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# Slow Exchange of Bidentate Ligands between Rhodium(I) Complexes: Evidence of Both Neutral and Anionic Ligand Exchange

**Si Chen,[a,b] Eric Manoury,[a,b] and Rinaldo Poli*[a,b,c]**

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The phosphine double exchange process involving [RhCl(COD)(TPP)] and [Rh(acac)(CO)(TMOPP)] (TPP = $PPh_3$, TMOPP = $P(C_6H_4$-4-$OMe)_3$) to yield [RhCl(COD)(TMOPP)] and [Rh(acac)(CO)(TPP)] is very rapid but is followed by a much slower process where the bidentate ligands are exchanged to yield [Rh(acac)(COD)] and a mixture of [RhCl(CO)$(TPP)_2$], [RhCl(CO)$(TMOPP)_2$], and [RhCl(CO)(TPP)(TMOPP)]. The exchange involving [RhCl(COD)(L)] and [Rh(acac)(CO)(L)] yields [Rh(acac)(COD)] and [RhCl(CO)$(L)_2$], where the reaction is much faster when L = TPP than when L = TMOPP. The mixed-metal system comprising [IrCl(COD)(TPP)] and [Rh(acac)(CO)(TPP)] yields all four complexes [M(acac)(COD)] and [MCl(CO)$(TPP)_2$], where M = Rh and Ir. This illustrates that both a neutral ligand exchange and an anionic ligand exchange occur. Possible pathways for these processes are discussed.

## Introduction

The mechanism of ligand exchange processes has long been a topic of interest for coordination chemists.[1–3] Ligand exchange dynamics is of importance in all catalytic processes, whether industrially or biologically relevant. Ligand exchange in square-planar $d^8$ complexes has occupied a dominant position, given the large number of catalytic reactions promoted by $d^8$ metal centers such as $Rh^I$, $Ir^I$, $Ni^{II}$, $Pd^{II}$, $Pt^{II}$ and $Au^{III}$. Most of the ligand exchange investigations have dealt with $Pt^{II}$ complexes, given their relative inertness and stereochemical stability that bring the reactions within a suitable half-life range for convenient studies by classical mixing and monitoring methods.[3]

Ligand exchange in $Rh^I$ complexes has been studied to a lesser extent. It is nevertheless well appreciated that it occurs predominantly via an associative pathway, as for the other $d^8$ systems and as anticipated by the "16 and 18 electron" rule.[4] For instance, NMR spectroscopic investigations on [RhCl(COD)(L)] (COD = 1,5-cyclooctadiene; L = $PPh_3$, $AsPh_3$) in the presence of excess L have revealed very fast (signal coalescence on the NMR timescale) and associative (first order in free L) ligand exchange. Furthermore, fast exchange still occurs in the absence of free L through monomer–monomer interactions.[5] A very rapid ligand exchange process was also observed for [Rh(acac)(CO)$(PPh_3)$] (acac = acetylacetonate) with free $PPh_3$.[6]

We have recently embarked on an investigation of the hydroformylation reaction catalyzed by Rh complexes supported on precise phosphine-functionalized macromolecular architectures built by controlled radical polymerization.[7–9] These polymers were obtained by copolymerization of styrene and 4-diphenylphosphinostyrene (DPPS); they can be considered as having polystyrene-linked triphenylphosphine ligands. With regard to these catalytic studies, we have explored double exchange processes where one phosphine ligand ($P_1$) bonded to one type of Rh complex ($Rh_1$) exchanges with a second phosphine ligand ($P_2$) bonded to a second type of Rh complex ($Rh_2$), see Equation (1).

$$Rh_1–P_1 + Rh_2–P_2 \rightleftarrows Rh_1–P_2 + Rh_2–P_1 \quad (1)$$

This double exchange process on polymer-supported phosphine ligands has provided important information in relation to polymer dynamics, which will be described separately in a specialized polymer journal. Here, we report our results on the model system using the regular (nonpolymer-supported) phosphine ligands since they provide interesting new information on the coordination chemistry of rhodium and notably on the mechanism of ligand exchange involving both neutral and anionic bidentate ligands.

## Results

### (a) The Phosphine Double Exchange Process

In order to conveniently follow the double exchange reaction in Equation (1) by NMR spectrometry, we searched for

[a] CNRS, Laboratoire de Chimie de Coordination (LCC), 205 route de Narbonne, BP 44099, 31077 Toulouse Cedex 4, France
E-mail: rinaldo.poli@lcc-toulouse.fr,
http://www.lcc-toulouse.fr/equipe_g/pages/poli/index.html
[b] Université de Toulouse, UPS, INPT, 31077 Toulouse Cedex 4, France
[c] Institut Universitaire de France, 103, Bd. Saint Michel, 75005 Paris, France

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a combination of two Rh systems and two phosphine ligands allowing us to individually detect all four compounds in the mixture, at least by the better resolved $^{31}P$ NMR spectroscopic technique. A suitable combination was identified as [RhCl(COD)(L)] and [Rh(acac)(CO)(L)] with L = $PPh_3$ and $P(C_6H_4$-4-OMe$)_3$. These ligands will be henceforth abbreviated as TPP for triphenylphosphine and TMOPP for tris(4-methoxyphenyl)phosphine. The double exchange process in Equation (1) was investigated by mixing [RhCl(COD)(TPP)] (**1**) and [Rh(acac)(CO)(TMOPP)] (**2**) in $CDCl_3$. Relevant $^{31}P$ NMR spectra were collected in Figure 1. They demonstrate that the double phosphine exchange process is very rapid; essentially an equilibrated 1:1:1:1 mixture of the four compounds is obtained rapidly after mixing [see Figure 1, spectrum (c)]. The observed chemical shifts and $J_{PRh}$ coupling constants for each compound are in agreement with those reported in the literature: **1**, $\delta$ = 30.68 ppm, $J_{PRh}$ = 150.6 Hz (ref. $\delta$ = 31.3 ppm, $J_{PRh}$ = 150.1 Hz);[10] **2**, $\delta$ = 44.14 ppm, $J_{PRh}$ = 173.7 Hz (ref. $\delta$ = 43.5 ppm, $J_{PRh}$ = 175.6 Hz);[11] [Rh(acac)(CO)(TPP)] (**3**), $\delta$ = 48.67 ppm, $J_{PRh}$ = 175.0 Hz (ref. $\delta$ = 48.6 ppm, $J_{PRh}$ = 179.7 Hz);[12] [RhCl(COD)(TMOPP)] (**4**), $\delta$ = 27.00 ppm, $J_{PRh}$ = 149.4 Hz (ref. $\delta$ = 27.7 ppm, $J_{PRh}$ = 148.7 Hz).[10] However, additional resonances were already visible after 30 min in the $\delta$ = 32–23 ppm region and increased slowly, indicating the formation of additional products.

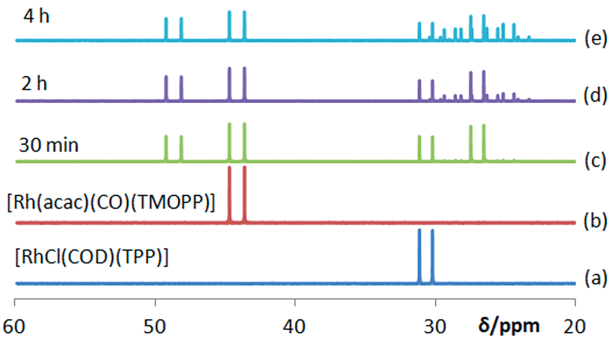



Figure 1. $^{31}P\{^1H\}$ NMR spectra recorded for the reaction between [RhCl(COD)(TPP)] (**1**) and [Rh(acac)(CO)(TMOPP)] (**2**) solvent = $CDCl_3$, room temperature. (a) complex **1**; (b) complex **2**; (c–e) 1:1 mixture, spectra recorded after the indicated time from mixing.

Further evolution of the mixture at room temperature led essentially to the complete disappearance of the resonances of the four above-mentioned complexes, indicating the irreversibility of the process. Crystallization of the final solution by pentane vapor diffusion led to the deposition of a crystalline solid that, after redissolution into $CDCl_3$, afforded the $^{31}P\{^1H\}$ NMR spectrum shown in Figure 2. Subsequent P–P COSY, P–Rh HMQC, and HSQC analyses (see Supporting Information, Figures S1–S3) indicated that the mixture consists of three different compounds: one (**5**) with $\delta$($^{31}P$ NMR) = 28.97 ppm, $\delta$($^{103}Rh$ NMR) = –8169 ppm, and $J_{PRh}$ = 124.7 Hz, a second one (**6**) with $\delta$($^{31}P$ NMR) = 24.76 ppm, $\delta$($^{103}Rh$ NMR) = –8149 ppm, and $J_{PRh}$ = 123.9 Hz, and a third one (**7**) characterized by an ABX ($P_2Rh$) system in the $^{31}P$ NMR spectrum with $\delta_1$($^{31}P$ NMR) = 29.90 ppm, $\delta_2$($^{31}P$ NMR) = 24.82 ppm, $\delta$($^{103}Rh$ NMR) = –8159 ppm, $J_{P1Rh}$ = 125.0, $J_{P2Rh}$ = 126.5 Hz, and $J_{P1P2}$ = 361.3 Hz.

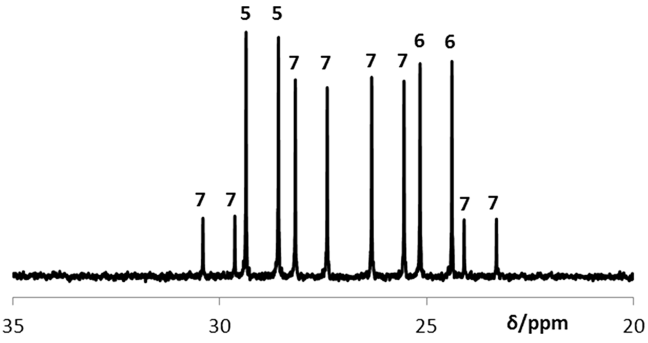



Figure 2. $^{31}P\{^1H\}$ NMR spectrum ($CDCl_3$) of the crystallized solid after the complete reaction between **1** and **2**.

### (b) Simpler Exchange Processes on Rh Complexes

In order to determine the precise nature of products **5**, **6**, and **7**, additional experiments were carried out by mixing, on the one hand, pairs of Rh complexes having the same ligand set except for the phosphine (i.e. **1** and **4** and separately **2** and **3**) and, on the other hand, pairs of Rh complexes having different ligand sets and the same phosphine (i.e. **1** and **3** and separately **2** and **4**). The former two experiments did not lead to any spectral evolution, whereas the latter ones led to the evolutions illustrated in Figure 3 and Figure 4. The reaction between **1** and **3** led selectively to the resonance of compound **5**, whereas that between **2** and **4** led selectively to the resonance of compound **6**. The formation of compound **5** (Figure 3) is already quantitative after mixing and immediately recording the spectrum, the resonance of **1** at $\delta$ = 30.68 ppm (Figure 1, a) having completely disappeared. A small residual resonance of **3** at $\delta$ = 48.67 ppm remains present because this compound was used in slight excess. The formation of compound **6** from **2** and **4** was much slower, since the resonances of both reagents are still visible with a small intensity after 24 h (Figure 4, b). Both **5** and **6** were isolated from the final mixtures by crystallization. Comparison with the literature[13,14] indicated that they correspond to *trans*-[RhCl(CO)$(TPP)_2$] and *trans*-[RhCl(CO)$(TMOPP)_2$], respectively. The identity of compound **6** was further confirmed by determination of the unit cell parameters of a single crystal, which matched those reported for [RhCl(CO)$(TMOPP)_2$].[15] The full spectral characterization of compounds **5** and **6** (see SI) also included the $^{13}C\{^1H\}$ NMR spectrum, which has apparently not been previously reported. The carbonyl C atom gives a complex multiplet from the coupling to the $^{103}Rh$ and the two equivalent $^{31}P$ nuclei. It is around $\delta$ = 187 ppm for both compounds but the spectra are distinguished by the different pattern for the aromatic C atoms and by the resonance of the OMe C atom at $\delta$ = 55.4 ppm for compound **6**.

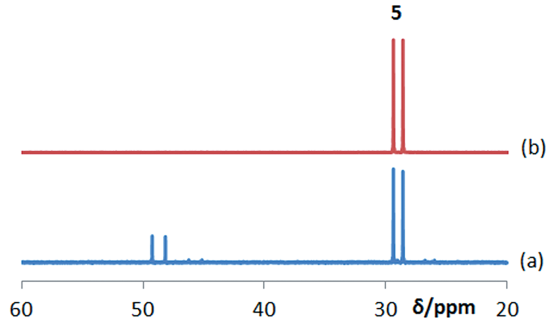



Figure 3. $^{31}P\{^{1}H\}$ NMR spectra recorded for the reaction between **1** and **3**; solvent: $CDCl_3$, room temperature. (a) immediately after mixing; (b) redissolved crystallized product.

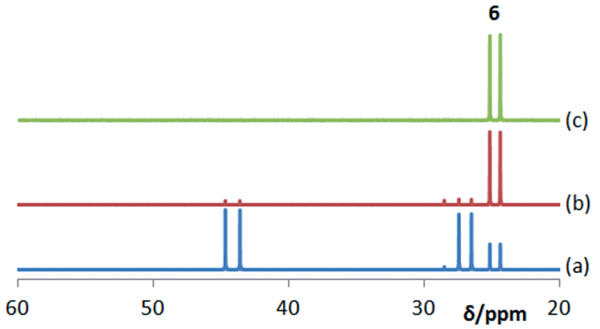


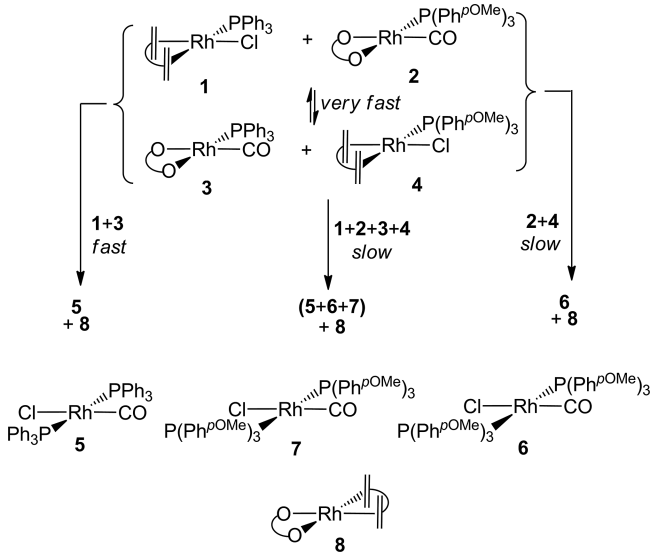

Figure 4. $^{31}P\{^{1}H\}$ NMR spectra recorded for the reaction between **2** and **4**; solvent: $CDCl_3$, room temperature. (a) Immediately; (b) after 24 h; (c) redissolved crystallized product.

The identification of the *trans*-[RhCl(CO)$L_2$] products in these simpler experiments implies the simultaneous formation of [Rh(acac)(COD)] (**8**), Equation (2) (L = TPP or TMOPP). Clearly, when the phosphine exchange is carried out on the mixture of both systems (L = TPP and TMOPP, as in Figure 1), the mixed phosphine derivative [RhCl(CO)(TPP)(TOMPP)] can also be generated by rapid phosphine exchange. This mixed phosphine complex must therefore correspond to product **7** [Equation (3)].

$$[RhCl(COD)(L)] + [Rh(acac)(CO)(L)] \rightarrow [RhCl(CO)(L)_2] + [Rh(acac)(COD)] \quad (2)$$

$$[RhCl(CO)(TPP)_2] + [RhCl(CO)(TMOPP)_2] \rightleftarrows 2\,[RhCl(CO)(TPP)(TMOPP)] \quad (3)$$

The experimental study of this phenomenon was completed with the identification of the byproduct **8**. Since this compound is phosphorus free, its detection could only be performed from its $^{1}H$ NMR and $^{13}C$ NMR properties. After removal of most of the less soluble [RhCl(CO)$(L)_2$] coproducts by crystallization, the residual solution indeed exhibited spectral properties in agreement with those reported for complex **8**.[16,17] The observed ligand exchange processes can be summarized as shown in Scheme 1.




Scheme 1. Summary of the observed exchange processes.

Several mixed phosphine complexes of type [RhCl(CO)$(L_1)(L_2)$], but not compound **7**, have previously been reported. Their $^{31}P$ NMR properties closely parallel those reported here for **7**. However, they have been obtained by phosphine exchange processes from [RhCl(CO)$(L_1)_2$] and free $L_2$, and therefore only in a phosphine-rich environment.[18] These reactions were described as very fast associative processes, whereas the process described here entails phosphine exchange from a Rh complex to another, accompanied by simultaneous rearrangement of the other ligands and notably bidentate ones. The exchange of bidentate ligands between two different $Rh^{I}$ complexes has not been the subject of extensive investigations.

### (c) Ligand Exchange between Rh and Ir Complexes

A final experiment consisted of running the same reaction as in Equation (2) (L = TPP) except that one complex contained Rh, **1**, whereas the other one contained Ir, [IrCl(COD)(TPP)] (**9**). The latter complex was generated in situ by adding 1 equiv. of TPP per Ir atom to compound [IrCl(COD)]$_2$. This experiment was expected to lead to either the products of Equation (4) or to those of Equation (4′), depending on the way in which the ligands are exchanged. Exchanging the neutral ligands (CO and TPP on the Rh complex with COD on the Ir complex) yields the products of Equation (4), whereas exchanging the anionic ligands (acac on the Rh complex with Cl and TPP on the Ir complex) leads to the products of Equation (4′). Thus, the results of this experiment provide useful information on the ligand exchange mechanism.

$$\underset{\mathbf{1}}{[Rh(acac)(CO)(TPP)]} + \underset{\mathbf{9}}{[IrCl(COD)(TPP)]} \rightarrow \underset{\mathbf{8}}{[Rh(acac)(COD)]} + \underset{\mathbf{10}}{[IrCl(CO)(TPP)_2]} \quad (4)$$

$$\underset{\mathbf{11}}{[Ir(acac)(COD)]} + \underset{\mathbf{5}}{[RhCl(CO)(TPP)_2]} \quad (4')$$

The $^{31}P\{^1H\}$ NMR results of this experiment are shown in Figure 5. After 7 h at room temperature, the starting material resonances are still visible (doublet at $\delta$ = 48.67 ppm with $J_{PRh}$ = 175.0 Hz for **1** and singlet at $\delta$ = 21.93 ppm for **9**). However, a new singlet at $\delta$ = 24.20 ppm can be assigned to **10**, the expected phosphine-containing products of Equation (4) (literature: 23.40 ppm in $CDCl_3$[14]) and a new small doublet at $\delta$ = 29.27 ppm ($J_{RhP}$ = 126.8 Hz) can be assigned to **5**, the expected product of Equation (4′). The integrated intensities of complex **10** and **5** are in a 2.8:1 ratio. The $^{13}C\{^1H\}$ spectrum confirmed the presence of both complex **8**[16,17] and **11**[19] through the characteristic resonances of the metal-bonded COD carbon atoms: doublet at $\delta$ = 76.24 ppm (literature: 76.76 ppm),[16,17] $J_{CRh}$ = 14 Hz for complex **8** and singlet at $\delta$ = 58.87 ppm (literature: 59.3 ppm)[19] for complex **11**. This spectrum is shown in the SI (Figure S4). The intensity of the resonance of complex **8** is greater than that of **11** by a factor of 15.0. Since the reaction was carried out with equimolar amounts, the **10/5** and **8/11** ratios should be identical, but identical intensity ratios are not to be expected because the NMR integration for the Overhauser-enhanced resonances of the slow-relaxing $^{31}P\{^1H\}$ and $^{13}C\{^1H\}$ nuclei does not carry quantitative information. The proton environment of these nuclei in the two compounds should be rather similar, suggesting that the Overhauser effect may not be significantly different in each pair of related compounds. However, the relaxation times could be significantly different because in one case the observed nucleus is bonded to a magnetically dipolar $^{103}Rh$ nucleus, whereas in the other case it is bonded to a quadrupolar ($I$ = 3/2) Ir nucleus. In order to reconcile the different observed intensities, we must consider that the resonance intensity of the Rh complex is underestimated relative to that of the Ir complex in at least one (but probably both) of the NMR spectra. At any rate, the two NMR measurements consistently indicate that the exchange proceeds preferentially through Equation (4), by a factor between the lower ($^{31}P\{^1H\}$ NMR) and upper ($^{13}C\{^1H\}$ NMR) integrated intensity limits of 2.8 and 15.0.[20]

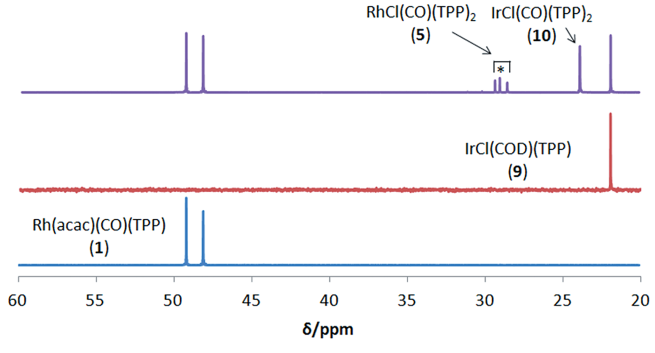



Figure 5. $^{31}P\{^1H\}$ NMR spectra recorded for the reaction between **1** (blue spectrum) and **9** (red spectrum). The violet spectrum was recorded 7 h after mixing; solvent: $CD_2Cl_2$, room temperature. The resonance marked with an asterisk corresponds to an impurity ($Ph_3PO$).

Hence, the reaction occurs through both mechanisms, but the neutral ligand exchange pathway prevails. It is worth noting that the exchange of Cl and acac ligands between different metals is not unprecedented, and is reported for the exchange between various $M(acac)_2$ and $M'Cl_2$ complexes (M, M′ = Mg, V, Fe, Co, Ni, Zn), leading in many cases to the observation of bimetallic intermediates.[21–23]

## Discussion

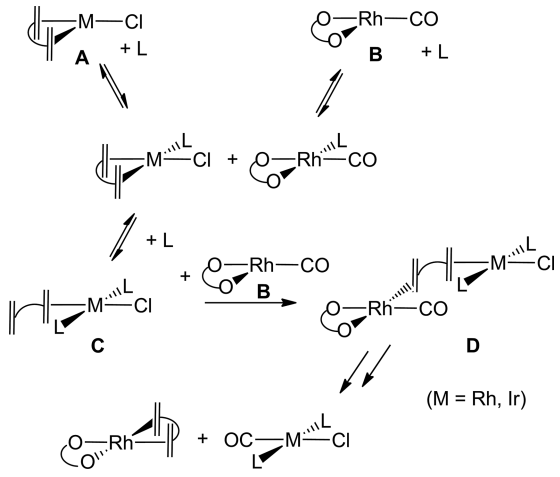
It is of interest to speculate on the mechanism of the slow bidentate ligand exchange processes of Equation (2) and Equation (4/4′). The redistribution, which will be represented for a generic phosphine ligand L, must involve either exchange of neutral ligands – a cyclooctadiene on the M complex, where M = Rh for Equation (2) or Ir for Equation (4), with L and CO on the other Rh complex – or exchange of the anionic ligands – acac on the Rh complex with Cl, accompanied by L, on the M complex, where M = Rh for Equation (2) or Ir for Equation (4′). In order to experimentally distinguish between the two possibilities for the Rh-only system of Equation (2), it would be necessary to carry out an isotope labeling experiment where the label is on the metal atom, which is impossible with naturally occurring isotopes because the metal is 100% $^{103}Rh$. However, a related reaction where one compound was labeled using the Ir congener [Equation (4) and Equation (4′)], showed the occurrence of both exchange pathways. Obviously, this result only proves that the mixed-metal system is able to follow both exchange pathways. The Rh-only system of Equation (2) could undergo the slow bidentate ligand exchange by only one of the two possible schemes. However, it seems reasonable to extrapolate the result of the mixed-metal system to the Rh-only system. Importantly, the operating mechanism must be able to rationalize the large rate difference observed when L = TPP or TMOPP.

We start by analyzing the "neutral ligand exchange" pathway. The system does not contain any free neutral ligand capable of triggering an associative exchange pathway, since the solvent chloroform has no significant coordinating properties. As mentioned in the introduction, ligand exchange processes in $Rh^I$ complexes are generally associative, but a few examples where the metal reactivity (ligand exchange or other) is triggered by ligand dissociation exist, including dissociation of $N_2$ *trans* to an aryl group[24] and $SiPr_2$ *trans* to an amido donor.[25] The dynamic behavior of complexes $[RhX(PPh_3)_3]$ (X = Cl, $CF_3$, H, $CH_3$, Ph) is a peculiar example of a dissociative self-exchange process.[26] However, all of these processes deal with monodentate ligands.

A reasonable dissociative pathway for the exchange of neutral ligands may be conceived as shown in Scheme 2. Given the known trends of *trans* effects of ligand bond dissociation energies, and of chelating effects, the most likely initial dissociation is that of L *trans* to one of the COD double bonds in complex [RhCl(COD)L] yielding intermediate **A**, but this process is unproductive. Dissociation of L *trans* to one acac O atom in the other reagent yields intermediate **B**. Next, compound [RhCl(COD)(L)] may re-

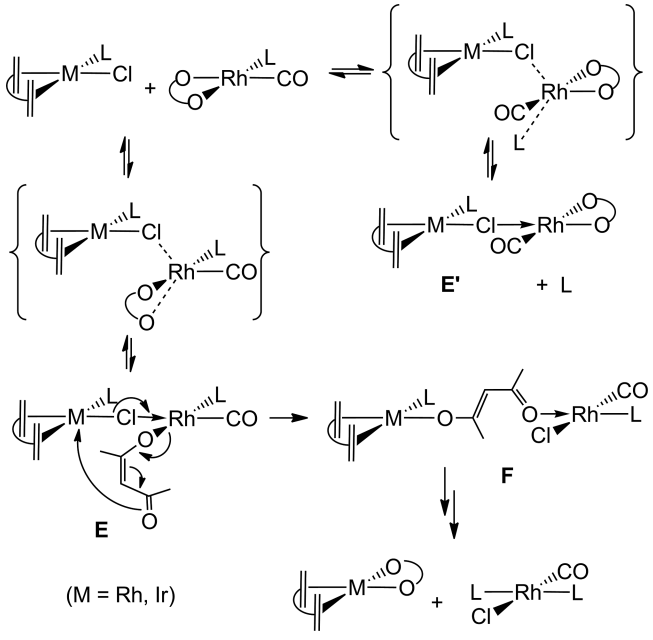
 

act with additional L, generated during the reversible formation of either **A** or **B**, to open the COD chelate and yield intermediate **C**, possibly via an associative pathway. This intermediate may then react with **B** to afford the COD-bridged bimetallic complex **D**. There are many ways in which this exchange may further proceed to the final products, but the important points are the formation of the bimetallic intermediate and the initial dissociation of L. On the basis of this hypothesis, the observed trend of reactivity (much faster rate when L = TPP) appears consistent with the literature. Indeed, through calorimetric studies, Nolan et al. reported that the reaction between $[RhCl(COD)]_2$ and L to yield [RhCl(COD)L] is more favorable for TMOPP ($58.7 \pm 0.3$ kcal/mol) than for TPP ($51.7 \pm 0.3$ kcal/mol).[27] Hence, the dissociation of TMOPP from $Rh^I$ is expected to be much slower than that of TPP.




Scheme 2. Possible mechanism for the neutral ligand exchange leading from [MCl(COD)L] and [Rh(acac)(CO)L] to [Rh(acac)-(COD)] and $[MCl(CO)L_2]$ (M = Rh, Ir).

Turning now to the "anionic ligand exchange" pathway, it is clear that a dissociative process involving dissociation of the anionic ligands and charge separation would be difficult, especially in a low polarity solvent such as chloroform. However, an associative process seems feasible. It is also possible that the association via formation of bridged dinuclear intermediates is triggered by dissociation of a neutral ligand. This is suggested by the literature report of rapid halide scrambling between $[RhBr(CO)(TPP)_2]$ and $[IrCl(CO)(TPP)_2]$.[18] The dynamic exchange on the $^1H$ NMR timescale of the two inequivalent halves of the COD ligand in compound [RhCl(COD)L] (L = $PPh_3$, $AsPh_3$), which is kinetically second order in the metal complex, might also involve halide-bridged intermediates.[5] We can thus propose that the first step of the reaction between [RhCl(COD)L] and [Rh(acac)(CO)L] is the formation of a dinuclear halide bridged complex (**E** in Scheme 3) with the elimination of a ketone group of the acetylacetonate, which rearranges to monodentate coordination. The associative pathway is shown in Scheme 3, but the dissociative variant would of course lead to the same result. The alternative exchange (associative or dissociative) of a phosphine ligand, leading to a similar dinuclear complex **E′**, also seems possible but would be unproductive. In order to satisfy first principles, all elementary steps envisaged for this "anionic ligand exchange" mechanism are such that they produce neutral systems (i.e. no charge separation) and maintain a square-planar configuration around each $Rh^I$ center in all intermediates. Although pentacoordination is possible in $Rh^I$ chemistry, square-planar complexes are preferred when potential π-donor atoms such as Cl or O are present in the coordination sphere. Thus, in intermediate **E**, for instance, the positive charge of the Rh atom on the left hand side is saturated by the covalent interaction with the bridging Cl atom, whereas the bond of this Cl atom to the Rh center at the right is dative. The charge of the Rh atom on the right hand side is saturated by the enolate of the monodentate acac ligand. In the next step, the Cl and acac ligand swap positions through an exchange reaction that involves attack of the Rh atom on the left hand side by the lone pair of the free acac carbonyl function, as suggested in Scheme 3, for a net charge change of zero and formation of intermediate **F**. From here onward, it is easy to see how the exchange may continue, with either associative or dissociative processes, to complete the ligand exchange.




Scheme 3. Possible mechanism for the anionic ligand exchange leading from [MCl(COD)L] and [Rh(acac)(CO)L] to [M(acac)-(COD)] and $[RhCl(CO)L_2]$ (M = Rh, Ir).

The O atoms in [Rh(acac)(CO)L] are also centers of nucleophilic reactivity. Therefore, it is possible in principle to envisage another anionic ligand exchange pathway, starting with attack of a rhodium complex by one O lone pair of the acac ligand in the second complex. However, compounds of the type [Rh(acac)(CO)L] have been reported not to lead to coalescence of the asymmetric acac resonances,[28] even upon warming, although the phenomenon is observed in the presence of excess L.[12] This suggests that ligand ex-

change by self-association, if it occurs, is a slower process for [Rh(acac)(CO)L] than for [RhCl(COD)L].

Note that the first two exchange processes in the "anionic ligand exchange" pathway up to intermediate **F** do not involve L dissociation, therefore they do not account for the marked reactivity difference in rate when L = TPP or TMOPP. L dissociation only occurs in the further steps going from **F** to the products. Therefore, the pathway of Scheme 3 can be reconciled with the experimentally observed trend only if intermediate **F** is generated by fast pre-equilibrium processes, relative to the L dissociation process that occurs in a later step and would be rate limiting.

Finally, it is necessary to comment on the difference in rate between Equation (2) when L = TPP, which is very rapid as shown in Figure 3, and Equations (4/4′) where the ligand is again TPP, which is on the other hand much slower. This difference may be explained by the stronger Ir–ligand bonds relative to the corresponding Rh–ligand bonds. For the "neutral ligand exchange" pathway of Scheme 2, the initial TPP dissociation would not be discriminating since it always occurs on the Rh complex, but the COD dissociation step leading to intermediate **C** is likely to be much slower for the iridium complex. For the "anionic ligand exchange" pathway of Scheme 3, it is the Ir–TPP bond dissociating in one of the later rate-limiting steps that would make the difference in the observed exchange rates.

## Conclusions

The rapid phosphine double exchange of Equation (1), using the **1**/**2** combination, has unveiled an unexpected side reaction consisting of the slow exchange of the bidentate ligands, leading to the formation of complex **8** and a statistical mixture of **5**, **6**, and **7**. Control experiments involving the reactions between [RhCl(COD)(L)] and [Rh(acac)(CO)(L)] for L = TPP or TMOPP, as well as involving the mixed-metal system **1** and **9** have provided useful information on the mechanism of this process. It has been demonstrated that both the neutral ligands (bidentate COD with CO and L) and the anionic ligands (bidentate acac with Cl and L) can be exchanged, at least for the mixed-metal system.

## Experimental Section

**General:** All manipulations were performed under an inert atmosphere of dry argon by using a vacuum line and Schlenk-tube techniques. Acetylacetonatodicarbonylrhodium(I), $[Rh(acac)(CO)_2]$ (99% Strem), chloro(1,5-cyclooctadiene)rhodium(I) dimer, $[Rh(COD)Cl]_2$ (98%, Strem), chloro(1,5-cyclooctadiene)iridium(I) dimer, $[Ir(COD)Cl]_2$ (99%, Strem), tris(4-methoxyphenyl)phosphine, TMOPP (>95%, TCI), and triphenylphosphine, $PPh_3$ (>98.5%, Fluka) were used as received. Solvents were dried by standard procedures and distilled under argon prior to use. 1D- and 2D-NMR spectra were recorded in 5-mm tubes at 297 K with Bruker Avance 400 and 500 spectrometers. $^1H$ NMR and $^{13}C$ NMR chemical shifts were determined using the residual peak of deuterated solvent as internal standard and are reported in ppm ($\delta$) relative to tetramethylsilane. $^{31}P$ NMR chemical shifts are reported relative to external 85% $H_3PO_4$. Peaks are labeled as singlet (s), doublet (d), triplet (t), multiplet (m), and broad (br). The proton and carbon assignments were assisted by $^1H$–$^{13}C$ HMQC experiments. Complexes [RhCl(COD)(TPP)],[29] [Rh(acac)(CO)(TPP)],[30] [Rh(acac)(CO)(TMOPP)],[11] and [IrCl(COD)(TPP)][31] were synthesized by procedures closely related to those reported in the literature (details in SI).

**Isolation of a Mixture of $[Rh(CO)Cl(TPP)_2]$ (5), $[Rh(CO)Cl(TMOPP)_2]$ (6), and $[Rh(CO)Cl(TPP)(TMOPP)]$ (7):** The two separately prepared solutions of **2** (35 mg, 0.06 mol) in $CH_2Cl_2$ (1 mL) and **1** (30.5 mg, 0.06 mmol) in $CH_2Cl_2$ (1 mL) were combined at room temperature. The resulting mixture was stirred overnight. The resulting solution was concentrated to ca. half the volume and then diffusion of pentane vapors yielded a crystalline solid, yield 29 mg. The solid was characterized by $^{31}P\{^1H\}$ NMR (see Figure 2) and by $^{31}P$–$^{31}P$ COSY, $^{31}P$–$^{103}Rh$ HMQC, and $^{31}P$–$^{103}Rh$ HSQC (see SI) in $CDCl_3$.

**Reaction Between [RhCl(COD)(TPP)] (1) and [Rh(acac)(CO)(TPP)] (3):** Generation of $[Rh(CO)Cl(TPP)_2]$ (**5**) and [Rh(acac)(COD)] (**8**). Two separately prepared solutions of **1** (30.5 mg, 0.06 mmol) in $CH_2Cl_2$ (1.5 mL) and **3** (29.5 mg, 0.06 mmol) in $CH_2Cl_2$ (1.5 mL) were combined and the resulting mixture was stirred at room temperature, progressively depositing a yellow precipitate. After 5 h, the solid was filtered, washed with pentane, and dried under vacuum. Pentane (20 mL) was added to the filtrate to yield an additional yellow crystalline precipitate, which was again filtered off, washed, and dried. These solids were identified as complex **5** by NMR spectroscopy (see below) and compared with those in the literature.[13] The residual yellow solution was evaporated under reduced pressure to yield a yellow-brown solid, identified by NMR spectroscopy as complex **8** by comparison of its $^1H$ and $^{13}C\{^1H\}$ NMR spectra (see below) with those in the literature.[16,17] $[Rh(CO)Cl(TPP)_2]$: $^1H$ NMR (400 MHz, $CDCl_3$): $\delta$ = 7.78–7.74 (m, 6 H, C$H_{Ar}$), 7.44–7.4 (m, 9 H, C$H_{Ar}$) ppm. $^{31}P\{^1H\}$ NMR (162 MHz, $CDCl_3$, 298 K): $\delta$ = 28.97 (d, $J_{P\text{-}Rh}$ = 126.4 Hz) ppm. $^{13}C\{^1H\}$ NMR (101.5 MHz, $CDCl_3$): $\delta$ = 187.6–186.6 (m, CO), 134.73, 132.96, 130.09, 128.12 (C$H_{Ar}$). [Rh(acac)(COD)] ppm. $^1H$ NMR (400 MHz, $CDCl_3$, 298 K): $\delta$ = 5.34 (s, 1 H, C$H_{acac}$), 4.09 (s, 4 H, C$H_{cod}$), 2.49–2.46 (m, 4 H, C$H_{2\,cod}$), 1.95 (s, 6 H, C$H_{3\,acac}$), 1.87–1.81 (m, 4 H, C$H_{2\,cod}$) ppm. $^{13}C\{^1H\}$ NMR (101.5 MHz, $CDCl_3$, 298 K): $\delta$ = 186.64 (s, C$O_{acac}$), 134.8–134.4, 131.2–131.9, 130.24, 128.49, 128 (C$H_{Ar}$), 99.76 (d, $J$ = 2.03 Hz, C$H_{acac}$), 76.47 (d, $J$ = 14.7 Hz, C$H_{cod}$), 30.24 (s, C$H_{2\,cod}$), 27.36 (s, C$H_{3\,acac}$) ppm.

**Reaction Between [RhCl(COD)(TMOPP)] (4) and [Rh(acac)(CO)(TMOPP)] (2):** Generation of $[Rh(CO)Cl(TMOPP)_2]$ (**6**) and [Rh(acac)(COD)] (**8**). This reaction was carried out according to the same protocol described in the previous section for the corresponding TPP complexes, starting from complex **2** (34.9 mg, 0.06 mmol) in $CH_2Cl_2$ (1.5 mL) and **4** (0.06 mmol) in $CH_2Cl_2$ (1.5 mL). The latter was generated in situ from $[Rh(COD)Cl]_2$ (14.8 mg, 0.03 mmol) and TMOPP (21.14 mg, 0.06 mmol). The recovered yellow precipitate (same workup as above) was identified as complex **6** by comparison of its NMR properties with those in the literature,[14] while the residue recovered from the solution corresponded again to complex **8**. $[Rh(CO)Cl(TMOPP)_2]$: $^1H$ NMR (400 MHz, $CDCl_3$): $\delta$ = 7.67–7.62 (m, 6 H, C$H_{Ar}$), 6.92 (d, 6 H, C$H_{Ar}$), 3.83 (s, 9 H, C$H_{3\,OMe}$) ppm. $^{31}P\{^1H\}$ NMR (162 MHz, $CDCl_3$, 298 K): $\delta$ = 24.8 (d, $J_{P\text{-}Rh}$ = 124.74 Hz) ppm. $^{13}C\{^1H\}$ NMR (101.5 MHz, $CDCl_3$): $\delta$ = 187.2 (m, CO), 160.86 (C$^q$), 136.09 (C$H_{Ar}$), 124.89 (C$^q$), 113.7 (C$H_{Ar}$), 55.4 (C$H_{3\,OMe}$) ppm. In ad-

dition, a single crystal of this compound was obtained from a dichloromethane solution by pentane vapor diffusion at room temperature. Its unit cell parameters correspond to those of the published structure of $[Rh(CO)Cl(TMOPP)_2]$.[15]

**Supporting Information** (see footnote on the first page of this article): Synthetic protocols and NMR spectra as specified in the text (four pages).

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