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Nickel-Catalyzed Asymmetric Hydrogenation of Cyclic Alkenyl Sulfones, Benzo[b]thiophene 1,1-Dioxides, with Mechanistic Studies

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nonlinear effect studies, deuterium-labeling experiments, and DFT calculation investigations, a reasonable catalytic mechanism for this nickel-catalyzed asymmetric hydrogenation was provided, which displayed that the two added hydrogen atoms of the hydrogenation products could be from H_2 through the insertion of Ni–H and subsequent hydrogenolysis.

symmetric hydrogenation of prochiral unsaturated A compounds has been regarded as a powerful and straightforward synthetic method to access chiral compounds that has been widely applied in the fields of agrochemicals and pharmaceuticals.¹ Most well-established asymmetric hydrogenation methods are based on heavy noble transition metal catalytic systems, mainly focused on ruthenium,² iridium,³ rhodium,⁴ and palladium.⁵ However, these precious metal catalysts could suffer from the difficulties of high costs, limited resources, strong toxicity, and negative environmental impact. Therefore, great attention has been devoted to the development of cheap, sustainable, earth-abundant transition metal catalytic systems for asymmetric hydrogenation.⁶ In recent years, some iron-, cobalt-, and nickel-catalyzed asymmetric (transfer) hydrogenation reactions of prochiral unsaturated compounds with C=C, C=O, or C=N bonds were reported⁷⁻⁹ that demonstrated the great potential advantages of the cheap first-row transition metals in asymmetric hydrogenation. Among these asymmetric catalytic systems, Ni-catalyzed asymmetric hydrogenation was relatively less investigated. Some pioneering and important research works on Ni-catalyzed asymmetric (transfer) hydrogenation of prochiral ketones, alkenes, ketimines, and enamides were developed by Hamada,^{9a,b} Chirik,^{9c} Zhou,^{9d-i} Zhang,^{9j} and our group.^{9k-p} In 2016, Chirik and co-workers developed the Nicatalyzed asymmetric hydrogenation of $\alpha_{,\beta}$ -unsaturated esters and performed deep research on the catalytic mechanism, proposing conjugate addition of the Ni-H species and nonselective protonation to release the product.^{9c} In 2017, our group realized Ni-catalyzed asymmetric hydrogenation of β -acylamino nitroolefins, and experimental and DFT computational investigations revealed that it also involved 1,4-hydride addition of Ni-H and subsequent protonation to generate the

desired product.^{9k} In addition, these Ni-catalyzed asymmetric hydrogenation reactions were mainly focused on acyclic functionalized olefins and imines as substrates, which were always in accordance with this catalytic mechanism. In contrast, cyclic functionalized olefin substrates have rarely been investigated.

Chiral cyclic sulfone scaffolds are widely distributed in many biologically active compounds and natural products with important applications.¹⁰ Because of the rigidity of cyclic functionalized olefins, the asymmetric hydrogenation of unsaturated cyclic sulfones to access these motifs is still in an early stage, and just a few examples involving unsaturated cyclic sulfones have been developed.^{21,30,p,4k} In 2012, Andersson and co-workers described Ir-catalyzed enantioselective hydrogenation of prochiral cyclic and acyclic unsaturated sulfones with excellent results.³⁶ In 2017, Pfaltz and coworkers successfully developed Ir-catalyzed asymmetric hydrogenation of cyclic alkenyl sulfones, benzo b thiophene 1,1dioxides.^{3p} Recently, our group realized Rh/N-methylated bisphosphine-thiourea ZhaoPhos-catalyzed asymmetric hydrogenation of these cyclic alkenyl sulfones with excellent results by the assistance of the possible hydrogen-bonding interaction between the substrate and the ligand.^{4k} However, cheap transition metals have never been applied to catalyze the asymmetric hydrogenation of cyclic alkenyl sulfones. On the

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basis of our long-standing efforts in asymmetric hydrogenation, we are interested in the development of cheap transition metal catalytic systems for the asymmetric hydrogenation of cyclic alkenyl sulfones. As a continuation of our investigations, in this work we successfully developed the first asymmetric hydrogenation of 3-substituted benzo[b]thiophene 1,1-dioxides catalyzed by the cheap transition metal Ni, affording a series





Table 1. Ligand Screening for Ni-Catalyzed Asymmetric Hydrogenation of 3-Phenylbenzo[b]thiophene 1,1-Dioxide $(1a)^{a}$



^{*a*}Conditions: 0.1 mmol of 1a, 1a:Ni(OAc)₂:ligand = 1:0.05:0.055 in 1.0 mL of CF₃CH₂OH (TFE) under 60 atm H₂ for 40 h. The catalyst was precomplexed in MeOH (0.5 mL for each reaction vial). ^{*b*}Determined by ¹H NMR analysis. NR = no reaction. ^{*c*}The ee values were determined by HPLC on a chiral phase. NA = not available.

of enantioenriched hydrogenation products with >99% conversion, 95-99% yield, and 90-99% ee (Scheme 1). Importantly, a reasonable catalytic mechanism was proposed for this hydrogenation on the basis of deuterium-labeling experiments and DFT calculations.

Initially, a variety of readily available chiral diphosphine ligands (Figure 1) were applied for the Ni(OAc)₂-catalyzed asymmetric hydrogenation of the model substrate 3-phenylbenzo[b]thiophene 1,1-dioxide (1a) under 60 atm H₂ at 70 °C for 40 h. The Ni(OAc)₂/chiral diphosphine ligand catalyst was generated in situ in MeOH to achieve good solubility. (*S,S*)-Me-DuPhos and JosiPhos displayed poor reactivities and enantioselectivities (8–50% conversion and 30–59% ee; Table 1, entries 1 and 3). Although 25% conversion was realized in the presence of (*S,S*)-Ph-BPE, excellent enantioselectivity of 95% ee was obtained (Table 1, entry 2). In addition, some other important chiral diphosphine



Figure 1. Structures of chiral diphosphine ligands.

Table 2. Optimization of the Reaction Conditions for Ni-Catalyzed Asymmetric Hydrogenation of 3-Phenylbenzo[b]thiophene 1,1-Dioxide (1a)^{*a*}



^{*a*}Conditions: 0.1 mmol of **1a**, **1a**:Ni(OAc)₂:(*S*,*S*)-Ph-BPE = 1:0.01:0.011 in 1.0 mL of TFE for 50 h. The catalyst was precomplexed in MeOH (0.1 mL for each reaction vial). THF is tetrahydrofuran. Solvent:MeOH ratios are v/v. ^{*b*}Determined by ¹H NMR analysis. ^{*c*}The ee values were determined by HPLC on a chiral phase. ^{*d*}0.1 mmol of **1a**, **1a**:Ni(OAc)₂:(*S*,*S*)-Ph-BPE = 1:0.05:0.055 in 1.0 mL of MeOH for 50 h. ^{*e*}0.1 mmol of **1a**, **1a**:Ni(OAc)₂:(*S*,*S*)-Ph-BPE = 1:0.01:0.011 in 2.0 mL of TFE for 50 h. The catalyst was precomplexed in MeOH (50 μ L for each reaction vial).

ligands, such as (S)-SegPhos, (Rc,Sp)-DuanPhos, (S)-BINAP, and ZhaoPhos could not promote this Ni-catalyzed asymmetric hydrogenation, and no desired hydrogenation product **2a** was detected (Table 1, entries 4–6 and 8). WalPhos provided good conversion (83%) with moderate enantiose-lectivity (56% ee) (Table 1, entry 7).

In order to improve the conversion of this Ni/(*S*,*S*)-Ph-BPEcatalyzed asymmetric hydrogenation of **1a**, the reaction conditions were further optimized in various solvents. It was found that there was no conversion in MeOH, which was an unfavorable solvent for this Ni-catalyzed asymmetric hydrogenation (Table 2, entry 1). The Ni(OAc)₂/(*S*,*S*)-Ph-BPE catalytic system was dissolved well in MeOH, and it was better to investigate other solvents in the presence of MeOH. Then the catalyst solution was decreased to a catalyst loading of 1.0 mol % to reduce the negative effect of MeOH, and the

Scheme 2. Scope Study for the Ni-Catalyzed Asymmetric Hydrogenation of 3-Substituted Benzo[b]thiophene 1,1-Dioxides^a



^{*a*}Conditions: 0.1 mmol of substrate, substrate 1:Ni(OAc)₂:(*S*,*S*)-Ph-BPE = 1:0.01:0.011 in 2.0 mL of CF₃CH₂OH for 50 h. The catalyst was precomplexed in MeOH (50 μ L for each reaction vial). Isolated yields are shown. The ee values were determined by HPLC on a chiral phase. ^{*b*}S:C = 50, 60 h.



conversion was greatly increased to 58% (Table 2, entry 2). Some other solvents were then examined, but no hydrogenation product was detected in EtOH, CH_2Cl_2 , THF, or toluene (Table 2, entries 3–6). The conversion was increased to 72% with 95% ee when the ratio of TFE was increased to TFE:MeOH = 40:1 (Table 2, entry 7). TFE may help to accelerate the addition of the Ni–H active species to the C==C bond. Good conversion (82%) and excellent enantioselectivity (96% ee) could be achieved at a higher reaction temperature of 80 °C (Table 2, entry 8). To our delight, >99% conversion



Figure 2. Linear effect study of the hydrogenation of model substrate **1a** using (S,S)-Ph-BPE with different ee values.

with 96% ee was obtained when the pressure of H_2 was increased to 70 atm (Table 2, entry 9).

Under the optimized reaction conditions, we began to explore the substrate generality of this Ni/(S,S)-Ph-BPEcatalyzed asymmetric hydrogenation of 3-substitued benzo[b]thiophene 1,1-dioxides. These results are summarized in

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Figure 3. DFT-calculated free energy profile for the proposed catalytic cycle of the Ni-catalyzed asymmetric hydrogenation.

Scheme 4. Gram-Scale Asymmetric Hydrogenation



Scheme 2. Substrates with electron-netural groups (1a and 1b), electron-donating groups (1c-k), and electron-withdrawing groups (11-p) on the phenyl ring reacted well, affording the corresponding hydrogenation products (2a-p) with excellent results (>99% conversion, 95-99% yields, 90-99% ee). In addition, the position of substituents on the phenyl ring had little influence on the reactivity and enantioselectivity. Remarkably, the challenging substrates 1c, 1d, 1f-h, 1j, 1m, 1n, and 1p bearing steric hindrance from the ortho and meta substituents on the phenyl ring were hydrogenated smoothly in 95-98% yield with 90-99% ee. Furthermore, the asymmetric hydrogenation of substrate 1q with a bulky 2-naphthyl group also proceeded smoothly to afford product 2q with full conversion in 95% yield with 92% ee. To our delight, the alkyl substrates with a 3-methyl group (1r), an ethyl group (1s), and an n-propyl group (1t) also worked well to provide the hydrogenation products 2r-t with excellent results (96-99% yield with 90-93% ee). In addition, the 2-substitued benzo[b]thiophene 1,1-dioxide substrates containing a phenyl group $(1\mathbf{u})$ or methyl group $(1\mathbf{v})$ were then hydrogenated, but unfortunately, there was no reaction in this Ni-catalyzed system.

In order to explore the possible mechanism, deuteriumlabeling experiments were carried out. The Ni-catalyzed asymmetric hydrogenation of model substrate **1a** was conducted under 50 atm D₂. Interestingly, we found that the deuterium atoms of the product **2a-D** were attached at both the α - and β -positions (Scheme 3a). In addition, this hydrogenation was conducted in the presence of H₂ with CF₃CH₂OD and CD₃OD as the reaction solvents, and no deuterium atoms were found in the hydrogenation product (Scheme 3b). These observations are quite different from the previous research work, ^{9c,k} indicating that this Ni-catalyzed asymmetric hydrogenation may go through a different catalytic pathway.

In addition, a nonlinear effect study of this Ni-catalyzed asymmetric hydrogenation was performed. A series of hydrogenations of model substrate **1a** were carried out with ligand (S,S)-Ph-BPE with different ee values. As shown in Figure 2, a linear effect could be observed in this asymmetric transformation, and it is possible that there is no catalyst self-aggregation or ligand—substrate agglomeration in this catalytic system.¹¹

According to these above reaction results, DFT calculations were performed to further reveal the possible catalytic mechanism of this hydrogenation. Our favored computed catalytic cycle (Figure 3) starts with the formation of Ni-H complex **B** through hydrogenolysis of the coordinated Ni(II) complex, which can be deemed as the active catalyst. Subsequent coordination of **B** with substrate 1a and then insertion of Ni-H into the double bond gives Ni-C complex D_{S} . The second hydrogenolysis delivers the desired *S*-configured hydrogenated product 2a and regenerates the active catalyst **B**. This catalytic process is quite different from the previously reported 1,4-hydride addition and subsequent

protonation to give the desired product.^{9c,k} In addition, the computational details of other catalytic pathways are provided in the Supporting Information.

To demonstrate the potential synthetic application of this Ni-catalyzed transformation, the gram-scale asymmetric hydrogenation on 4 mmol of model substrate **1a** was conducted in the presence of just 0.5 mol % catalyst. The desired product **2a** was obtained in full conversion and 99% yield with 97% ee, showing the potential synthetic importance of our catalytic system (Scheme 4).

In summary, we successfully developed the first asymmetric hydrogenation of challenging cyclic alkenyl sulfones, 3substituted benzo[b]thiophene 1,1-dioxides, catalyzed by the cheap transition metal Ni. A series of hydrogenation products, chiral 2,3-dihydrobenzo[b]thiophene 1,1-dioxides, were obtained in 95–99% yield with 90–99% ee. A reasonable catalytic cycle was proposed for this Ni-catalyzed hydrogenation on the basis of combined deuterium-labeling experiments and DFT calculations. Also, we found that the two added hydrogen atoms in the hydrogenation product could be from H₂ through insertion of Ni–H and subsequent hydrogenolysis, which is different from the previously reported 1,4-hydride addition and subsequent protonation.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at https://pubs.acs.org/doi/10.1021/acs.orglett.0c03723.

Experimental procedures and compound characterization (PDF)

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Notes

The authors declare no competing financial interest.

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