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ARTICLE

Methylbismuth: an Organometallic Bismuthinidene Biradical

Deb Pratim Mukhopadhyay,^a Domenik Schleier,^a Sara Wirsing,^a Jacqueline Ramler,^b Dustin Kaiser,^a Engelbert Reusch,^a Patrick Hemberger,^c Tobias Preitschopf,^a Ivo Krummenacher,^b Bernd Engels,^{*a} Ingo Fischer^{*a} and Crispin Lichtenberg^{*b}

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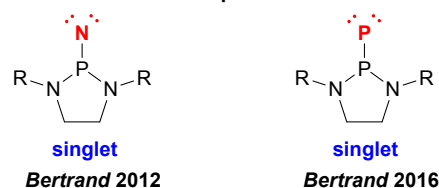
DOI: 10.1039/x0xx00000x

We report the generation, spectroscopic characterization, and computational analysis of the first free (non-stabilized) organometallic bismuthinidene, BiMe. The title compound was generated in situ from BiMe₃ by controlled homolytic Bi–C bond cleavage in the gas phase. Its electronic structure was characterized by a combination of photoion mass-selected threshold photoelectron spectroscopy and DFT as well as multi-reference computations. A triplet ground state was identified and an ionization energy (IE) of 7.88 eV was experimentally determined. Methyl abstraction from BiMe₃ to give [BiMe₂][•] is a key step in the generation of BiMe. We reveal a bond dissociation energy of 210 ± 7 kJ mol⁻¹, which is substantially higher than the previously accepted value. Nevertheless, the homolytic cleavage of Me–BiMe₂ bonds could be achieved at moderate temperatures (60–120 °C) in the condensed phase, suggesting that [BiMe₂][•] and BiMe are accessible as reactive intermediates under these conditions.

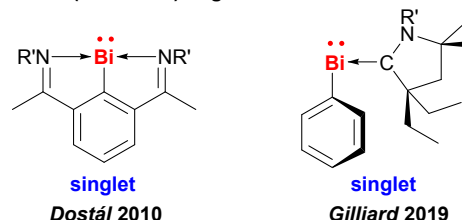
Introduction

Low-valent molecular compounds of group 15 elements, E–R, with the central atom E in the oxidation state of +1 are highly reactive, electron-deficient species (E = N–Bi; R = monoanionic ligand).¹ In principle, they can adopt either singlet or triplet electronic ground states, which strongly influences their physical and chemical properties (Scheme 1). Thus, understanding their (electronic) structures is key to rationalizing the reactivity of these species, which represent important intermediates in fundamental transformations such as insertion and ring expansion reactions.² The isolation of such compounds, in which the group 15 atom is bound to only one additional atom, is extremely challenging and has only recently been achieved for the lightest congeners in landmark contributions (E = N, P, Scheme 1a).^{2,3} Access to the heavier homologues is increasingly difficult due to their tendency to undergo degradation reactions such as oligomerization, disproportionation, and bond activation processes. For the heaviest group 15 element bismuth, mononuclear organometallic compounds with the metal atom in the

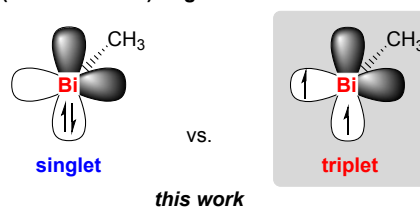
a) Isolable Nitrene and Phosphinidene



b) Isolable (Stabilized) Organometallic Bismuthinidenes



c) Free (Non-stabilized) Organometallic Bismuthinidene



Scheme 1. Low-valent group 15 compounds: a) isolable nitrene and phosphinidene (singlet ground state). b) stabilized organometallic bismuthinidenes (singlet ground state; occupied bismuth 6s-orbital omitted for clarity); c) free (non-stabilized) singlet and triplet species that may be envisaged for BiMe (only two bismuth p-orbitals shown; p-orbital used for Bi–Me bonding and occupied s-orbital are omitted for clarity). R = 2,6-Me₂-C₆H₃; 2,6-*i*-Pr₂-C₆H₃; R' = 2,6-[(4-*t*Bu-C₆H₄)₂CH]₂-4-Me-C₆H₃.

oxidation state of +1 (so-called bismuthinidenes) could so far only be accessed by stabilization through adduct formation with Lewis bases (Scheme 1b).^{4–9} Very recently, the intriguing

^a Institute of Physical and Theoretical Chemistry
University of Würzburg
Am Hubland, D-97074 Würzburg
E-mail: bernd.engels@uni-wuerzburg.de, ingo.fischer@uni-wuerzburg.de

^b Institute of Inorganic Chemistry
University of Würzburg
Am Hubland, D-97074 Würzburg
E-mail: crispin.lichtenberg@uni-wuerzburg.de

^c Laboratory for Femtochemistry and Synchrotron Radiation
Paul Scherrer Institut (PSI), CH-5232 Villigen, Switzerland
E-mail: patrick.hemberger@psi.ch

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properties of such bismuthinidenes have been exploited for (electro-)catalytic and photophysical applications: a Bi^I/Bi^{III} redox couple has been proposed to be a key component in bismuth-catalyzed transfer hydrogenation reactions with ammonia borane.^{7,10} In addition, Bi^I oxalates have been suggested to be involved in the electrochemical reduction of carbon dioxide at bismuth electrodes.¹¹ When embedded in host materials such as zeolites, [Bi^I]⁺ centers show intriguing photophysical properties such as ultrabroad near-infrared emission.^{12,13} Without any exceptions, the isolable, Lewis-base-stabilized bismuthinidenes reported to date show singlet ground states. This situation changes, when non-stabilized, (so far) non-isolable, electron-deficient molecular bismuthinidenes, Bi-R, are targeted. To date, only inorganic species Bi-X have been reported and were accessed by comproportionation (e.g. at high temperature in the melt) or by reaction of Bi atoms with reactive species such as F₂ in the gas phase (X = H, F-I, AlCl₄).¹⁴⁻¹⁸ Importantly, these synthetic routes did not allow to study the reactivity of bismuthinidenes, they did not allow to work in conventional solvents at moderate temperature, and they did not allow to access compounds with tunable substituents R, such as organometallic bismuthinidenes Bi-R (R = alkyl, aryl; Figure 1c).

Here we report the generation and characterization of the first free (i.e. non-stabilized) organometallic bismuthinidene methylbismuth, BiMe, in a gas-phase reaction with implications for related reactions in condensed phase at moderate temperature. In addition, dimethylbismuth, [BiMe₂][•] was studied to determine the Bi-CH₃ bond dissociation energy in BiMe₃.

Results and discussion

In order to access the elusive class of non-stabilized organometallic bismuthinidenes, we aimed at a “top-down approach”, more specifically the homolytic cleavage of Bi-C bonds from the well-defined organometallic precursor BiMe₃ (**1**) as shown in Scheme 2. The long-known BiMe₃ has been investigated from many different perspectives,¹⁹⁻²¹ including its application in chemical vapor deposition, where the abstraction of all its methyl groups is targeted.^{22,46}

In a recent study, it was demonstrated that the abstraction of a single methyl group can be achieved by dissociative

Scheme 2. Controlled, stepwise abstraction of CH₃ radicals from **1** in the gas phase by flash pyrolysis.

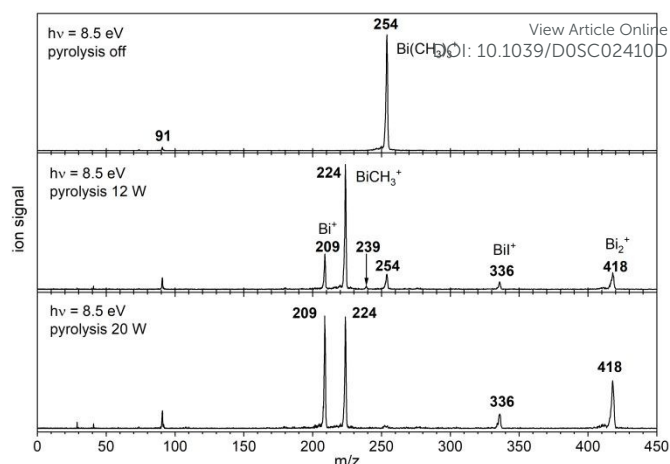
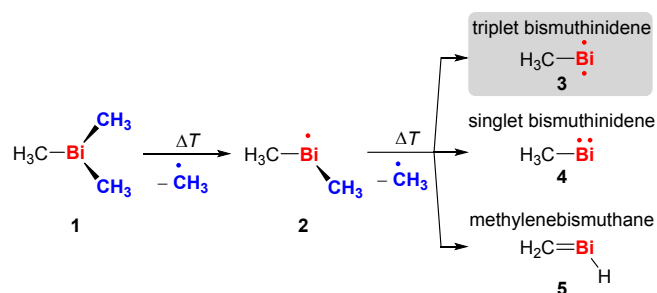
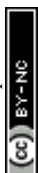


Figure 1. Photoionization mass spectra recorded at 8.5 eV without pyrolysis (top trace), low pyrolysis power ($T \approx 470$ K, centre trace) and medium pyrolysis power ($T \approx 600$ K, bottom trace). The signal at $m/z=91$ is due to a background signal from previous experiments.

photoionization of **1**, i.e. the Bi-C bond cleavage of the [BiMe₃]^{•+} cation in the gas phase.²³ Under these conditions, only a single methyl group was abstracted, yielding the cation [BiMe₂]⁺. Here we attempt a thermally-induced, controlled and stepwise abstraction of methyl groups from neutral BiMe₃. Thus, a sample of BiMe₃ diluted in Ar was pyrolysed in a microreactor and analyzed by photoelectron-photoion coincidence spectroscopy (PEPICO) using synchrotron radiation.²⁴ This method permits to record photoion mass-selected threshold photoelectron spectra (ms-TPE) for each species by correlating ions and electrons produced in a single photoionization event. Isomer-selective information is then obtained from an analysis of the photoelectron spectrum based on computations.

Figure 1 shows mass spectra under various pyrolysis conditions. Without pyrolysis (top trace) only the parent ion **1**^{•+} is visible, thus dissociative photoionization is irrelevant under our experimental conditions. Already at low pyrolysis power (center trace) a stepwise methyl loss down to atomic Bi occurs, associated with formation of **2** and one of the products **3-5**. The small intensity of $m/z=239$ ([Bi(CH₃)₂]⁺) compared to $m/z=224$ ([BiCH₃]^{•+}) indicates that cleavage of the second methyl group is more facile than the first. In addition, Bi₂ is visible at $m/z=418$, due to dimerization of bismuth atoms (see supplementary information (SI), Figure S4). In some experiments a further peak appeared at $m/z=478$ and is most likely due to Me₂Bi-BiMe₂. Note that CH₃ is not observed due to its ionization energy of 9.83891 eV.^{25,26} Traces of BiI from the synthesis are also present in the spectrum.

When the pyrolysis power is further increased (bottom trace) the precursor is fully converted and experimental conditions are suitable for studying the molecule formed after loss of two methyl groups. Three structures are possible for $m/z=224$, the bismuthinidenes **3** (triplet) / **4** (singlet) or the methylenebismuthane **5**. Threshold photoelectron spectroscopy provides structural isomer-selective information through comparison with Franck-Condon-simulated or reference spectra. Figure 2 represents the ms-TPE spectrum of a species with the composition BiCH₃ at $m/z = 224$. The first



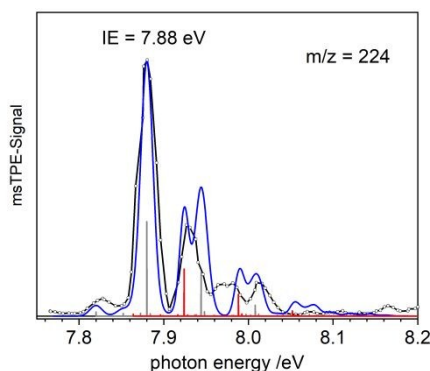


Figure 2. Mass-selected threshold photoelectron spectrum of $m/z=224$. The simulation based on **3** as the carrier (blue line) fits the experiment well. Transitions into the $X^+ {}^2A$ ground state of the ion are given as grey bars, transitions into the $A^+ {}^2A$ state as red bars.

major band at 7.88 eV is assigned to the adiabatic ionization energy (IE_{ad}). It is followed by several smaller bands that are ca. 40–50 meV apart. Simulations based on DFT and multi-reference calculations were carried out for **3**, **4** and **5**. While DFT often provides very accurate geometries and frequencies even for molecules with complicated electronic structures,⁴⁸ it is in many cases less accurate for the computation of energy surfaces or excitation energies.⁴⁹ Hence, the $\omega B97X-D3^{50}$ functional was employed for geometry optimizations and frequency computations, but the multi-reference NEVPT2⁵¹ approach was used for optimizations and to compute the ionization energies and the energy difference between the two relevant states of the $[BiCH_3]^+{}^*$ cation (*vide infra*). Both methods (DFT and multi-reference) were combined with a scalar relativistic approach and with the SARC-ZORA-TZVP basis set which allows for an all-electron treatment of bismuth⁵² (for computational details, see SI). Bismuthinidene **3** with a C_{3v} symmetry and an $X^3E (T_0)$ triplet ground state is the lowest-energy structure ($\Delta E = 0$ kJ·mol⁻¹). The computations on the NEVPT2-level show a very good agreement with the experimentally determined IE of 7.88 eV ($IE_{calc} = 7.98$ eV for the $X^+ {}^2A'' \leftarrow X^3E$ transition (*vide infra*)), as compared to singlet bismuthinidene **4** ($\Delta E = +0.78$ eV / +75 kJ·mol⁻¹; $IE_{calc} = 7.21$ eV) and methylenebismuthane **5** ($\Delta E = +0.91$ eV / +88 kJ·mol⁻¹; $IE_{calc} = 8.68$ eV; for energy values obtained through DFT calculations see SI). This shows, that the ground state (electronic) structure of species with the sum formula ECH_3 are fundamentally different, depending on the choice of the element E. For E = Bi the triplet bismuthinidene **3** is energetically favored and observed (*vide supra*), whereas for the lighter congeners (E = N, P), the formation of the methylene species $HN=CH_2$ and $HP=CH_2$ has been determined to be more favorable.^{27–32,53}

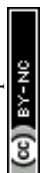
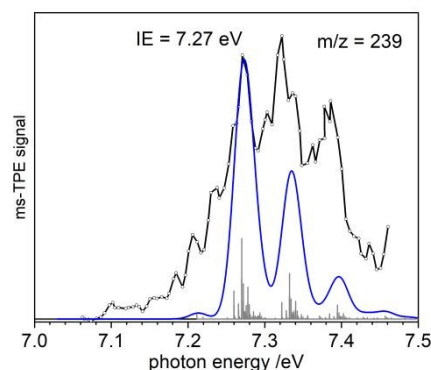
Upon photoionization of **3**, one electron is removed from either of the two degenerate SOMOs, which correspond in first approximation to the p_x and p_y orbital on the Bi center. The computations for the ionic ground state, $[BiCH_3]^+{}^*$ (**3**⁺), yielded a shortening of the Bi–C bond (from 2.27 Å in **3** to 2.21 Å in **3**⁺) and a deformation of the methyl group with a tilt angle of 4° relative to the Bi–C axis. This was ascribed to antibonding interactions between the unpaired electron and the bonding electrons of the two C–H groups, which are approximately aligned with the singly occupied p-type orbital of bismuth (Table

S18). This leads to a loss of the C_3 axis and a reduction to C_s symmetry. As a consequence of this Jahn–Teller distortion, the 2E state in the C_{3v} symmetric cation splits into a $X^+ {}^2A''$ and a $A^+ {}^2A'$ component. The computations indicate an energy difference of only 50 meV between the two states at the geometry of the $X^+ {}^2A''$. Thus, transitions into both states contribute to the spectrum and have to be included in the simulation in Figure 2 (blue curve). Transitions from the T_0 ground state of neutral **3** into the $X^+ {}^2A''$ ionic ground state are given as grey bars, while transitions into the $A^+ {}^2A'$ excited state are shown as red bars. In addition, there is a vibrational structure evident in the spectrum with a spacing of around 50 meV, including a hot band transition at 7.83 eV, which is assigned to the Bi–C stretching motion (corresponding to $v'' \approx 50$ meV). Vibrational activity is expected due to the reduction of the Bi–C bond lengths in the cation by >0.06 Å (see SI for all geometry parameters). Franck–Condon simulations based on computations of isomer **5** further support the triplet bismuthinidene **3** as the carrier of the spectrum (Figure S3). First, the computed IE of **5** is 0.80 eV higher than the experimental value, and second, a more pronounced vibrational progression with a maximum intensity for a transition into an excited vibrational state would be expected for **5**. Thus a contribution of **5** to the spectrum can be ruled out.

The computed IE of **3** (for the lowest state of **3**⁺) and the relative energies for the two states of the cation **3**⁺ are in excellent agreement with the corresponding experimental data. This indicates that the neglected spin-orbit effects³³ do not play a key role for the determination of the ionization energies, possibly because they are similar in magnitude for all involved states. While the experimental and calculated IEs nicely agree, deviations were observed in the shape of the spectra and were ascribed to the flatness of the potential energy surface (PES) of the cation. Two factors mainly contribute to the flatness of this PES: i) the $X^+ {}^2A''$ state is threefold degenerate due to facile rotation around the Bi–C bond. ii) The shape of the PES going from the equilibrium geometry of the $X^+ {}^2A''$ towards the equilibrium geometry of the $A^+ {}^2A'$ state is expected to be non-harmonic. Efforts to obtain a better description of the surface were so far unsuccessful due to strong correlations of the various internal coordinates, so that high dimensional surfaces would be necessary for an appropriate description.

To gain additional information on the formation of BiMe (**3**)

Figure 3. Ms-TPE spectrum of $Bi(CH_3)_2$, $m/z=239$. The vibrational progression is due to the symmetric bismuth-carbon stretch and its combination with torsional motion. The blue line represents the simulated spectrum based on **2**.



by stepwise abstraction of methyl groups from BiMe_3 (**1**), an ms-TPE spectrum of $[\text{BiMe}_2]^+$ (**2**) ($m/z(\mathbf{2}^+) = 239$) was recorded at a low pyrolysis temperature (Figure 3; cf. Figure 1). Simulations of the spectrum based on DFT calculations indicate C_{2v} symmetry for both **2** and $\mathbf{2}^+$ as well as a $X^+ \ ^1A_1 \leftarrow X \ ^2B_1$ transition in $[\text{BiMe}_2]^+$ (**2**). The first major band at 7.27 eV is assigned to the IE, in excellent agreement with the computed value of 7.35 eV at the NEVTP2 level of theory. A vibrational progression with a spacing of 60 meV is visible and is dominated by the symmetric Bi–C stretching mode in the cation. Additional torsional modes of the CH_3 groups may lead to a broadening of the bands. While there is a good agreement between simulation and experiment, the vibrational intensities (including hot bands) are somewhat underestimated in the simulations. This indicates a slightly larger change of the Bi–C bond length upon ionization than the computed shortening of 0.05 Å (see SI for all geometry parameters).

Bond dissociation energies (BDEs) can be determined via thermochemical cycles that combine appearance energies (AE) and ionization energies. The zero Kelvin appearance energy $\text{AE}_{0\text{K}}$ for the abstraction of the first methyl radical from **1** has been determined with very high accuracy, $\text{AE}_{0\text{K}}(\text{Bi}(\text{CH}_3)_3, \text{Bi}(\text{CH}_3)_2^+) = 9.445 \pm 0.064$ eV.²³ According to computations, the methyl loss in the cation is a simple homolytic bond cleavage without a reverse barrier. Combined with the IE of **2**, the $\text{Me}_2\text{Bi}-\text{CH}_3$ bond dissociation energy in **1**, $\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3)$, can be derived:

$$\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3) = \text{AE}_{0\text{K}}(\text{BiMe}_3, [\text{BiMe}_2]^+) - \text{IE}([\text{BiMe}_2]^+) \quad (\text{eq 1}).$$

From eq 1 a value of 210 ± 7 kJ/mol is obtained for $\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3)$. The bond dissociation energy of the first Bi– CH_3 bond in BiMe_3 ($\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3)$) can be expected to be the highest of the three Bi– CH_3 BDEs in this molecule³⁴ and is thus crucial for any type of reaction initiation via Bi– CH_3 homolysis. However, the value of $\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3)$ has never been determined explicitly in the primary literature. Based on previous investigations into the thermal decomposition of **1**,^{34–36} an estimation of $\text{BDE}(\text{Me}_2\text{Bi}-\text{CH}_3)$ can be made, which yields a value of 182 kJ mol^{-1} as the upper limit that would be possible for this parameter.³⁷ Thus, our results substantially revise the bond dissociation energy of the first Bi–Me bond in **1**, which is key to the radical chemistry of **1** and related bismuth compounds. The calculated isodesmic reaction provides a value of 226 kJ mol^{-1} in good agreement with the experimental findings.

Figure 4. Abstraction and spin trapping of methyl radicals from BiMe_3 .

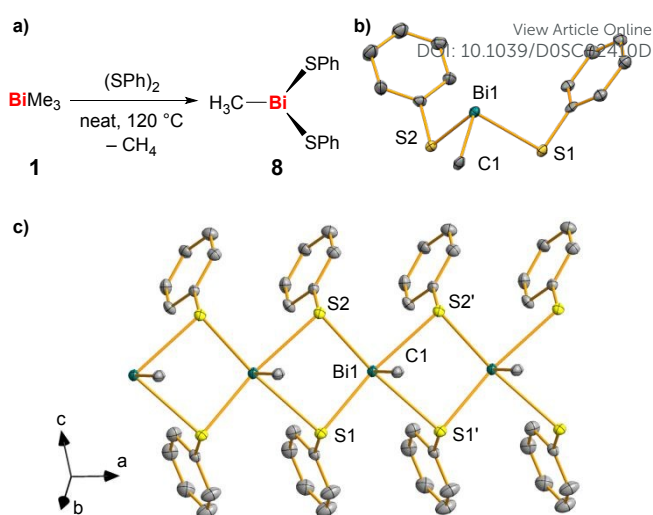
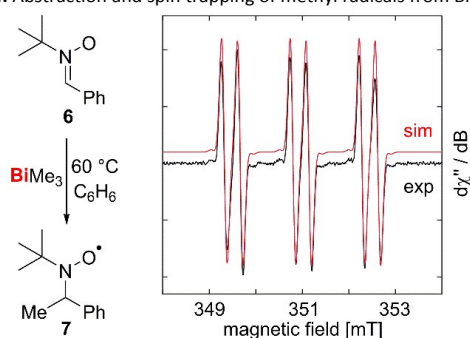


Figure 5. a) Synthesis of $\text{BiMe}(\text{SPh})_2$ (**8**) from BiMe_3 and $(\text{SPh})_2$ with methane as a detected by-product. b,c) Molecular structure of **8** in the solid state with one formula unit shown in b) and a cutout of the coordination polymer shown in c). Displacement ellipsoids are shown at the 50% probability level. Hydrogen atoms and one set of split positions of disordered atoms are omitted for clarity. For detailed discussion of structural parameters see Supp. Inf.. Selected bond lengths (Å) and angles ($^\circ$): Bi1–C1, 2.208(10); Bi1–S1, 2.736(2); Bi1–S2, 2.699(2); C1–Bi1–S1, 89.6(3); C1–Bi1–S2, 89.4(3); S1–Bi1–S2, 93.18(7).

Our correction of the BDE in a fundamentally important organometallic compound such as BiMe_3 raised the question, whether homolytic Bi–C bond cleavage is possible at moderate reaction temperatures in the condensed phase, making this process relevant for synthetic chemistry under conventional experimental conditions. To test for methyl radical abstraction from BiMe_3 , a benzene solution of BiMe_3 and the radical trap **6** was heated to 60°C and subsequently analyzed by EPR spectroscopy (Figure 4). Indeed, a resonance was detected with $g_{\text{iso}} = 2.006$ and $a(^{14}\text{N}) = 41.6$ MHz, $a(^1\text{H}) = 9.63$ MHz, $a(^{13}\text{C}) = 12.9$ MHz, indicating the formation of **7** by methyl radical transfer.^{38,54,55}

In order to gain further hints at the generation and subsequent trapping of BiMe in the condensed phase, neat BiMe_3 was reacted with stoichiometric amounts of $(\text{PhS})_2$ at 120°C (Figure 5a). $\text{BiMe}(\text{SPh})_2$ (**8**)^{39–41} was isolated as the product of this reaction in 41% yield and fully characterized (Figure 5b,c; for details see SI). Methane was detected in the headspace of the reaction by IR spectroscopy (Figure S5), suggesting the appearance of methyl radicals in this reaction.

This – together with literature reports⁴² – supports the potential of BiMe to act as a transient reactive species in this reaction. The thioether MeSPh was also detected, suggesting that methyl radical attack at sulfur or σ -bond metathesis / disproportionation sequences may also be operative as a parallel reaction pathway.⁴¹ In the solid state, compound **8** forms a coordination polymer through bridging coordination modes of the thiolate ligands (Figure 5c). This is a unique structural feature within the small number of literature-known compounds $\text{BiR}(\text{SR}')_2$ ($\text{R}, \text{R}' = \text{aryl, alkyl}$; for details see SI).^{42–45}



Conclusions

In conclusion, we have generated methylbismuth (BiMe), a fundamental organometallic compound and the first example of an organometallic non-stabilized bismuthinidene. BiMe was accessed via controlled thermal homolysis in the gas phase. The title compounds shows a triplet (biradical) ground state and an ionization energy of 7.88 eV, as revealed by combination of photoelectron spectroscopy and computations. The homolytic dissociation of the first Me₂Bi–CH₃ bond in BiMe₃ is crucial to the radical chemistry of this compound (and related species) and is the initiating step in the formation of BiMe. Our results reveal a Me₂Bi–CH₃ homolytic bond dissociation energy of 210±7 kJ/mol, i.e. the previously reported value is revised by more than +15% (+28 kJ·mol⁻¹). Nevertheless, reactions in the condensed phase demonstrate that the abstraction of methyl radicals from BiMe₃ is possible at moderate reaction conditions and suggest that BiMe may act as an intermediate in reactions with suitable trapping reagents. Future research will be directed towards the generation of non-stabilized bismuthinidenes and their exploitation in synthetic chemistry.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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Author contributions

DPM, DS, ER and PH carried out the synchrotron radiation experiments, JR carried out all synthetic work, SW and DK performed the computations, IK performed the ESR and TP the IR spectroscopy. IF, BE and CL planned the project, contributed to the interpretation of the results and wrote the manuscript.

Methods and Experimental

Details of experimental conditions for synchrotron experiments, for calculations with DFT and multi-reference methods, for the preparation of compounds **1** and **8** (including two methods for the preparation of **1**¹⁹), and for IR and EPR spectroscopic experiments are given in the Supporting Information.

All calculations were performed with the ORCA program package, version 4.1.1 and 4.2.⁵⁶

The spectroscopic experiments were carried out at the VUV beamline of the Swiss Light Source (SLS) at the Paul Scherrer Institute, Villigen/CH. In most experiments the photon energy was scanned in 5 meV steps and calibrated using autoionization resonances in Ar. The ionization energies reported in the main paper are accurate to within ±20 meV and were corrected for the Stark-shift by the extraction field (8–9 meV). Note that in some experiments 10 meV steps were used. A detailed description of the beamline is given in the literature.⁴⁷

[MeBi(SPh)₂] (8**).**³⁹ Neat trimethyl bismuth (100 mg, 0.394 mmol) and diphenyl disulfide (85.9 mg, 0.393 mmol) were heated to 120 °C for 3 d. After cooling to ambient temperature, a green solid was obtained, which was dissolved in benzene (1.5 mL) and layered with *n*-pentane (1.5 mL). After 16 h at ambient temperature, the product could be isolated by filtration and dried *in vacuo* as yellow needles. Yield: 70.8 mg, 0.160 mmol, 41%.

¹H NMR (500 MHz, C₆D₆): δ = 0.94 (s, 3H, CH₃), 6.86 (dd, 2H, ³J_{HH} = 7.4 Hz, ³J_{HH} = 7.5 Hz, *p*-C₆H₅), 7.03 (t, 4H, ³J_{HH} = 7.6 Hz, *m*-C₆H₅), 7.48 (d, ³J_{HH} = 7.8 Hz, *o*-C₆H₅) ppm.

¹³C NMR (126 MHz, C₆D₆): δ = 40.40 (br, CH₃), 127.29 (s, *p*-C₆H₅), 128.60 (s, *m*-C₆H₅), 135.62 (s, *o*-C₆H₅), 136.06 (s, *ipso*-C₆H₅) ppm.

Elemental analysis: Anal. calc. for: [C₁₃H₁₃BiS₂] (442.35 g/mol): C 35.30, H 2.96, S 14.50; found: C: 35.12, H 2.90, S 14.40.

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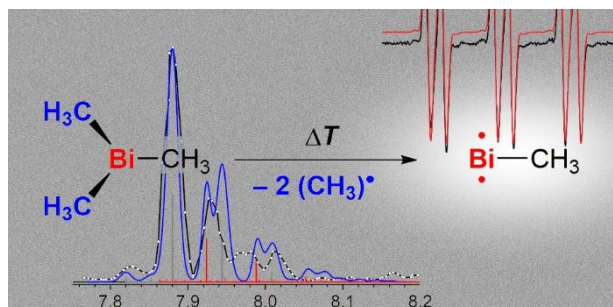


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Methylbismuth, the first organometallic non-stabilized bismuth(I) compound, was generated in the gas phase and characterized. Implications for its chemistry in the condensed phase were investigated.

