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An ESIPT-based fluorescent probe with fast-response for detection of hydrogen sulfide in mitochondria



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Yuting Du^{*}, Hongliang Wang, Ting Zhang, Wei Wen, Zhiying Li, Minjie Bi, Juan Liu

Department of Chemistry, Xinzhou Teachers University, Xinzhou, Shanxi 034000, China

HIGHLIGHTS

• HBTP-H₂S reacts with H₂S via nucleophilic addition reaction.

• HBTP-H₂S is an excited-state intramolecular proton transfer (ESIPT) fluorescent probe for rapidly distinguish H₂S.

• HBTP-H₂S exhibits ultrafast-response (within 30 s), good selectivity and large Stokes shift (188 nm) for detection of H₂S.

• HBTP-H₂S has been successfully applied for imaging of mitochondrial H₂S in live cells.

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ABSTRACT

Excited-state intramolecular proton transfer (ESIPT) has recently received considerable attention due to its dual fluorescent changes and large Stokes shift. Hydrogen sulfide (H_2S) is a gas signal molecule that plays important roles in modulating the functions of different systems. Herein, by modifying 2-(2-hydroxyphenyl) benzothiazole (HBT) scaffold, a novel near-infrared mitochondria-targeted fluorescent probe **HBTP-H₂S** has been rationally designed based on excited-state intramolecular proton transfer (ESIPT) effect. The nucleophilic addition reaction of the H₂S with probe **HBTP-H₂S** caused the break of the conjugated skeleton, resulting the shifting of maximum emission peak from 658 nm to 470 nm. **HBTP-H₂S** showed fast-response response time, good selectivity and a large Stokes shift (188 nm) toward H₂S. Most importantly, inspired by the inherent advantages of the probe, **HBTP-H₂S** was successfully employed to monitor mitochondrial H₂S in HepG2 cells.

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1. Introduction

Hydrogen sulfide (H_2S) is a colorless gas with foulodor of rotten egg and has been confirmed as gas signal molecule[1-3]. In mammals, mostly H_2S is generated by using sulfur-containing amino acids as a substrate that is catalyzed by enzymes[4]. Moreover, H_2S has been identified as a gasotransmitter, which regulats numerous important physiological functions such as nervous, cardiovascular and immune systems[5-7]. However, the abnormal levels of H_2S lead to many human diseases including diabetes[8], Alzheimer's disease[9], liver cirrhosis[10] and the symptoms of Down's syndrome [11-13]. In addition, higher levels of H_2S promote tumor proliferation and growth in cancer cells[14-17].

Mitochondria is the most capital organelle in eukaryotic cells and a major source of intracellular reactive sulfur species [18,19]. Mitochondrial H_2S plays a critical role in the physiological and

* Corresponding author. E-mail address: yutingdu123@163.com (Y. Du). pathological processes[20-23]. Hence, developing efficient tools to monitor mitochondrial H_2S levels will be promising for further exploring its pathophysiological roles.

The traditional methods of H₂S mainly contain sulfide selective electrodes^[24,25], gas chromatography^[26,27], colori-metric [28,29] and methylene blue assay[30]. However, the processes of treating samples are complicated and these technology are not capable of real-time monitor the H₂S. Therefore, various organic fluorescent probes have been explored in recent years owing to the advantages such as handy staining processes, flexible molecular design strategies and high sensitivity[31-33]. Accordingly, a lot of fluorescent probes for H₂S have been developed [34-36], most of which are based on the chemical properties of H₂S, such as its reducing properties[37-47], high binding affinity towards copper ions[48,49], specific nucleophilicity[50-58], thiolysis of 7-nitro-1,2,3-benzoxadiazole (NBD) moiety[59-62] and azide reduction[63-66]. Most probes cover either intramolecular charge transfer (ICT) [67] or photoinduced electron transfer (PET) processes, which have displayed only small stokes shifts or relatively weak fluorescence. Compare with ICT and PET processes, excited state intramolecular proton transfer (ESIPT) process recently received considerable attention due to its large stokes shift[68,69], improved sensitivity[70,71], high fluorescence quantum yield and dual fluorescence intensity changes[72,73]. Therefore, the development of ESIPT-based fluorescent probes for H₂S is valuable. Moreover, because the ESIPT process is often inhibited in polar and hydrogen bond donating solvents. Thus, mosts of fluorescent probes which based on ESIPT usually works with the assistant of some surfactants such as cetyltrimethylammonium bromide (CTAB) that creates a sufficiently hydrophobic environment[68]. However, due to the toxicity of surfactants at high concentrations, thus, the use of surfactants may limit the biological application of ESIPT-based probes.

A few fluorescent probes that based on ESIPT for detection of H_2S have been reported[74-77] (see the Supplementary Table S1). However, it remains a challenge to develop fluorescent probes with excellent sensitivity and high selectivity for detection of mitochondrial H_2S . Additionally, since H_2S has a strong nucle-ophilic addition reaction property, it can change the fluorescence after reacting with the probe **HBTP-H_2S(**(E)-4-(3-(benzo[d]thia zol-2-yl)-2-hydroxy-5-methylstyryl)-1-(2-hydroxyethyl)pyridin-1 -ium) (Scheme 1). However, there are hardly probes to detect H_2S by the reaction mechanism of nucleophilic addition[78].

Accordingly, inspired by previous works^[74-77], we designed and synthesized HBTP-H₂S to detect H₂S which used HBT (2-(2'-Hydroxyphenyl) benzothiazole) as an ESIPT fluorophore[79] and pyridinium scaffold as mitochondrial targetable group without the need of introducing CTAB (Scheme 1). Due to the strong electron-pulling ability of the pyridine salt, the fluorescence of HBTP-H₂S showed red emission (658 nm). Meanwhile, pyridinium scaffold has strong electrophilic ability. While HBTP-H₂S undergoes a nucleophilic addition reaction with H₂S (a stronger nucleophile), which resulting the conjugate structure of HBTP-H₂S destroyed and leads to the shifting of maximum emission peak from 658 nm to 470 nm. Furthermore, the probe was with fastresponse response time (within 30 s) in monitoring mitochondrial H₂S and excellent selectivity. Most importantly. HBTP-H₂S exhibites near-infrared fluorescence characteristics, which is advantageous in biological applications.

2. Experiment section

2.1. Instruments

Deionized water was used during all experiments. ¹H NMR and ¹³C NMR spectra were carried out on a 600 MHz Bruker ADVANCE III spectrometer. High-resolution mass spectra were collected on a High-Resolution Mass Spectrometer (HRMS, Bruker APEX II 47e mass spectrometer). The UV–vis absorption spectra were measured by using an Evolution 220 spectrometer (Thermo Fisher Scientific). Fluorescent spectra and the absolute quantum yield ($\Phi_{\rm fl}(\%)$) of **HBTP-H₂S** were captured on FS5 spectrofluorometer. Unless otherwise stated, all excitation and emission bandwidths

were both set at 2 nm, respectively. Images were acquired using Leica Fluorescent Microscope (DM4000B) with \times 20 objective lens.

2.2. Synthesis of HBTP-H₂S

The probe **HBTP-H₂S** was readily synthesized by using simple and cheap raw materials. The details of synthesis procedure and the characterization of compounds were included in the Supporting Information.

HBTP-H₂S, Greenyellow solid; ¹H NMR (600 MHz, DMSO *d*₆, ppm) δ 13.05 (s, 1H), 8.85 (d, *J* = 6.9 Hz, 2H), 8.23 (d, *J* = 7.0 Hz, 2H), 8.21–8.17 (m, 1H), 8.12 (d, *J* = 16.4 Hz, 1H), 8.08 (dt, *J* = 8.1, 0.8 Hz, 1H), 7.79–7.74 (m, 2H), 7.66 (d, *J* = 16.4 Hz, 1H), 7.58 (ddd, *J* = 8.3, 7.1, 1.3 Hz, 1H), 7.50 (ddd, *J* = 8.2, 7.2, 1.2 Hz, 1H), 5.24 (t, *J* = 5.3 Hz, 1H), 4.56 (t, *J* = 5.0 Hz, 2H), 3.84 (q, *J* = 5.1 Hz, 2H), 2.36 (s, 3H); ¹³C NMR (150 MHz, DMSO *d*₆, ppm): δ 168.66, 154.46, 153.46, 151.26, 145.17, 135.42, 132.99, 132.88, 131.16, 129.55, 127.60, 126.51, 124.97, 124.08, 123.94, 122.87, 122.40, 117.60, 62.54, 60.50, 20.40; HRMS (ESI) Calcd. for C₂₃H₂₁O₂N₂S (M + H)⁺: 389.1314, Found 389.1318.

2.3. Spectral properties of HBTP-H₂S

The stock solutions of **HBTP-H₂S** (10 mM) were prepared in DMSO. Sodium sulfide (Na₂S·9H₂O) was diluted to 100 mM by double distilled water as the source of H₂S. The fluorescence and UV/ vis spectra were obtained after a certain amount of Na₂S or another analyte was added to a solution of probe. The test solutions of **HBTP-H₂S** (5 μ M) were prepared by diluting the stock solution and all the spectra experiments were measured in CH₃CN / PBS buffer (v/v = 1:1, 10 mM, pH = 7.4). The selectivity of **HBTP-H₂S** for biothiols (Cys, GSH and Hcy), Thioacetamide (TAA) and different anions (Br⁻, Cl⁻, F⁻, I⁻, SO₄²⁻, SO₃²⁻, NO₃⁻, NO₂⁻, S₂O₃²⁻, CN⁻) were evaluated by the fluorescent response with addition of excess biological species (3 mM). All solutions of the anions were produced in deionized water.

2.4. Cytotoxicity of HBTP-H₂S

HepG2 (4 \times 10⁴ cells/well) cells were seeded in 96-well plates for 24 h followed by replacement of the culture medium with fresh medium and then incubation of the cells with **HBTP-H₂S** for another 24 h at 37 °C. Subsequently, 0.5 mg/mL of MTT solution was added, and the cells were incubated for another 4 h at 37 °C. The OD values were measured by microplate reader (Bio-Rad M680, USA) at 570 nm.

2.5. Cell culture and fluorescent imaging

HepG2 cells were obtained from the Shanghai Biochemistry and Cell Biology Cell Bank of the Chinese Academy of Sciences. HepG2 cells were cultured in RPMI medium 1640 supplemented with 10% FBS (fetal bovine serum), 1% penicillin and 1% streptomycin in an atmosphere of 5% CO₂ at 37 °C. For cells imaging studies, HepG2



Scheme 1. Design of fluorescent probe HBTP-H₂S and its sensing mechanism for H₂S.

cells were seeded in 6-well plates (4 \times 10⁵ cells/well) for 24 h and replaced the old medium with fresh medium. The cells were preincubated with Na₂S (250, 500 or 1000 μ M, respectively) for 30 min and treated with **HBTP-H_2S** (10 μ M) for another 30 min. For imaging of H₂S in mitochondria, the cells were incubated with 500 nM Mito-Tracker Green for 30 min, then incubated cells with 5 μ M **HBTP-H_2S** for another 30 min. The cells were washed with PBS three times to wash away the excess probe and imaged with fluorescent microscope. Fluorescent imaging was acquired by fluorescent microscope (Leica DM 4000B, United States).

3. Result and discussion

3.1. Synthesis of HBTP-H₂S

HBTP-H₂S was prepared by following the synthesis procedure shown in Scheme 2[80-82]. In brief, Compound 1 was synthesized through cyclization of 5-methylsalicylaldehyde and 2-aminobenzenethiol. Then compound 1 was used to react with hexamethylenetetramine to obtain compound 2. Finally, compound 2 was reacted with 1-(2-hydroxyethyl)-4-methylpyridin-1-ium to get the probe **HBTP-H₂S**.

3.2. Fluorescent response of HBTP-H₂S to H₂S and the detection mechanism

Firstly, we examined the photostability of the probe HBTP-H₂S. As shown in Fig. S1, the probe was stable for 30 min. Simultaneously, under the visible light or UV-light (254 nm) for 30 min, HBTP-H₂S was always stable, suggesting that HBTP-H₂S is photostability extracellularly. Then, the response of the HBTP-H₂S to H_2S were studied in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/ v). As seen in the Fig. 1A, the probe **HBTP-H₂S** exhibited two peaks at about 314 and 400 nm. When the probe reacts with Na₂S, it causes the two peaks red shifted at about 380 and 584 nm respectively. For the macroscopic performance, the color changed from faint yellow to peach, indicating that a new product might have been generated. At the same time, an sharply decreased was observed at 658 nm while a new emission band at 470 nm was appeared in the case of HBTP-H₂S add Na₂S (Fig. 1B). Notably, Na₂S trigged a large Stokes shift (188 nm), which enabled the ratiometric detection of by measuring the fluorescence intensity ratio F_{658}/F_{470} . As indicated above, **HBTP-H₂S** could serve as an efficient fluorescent probe for sensing of H₂S.

Considering the solvent effect of ESIPT mechanism, the fluorescent spectra and the absolute quantum yield ($\Phi_{\rm fl}(\%)$) of HBTP-H₂S

were also examined in different solvent such as acetonitrile, dimethyl sulfoxide, methanol and tetrahydrofuran. These results illustrated that **HBTP-H₂S** exhibited a strong keto-form emission in acetonitrile based on an ESIPT process, but a weakened enolform emission in ethanol and methanol (Fig. S2). Because methanol can form intermolecular hydrogen bonds with **HBTP-H₂S**. Meanwhile, the absolute quantum yield ($\Phi_{\rm fl}(\%)$) of **HBTP-H₂S** in polar solvent is high nonpolar solvent except dimethyl sulfoxide (Table S2), suggesting that nonpolar solvent and dimethyl sulfoxide showed that disadvantage of ESIPT process because they were no hydrogen protons. Additionally, due to the strong electronpulling ability of the methyl pyridine salt, the fluorescence of **HBTP-H₂S** showed only keto tautomer emission (658 nm) in PBS/ CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v).

To further verify our putative reaction mechanism. 5 uM probes were incubated with 1 mM of Na₂S in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v) for 10 min and detected by using HRMS mass and ¹H NMR spectrometry. We discovered that a new peak appeared at M/Z 423.11 [M]⁺ (Fig. S3), proving that the mass peak was owing to the molecular of the compound HBTP-SH. Moreover, the main product of the structure was characterized via NMR spectra (Fig. S4). When we add excessive Na₂S, almost all proton signals from the pyridinium scaffold conjugated system underwent an obvious up-field shift, this shift is consistent with the proposed product where the electron-withdrawing ability of indolenium N⁺ has disappeared. Meanwhile, we assumed that HS⁻ addition to the para-position of $C = N^+$ and the vinyl protons at 8.099 (Hb) and 7.537 (Ha) were upfield shifted to 8.082 (Hb) and 7.441 (Ha) upon Na₂S addition at room temperature. This mechanism is similar to the previously reported detection of H₂S[55,78,83] (Scheme S2). The above results support the original intention of our design (Scheme 1).

3.3. Sensitivity and response time of the probe HBTP-H₂S towards H₂S

The capability of **HBTP-H₂S** to sense H₂S were measured in PBS/ CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v). Studies on titration experiments showed that with the rise concentration of Na₂S, the ratio F_{658}/F_{470} showed an excellent linear relation to H₂S (Fig. 2A) and the linear equation was concluded as y = -0.1392x + 36.1807 (R² = 0.99977). The detection limit for H₂S was determined to 8.5 µM according to the 3 σ /k method. Additionally, pyridinium scaffold of **HBTP-H₂S** has strong electrophilic ability. Simultaneously, H₂S has a lower pKa value than other biothiols (H₂S: ~7.0 [84]; Cys: ~8.3; GSH: ~9.2), therefore, it is a stronger nucleophile and small steric hindrance. Hence, the detection limit for H₂S was



Scheme 2. Synthetic route of the probe HBTP-H₂S.



Fig. 1. Optical response of **HBTP-H₂S** to H₂S. (A) UV-vis absorption spectra of **HBTP-H₂S** (5 μ M) before and after reaction with Na₂S (1 mM) for 10 min in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v). Insert images were taken under sunlight. (B) Fluorescent spectra of **HBTP-H₂S** (5 μ M) before and after reaction with Na₂S (1 mM) for 1 min in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v). Insert images were taken under sunlight. (B) Fluorescent spectra of **HBTP-H₂S** (5 μ M) before and after reaction with Na₂S (1 mM) for 1 min in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v). Aex = 380 nm, slit width (ex/em) = 2/2 nm.



Fig. 2. (A) The linear relationship between the ratio F_{658}/F_{470} of **HBTP-H₂S** (5 μ M) and Na₂S (0–80 μ M). (B) Time-dependent fluorescent spectra of **HBTP-H₂S** (5 μ M) in the presence of 1 mM Na₂S. All spectra were measured in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v). λ ex = 380 nm, slit width (ex/em) = 2/2 nm.

as low as 8.5 μ M. It was reported that probes can detect H₂S with fast response time[85,86]. Therefore, the time-dependent fluorescent spectrum were performed to determine the reaction time of -**HBTP-H₂S** with Na₂S in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/ v). As the reaction progressed, **HBTP-H₂S** displayed emission at maximum of 470 nm with fast response which reached to the saturated intensities within 30 s and finished within 1 min, suggesting the feasibility for rapid detection of H₂S (Fig. 2B).

3.4. Selectivities and pH response of the probe HBTP-H₂S towards H_2S

Subsequently, the selectivity of **HBTP-H₂S** towards H₂S were investigated by the fluorescent changes in the presence of various interfering species including thiol compounds (Cys, GSH, and Hcy), TAA and different anions (Br⁻, Cl⁻, F⁻, I⁻, SO²₃⁻, NO₃, NO₂, S₂O²₃⁻, CN⁻) at 3 mM. According to the results Fig. 3A, except SO²₃⁻ caused negligible spectral changes, none of the species causes an obvious fluorescent change, suggesting that the probe can detect H₂S with high selectivity. Notably, the classical nucleophile CN⁻ induces almost no change of the ratio; the protonation of CN⁻ in neutral medium might be the origin, since HCN possess a pKa about 9.2 [87]. Moreover, endogenous H₂S has been found in high concentrations (10 to 600 μ M) in the brain of bovine, rat, and human[88]. Next, we attempted to examine the effect of pH on the fluorescent properties of **HBTP-H₂S**. It was shown that Fig. 3B, the probe itself was not significantly varied the ratio F_{658}/F_{470} in the pH range of 5–

10, illustrating that **HBTP-H₂S** is stable in this pH range. Meanwhile, because the mitochondrial pKa is approximately 8.8, therefore OH⁻ would not react with the probe at the same condition. Whereas the pH ranges from 5 to 10 in the presence of Na₂S, the ratio F_{658}/F_{470} showed decreasing. This is because that ESIPT mechanism is sensitive to the changes of pH. In brief, the probe under physiological conditions can ensure the ratio F_{658}/F_{470} when detecting of H₂S, suggesting the potential of the probe in biological research.

3.5. Biological applications

Inspired by the in vitro results described above, **HBTP-H₂S** was applied for the detection of H₂S in HepG2 cells. Firstly, the cytotoxicity of **HBTP-H₂S** was examined by the MTT assay, we discovered that 10 μ M probe was non-cytotoxic and thus used in the following experiments (Fig. S5). As shown in the Fig. 4A, HepG2 cells were directly incubated with **HBTP-H₂S** for 30 min, we observed that the light red fluorescence of **HBTP-H₂S**, suggesting that the content of H₂S is very little in cells. Next, the exogenous H₂S donor sodium sulfide (250, 500 and 1000 μ M) was added, we found that with the increase in the concentration of Na₂S, the fluorescent intensity decreased in a certain concentration dependent.

Because $HBTP-H_2S$ contains a positively charged pyridinium moiety and the mitochondrial membrane potential is negative, we further tested its ability to localize in mitochondria by means



Fig. 3. (A) Changes in the ratioratio F_{658}/F_{470} of **HBTP-H₂S** (5 μ M) in the presence of various relevant anions (3 mM) or biothiols (Cys, Hcy, GSH, 3 mM) after incubating for 10 min in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v) at room temperature. (1) blank (the ratioratio F_{658}/F_{470} without the probe), (2) Na₂S, (2) Cys, (3) GSH, (4) Hcy, (5) Cl⁻, (6) F⁻, (7) Br⁻, (8) I⁻, (9) NO₂⁻, (10) NO₃⁻, (12) S₂O₃⁻, (12) S₂O₃⁻, (12) CN⁻, (18) TAA. (B) pH effect on the ratio of F_{658}/F_{470} of **HBTP-H₂S** (5 μ M) in the absence and presence of Na₂S (1 mM). All experiments were performed in PBS/CH₃CN buffer (10 mM, pH = 7.4, 1/1, v/v) at room temperature.



Fig. 4. Fluorescent images of H₂S in HepG2 cells. (A) Cells were stained with 10 μM **HBTP-H₂S**, and then incubated with 250, 500 and 1000 μM Na₂S for another 30 min before imaging. (B) The cells were incubated with 500 nM Mito-Tracker Green for 30 min after incubated cells with 5 μM **HBTP-H₂S** for another 30 min. Scale bar: 50 μm.

of the co-localization experiments. The cells were preincubated with 500 nM Mito-Tracker Green (a commercially available mitochondrial dye) for 30 min and then incubated with 5 μ M **HBTP-H**₂**S** for another 30 min. Fig. 4B showed that the strong yellow fluorescence of **HBTP-H**₂**S** overlapped very well with the Mito-Tracker Green. The overlap coefficient is being 0.88. These results strongly support that the probe **HBTP-H**₂**S** can selectively and efficiently localize in to mitochondria.

4. Conclusion

To sum up, a new NIR fluorescent probe **HBTP-H₂S** was developed, which is capable of imaging H₂S of mitochondrial in HepG2 cells. The probe was composed of HBT (2-(2'-Hydroxyphenyl) benzothiazole) and pyridinium scaffold and detected mitochondrial H₂S in HepG2 cells. Under the physiological conditions, **HBTP-H₂S** exhibits fast-response (within 30 s), large Stokes shift (188 nm), higher sensitivity and higher selectivity towards H₂S than other levels of thiol compounds. Furthermore, **HBTP-H₂S** is also used to bioimage mitochondrial H₂S in living cells.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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Appendix A. Supplementary material

Supplementary data to this article can be found online at https://doi.org/10.1016/j.saa.2021.120390.

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