

Communication

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Reticular Chemistry at its Best: Directed Assembly of Hexagonal Building Units into the Awaited MOF with the Intricate Polybenzene Topology, pbz-MOF

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Supporting Information Placeholder

ABSTRACT: The ability to direct the assembly of hexagonal building units offers great prospective to construct the awaited and looked-for hypothetical polybenzene (**pbz**) or “cubic graphite” structure, described seventy years ago. Here, we demonstrate the successful use of Reticular Chemistry as an appropriate strategy for the design and deliberate construction of a zirconium based MOF with the intricate **pbz** underlying net topology. The judicious selection of the requisite hexagonal building units, six connected organic and inorganic building blocks, allowed the formation of the **pbz**-MOF-1, the first example of a Zr(IV) based MOF with **pbz** topology. Prominently, **pbz**-MOF-1 is highly porous, with associated pore size and pore volume of 13 Å and 0.99 cm³ g⁻¹ respectively, and offers high gravimetric and volumetric methane storage capacities (0.23 g g⁻¹ and 210.4 cm³STP cm⁻³ at 80 bar). Notably, the **pbz**-MOF-1 pore system permits the attainment of one of the highest CH₄ adsorbed phase density enhancement at high pressures (0.15 and 0.21 g cm⁻³ at 35 and 65 bar, respectively) as compared to benchmark microporous MOFs.

Metal-Organic Frameworks (MOFs), a relatively new class of functional solid-state materials, have demonstrated impressive capabilities in various key applications such as gas storage/separation, catalysis, sensing and drug release.¹ The crystalline nature of MOFs makes this new generation of porous solids highly appealing due to the great prospective to correlate structure to property. The crystal structure information allows for the precise analysis of these materials and most importantly gaining insights on the connectivity of their basic building units. Indeed, the topology, defined as the underlying connectivity of a material, is nowadays of prime importance in MOF chemistry as it enables the successful practice of Reticular Chemistry and the subsequent construction of the targeted MOF with the desired structural features.² It is to note that important topologies (enumerat-

ed/predicted or experimentally evidenced) are catalogued and available online in the Reticular Crystal Structure (RCSR) Database.³

The introduction of Reticular Chemistry has permitted to

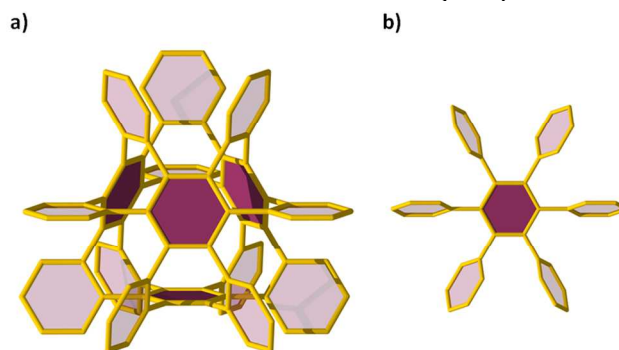


Figure 1. Part of the polybenzene (**pbz**) structure emphasizing the tetrahedral cavities (a) and the connectivity of adjacent phenyl rings (b).

enlarge and enrich the discovery and the synthesis of functional solid-state materials in general and MOFs in particular. Its successful implementation relies on the combination of predesigned inorganic building units and organic linkers, taking into account both the resultant net underlying topology and the building units geometrical constrains. One way to achieve this goal is via the deployment of the molecular building block (MBB) approach for the design and synthesis of MOFs based on the reticulation of edge-transitive nets (nets with all edges related by symmetry) as blueprints.⁴ The viability of the MOF design depends on the availability of the organic and inorganic MBBs, and subsequently the design complexity increases proportionally depending on the requisite control of the MBBs connectivity and their vis-à-vis positioning. For simplicity we refer to the design complexity levels as level 1 when it is relatively simple and level 4 when it is intricate. For level 1, basic level, the targeted MOF is mainly based on readily accessible building blocks and their sub-

sequent connectivity with no specific requirement for their *vis-a-vis* positioning; the common example is diamond topology which is the most abundant topology in MOF chemistry.^{4a} For level 2, the medium level, a specific positioning of the building units is required and an illustrative example is the **nbo**-MOFs.^{1a} For level 3, elaborate level, the limiting step is the ability to access highly-connected building units and illustrative examples are UiO-66 and **fcu**-MOFs.⁵ Level 4, the intricate level, requires the use of scarce and well defined building units and their subsequent precise/particular positioning in order to achieve the targeted structure.

In this context, highly connected symmetrical metal clusters are of particular interest in targeting new fascinating MOFs, because of the limited number of possible nets that are compatible with their connectivity.⁶ Representative examples are the highly connected hexanuclear Zr-, Hf- and RE-based MBBs, which can be employed as a 12-, 10-, 8- or even 6-connected secondary building units (SBUs).^{2d,7} Markedly, the 6-connected Zr₆ cluster, [M₆(μ₃-O)_{8-x}(μ₃-OH)_x(O₂C)₆(O₂C)₆], can express either a trigonal antiprism⁸ or a hexagonal planar SBU depending on the relative position of the 6-points of extension.⁷

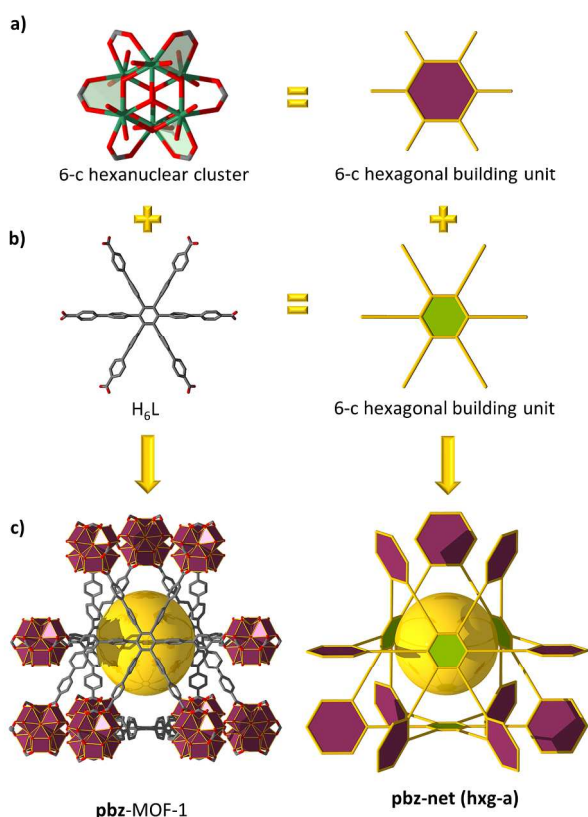


Figure 2. The combination of Zr₆ clusters (a), with the hexatopic H₆L ligand (b), both with hexagonal planar geometry, lead to the construction of **pbz**-MOF-1 (c), displaying the augmented **hxg**-a net, originally described in the structure of polybenzene (**pbz**).

Noticeably, the 6-connected Zr₆ cluster, having the perfect *D*_{6h} symmetry and the points of extension matching the 6 vertices of a hexagonal planar SBU, is an ideal inorganic building unit to target the synthesis of MOFs with the intricate underlying 6-c **hxg** topology. Employing this topology as a blueprint to design MOF requires precise control of the positioning of the hexagonal building units, the intricate design level (level 4). This type of net can be found in the hypothetical structure of polybenzene (**pbz**) or “cubic graph-

ite”¹⁰, first proposed by Gibson et al. in 1946¹¹ and portrayed as a plausible very stable polymorph of three-coordinated (sp²) carbon. In 1992¹² this structure was predicted to be by far more stable energetically than C₆₀ and can be best viewed as aromatic C₆ rings linked by single bonds (see Figure 1). Therefore, there is a great effort to synthesize this unique material as well as augmented analogues described with the general term, “polyphenylenes”.^{15,18}

In MOF chemistry, the augmented **hxg** net (**hxg**-a and also referred to as **pbz** due to its interest in other contexts) has been described and presented as a very rare example of a structure that its construction would require a true design, where the dictated particular and directional information of the organic and inorganic building blocks will permit the assembly of this looked-for and specific predetermined structure, which is based on the only edge-transitive net for linking 6-c hexagonal building units.¹³ Valuably, the successful synthesis of this intricate **pbz**-MOF will provide access to a topologically very distinct pore system, allowing its exploration in gas adsorption storage and gaining insights on the topology-property relationship. It is to note that the **hxg** topology has been observed in the case of NH₄[Cu₃-(μ₃-OH)(μ₃-4-carboxypyrazolato)₃]¹⁴, [Cu₃-(μ₃-OH)-(4-pyridyltetrazolato)₃(OH) (DMF)₄]¹⁵ and NJU-Bai⁹¹⁶, but these do not represent a **pbz** net as their basic organic and inorganic building units are not hexagons.

Here, we report the successful implementation of reticular chemistry, where the 6-c hexagonal zirconium cluster was used for the construction of a porous Zr-MOF based on the edge transitive **hxg** net. To the best of our knowledge, this is the first report disclosing the assembly of the 6-c hexagonal Zr-clusters with 6-c hexagonal organic linker into the default 6-c **hxg**-MOF, in which the augmented **hxg** net matches the **pbz**-net (polybenzene type structure).

Fundamentally, as has been described previously¹² and shown in Figure 1b, in the 3-periodic structure of **pbz**, the dihedral angle between two phenyl rings is 70.5°, while the neighboring rings attached to the same ring have complementary dihedral angles of 70.5° and 109.5° (see Figure S9), resembling the blades in contra-rotating propellers. Careful structural analysis of the hexagonal Zr₆ cluster and in view of the hexagon formed by the carbon atoms of the six equatorial carboxylate groups, a very similar propeller-like arrangement is observed, see Figure 2a. However, in this case the complementary dihedral angles are 49.6° and 130.4°. These simple geometrical considerations infer that in order for the looked-for **pbz**-MOF to be materialized, using the hexagonal Zr₆ cluster as the inorganic building unit, it is necessary to employ a hexagonal 6-connected carboxylate-based organic linker that will be able to provide and accommodate the additional twist in the dihedral angle and thus affords the 70.5° found in the **pbz** type structure.

Based on this careful and insightful geometrical analysis, we designed and synthesized the rigid and nanosized hexatopic carboxylate-based organic ligand H₆L (Figure S1-S5), depicted in Figure 2b. As anticipated, the solvothermal reaction of H₆L with ZrCl₄ in *N,N'*-dimethylformamide (DMF) in the presence of acetic acid afforded colorless octahedral-like single crystals (Figure S6), suitable for structural determination. The resultant MOF, denoted here as **pbz**-MOF-1, crystallizes in the cubic system, space group *Fd-3m* with a large unit cell, *a* = 44.856(2) Å, *V* = 90252(9) Å³. The 3-periodic framework consists of 6-c Zr₆ clusters linked by the fully deprotonated hexacarboxylate ligands L⁶⁻. The single crystal

X-ray data (Table S1) show that each Zr_6 cluster encompasses six crystallographically equivalent μ_3-O^{2-} anions (Zr-O3 bond length, 2.133 Å), placed at the equatorial positions of the hexagon which are located 0.682(6) Å above the plane containing the three Zr atoms. In addition, there are two crystallographically equivalent μ_3-OH^- anions (Zr-O2 bond length, 2.187 Å) at the faces of the hexagon, placed 0.840(1) Å above the plane of the adjacent, three Zr atoms. Interestingly, the observed chemical formula, $[Zr_6(\mu_3-O)_6(\mu_3-OH)_2]^{10+}$, representing the Zr_6 core, is different from the previously reported hexagonal Zr_6 clusters, formulated as $[Zr_6(\mu_3-OH)_8]^{16+}$ (PCN-224)^{9a}, $[Zr_6(\mu_3-O)_4(\mu_3-OH)_4]^{12+}$ ^{9b} and $[Zr_6(\mu_3-O)_8(OH)_2(H_2O)_{10}]^{6+}$ ^{9c}. The Zr atoms at the faces of the Zr_6 hexagon complete their coordination sphere with acetate anions (Ac), OH^- and H_2O species. ¹H NMR measurements in acetone exchanged samples, digested with NaOH, showed that the Ac:L ratio is 5 (see Figure S12). Accordingly, the proposed charge balanced chemical formula for the **pbz**-MOF-1 is $[Zr_6(\mu_3-O)_6(\mu_3-OH)_2(Ac)_5(OH)(H_2O)]L$. The formula is also confirmed by thermogravimetric (TGA) and elemental analysis.

Further structural analysis of the **pbz**-MOF-1, revealed that the observed dihedral angle between the central phenyl ring of the organic linker and the hexagonal plane of the Zr_6 cluster is 70.53° (see Figure S9), as expected and imposed by this topology. In point of fact, given that there are no pre-assembled Zr_6 clusters as a starting entity in the initial reaction mixture, it is conceivable to regard the organic ligand as a directing-cluster agent and a template for the in situ formation of these 6-c clusters, prompted by the suitable geometrical features of the ligand favoring the assembly of the stable **pbz**-MOF, a highly symmetric cubic lattice.

The phase purity of **pbz**-MOF-1 was confirmed by comparison of the powder X-ray diffraction (PXRD) patterns. The compound was found to be stable in common organic solvents including acetonitrile and acetone (Figure S7).

As shown in Figure 2c, the structure of **pbz**-MOF-1 contains tetrahedral cavities of approximately 13 Å in diameter, delimited by four organic linkers, occupying the faces of the tetrahedron, and twelve Zr_6 clusters. These cavities are interconnected through triangular-like windows and smaller cavities of approximately 6.5 Å and 9.5 Å, respectively, running along the [100] crystallographic direction (Figure S10, S11). The corresponding solvent-accessible volume in the **pbz**-MOF-1, calculated using the PLATON software¹⁷, is estimated to be 69%. Looking along the [100] crystallographic direction, the distance between opposite Zr_6 clusters is approximately 3.8 Å, measured from the methyl carbon atoms of the coordinate acetate anions and without taking into account the hydrogen atoms. Therefore, it is highly likely that the size of the hexagonal ligand in the **pbz**-MOF-1 represents a size cutoff for this type of linkers, below which the **pbz** net cannot be formed due to steric effects.

Argon adsorption isotherm recorded at 87 K confirmed the permanent porosity of **pbz**-MOF-1, showing a fully reversible type-I isotherm characteristic of a microporous material, see Figure 3a. The apparent BET (Langmuir) surface area, estimated using the pressure range 0.0065-0.0051 P/P₀ (seven points), was estimated to be 2415 m² g⁻¹ (2556 m² g⁻¹), in excellent agreement with the calculated BET surface area, 2432 m² g⁻¹, obtained by Poreblazer.¹⁸ The pore volume, calculated at 0.95 P/P₀ is 0.99 cm³ g⁻¹, is in excellent agreement with the crystallographically determined value of 1.0 cm³ g⁻¹. In addition to gas adsorption results, the structural integrity of the activated **pbz**-MOF-1 was supported by PXRD meas-

urements (Figure S7). The pore size distribution was calculated from the Ar adsorption isotherm using NLDFT, revealing three distinct peaks centered at 13 Å, 9.5 Å and 6.4 Å (inset in Figure 3a), in full agreement with the crystallographic analysis.

In light of its interesting structural features containing a high density of phenyl rings per accessible volume (2.3 per nm³), combined with a well-defined micropore system and high BET surface area (2415 m² g⁻¹), **pbz**-MOF-1 may be considered as an excellent adsorbent candidate for gas storage applications and in particular for CH₄. Accordingly, CH₄ adsorption measurements at 112 K, revealed a reversible type-I isotherm (Figure S13), from which the

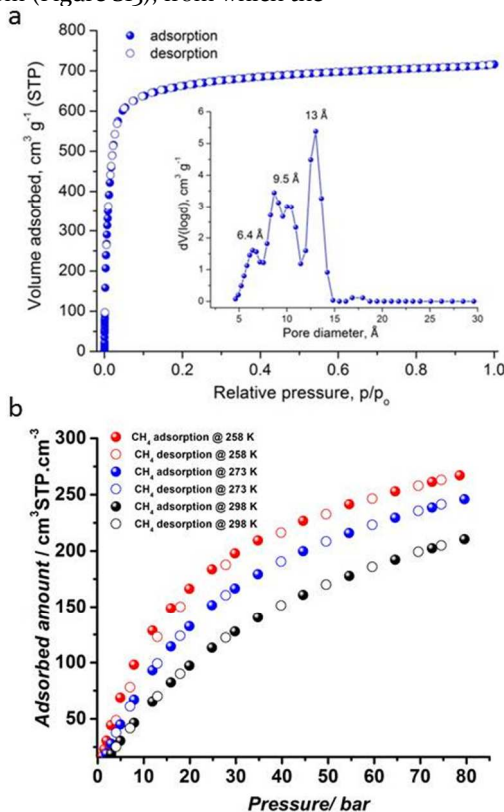


Figure 3. (a) Argon sorption isotherm of **pbz**-MOF-1 recorded at 87 K and the corresponding pore size distribution curve calculated using NLDFT (inset). (b) CH₄ sorption isotherms of **pbz**-MOF-1 at 298 K, 273 K and 258 K, up to 80 bar.

total pore volume at 0.99 P/P₀ was calculated to be 0.95 cm³ g⁻¹, indicating that the full available pore space in the **pbz**-MOF-1 is readily accessible by CH₄ molecules. High-pressure CH₄ adsorption measurements were carried out at different temperatures up to 80 bar (Figure 3b). The observed gravimetric and volumetric uptake at 80 bar and 298 K is 0.23 g g⁻¹ and 210.4 cm³(STP) cm⁻³, placing **pbz**-MOF-1 among the high performing MOFs for CH₄ storage (Table S2 and Figure S14). The corresponding CH₄ capacities at 273 K and 258 K are 0.27 g g⁻¹/246 cm³ cm⁻³ and 0.29 g g⁻¹/267 cm³ cm⁻³, respectively. Further analysis of the adsorption uptakes at different pressures and room temperature showed that **pbz**-MOF-1 exhibits one of the largest increase in the CH₄ adsorbed phase density between 35 and 65 bar (38.3%), as compared to other well-performing MOFs for CH₄ storage (Table S3 and Figure S15). Reasonably, this enhanced CH₄ density in the **pbz**-MOF-1 could be associated with the favorable and very high efficiency in filling the available space in the pore network. Equivalent values for other existing

MOFs such as HKUST-1 (17.4 %) ¹⁹, Al-soc-MOF-1 (62.8 %) ²⁰, PCN-14 (16.6 %) ²¹, MOF-520 (38.7 %) ²² and MOF-905 (44.6 %) ²³ supports and affirms the high degree of CH₄ packing in the **pbz**-MOF-1 between 35 and 65 bar. The gravimetric and volumetric working capacity between 5-80 bar and 298 K is estimated to be 22.7 wt% and 180 cm³ cm⁻³, respectively. Interestingly, in contrast to other MOFs with intermediate porosity such as PCN-14 ²¹, HKUST-1 ¹⁹, Ni-MOF-74 ¹⁹ and UTSA-20 ²⁴, the decrease of the storage temperature to 273 K led to an increase in the CH₄ working capacity. For instance, in the case HKUST-1 and when going from 298 K to 273 K, the volumetric working capacity between 5-65 bar decreased from 190 cm³ cm⁻³ to 155 cm³ cm⁻³. Whereas for **pbz**-MOF-1, under the same conditions, the opposite occurred and an increase from 162 cm³ cm⁻³ to 185 cm³ cm⁻³ was observed. This manifest unique feature in the microporous **pbz**-MOF-1 could be attributed to the favorable synergistic effect between the pore system dimensions and the localized high density of aromatic rings per accessible unit volume.

In summary, we have demonstrated for the first time the successful use of reticular chemistry and the MMB approach for the construction of a highly connected Zr-based MOF with the intricate **pbz** topology. The synthesis of the **pbz**-MOF-1, constructed from 6-connected hexagonal Zr₆ clusters and bridged by a novel nanosized hexatopic linker, represents the first example of a crystalline solid-state material displaying the 6-c **pbz** net. High pressure CH₄ adsorption measurements on the highly porous **pbz**-MOF-1 revealed a high storage capacity of 0.23 g g⁻¹ and 210.4 cm³(STP) cm⁻³ at 80 bar. Moreover, **pbz**-MOF-1 exhibits one of the highest enhancement in the adsorbed phase density when going from 35 to 65 bar. The successful synthesis of the first **pbz**-MOF paves the way to expand this approach to other ligands and inorganic building blocks with the aim to tune the pore system and functionality.

ASSOCIATED CONTENT

Supporting Information

Synthesis, NMR, PXRD, TGA, and additional sorption data. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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