles on the SEM micrograph. On this basis, if one wants to recover individual particles, for example for analytical purposes, heating above 800/1000 °C should be avoided to limit coalescence.

4. Conclusion

An easy wet chemistry route to prepare submicrometric spherical UO₂ particles was designed through the

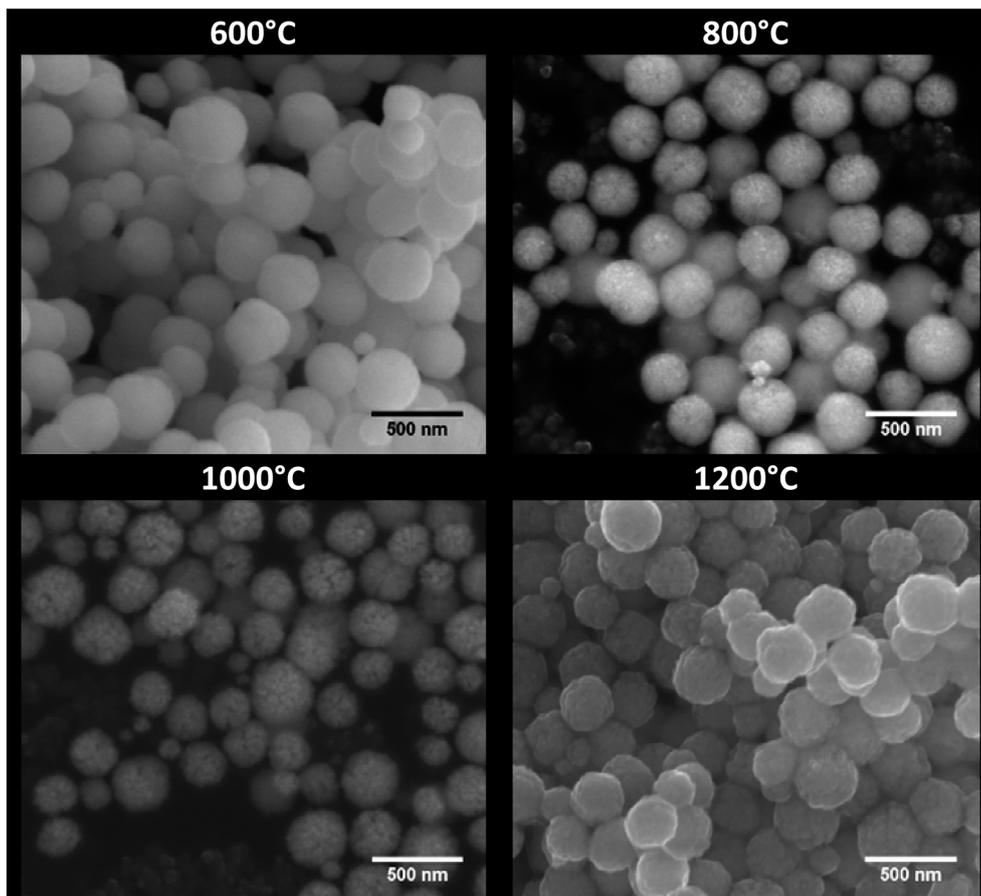


Fig. 8 SEM observations of the UO_2 spherical particles prepared after hydrothermal conversion of uranium(IV) aspartate ($T = 160\text{ }^\circ\text{C}$; $t = 30$ hours; $M/L = 1/3$) then heating between 600 and 1200 $^\circ\text{C}$ under an argon atmosphere.

hydrothermal conversion of uranium(IV) aspartate. An initial amorphous precursor was obtained by mixing a hydrochloric solution containing tetravalent uranium with aspartic acid, resulting in quantitative precipitation. The conversion towards UO_2 precipitates was then operated *via* heat treatment under mild hydrothermal conditions at $T = 160\text{ }^\circ\text{C}$. A multiparametric study then allowed us to specify the operating conditions leading to monodisperse and size-controlled particles. 6 hours of hydrothermal treatment were found necessary to completely convert the precursor into dioxide. Moreover, additional heating up to 30 hours appeared to be mandatory in order to recover only spherical particles. Their size distribution appeared to be driven by the amount of organics in the solution, *i.e.* by the initial uranium/aspartic acid molar ratio. On this basis, a minimal excess of 50% of ligand ($M/L = 1/3$) was required to obtain monodisperse powders, with a typical variation of about $\pm 10\%$ on the average diameter. Finally, it was possible to control the particle size by applying magnetic stirring during the hydrothermal treatment. In a turbulent flow regime, *i.e.* impeller Reynolds numbers above 10^4 , a linear variation of the particle diameter *versus* the stirring speed was observed, with sizes ranging from 400 nm (1050 rpm) to 2500 nm (900 rpm), which paves the way to applications of such materials in various scientific areas. Also,

the protocol was found to be robust and reproducible, with only very limited size variation from one batch to another.

Such a one-pot wet chemistry route appears to be very promising for the production of spherical UO_2 particles and can be simply implemented in any nuclear chemistry lab. Additional experiments are now in progress to extend this methodology to other systems of interest, including mixed oxides such as $(\text{U,Th})\text{O}_2$, $(\text{U,Ce})\text{O}_2$ and $(\text{U,Ln}^{\text{III}})\text{O}_{2-x}$.

Conflicts of interest

There are no conflicts to declare.

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References

- M. Pradhan, S. Sarkar, A. K. Sinha, M. Basu and T. Pal, Morphology controlled uranium oxide hydroxide hydrate for catalysis, luminescence and SERS studies, *CrystEngComm*, 2011, **13**, 2878–2889.
- J. E. Kelly, Generation IV International Forum: A decade of progress through international cooperation, *Prog. Nucl. Energy*, 2014, **77**, 240–246.
- H. H. S. Santana, G. Maier and J. Rodenas, Flowability analysis of uranium dioxide powder at different temperatures containing different lubricants, *Appl. Radiat. Isot.*, 2011, **69**, 1162–1164.
- E. Remy, S. Picart, S. Grandjean, T. Delahaye, N. Herlet, P. Allegri, O. Dugne, R. Podor, N. Clavier, P. Blanchart and A. Ayrat, Calcined resin microsphere pelletization (CRMP): A novel process for sintered metallic oxide pellets, *J. Eur. Ceram. Soc.*, 2012, **32**, 3199–3209.
- R. Middendorp, M. Durr, A. Knott, F. Pointurier, D. F. Sanchez, V. Samson and D. Grolimund, Characterization of the Aerosol-Based Synthesis of Uranium Particles as a Potential Reference Material for Microanalytical Methods, *Anal. Chem.*, 2017, **89**, 4721–4728.
- R. Middendorp, M. Durr and D. Bosbach, The stability of uranium microspheres for future application as reference standard in analytical measurements, *Procedia Chem.*, 2016, **21**, 285–292.
- Q. Wang, G. D. Li, S. Xu, J. X. Li and J. S. Chen, Synthesis of uranium oxide nanoparticles and their catalytic performance for benzyl alcohol conversion to benzaldehyde, *J. Mater. Chem.*, 2008, **18**, 1146–1152.
- G. I. Nkou Bouala, N. Clavier, S. Martin, J. Lechelle, J. Favrichon, H. P. Brau, N. Dacheux and R. Podor, From in Situ HT-ESEM Observations to Simulation: How Does Polycrystallinity Affects the Sintering of CeO₂ Microspheres?, *J. Phys. Chem. C*, 2016, **120**, 386–395.
- G. I. Nkou Bouala, N. Clavier, J. Lechelle, J. Monnier, C. Ricolleau, N. Dacheux and R. Podor, High-temperature electron microscopy study of ThO₂ microspheres sintering, *J. Eur. Ceram. Soc.*, 2016, **37**, 727–738.
- D. Hudry, C. Apostolidis, O. Walter, T. Gouder, E. Courtois, C. Kubel and D. Meyer, Non-aqueous Synthesis of Isotropic and Anisotropic Actinide Oxide Nanocrystals, *Chem. – Eur. J.*, 2012, **18**, 8283–8287.
- D. Hudry, C. Apostolidis, O. Walter, T. Gouder, A. Janssen, E. Courtois, C. Kubel and D. Meyer, Synthesis of transuranium-based nanocrystals via the thermal decomposition of actinyl nitrates, *RSC Adv.*, 2013, **3**, 18271–18274.
- D. Hudry, C. Apostolidis, O. Walter, A. Janssen, D. Manara, J. C. Griveau, E. Colineau, T. Vitova, T. Prussmann, D. Wang, C. Kubel and D. Meyer, Ultra-Small Plutonium Oxide Nanocrystals: An Innovative Material in Plutonium Science, *Chem. – Eur. J.*, 2014, **20**, 10431–10438.
- H. Daniels, S. Neumeier, A. A. Bukaemskiy, G. Modolo and D. Bosbach, Fabrication of oxidic uranium-neodymium microspheres by internal gelation, *Prog. Nucl. Energy*, 2012, **57**, 106–110.
- A. Kumar, J. Radhakrishna, N. Kumar, R. V. Pai, J. V. Dehadrai, A. C. Deb and S. K. Mukerjee, Studies on preparation of (U_{0.47}Pu_{0.53})O₂ microspheres by internal gelation process, *J. Nucl. Mater.*, 2013, **434**, 162–169.
- S. Picart, P. Parant, M. Caisso, E. Remy, H. Mokhtari, I. Jobelin, J. P. Bayle, C. L. Martin, P. Blanchart, A. Ayrat and T. Delahaye, Porous metal oxide microspheres from ion exchange resin, *Eur. Phys. J.: Spec. Top.*, 2015, **224**, 1675–1687.
- G. I. Nkou Bouala, N. Clavier, R. Podor, J. Cambedouzou, A. Mesbah, H. P. Brau, J. Lechelle and N. Dacheux, Preparation and characterisation of uranium oxides with spherical shapes and hierarchical structures, *CrystEngComm*, 2014, **16**, 6944–6954.
- L. Wang, R. Zhao, C. Z. Wang, L. Y. Yuan, Z. J. Gu, C. L. Xiao, S. A. Wang, X. W. Wang, Y. L. Zhao, Z. F. Chai and W. Q. Shi, Template-Free Synthesis and Mechanistic Study of Porous Three-Dimensional Hierarchical Uranium-Containing and Uranium Oxide Microspheres, *Chem. – Eur. J.*, 2014, **20**, 12655–12662.
- R. Zhao, L. Wang, Z. J. Gu, L. Y. Yuan, C. L. Xiao, Y. L. Zhao, Z. F. Chai and W. Q. Shi, A facile additive-free method for tunable fabrication of UO₂ and U₃O₈ nanoparticles in aqueous solution, *CrystEngComm*, 2014, **16**, 2645–2651.
- L. Wang, R. Zhao, X. W. Wang, L. Mei, L. Y. Yuan, S. A. Wang, Z. F. Chai and W. Q. Shi, Size-tunable synthesis of monodisperse thorium dioxide nanoparticles and their performance on the adsorption of dye molecules, *CrystEngComm*, 2014, **16**, 10469–10475.
- Y. Andres, H. J. MacCordick and J. C. Hubert, Selective biosorption of thorium ions by an immobilized mycobacterial biomass, *Appl. Microbiol. Biotechnol.*, 1995, **44**, 271–276.
- D. Humelnicu, G. Drochioiu and K. Popa, Bioaccumulation of thorium and uranyl ions on *Saccharomyces cerevisiae*, *J. Radioanal. Nucl. Chem.*, 2004, **260**, 291–293.
- N. Sato, A. T. Quitain, K. Kang, H. Daimon and K. Fujie, Reaction kinetics of amino acid decomposition in high-temperature and high-pressure water, *Ind. Eng. Chem. Res.*, 2004, **43**, 3217–3222.
- N. Dacheux, V. Brandel and M. Genet, Synthesis and Properties of Uranium Chloride Phosphate Tetrahydrate - UClPO₄·4H₂O, *New J. Chem.*, 1995, **19**, 1029–1036.
- N. Clavier, J. Maynadie, A. Mesbah, J. Hidalgo, R. Lauwerier, G. I. Nkou Bouala, S. Parres-Maynadie, D. Meyer, N. Dacheux and R. Podor, Thorium aspartate tetrahydrate precursor to ThO₂: Comparison of hydrothermal and thermal conversions, *J. Nucl. Mater.*, 2017, **487**, 331–342.
- N. Dacheux and J. Aupiais, Determination of uranium, thorium, plutonium, americium, and curium ultratraces by photon electron rejecting alpha liquid scintillation, *Anal. Chem.*, 1997, **69**, 2275–2282.
- N. Lamarque, B. Zoppe, O. Lebaigue, Y. Dolias, M. Bertrand and F. Ducros, Large-eddy simulation of the turbulent free-

- surface flow in an unbaffled stirred tank reactor, *Chem. Eng. Sci.*, 2010, **65**, 4307–4322.
- 27 C. A. Schneider, W. S. Rasband and K. W. Eliceiri, NIH Image to ImageJ: 25 years of image analysis, *Nat. Methods*, 2012, **9**, 671–675.
- 28 O. Walter, K. Popa and O. Dieste Blanco, Hydrothermal decomposition of actinide(IV) oxalates: a new aqueous route towards reactive actinide oxide nanocrystals, *Open Chem.*, 2016, **14**, 170–174.
- 29 D. Rai, A. R. Felmy and J. L. Ryan, Uranium(IV) Hydrolysis Constants and Solubility Product of $\text{UO}_2 \cdot x\text{H}_2\text{O}(\text{am})$, *Inorg. Chem.*, 1990, **29**, 260–264.
- 30 V. Neck and J. I. Kim, Solubility and hydrolysis of tetravalent actinides, *Radiochim. Acta*, 2001, **89**, 1–16.
- 31 M. Faisal, N. Sato, A. T. Quitain, H. Daimon and K. Fujie, Reaction kinetics and pathway of hydrothermal decomposition of aspartic acid, *Int. J. Chem. Kinet.*, 2007, **39**, 175–180.
- 32 J. Martinez, N. Clavier, A. Mesbah, F. Audubert, X. F. Le Goff, N. Vigier and N. Dacheux, An original precipitation route toward the preparation and the sintering of highly reactive uranium cerium dioxide powders, *J. Nucl. Mater.*, 2015, **462**, 173–181.
- 33 M. Nazir and I. I. Naqvi, Synthesis and characterization of uranium (IV) complexes with various amino acids, *J. Saudi Chem. Soc.*, 2010, **14**, 101–104.
- 34 S. Deguchi, S. K. Ghosh, R. G. Alargova and K. Tsujii, Viscosity measurements of water at high temperatures and pressures using dynamic light scattering, *J. Phys. Chem. B*, 2006, **110**, 18358–18362.
- 35 M. Bertrand, E. Plasari, O. Lebaigue, P. Baron, N. Lamarque and F. Ducros, Hybrid LES-multizonal modelling of the uranium oxalate precipitation, *Chem. Eng. Sci.*, 2012, **77**, 95–104.
- 36 M. H. Sung, I. S. Choi, J. S. Kim and W. S. Kim, Agglomeration of yttrium oxalate particles produced by reaction precipitation in semi-batch reactor, *Chem. Eng. Sci.*, 2000, **55**, 2173–2184.
- 37 G. K. Cilgi, H. Cetisli and R. Donat, Thermal and kinetic analysis of uranium salts, *J. Therm. Anal. Calorim.*, 2014, **115**, 2007–2020.
- 38 R. M. Orr, H. E. Sims and R. J. Taylor, A review of plutonium oxalate decomposition reactions and effects of decomposition temperature on the surface area of the plutonium dioxide product, *J. Nucl. Mater.*, 2015, **465**, 756–773.
- 39 J. Martinez, N. Clavier, T. Ducasse, A. Mesbah, F. Audubert, B. Corso, N. Vigier and N. Dacheux, From uranium(IV) oxalate to sintered UO_2 : consequences of the powders' thermal history on the microstructure, *J. Eur. Ceram. Soc.*, 2015, **35**, 4535–4546.
- 40 C. Chambon, S. Vaudez and J. M. Heintz, De-densification mechanisms of yttria-doped cerium oxide during sintering in a reducing atmosphere, *J. Am. Ceram. Soc.*, 2018, **101**, 4956–4967.
- 41 G. I. Nkou Bouala, N. Clavier, J. Lechelle, A. Mesbah, N. Dacheux and R. Podor, In situ HT-ESEM study of crystallites growth within CeO_2 microspheres, *Ceram. Int.*, 2015, **41**, 14703–14711.
- 42 L. Claparede, N. Clavier, N. Dacheux, A. Mesbah, J. Martinez, S. Szenknect and P. Moisy, Multiparametric Dissolution of Thorium-Cerium Dioxide Solid Solutions, *Inorg. Chem.*, 2011, **50**, 11702–11714.
- 43 L. Claparede, N. Clavier, N. Dacheux, P. Moisy, R. Podor and J. Ravaux, Influence of Crystallization State and Microstructure on the Chemical Durability of Cerium-Neodymium Mixed Oxides, *Inorg. Chem.*, 2011, **50**, 9059–9072.