

Probing 'Spin-Forbidden' Oxygen-Atom Transfer: Gas-Phase Reactions of Chromium–Porphyrin Complexes

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Abstract: Oxygen-atom transfer reactions of metalloporphyrin species play an important role in biochemical and synthetic oxidation reactions. An emerging theme in this chemistry is that spin-state changes can play important roles, and a 'two-state' reactivity model has been extensively applied especially in iron porphyrin systems. Herein we explore the gas-phase oxygen-atom transfer chemistry of *meso*-tetrakis(pentafluorophenyl)porphyrin (TPFPP) chromium complexes, as well as some other tetradentate macrocyclic ligands. Electrospray ionization in concert with Fourier transform ion cyclotron resonance (FT-ICR) mass spectrometry has been used to characterize and observe reactivity of the ionic species [(TPFPP)Cr^{III}]⁺ (**1**) and [(TPFPP)Cr^{VO}]⁺ (**2**). These are attractive systems to examine the effects of spin-state change on oxygen-atom transfer because the d¹ Cr^V species are doublets, while the Cr^{III} complexes have quartet ground states with high-lying doublet excited states. In the gas phase, [(TPFPP)Cr^{III}]⁺ forms adducts with a variety of neutral donors, but O-atom transfer is only observed for NO₂. Pyridine *N*-oxide adducts of **1** do yield **2** upon collision-induced dissociation (CID), but the ethylene oxide, DMSO, and TEMPO analogues do not. [(TPFPP)Cr^{VO}]⁺ is shown by its reactivity and by CID experiments to be a terminal metal–oxo with a single, vacant coordination site. It also displays limited reaction chemistry, being deoxygenated only by the very potent reductant P(OMe)₃. In general, [(TPFPP)Cr^{VO}]⁺ species are much less reactive than the Fe and Mn analogues. Thermochemical analysis of the reactions points toward the involvement of spin issues in the lower observed reactivity of the chromium complexes.

1. Introduction

The role of spin states in reactions of transition metal complexes is being increasingly discussed, specifically how reactions are affected when a change in spin state occurs between reactants and products.¹ This fundamental question appears to be particularly important in the oxidative chemistry of metalloporphyrin-containing enzymes and complexes. Heme-containing monooxygenase enzymes, such as cytochromes P450, nitric oxide synthases, and peroxidases, mediate myriad transformations of biological importance.^{2–4} In particular, the reaction chemistry of the cytochromes P450 has been intensively studied, both in enzymatic systems and using small molecule

models, because this enzyme superfamily is primarily responsible for processing xenobiotic substances *in vivo*. In addition, the enzymes and models accomplish very interesting reaction chemistry, including the hydroxylation of alkanes and epoxidation of olefins. The mechanisms of these reactions have therefore been of much interest, and have been the subject of some controversy.⁵

For all these reasons, there have been many computational studies of metalloporphyrin oxidation reactions and their mechanisms. Shaik, Schwarz and their co-workers have proposed that multiple spin states play an important role in these reactions, with many of them involving what they have called "two-state reactivity (TSR)."^{6,7} They state, "A thermal reaction which involves spin crossover along the reaction coordinate from

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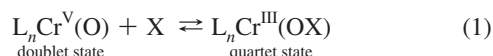
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reactants to products needs to be described in terms of two-state reactivity if product formation arises from an interplay of spin inversion and the respective barrier heights on both spin surfaces.²⁷ Perhaps the clearest experimental evidence for the TSR model comes from the gas-phase reactions of FeO^+ or OCrO^+ studied by Schwarz and co-workers.⁸ The interpretation of the experiments and computations for the iron-porphyrin species in oxidation reactions is complicated, however, by the multiplicity of accessible spin states. The ‘Compound I’ intermediate, two oxidation levels above Fe^{III} , is most simply described as an Fe^{IV} –oxo (ferryl) species coupled to a porphyrin radical cation and can have closely spaced doublet, quartet, and sextet states.⁹ The Fe^{III} state can similarly have accessible doublet, quartet, and sextet states.⁶ Experimental and computational studies also indicate that spin-state effects play important roles in the reactions of oxomanganese porphyrin complexes,¹⁰ other metal–oxo complexes,¹¹ and in oxygen atom abstractions by $\text{M}(\text{OCBu}_3)_3$ [$\text{M} = \text{V}, \text{Nb}(\text{PME}_3)$, and Ta].¹²

The chromium(V)–oxo/chromium(III) redox couple could provide a more direct test of the importance of spin-state changes on oxygen-atom transfer reactions because there is one predominant spin state present at each oxidation level in such a system (eq 1). Cr^{V} has a d^1 configuration and therefore is an obligate doublet state.¹³ A number of Cr^{V} oxo species have



been prepared, including with porphyrin, salen, and other ligands.¹⁴ Some of these compounds undergo oxygen-atom transfer reactions such as epoxidation, as found by Groves,¹⁵ Kochi,¹⁶ and others.¹⁷ Cr^{III} coordination complexes essentially always have quartet ground states (except for organometallic

species with π -acid ligands).^{18,19} In octahedral symmetry, the ground state for Cr^{III} is $^4\text{A}_{2g}$ and the lowest lying excited state is ^2E .²⁰ Since both states have a t_{2g}^3 configuration, the energy of the ^2E state is much less dependent in the nature of the ligands than most ligand-field excited states and is at $\sim 13,000 \text{ cm}^{-1}$ (37 kcal mol^{-1}) for $\text{Cr}(\text{NH}_3)_6^{3+}$.²⁰ When X and XO are closed shell (singlet) species, oxygen-atom transfer from doublet ground state Cr^{VO} to a quartet ground state Cr^{III} is thus formally spin forbidden. Herein we report our investigations of the gas-phase reactivity of ligated chromium(III) and chromium(V)–oxo complexes with a variety of oxygen atom donors and acceptors, using electrospray ionization (ESI) coupled to mass spectrometry (MS).

Earlier studies based on Fourier transform ion cyclotron resonance (FT-ICR) mass spectrometry combined with ESI have provided valuable information on functional models of the Compound I intermediates of monooxygenase enzymes, namely the $[(\text{TPFPP})^+\text{Fe}^{\text{IV}}\text{O}]^+$ and $[(\text{TPFPP})\text{Mn}^{\text{VO}}]^+$ ($\text{TPFPP} = \text{meso-tetrakis(pentafluorophenyl)porphyrinato dianion}$) complexes.²¹ These species have been obtained as naked ions and their reactivity has been investigated in an environment lacking any proximal or distal ligand, solvent or counterion. In this way it could be ascertained that the $[(\text{TPFPP})\text{Mn}^{\text{VO}}]^+$ complex is responsible for catalytic activity, while a strong trans axial ligand (an additional oxygen) can switch off the reactivity, an outcome consistent with barriers due to TSR effects.^{21c} By gas-phase synthesis using ozone as oxidant, the naked core of Compound I, $[(\text{PPIX})^+\text{Fe}^{\text{IV}}\text{O}]^+$ ($\text{PPIX} = \text{protoporphyrin IX dianion}$) has been obtained for the first time.²² The dilute gas phase has ensured that these species have a lifetime long enough to reveal both structural details and elementary steps of the catalytic activity. High-valent manganese(V)–oxo complexes of non-heme ligands have also been examined in the gas phase by Plattner, showing that ESI-tandem MS techniques are a convenient approach for the study of the coordination chemistry of highly reactive solution-phase species.²³ On these premises an attempt to answer the spin-forbidden question for oxo–chromium complexes was undertaken.

2. Experimental Section

2.1. Materials. 5,10,15,20-Tetrakis(pentafluorophenyl)porphyrinato chromium(III) chloride, $(\text{TPFPP})\text{Cr}^{\text{III}}\text{Cl}$ (**1-Cl**) was purchased from Frontier Scientific (Logan, UT, U.S.A.). 5,10,15,20-meso-Tetraphenylporphyrin (TPP, Strem) was purified by the method of Adler.²⁴ $(\text{TPP})\text{Cr}^{\text{III}}\text{Cl}$ was prepared from TPP and anhydrous $\text{Cr}^{\text{II}}\text{Cl}_2$ (Strem) according to the literature.²⁵ 3,3'-5,5'-Tetra-*tert*-butyl-salen,

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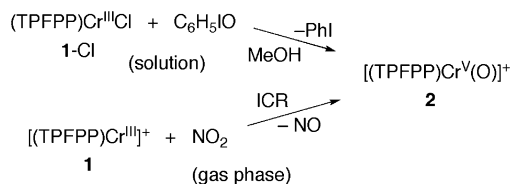
and salphen were prepared by condensation of the appropriate diamine and benzaldehyde (both from Sigma), and recrystallized from ethanol (J. T. Baker).²⁶ (salen)Cr^{III}Cl Complexes were synthesized from Cr^{III}Cl₂ according to known procedures.²⁷ The corresponding triflate salts were prepared by addition of stoichiometric AgOTf (Strem) and filtration. Iodosylbenzene (C₆H₅IO) was prepared according to the literature procedure²⁸ and stored at -20 °C. Pyridine-*N*-oxide (pyO) was synthesized by reaction of pyridine with a stoichiometric amount of *m*-chloroperbenzoic acid (*m*-CPBA) in MeOH at room temperature. All the solvents were analytical grade. H₂¹⁸O (95%), and all other chemicals were research grade products obtained from commercial sources (Sigma-Aldrich; Matheson Gas Products Inc.) and used as received.

2.2. Instrumental. Experiments were performed either with a Bruker BioApex 4.7 T Fourier transform ion cyclotron resonance (FT-ICR) mass spectrometer, or with a triple quadrupole linear ion trap mass spectrometer (Applied Biosystems 2000 Q TRAP). The first instrument is equipped with an Apollo I electrospray ionization (ESI) source. Ions are led into a cylindrical “infinity” cell where neutrals may be admitted by way of either pulsed valves or leak valves. The second device consists of an ESI source combined with a tandem mass spectrometer (QHQ, quadrupole–hexapole–quadrupole configuration).

2.3. Generation of Oxochromium(TPFPP)⁺ [(TPFPP)Cr^V(O)]⁺ Ions. Two routes were followed in order to generate the [(TPFPP)Cr^V(O)]⁺ ions of interest. In the first one, a 10 μM solution of 1-Cl in methanol was submitted to ESI, and the so formed [(TPFPP)Cr^{III}]⁺ ions (**1**) were exposed to a stream of NO₂ in the ionic path driving the ions into the FT-ICR cell. This procedure leads to incorporation of an O-atom into **1**, presumably resulting in formation of [(TPFPP)Cr^V(O)]⁺ (**2**). In the second method, iodosylbenzene (C₆H₅IO, 0.044 mg, 2 × 10⁻⁴ mol) was added to 1-Cl (10 μM) in methanol (2 mL) at room temperature, yielding a red colored solution, stable for several hours at room temperature. When submitted to ESI, this solution delivered the same species described as [(TPFPP)Cr^V(O)]⁺ (**2**). In both cases, the mass spectra showed two prominent species, both characterized by the same characteristic isotopic pattern. The first one is assigned to ion **1**, [(TPFPP)Cr^{III}]⁺, centered at *m/z* = 1024, while the second one is centered at *m/z* = 1040, consistent with ion **2** [(TPFPP)Cr^V(O)]⁺.

2.4. Reactions of [(TPFPP)Cr^V(O)]⁺. The solutions were continuously sprayed at a 2 μL min⁻¹ flow rate controlled by a syringe pump. A countercurrent flow of heated dry gas (nitrogen at 120 °C) was used to desolvate the ions. After an accumulation interval of 0.6 s in a rf-only hexapole, the ion population was pulsed into the ICR cell placed within a 4.7 T superconducting magnet. In the cell, maintained at room temperature (*ca.* 300 K), the neutral reagent was leaked at a constant pressure in the range of 0.5–10 × 10⁻⁸ mbar. The pressure readings, obtained from a cold-cathode ionization gauge (IKR Pfeiffer Balzers S.p.A., Milan, Italy), were calibrated by using the rate constant *k* = 11.7 × 10⁻¹⁰ cm³ molecule⁻¹ s⁻¹ for the reference reaction CH₄⁺⁺ + CH₄ → CH₅⁺ + CH₃[•] and weighted by using individual response factors.²⁹ The ion of interest was isolated by ejection routines as commonly performed in FT-ICR mass spectrometry and allowed to react with the neutral leaked by a needle valve. To avoid unintended off-resonance excitation of the reactant ion, the whole isotopic cluster corre-

Scheme 1. Generation of (TPFPP)Cr^V(O)⁺ (**2**)



sponding to ion **1** or **2**, as desired, was isolated and allowed to react with the selected neutral. Pseudo-first-order rate constants were obtained from the semilogarithmic decrease of the reactant ion abundance over time at each selected pressure and were divided by the concentration of the neutral species to attain second-order rate constants (*k*_{exp}). The rate constants were also expressed as percentages (efficiencies) of the collision rate constant (*k*_{coll}) calculated by the parametrized trajectory theory.³⁰ Whereas the reproducibility of *k*_{exp}, reported as the average of usually three experiments run at different neutral pressures, was good, the estimated error in their absolute values (±30%) is mainly due to uncertainty in pressure measurements.

2.5. Collision-Induced Dissociation (CID) Experiments. For CID performed in FT-ICR, the mass-selected reagent ions were accelerated in the presence of argon gas pulsed into the cell (peak pressure *ca.* 4 × 10⁻⁶ mbar) for 1 s. In ESI-MS/MS effected in the Q TRAP, the ions of interest were isolated using Q1, accelerated at variable collision energy (CE) using N₂ as a collision gas (normally at 2 × 10⁻⁴ mbar) in the hexapole H and detected together with their ionic fragments in Q2. The energy dependence of the dissociation process has been analyzed yielding the phenomenological appearance energies (*E*_{app}) for the fragmentation processes under study.

3. Results

We have primarily focused on the tetrakis(pentafluorophenyl)porphyrin (TPFPP) complexes of chromium because this ligand has proven very amenable to ESI-FT-ICR studies with the related iron^{21a,b} and manganese^{21c} complexes. Studies with other porphyrin and related tetradentate ligands are described after the TPFPP results.

3.1. Generation of [(TPFPP)Cr^V(O)]⁺. Treatment of chromium(III)–porphyrin chloride complexes (P)CrCl with oxygen atom donors are reported to give Cr^Voxo and/or Cr^Voxo compounds, depending on the reaction conditions; in the presence of limited oxidant, Cr^V and Cr^{III} disproportionate to Cr^{IV}.^{14,31} With the TPFPP ligand, there has been a brief comment about the formation of (TPFPP)Cr^{IV}(O) from (TPFPP)Cr^{II} and O₂, but no mention of the Cr^Voxo complex.³¹ We find that treatment of (TPFPP)Cr^{III}Cl (abbreviated 1-Cl) with iodosylbenzene (C₆H₅IO) in methanol solution gives substantial amounts of [(TPFPP)Cr^V(O)]⁺ (**2**), as indicated by electrospray ionization (ESI) mass spectrometry of these mixtures (Scheme 1). The same product is formed by the gas-phase reaction of [(TPFPP)Cr^{III}]⁺ (**1**) and NO₂ in the ICR (Scheme 1). Other reactions of **1** are described below.

Figure 1 shows a high-resolution ESI-FT-ICR mass spectrum of **2** generated in solution from PhIO + [(TPFPP)Cr^{III}]⁺ (**1**), showing ions at *m/z* = 1039.9806 (calculated value 1039.9778). Complex **1** is also seen in such mass spectra, at *m/z* = 1023.9857 (calculated value 1023.9829). If no oxidizable species are added to the solution, **2** is remarkably inert toward side reactions, as has been found for iron–oxo cations containing porphyrin

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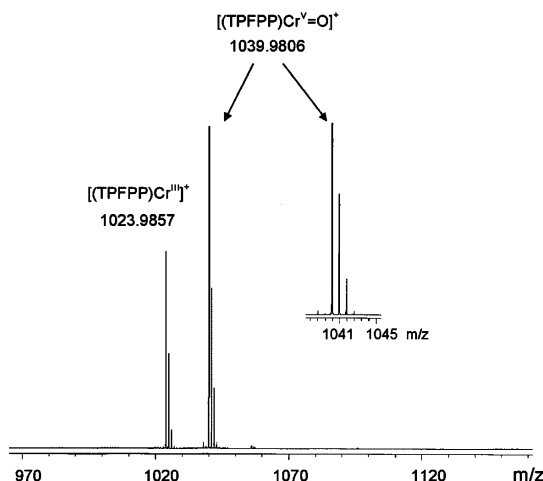
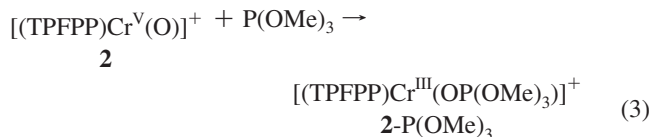
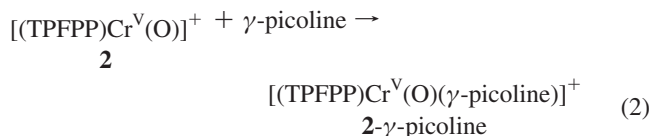


Figure 1. ESI (+) FT-ICR mass spectrum of a methanol solution of (TPFPP)Cr^{III}Cl (**1**-Cl, 10 μM) and a 10-fold excess of iodosylbenzene, showing [(TPFPP)Cr^{III}]⁺ (**1**) and [(TPFPP)Cr^V(O)]⁺ (**2**). Inset: enlargement of the isotopic cluster at *m/z* 1040.

ligands with electron-withdrawing groups.^{21b} When the oxidation of **1**-Cl with iodosylbenzene is effected in the presence of excess H₂¹⁸O (95% ¹⁸O content) the ion at *m/z* = 1040 is shifted to *m/z* = 1042 (**2**-¹⁸O). This indicates that compound **2** contains an added oxygen atom, likely a formal Cr^Voxo complex (Figure 1), as discussed in further detail below.

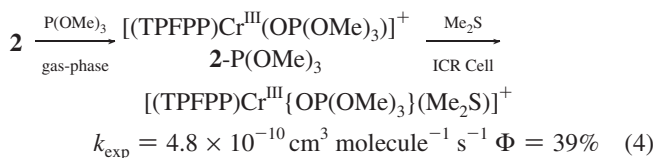
3.2. Gas-Phase Reactions of [(TPFPP)Cr^V(O)]⁺. The nature and reactivity of [(TPFPP)Cr^V(O)]⁺ (**2**), has been assayed by FT-ICR mass spectrometry. In all cases, there is no difference between **2** formed either in solution with PhIO or in the ICR with NO₂. When mass-selected and exposed in the ICR cell to a stationary concentration of potential reducing agents (L), **2** proved to be quite unreactive. No change was observed on exposure to NO, NO₂, Me₂S, β-pinene, pyridine or thioanisole. This indicates that reactions with these substrates are disfavored either by a positive free energy change,³² or by rate constants less than *k*_{exp} < 10^{−13} cm³ molecule^{−1} s^{−1}. Reactions do occur with γ-picoline (*p*-CH₃C₅H₄N) and trimethylphosphite (P(OMe)₃) to give addition products [(TPFPP)CrOL]⁺ (**2**-L) and, for P(OMe)₃, the bis-addition product [(TPFPP)CrOL₂]⁺ (**2**-L₂). These adduct ions likely owe their stabilization to the dissipation of any excess energy released in the association process by radiative and/or collisional cooling.³³ The bimolecular rate constants (*k*_{exp}) and the relative efficiencies (Φ) are summarized in Table 1. Note that the abbreviations used in this paper indicate how the ions were made, in this case “**2**-L” and “**2**-L₂” indicating that the complexes were made from **2** and L. The chemical structure of these ions is discussed below.

When assayed by low-energy collision-induced dissociation (CID), the primary adduct [(TPFPP)Cr(O)(γ-picoline)]⁺ yields ion **2** and γ-picoline as dissociation products. This suggests that γ-picoline has simply bound to **2** as a ligand to chromium, namely that this material is a picoline-ligated Cr^V–oxo complex, **2**-γ-picoline (eq 2). In contrast, the CID spectra of [(TPFPP)CrOP(OMe)₃]⁺ (**2**-P(OMe)₃), formed as shown in eq 3)



display the release of **1** by concomitant loss of OP(OMe)₃ (Figure 2). The CID spectra show no significant fragmentation of the phosphite ligand in **2**-P(OMe)₃, irrespective of the origin of **2** (either from the solution or the gas-phase synthesis). These products imply a net transfer of oxygen from **2** to P(OMe)₃. The same fragment ion **1** is observed when the CID experiment is performed on labeled [(TPFPP)Cr¹⁸OP(OMe)₃]⁺, from **2**-¹⁸O + P(OMe)₃, indicating loss of ¹⁸OP(OMe)₃. These results support a Cr^{III}-trimethylphosphate coordination structure for the ion, [(TPFPP)Cr^{III}(OP(OMe)₃)]⁺. The observed reactivity with P(OMe)₃ also shows that **2** is a high-valent chromium–oxo complex. Alternative structures that could be conceived for the elemental composition of the sampled ion **2**, for example isomers bearing an O-atom on the porphyrin ligand, would not account for the observed net O-transfer to P(OMe)₃.

Compound **2**-P(OMe)₃ also undergoes a subsequent addition process to give [(TPFPP)Cr^{III}{OP(OMe)₃}₂{P(OMe)₃}]⁺. This is evidence that the complex has a vacant sixth coordination site. In the presence of Me₂S, **2**-P(OMe)₃ undergoes a second addition step, yielding what must be a six-coordinate species with OP(OMe)₃ and Me₂S ligands, [(TPFPP)Cr^{III}{OP(OMe)₃}₂(Me₂S)]⁺ (eq 4).



The reactivity of **2**-P(OMe)₃ has been compared with that of the isobaric species **1**-OP(OMe)₃ generated by electrospray in situ mixtures of **1**-Cl + OP(OMe)₃. Interestingly, this cation **1**-OP(OMe)₃ is also able to undergo addition of one molecule of Me₂S with a reaction efficiency similar to that of reaction 4. Further analogy is found when comparing the CID behavior of

Table 1. Gas-Phase Reactivity of [(TPFPP)Cr^V(O)]⁺ Ions (**2**)^a with Substrates (L) in the Gas Phase

L	product ^b	<i>k</i> _{exp} ^c	Φ (efficiency, %) ^d
NO	NR	—	—
NO ₂	NR	—	—
Me ₂ S	NR	—	—
β-pinene	NR	—	—
pyridine	NR	—	—
γ-picoline	[(TPFPP)Cr(O)(γ-picoline)] ⁺	1.2	8.7
P(OMe) ₃	[(TPFPP)Cr(O)P(OMe) ₃] ⁺ ^e	3.9	35

^a [(TPFPP)Cr^V(O)]⁺ ions were produced either in solution, by treatment of **1**-Cl with C₆H₅IO, or in the gas phase by reaction of **1** with NO₂. ^b NR = no observed reaction. ^c Second-order rate constants in units of 10^{−10} cm³ molecule^{−1} s^{−1}. The estimated error is ±30%, while the internal consistency of the data is within ±10%. ^d Reaction efficiency, Φ = *k*_{exp}/*k*_{coll} × 100. Collision rate constants (*k*_{coll}) evaluated with the parametrized trajectory theory.³⁰ ^e In a consecutive second addition step the [(TPFPP)Cr(O)P(OMe)₃]₂⁺ ion is formed.

(32) Bouchoux, G.; Salpin, J. Y.; Leblanc, D. *Int. J. Mass. Spectrom. Ion Processes* **1996**, *153*, 37–48.

(33) (a) Angelelli, F.; Chiavarino, B.; Crestoni, M. E.; Fornarini, S. *J. Am. Soc. Mass Spectrom.* **2005**, *16*, 589–598. (b) Dunbar, R. C. *Mass Spectrom. Rev.* **2004**, *23*, 127–158.

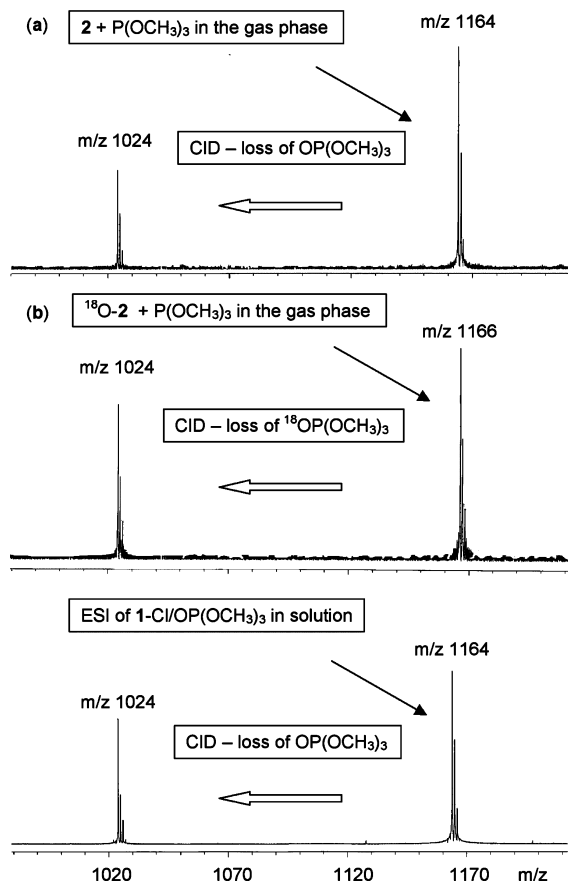


Figure 2. CID spectrum of mass-selected $[(\text{TPFP})\text{Cr}^{\text{III}}(\text{OP}(\text{OMe})_3)]^+$ ($m/z = 1164$) prepared in a FT-ICR mass spectrometer by ion–molecule reactions of **2** (a) and **2**- ^{18}O (b) with $\text{P}(\text{OMe})_3$ and by electrospray ionization of a **1**-Cl + $\text{OP}(\text{OMe})_3$ mixture in acetonitrile (c).

the complex at m/z 1226 obtained by the gas-phase addition of Me_2S to either complex **2**- $\text{P}(\text{OMe})_3$ or **1**- $\text{OP}(\text{OMe})_3$ formed in solution. In both cases, CID of these mixed adducts release **1**- Me_2S by loss of neutral $\text{OP}(\text{OMe})_3$. This is further indication that the m/z 1226 ion is best described as a $[(\text{TPFP})\text{Cr}^{\text{III}}]^+$ ion bound to $\text{OP}(\text{OMe})_3$ and Me_2S axial ligands.

Likewise, when **2** is exposed to a gaseous mixture of $\text{P}(\text{OMe})_3$ and $\text{OP}(\text{OMe})_3$, only the primary addition of the former ligand exhibits a second addition step, whereas $\text{OP}(\text{OMe})_3$ displayed the addition of a just one molecule to ion **2**, to form **2**- $\text{O}-\text{P}(\text{OMe})_3$. This result is inconsistent with the presence of distinct oxo and $\text{P}(\text{OMe})_3$ axial ligands within **2**- $\text{P}(\text{OMe})_3$. Rather, it implies that an O-coupling occurs upon reaction of **2** with $\text{P}(\text{OMe})_3$, such that a sixth coordination site within **2**- $\text{P}(\text{OMe})_3$ is still vacant and may further coordinate a second ligand molecule. However, when the neutral ligand is not a good oxygen atom acceptor, when $\text{L} = \text{OP}(\text{OMe})_3$ or γ -picoline, a six-coordinate addition complex is formed with **2** in which the oxo group and L are distinct axial ligands. Consistently, these complexes do not display any tendency to give a second addition step.

As a further comparison, separately prepared **2**- $\text{P}(\text{OMe})_3$ and **1**- $\text{OP}(\text{OMe})_3$ were each reacted with a 70:30 mixture of $\text{P}(\text{OMe})_3$ and $\text{OP}(\text{OMe})_3$. Once again, **2**- $\text{P}(\text{OMe})_3$ and **1**- $\text{OP}(\text{OMe})_3$ display the same reactivity. Reversible addition reactions are observed, which attain the same equilibrium when starting from either reactant ion (Figure 3). One may thus infer that a sixth coordination site with indistinguishable binding

properties is present in both **2**- $\text{P}(\text{OMe})_3$ and **1**- $\text{OP}(\text{OMe})_3$ under these experimental conditions, and therefore that **2**- $\text{P}(\text{OMe})_3$ and **1**- $\text{OP}(\text{OMe})_3$ are actually the same species.

3.3. Reactions of $[(\text{TPFP})\text{Cr}^{\text{III}}]^+$ (1**) with Oxygen Atom Donors and CID Experiments.** $[(\text{TPFP})\text{Cr}^{\text{III}}]^+$ (**1**) reacts with a variety of small molecules to form addition products $[(\text{TPFP})\text{Cr}^{\text{III}}(\text{L})]^+$ and, in some cases, $[(\text{TPFP})\text{Cr}^{\text{III}}(\text{L})_2]^+$, as summarized in Table 2. Some of these ions are formed by electrospraying an equimolar solution (5 μM) of **1**-Cl and L in MeOH, while others have been formed in the ICR by gas-phase ion–molecule reactions. Many of the ligands are potential oxygen atom donors, such as ethylene oxide which can be deoxygenated to ethylene, but in only two cases was the oxo complex **2** formed. These are the two reactions described above, the solution reaction of **1**-Cl + iodosylbenzene and the gas-phase reaction of **1** + NO_2 (Scheme 1). The latter reaction is characterized by a notable efficiency of 22% (note *d* in Table 2). The potential oxygen atom donors XO are given in the Table in increasing order of X–O bond dissociation enthalpy (BDE), as discussed below.

To understand the nature of the ligand adducts, a number of them were analyzed by collision-induced dissociation, CID. Upon CID the adducts of ethylene oxide, DMSO (Me_2SO), OPPh_3 , and $\text{OP}(\text{OMe})_3$ all dissociated to give the neutral ligand and **1**. A representative CID spectrum of the mass-selected ion **1**- $\text{OP}(\text{OMe})_3$, showing the loss of intact trimethylphosphate, is presented in Figure 2c. In contrast, pyridine *N*-oxide and γ -picoline *N*-oxide undergo CID to yield the oxo complex **2** and (presumably) the deoxygenated ligand. The phenomenological appearance energies for these CID processes are 0.41 and 0.48 eV, respectively. The adduct of **1** with TEMPO (2,2'-6,6'-tetramethylpiperidin-1-oxyl), dissociates into TEMPO^+ and, presumably, neutral $(\text{TPFP})\text{Cr}^{\text{II}}$ as the only product observed.

3.4. Reactivity of Other $[(\text{L}_4)\text{Cr}^{\text{V}}(\text{O})]^+$ and $[(\text{L}_4)\text{Cr}^{\text{III}}]^+$ Cations. Several chromium complexes with tetradentate macrocyclic ligands L_4 other than TPFPP have been tested for potential oxygen transfer reactivity using ESI-FT-ICR. These have involved $\text{L}_4 = 5,10,15,20$ -*meso*-tetraphenylporphyrin (TPP), 3,3'-5,5'-tetra-*tert*-butyl-salen (salen), and *N,N'*-phenylenebis(salicylideneimine) (salphen); see Table S1 in the Supporting Information. As in the TPFPP case, addition of ligands to $[(\text{L}_4)\text{Cr}^{\text{III}}]^+$ results in sequential formation of the mono- and bis-adducts, without formation of the corresponding Cr^{V} oxo complex. The one exception is the reaction of $[(\text{Bu}_4\text{salen})\text{Cr}^{\text{III}}]^+$ with NO_2 to give $[(\text{Bu}_4\text{salen})\text{Cr}^{\text{V}}(\text{O})]^+$, the product of a net O-atom transfer. However, neither $[(\text{salphen})\text{Cr}^{\text{III}}]^+$ nor $[(\text{TPP})\text{Cr}^{\text{III}}]^+$ reacts with NO_2 .

The $[(\text{L}_4)\text{Cr}(\text{O})]^+$ complexes (see Supporting Information) showed no reactivity with potential reductants including Me_2S , $\text{P}(\text{OMe})_3$, γ -picoline, NO, *p*-Me- $\text{C}_6\text{H}_4\text{SMe}$, and β -pinene. The lack of reactions with $\text{P}(\text{OMe})_3$ or alkenes contrasts with the reactivity of these oxo complexes in solution, such as the reported epoxidations.^{15,16} We have also found that, in solution, the salen complex is very reactive with phosphines, rapidly forming Cr^{III} products.³⁴ A distinct effect by donor ligands such as (OPET_3) is also reported to enhance the epoxidation rates and yields displayed by $(\text{salen})\text{Cr}^{\text{V}}(\text{O})$ complexes in solution.^{16a} The reaction of $[(\text{Bu}_4\text{salen})\text{Cr}^{\text{V}}(\text{O})(\text{OPPh}_3)]^+$ probed in the gas phase with a powerful reductant such as *p*- $\text{CH}_3\text{C}_6\text{H}_4\text{SCH}_3$ did not lead to any enhanced reactivity, however, and this complex remains

(34) Warren, J. J.; Mayer, J. M. Unpublished results.

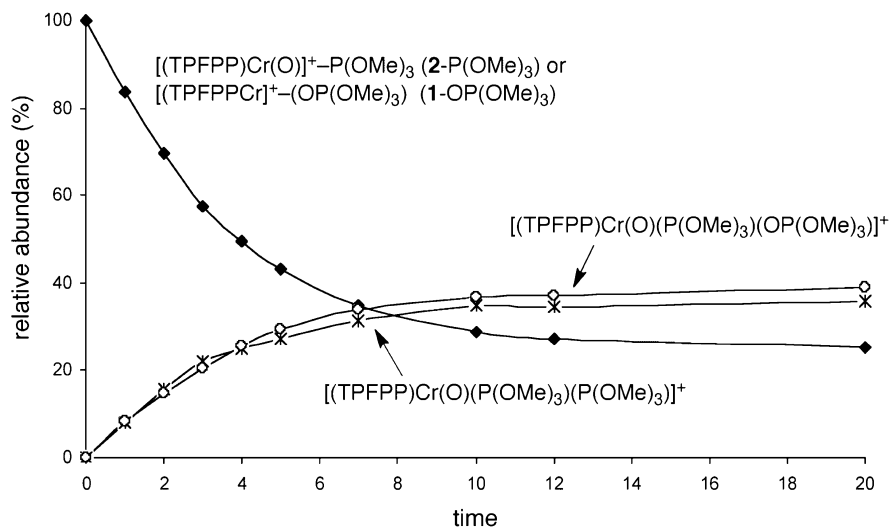


Figure 3. Time dependence of relative ion abundances when either **2**-P(OMe)₃ or 1-OP(OMe)₃ ions are allowed to react with a 70:30 mixture of P(OMe)₃ and OP(OMe)₃ at 7×10^{-8} mbar in the FT-ICR cell. The reactant ion is seen to undergo addition of either one P(OMe)₃ or one OP(OMe)₃ molecule.

Table 2. Products from [(TPFPP)Cr^{III}]⁺ (**1**) + Ligands and Their Behavior upon CID

synthesis ^a	L	product	CID reaction	<i>E</i> _{app} ^b
S ^c	PhIO	2	—	—
S	γ-picoline-O	+ L	1-L → 2 + γ-picoline (90%)	0.48
			1-L → 1 + γ-picoline-O (10%)	0.58
S,G	γ-picoline	+ L	1-L → 1 + γ-picoline	0.55
S	py-O	+ L	1-L → 2 + py (75%)	0.41
			1-L → 1 + py-O (25%)	0.38
S,G	py	+ L	1-L → 1 + py	0.56
G	NO ₂	2 ^d	— ^c	—
G	C ₂ H ₄ O	+ L, + 2 L	1-L → 1 + L	— ^e
			1-L ₂ → 1 -L + L	— ^e
S	DMSO	+ L	1-L → 1 + L	0.43
G	Me ₂ S	+ L, + 2 L	1-L → 1 + L	— ^e
			1-L ₂ → 1 -L + L	— ^e
S,G	TEMPO	+ L	1-L → TEMPO ⁺ + [(TPFPP)Cr ^{III}]	— ^e
S	OPPh ₃	+ L, + 2 L	1-L → 1 + L	0.56
S,G	OP(OMe) ₃	+ L, + 2 L	1-L → 1 + L	0.58
			1-L ₂ → 1 -L + L	—
S,G	P(OMe) ₃	+ L	1-L → 1 + L	0.31

^a S = prepared in solution from 5 μM **1**-Cl + 5 μM LO in MeOH, followed by electrospray; G = formed in the gas phase from **1** + L.

^b Phenomenological appearance energies in eV. ^c The so-formed ion **2** is stable toward CID in the energy range explored. ^d Efficiency 22%. ^e Not available.

unreactive just as [(^tBu₄salen)Cr^V(O)]⁺ (See Table S1, Supporting Information).

4. Discussion

The gas-phase reactivity of several porphyrin- and salen-ligated complexes of chromium have been explored using gas-phase ICR-MS. Most of the work has focused on the tetrakis(pentafluorophenyl)porphyrin (TPFPP) complexes because of their higher reactivity. The results, particularly the reactions with P(OMe)₃ and CID studies of the products, show that complex **2** is the metal–oxo cation, [(TPFPP)Cr^V≡O]⁺. For example, the addition of P(OMe)₃ to **2** gives an adduct that loses OP(OMe)₃ on collisional activation. This adduct is essentially identical to the complex formed from [(TPFPP)Cr^{III}]⁺ (**1**) and OP(OMe)₃; for instance, both species bind P(OMe)₃ and

Table 3. Homolytic X–O Bond Dissociation Enthalpies (BDEs)^a

reaction	Δ <i>H</i> ^o (kcal mol ^{−1}) ^b
PhIO → PhI + O	~26 ^c
γ-picoline → O → γ-picoline + O	62.6
pyO → py + O	63.3
NO ₂ → NO + O	73.2
C ₂ H ₄ O → C ₂ H ₄ + O	84.7
Me ₂ SO → Me ₂ S + O	86.8
Me ₂ NO [•] → Me ₂ N [•] + O	86 ^d
H ₂ NO [•] → H ₂ N [•] + O	91 ^d
Ph ₃ PO → Ph ₃ P + O	139
(EtO) ₃ PO → (EtO) ₃ P + O	148

^a Data from ref 44 unless otherwise noted. ^b Derived from solution data or calculated gas-phase heats of formation. ^c Estimated from data in ref 45. ^d See Supporting Information.

OP(OMe)₃ with the same equilibrium constants. This assignment is similar to chromium(V)–terminal oxo complexes made in solution with porphyrin- and salen-supporting ligands.^{15–17}

Net oxygen-atom transfers have been observed, from NO₂ to **1**, and from **2** to P(OMe)₃. However, **2** shows no observable reactivity with NO, NO₂, Me₂S, β-pinene, or pyridine. This chromium–oxo cation is thus much less reactive than the analogous iron cation, [(TPFPP⁺)Fe^{IV}O]⁺, which reacts rapidly with all of these substrates.^{21a} The difference in reactivity is consistent with the generally much higher oxidizing power of high-valent iron versus chromium compounds as observed in solution.³⁵

4.1. Thermochemistry of Oxygen-Atom Transfer. The oxygen atom donors/acceptors XO/X used in this study have a wide range of homolytic bond dissociation enthalpies (BDEs, Δ*H*^o for XO → X + O), as shown in Table 3.

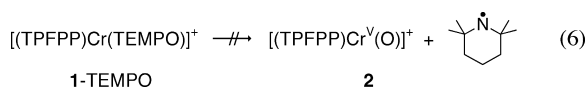
The reaction of NO₂ with [(TPFPP)Cr^{III}]⁺ (**1**) to give [(TPFPP)Cr^V≡O]⁺ (**2**) and (presumably) NO (eq 5) is a particularly simple reaction that provides a lower limit for the Cr^V≡O bond strength. This reaction proceeds with a moderately high efficiency (22% at room temperature, Table 2, footnote *d*) consistent with a negative value of Δ*G*^o.^{32,36} Indeed, definitely exoergic reactions with Δ*G*^o less than −10 kcal mol^{−1} and negligible intrinsic barriers (such as proton-transfer processes)

(35) Sheldon, R. A., Ed. *Metalloporphyrins in Catalytic Oxidations*; Marcel-Dekker: New York, 1994.

proceed with high efficiency (>70%).³² We may then consider that if ΔG° for the reaction of **1** with NO_2 is <0, the bond dissociation free energy (BDFE) should be higher than that of NO_2 , which has a BDFE of 64 kcal mol⁻¹ (see Supporting Information).³⁷ Since the entropy change for reaction 5 is likely to be small,³⁸ we can conclude that the more commonly discussed bond dissociation enthalpy (BDE) of $\text{Cr}^{\text{V}}\equiv\text{O}$ should be greater than that of NO_2 , BDE = 73 kcal mol⁻¹.³⁹ For comparison, the Cr–O BDE in the gas-phase diatomic CrO^+ has been measured to be 86 ± 3 kcal mol⁻¹.⁴⁰

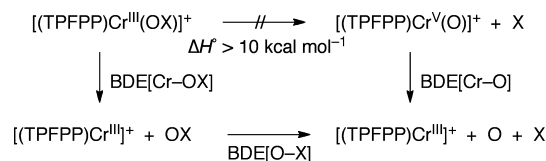


Rough upper limits to the $\text{Cr}^{\text{V}}\equiv\text{O}$ bond strength can be derived from the reduction of **2** by $\text{P}(\text{OMe})_3$ and by the absence of O-atom transfer from TEMPO (2,2',6,6'-tetramethylpiperidin-1-oxyl) to chromium (eq 6), with the assumption that the lack of reaction has a thermochemical rather than kinetic origin. The analyses of these reactions are more complicated than that of eq 5 because the ligand, $\text{OP}(\text{OMe})_3$ or TEMPO, is bound to the Cr^{III} so that the enthalpy of binding plays a role.



The Cr– $\text{OP}(\text{OMe})_3$ binding energy is estimated to be near 13 ± 5 kcal mol⁻¹ based on the CID experiments at variable collision energy (Table 2). The Cr–O bond strength in **1**-TEMPO is taken as equal to the Ti–O BDE in $\text{Cp}_2\text{TiCl}(\text{TEMPO})$, 25 kcal mol⁻¹.⁴¹ This should be an upper limit because Cr is less oxophilic than Ti and because Cr^{IV} is more oxidizing than Ti^{IV} [these compounds are formally $\text{M}^{4+}(\text{TEMPO}^-)$]. In addition, the Cr^{V} sides of these reactions are likely to be entropically favorable by ~ 10 kcal mol⁻¹ ($\Delta S^\circ \approx 34$ cal K⁻¹ mol⁻¹).⁴² The reactions at the top of Scheme 2, with $\text{X} = \text{OP}(\text{OMe})_3$ or TEMPO, should be at least 10 kcal mol⁻¹ enthalpically unfavorable. Finally, the BDEs of $\text{OP}(\text{OMe})_3$ and TEMPO are taken to be the same as those of $\text{OP}(\text{OEt})_3$ and $\text{Me}_2\text{NO}^\bullet$, respectively (Table 3). With these various values, Scheme 2

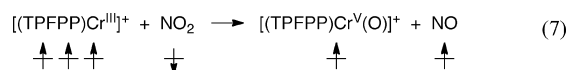
Scheme 2. Thermochemical Analysis of the Lack of Oxygen Atom Transfer from **1**- $\text{OP}(\text{OMe})_3$ or **1**-(TEMPO)



implies that the $\text{Cr}^{\text{V}}\equiv\text{O}$ BDE must be less than ~ 143 kcal mol⁻¹ to account for the observed reaction of **2** with $\text{P}(\text{OMe})_3$. Similarly, the lack of deoxygenation of TEMPO[•] by **1** suggests that $\text{BDE}(\text{Cr}^{\text{V}}\equiv\text{O}) \lesssim 101$ kcal mol⁻¹. These thermochemical considerations are consistent with the estimated BDE for $\text{Cr}^{\text{V}}\equiv\text{O}$ in ion **2** to be higher than 73 kcal mol⁻¹ based on the reaction of **1** with NO_2 .

4.2. Effect of Spin on the Oxygen-Atom Transfer Reactions. It is surprising that none of the XO or X reagents with BDEs between NO_2 and $\text{OP}(\text{OMe})_3$ (73.2 and ~ 148 kcal mol⁻¹, Table 3) are observed to interconvert **1** and **2**. The loss of ethylene from **1**• $\text{C}_2\text{H}_4\text{O}$ should be only slightly less favorable than the loss of NO from the putative **1**• NO_2 intermediate in the NO_2 reaction. However, not only does this process not occur, but **2** is also unable to release an O atom to an activated olefin such as β -pinene. Similarly, the **1**•DMSO adduct does not fragment into **2** and Me_2S when activated to by CID, displaying no evidence for an O-atom transfer from sulfur to **1**. Once again the reverse reaction is not observed either, because **2** appears to be unreactive toward gaseous Me_2S . The pyridine *N*-oxides py-O and γ -picoline-O are among the strongest oxidizing agents in the set of species that were considered (Table 3). Both pyridine *N*-oxide complexes **1**•Opy and **1**•O- γ -picoline share similar E_{app} values upon CID at variable energy, leading to either **2** (75–90%) or **1** (25–10%) in comparable amounts. This finding of only partial oxidation of **1** by relatively powerful oxidizing agents, stronger than NO_2 which does efficiently oxidize **1** in the gas phase, suggests that kinetic effects play an important role in the partitioning of the O atom between **1** and the pyridine nitrogen, and, more generally, in O-atom transfer reactions involving **1** and **2**.

As noted above, the interconversion of **1** + XO with **2** + X is formally spin-forbidden with closed-shell (singlet) reagents X and XO. All but two of the X/XO reagents used here are closed-shell species, including ethylene oxide/ethylene, pyO/py, $\text{Me}_2\text{SO}/\text{Me}_2\text{S}$, and $\text{OP}(\text{OMe})_3/\text{P}(\text{OMe})_3$. The two exceptions are NO_2 and TEMPO, which are both free radicals with an unpaired electron and are therefore doublet states. Reaction of quartet **1**¹⁹ with doublet NO_2 can therefore give doublet **2** + doublet NO in a spin-allowed reaction, as shown in eq 7 (the orbitals underneath the reagents show only the unpaired electrons). Reaction 7 formally involves transfer of a triplet oxygen atom NO_2 to Cr^{III} . This analysis is consistent with the measured efficiency of 22% for reaction 7.



TEMPO can react with **1** in a similar spin-allowed fashion.⁴³ Therefore the absence of O-atom transfer from TEMPO to

- (36) (a) Meot-Ner, M. *J. Am. Chem. Soc.* **1982**, *104*, 5–10. (b) Bucher, H.-H.; Grutzmacher, H.-F. *Int. J. Mass Spectrom. Ion Processes* **1991**, *109*, 95–104. (c) Baciocchi, E.; Bietti, M.; Chiavarino, B.; Crestoni, M. E.; Fornarini, S. *Chem.-Eur. J.* **2002**, *8*, 532–537.
- (37) Afeefy, H. Y.; Liebman, J. F.; Stein, S. E. Neutral Thermochemical Data. In NIST Chemistry WebBook, NIST Standard References Database Number 69, Linstrom, P. J.; Mallard, W. G., Eds.; National Institutes of Standards and Technology, Gaithersburg, MD, 20899, <http://webbook.nist.gov/chemistry/> (accessed 9 October 2009).
- (38) Gas-phase $\text{A} + \text{B} \rightarrow \text{C} + \text{D}$ reactions typically have small entropy changes. In this case, the S° values for NO_2 and NO (57.37, 50.37 cal K⁻¹ mol⁻¹) differ only by 7 cal K⁻¹ mol⁻¹. For a discussion of entropies in hydrogen-atom transfer reactions, including cases where $\Delta S^\circ \neq 0$, see: (a) Mader, E. A.; Manner, V. W.; Markle, T. F.; Wu, A.; Franz, J. A.; Mayer, J. M. *J. Am. Chem. Soc.* **2009**, *131*, 4335–4345. (b) Mader, E. A.; Davidson, E. R.; Mayer, J. M. *J. Am. Chem. Soc.* **2007**, *129*, 5153–5166.
- (39) For $\text{XO} \rightarrow \text{X} + \text{O}$, when $S^\circ(\text{XO}) = S^\circ(\text{X})$ the gas-phase BDFE will be smaller than the gas-phase BDE by $TS^\circ(\text{O}) = (298 \text{ K}) (38.49 \text{ cal K}^{-1} \text{ mol}^{-1}) = 11.5 \text{ kcal mol}^{-1}$.
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- (41) Acree, W. E., Jr.; Pilcher, G.; Ribeiro da Silva, M. D. M. C. *J. Phys. Chem. Ref. Data* **2005**, *34*, 553–572.
- (42) Conry, R. R.; Mayer, J. M. *Inorg. Chem.* **1990**, *29*, 4862–4867. The BDE(PhI–O) of 26 kcal is estimated from the solution thermodynamic oxygen-transfer potentials (TOP) in this reference, assuming constant solvation effects on the XO/X couples.

- (43) Huang, K.-W.; Han, J. H.; Cole, A. P.; Musgrave, C. B.; Waymouth, R. M. *J. Am. Chem. Soc.* **2005**, *127*, 3807–3816.
- (44) Conry, R. R.; Mayer, J. M. *Inorg. Chem.* **1990**, *29*, 4862–4867.
- (45) The initial adduct **1**-TEMPO is probably best described as $\text{Cr}^{\text{IV}}(\text{TEMPO}^-)$ and is probably a spin triplet. This species can then dissociate a doublet tetramethylpiperidiny radical to give **2**.

chromium (eq 6) is likely due to the reaction being endoergic, rather than due to a spin-based spin barrier. The thermodynamic values given above support this conclusion: the Cr≡O BDE is indicated to be only a little larger than that the N–O BDE in NO₂ (73 kcal mol^{−1}), and the N–O BDE in TEMPO is estimated as being significantly stronger, close to that in Me₂NO[•], namely 86 kcal mol^{−1} (Table 3). Hence, the reaction of **1** with TEMPO is likely to be unfavorable based on thermodynamic considerations.

The reactivity pattern of **1** and **2** discussed above appears to be a case where the spin state change between Cr^{III} and Cr^V affects the reaction chemistry. The most striking feature of these chromium complexes is their low reactivity in either reaction direction. One notable exception is the facile and spin-allowed reaction of **1** + NO₂. The other exception is the reaction of **2** + P(OMe)₃ to give (TPFPP)Cr^{III}(OP(OMe)₃)⁺. This reaction is so highly exoergic that it could proceed in a spin-allowed fashion to initially form the ²E excited state of the Cr^{III} product. The P(OMe)₃ reaction is >60 kcal mol^{−1} downhill, using the conservative upper limit on BDE(Cr^V≡O) from the lack of deoxygenation of TEMPO. This is much larger than the estimate of the energy of the ²E state by analogy with [Cr^{III}(NH₃)₆]³⁺, ~37 kcal mol^{−1} (as noted above). Attempts in solution to observe emission from such a potential ²E intermediate (chemiluminescence), potentially generated thermally upon reaction of the (salen)Cr^V(O)⁺ derivatives with phosphines, have unfortunately not been successful.³⁴ Consistent with the suggestion that the P(OMe)₃ reaction could proceed via a doublet Cr^{III} excited state, this reaction has a rate constant very similar to that of the analogous reaction of gas-phase (TPFPP⁺)Fe^{IV}O with trimethylphosphite. The iron reaction proceeds with somewhat higher efficiency and proceeds with full oxygen atom transfer to give [(TPFPP)Fe^{III}]⁺, while **2** forms the Cr^{III} adduct (TPFPP)Cr^{III}(OP(OMe)₃)⁺, perhaps because of the more substitution-inert character of Cr^{III} complexes.

5. Conclusions

High-valent chromium(V)–oxo complexes have been obtained from chromium(III) precursors both by an oxidation reaction in solution, and by a gas-phase reaction using NO₂ as an oxygen-atom donor. Among several sampled species, the [(TPFPP)Cr^VO]⁺ ion (**2**) from the oxidation of the [(TPFPP)Cr^{III}]⁺ (**1**) has been the most thoroughly studied. The presence of a Cr^VO unit in **2** has been probed by a variety of processes in the gas phase, including O-atom transfer to potential reductants and the assessment of association equilibria. These tools allow us to establish that the complex formed in solution using PhIO and then transferred to the gas phase by ESI, and the species obtained by the gas-phase reaction of **1** with NO₂, are actually the same chromium(V)–oxo cation [(TPFPP)Cr^VO]⁺ (**2**). Gas-phase FT-ICR mass spectrometry has allowed us to investigate the gas-phase reactivity of these cationic Cr species.

The major purpose of this study was aimed at exploring the role of spin state constraints on the O-atom transfer reactivity connecting [(TPFPP)Cr^{III}]⁺ (**1**) and [(TPFPP)Cr^VO]⁺ (**2**). To this end, gas-phase reactivity data have been collected regarding both the bimolecular kinetics for the reaction with selected gaseous neutrals, and the fragmentation patterns observed when adducts of **1** and selected ligands undergo CID at variable energy. The

reactivity behavior of the **1/2** couple has been interpreted with reference to the thermodynamic parameters of the reactions, namely to the X–O bond dissociation free energies of the XO/X neutral reaction partners. The gas-phase reaction of **1** with NO₂ sets an approximate value for the Cr^V≡O BDE of **2**. The efficiency of this reaction and the absence of the reverse reaction (**2** + NO) indicate a Cr^V≡O BDE somewhat larger than the O–NO BDE. In agreement with this evaluation, no O-atom transfer is observed with TEMPO, whose N–O BDE is definitely higher. Both NO₂ and TEMPO are free radicals with one unpaired spin, and therefore they can transfer an oxygen atom to **1** to give **2** in a spin-allowed process.

Oxygen-atom transfer reactions of the **1/2** couple with closed-shell XO/X neutrals, however, are formally spin forbidden. In these cases, the O-atom transfer reactivity is found to be abnormally low. Compound **2** does not oxidize olefins, sulfides, or pyridines, and the corresponding epoxides, sulfoxides, and pyridine *N*-oxides are not deoxygenated by **1**. Thus, there are substantial kinetic impediments to these reactions, and it seems likely that this is due to the requirement of a spin change from doublet **2** to quartet **1**. Of the various closed-shell reagents explored, only trimethylphosphite reacts with **2** to give a Cr^{III} product ([(TPFPP)Cr^{III}(OP(OMe)₃)]⁺ (**1**-OP(OMe)₃), albeit with a relatively modest efficiency of 35%. P(OMe)₃ has a very high affinity for an oxygen atom, with an estimated (MeO)₃P–O BDE of >140 kcal mol^{−1}, so that its reaction with **2** could proceed in a spin-allowed fashion to form the ²E excited state of the Cr^{III} product. Thus, all of the data are consistent with spin-forbidden reactions of **1** and **2** being inhibited while thermodynamically favorable spin-allowed reactions occur with reasonable efficiency.

As a final remark, the relative reactivity of [(TPFPP)–Metal(O)]⁺ complexes in the gas phase may now be compared with regard to the oxygen-atom transfer reaction with olefins. While the iron complex reacts with 2-butenes with ~1% efficiency^{21a} and the manganese complex with 0.8–8% efficiency,^{21c} depending on the configuration at the double bond, the corresponding chromium complex is unreactive, even with a highly activated olefin such as β-pinene. Dissecting the thermodynamic and kinetic contribution to the observed reactivity behavior is not straightforward. Moreover the mechanism of the alkene epoxidation by compound **1** and model species involves intermediates of different nature and presents TSR and multistate reactivity scenarios.⁶ A more complete analysis of these issues will require computational examination, as well as more experimental studies.

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Supporting Information Available: Derivation of hydroxylamine BDE and BDFE, analysis of reactions with TEMPO, Table S1. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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