

Practical Catalytic Asymmetric Synthesis of Diaryl-, Aryl Heteroaryl-, and Diheteroarylmethanols

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Abstract: Enantioenriched diaryl-, aryl heteroaryl-, and diheteroarylmethanols exhibit important biological and medicinal properties. One-pot catalytic asymmetric syntheses of these compounds beginning from readily available aryl bromides are introduced. Thus, lithium-bromide exchange with commercially available aryl bromides and *n*-BuLi was followed by salt metathesis with ZnCl₂ to generate ArZnCl. A second equivalent of *n*-BuLi was added to form the mixed organozinc, ArZnBu. In the presence of enantioenriched amino alcohol-based catalysts, ArZnBu adds to aldehydes to afford essentially racemic diarylmethanols. The low enantioselectivities were attributed to a LiCl-promoted background reaction. To inhibit this background reaction, the chelating diamine TEEDA (tetraethylethylene diamine) was introduced prior to aldehyde addition. Under these conditions, enantioenriched diarylmethanols were obtained with >90% ee. Arylations of enals generated allylic alcohols with 81–90% ee. This procedure was unsuccessful, however, when applied to heteroaryl bromides, which was attributed to decomposition of the heteroaryl lithium under the salt metathesis conditions. To avoid this problem, the metathesis was conducted with EtZnCl, which enabled the salt metathesis to proceed at low temperatures. The resulting EtZn(Ar_{Hetero}) intermediates (Ar_{Hetero} = 2- and 3-thiophenyl, 2-benzothiophenyl, 3-furyl, and 5-indolyl) were successfully added to aldehydes and heteroaryl aldehydes with enantioselectivities between 81–99%. These are the first examples of catalytic and highly enantioselective syntheses of diheteroarylmethanols. In a similar fashion, ferrocenyl bromide was used to generate FcZnEt and the ferrocenyl group added to benzaldehyde and heteroaromatic aldehydes to form ferrocene-based ligand precursors in 86–95% yield with 96–98% ee. It was also found that the arylation and heteroarylation of enals could be followed by diastereoselective epoxidations to provide epoxy alcohols with high enantio- and diastereoselectivities in a one-pot procedure.

1. Introduction

Enantioenriched diaryl-, aryl heteroaryl-, and diheteroarylmethanols are important intermediates and structural motifs in medicinal chemistry. Diarylmethanols form the core of several biologically active compounds, including (*R*)-neobenodine, (*R*)-orphenadrine, and (*S*)-cetirizine.^{1–6} The 1-benzofuran derivatives **1a–c** are intermediates in the synthesis of chiral azoles (**2a–c**, Figure 1). Compounds **2a–c** were initially examined as antifungal agents⁷ and have been found to be powerful nonsteroidal aromatase inhibitors.^{8–13} They are indicated in the

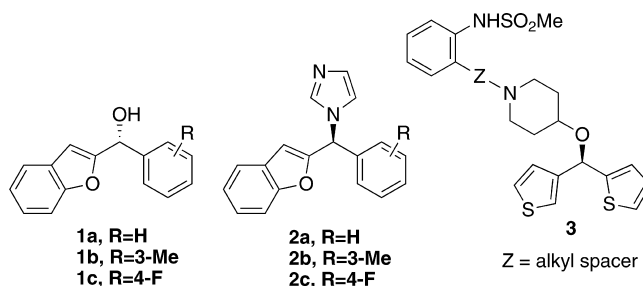


Figure 1. Biologically active heteroaryl- and diheteroarylmethanols.

treatment of hormone-dependent breast cancer.¹⁴ Likewise, diheteroarylmethanols have received recent attention. For example, chiral dithienylmethanols have been evaluated as antiallergic and antiischemic agents (**3**, Figure 1).¹⁵ Furthermore, enantioenriched diarylmethanols can be converted into diaryl-

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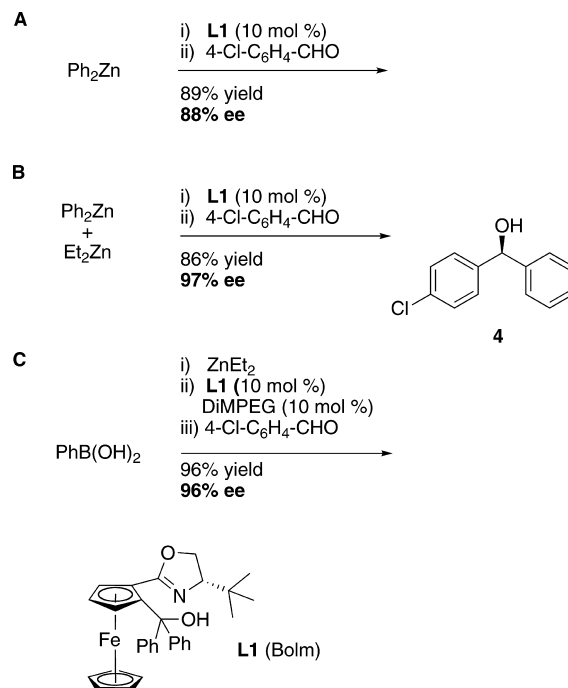
methane derivatives via S_N2 substitution at the C–O bond without loss of ee.¹ The diarylmethane motif is found in antimuscarinics,¹⁶ antidepressants,¹⁷ and endothelin antagonists.¹⁸

The catalytic enantioselective synthesis of diarylmethanols has been the focus of many studies.¹⁹ The most efficient approach to their preparation is the arylation of aromatic aldehydes to generate a C–C bond and stereocenter in a single step. Early studies by Seebach and co-workers employed Ph–Ti(O-*i*Pr)₃, generated from Cl–Ti(O-*i*Pr)₃ and PhLi, in combination with TADDOL-based titanium catalysts.^{20–22} To achieve high enantioselectivity, it was necessary to remove the LiCl byproduct formed during salt metathesis by centrifugation.

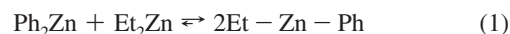
Pioneering studies by Fu and co-workers²³ with a planar-chiral catalyst, diphenylzinc, and 4-chlorobenzaldehyde furnished the diarylmethanol product with 57% ee. A highly enantioselective catalyst was reported by Pu shortly thereafter.²⁴ These works inspired many subsequent investigations using diphenylzinc, as exemplified in Scheme 1A.^{25–30}

Despite the large number of studies with diphenylzinc, significant drawbacks remained. Diphenylzinc is prohibitively expensive and limited to phenyl transfer.^{25,31–34} Furthermore, unlike dialkylzinc additions to aldehydes, which exhibit slow background reactions, the uncatalyzed addition of diphenylzinc to aldehydes is sufficiently rapid to compete with most catalyzed additions.^{34–36} The latter problem was addressed by Bolm and co-workers, who discovered that the mixed reagent EtZnPh³³ exhibited a slower background reaction than diphenylzinc. EtZnPh also resulted in higher enantioselectivities with the same

Scheme 1. Asymmetric Arylation of 4-Chlorobenzaldehyde with Ph₂Zn (A) and PhZnEt Generated from Ph₂Zn and Et₂Zn (B) or from PhB(OH)₂ and Et₂Zn (C)



catalysts (Scheme 1B), in part from the reduced contribution of the background reaction.^{37–42} EtZnPh is easily generated by combining Ph₂Zn and Et₂Zn (eq 1).



The development of methods for the enantioselective transfer of substituted aryl groups to aldehydes remained a challenge for several years. Although diarylzinc reagents can be easily prepared from ArLi or ArMgX and zinc halides (ZnX₂), the salt byproducts LiX and MgX₂ are Lewis acidic and readily promote the background reaction in the presence of enantioenriched catalysts, resulting in diarylmethanols with little or no ee. To bypass this problem, Bolm and co-workers introduced a method whereby aryl boronic acid derivatives underwent transmetalation with dialkylzinc reagents to provide access to salt-free arylzinc reagents, ArZnEt (Scheme 1C).^{30,42–48} In this

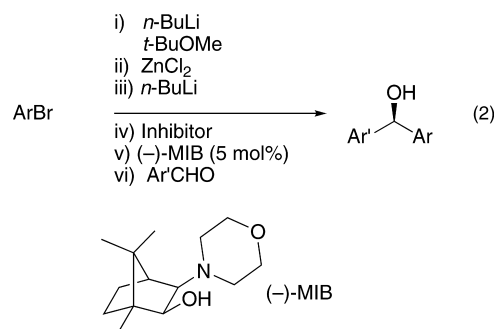
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procedure,⁴⁴ 2.4 equiv of the aryl boronic acid was heated with 7.2 equiv of diethylzinc for 12 h to generate ArZnEt. Dimethyl (polyethylene glycol) (DiMPEG, 10 mol %, MW $\approx$ 2000) was used as an additive to inhibit the background reaction caused by achiral Lewis acidic ZnPh₂ or ZnBr₂.^{43,49,50} Further study and optimization by the groups of Pericàs and Magnus significantly reduced the transmetalation time.^{51,52} Braga demonstrated that the acceleration could be achieved by microwave irradiation.⁵³ Based on Bolm's breakthrough, several aryl boron derivatives, such as triarylboranes^{42,43,49,54–57} and boroxines,^{52,58,59} were successfully employed in the asymmetric arylation of aldehydes. Highly enantioselective late transition metal-based catalysts can also be employed with boronic acid derivatives.^{60–62}

Main group metals other than zinc have been applied to the asymmetric arylation of aldehydes. In 2008 Muramatsu and Harada⁶³ introduced a method wherein Grignard reagents (1.2 equiv) could be added to titanium tetraisopropoxide (3 equiv) in the presence of 2 mol % 3-(3,5-Ph₂-C₆H₃)-H₈-BINOL and aldehydes to provide diarylmethanols with high ee. In a similar vein, Gau and co-workers employed salt-free Ar₃Al·THF reagents in combination with titanium tetraisopropoxide and either H₈-BINOL⁶⁴ or sulfonamide alcohol-based ligands,⁶⁵ which led to diarylmethanols with high enantioselectivities. These reactions may proceed via an Ar–Ti intermediate.⁶⁶

At the outset of our research into the arylation of aldehydes in 2005, two major limitations existed. The first was the necessity for salt-free aryl zinc reagents. The second was the use of costly aryl sources such as Ph₂Zn and aryl boronic acids, which are synthesized from aryl halides. To address these problems, we deemed the following criteria essential to a practical, cost-effective and scalable protocol: (1) to use readily available aryl bromides, and (2) to avoid filtration or centrifugation^{20–22,43,67} of metal halide byproducts from the aryl organometallic reagent. Herein we report the full details of the successful development of a method that fulfills these criteria.⁶⁸



Thus, metalation of an aryl bromide with *n*-BuLi, transmetalation to zinc, and enantioselective addition to aldehydes in the presence of the MIB-based^{69,70} catalyst can now be performed in a one-pot procedure (eq 2).⁶⁸ To circumvent the need for tedious sublimation, filtration, or centrifugation of the intermediate arylzinc reagents, we introduced a method to sequester the LiCl byproduct, enabling the generation of diarylmethanols with high levels of enantioselectivity in the presence of lithium chloride. Unfortunately, this procedure was unsuccessful when applied to the generation of enantioenriched diheteroarylmethanols. Therefore, an alternative procedure for heteroaryl additions to aldehydes was developed. To our knowledge, these studies represent the first highly enantioselective catalytic asymmetric synthesis of diheteroarylmethanols.

Table 1. Solvent Screen in the Asymmetric Phenyl Addition to 2-Naphthaldehyde

entry	solvent	ee (%)
1	toluene	94
2	Et ₂ O	60
3	<i>t</i> -BuOMe	88
4	<i>t</i> -BuOMe/Hex (1:3)	89
5	<i>t</i> -BuOMe/Hex (1:3)	92 ^a
6	2 PhLi + ZnCl ₂ in <i>t</i> -BuOMe/Hex (1:3)	2

^a Reaction conducted at 0 °C.

2. Results and Discussion

The ultimate goal of these investigations was to develop a practical method for the addition of aryl and heteroaryl groups to aldehydes using readily available aryl and heteroaryl bromides. Asymmetric additions with commercial diphenylzinc were used to evaluate catalyst enantioselectivity and for comparison with reactions using bromobenzene.

2.1. Phenylation with Ph₂Zn and MIB. Our first priority was to determine the enantioselectivity of the (–)-MIB-based catalyst in phenyl additions to aldehydes. The substrate selected for these studies was 2-naphthylaldehyde (Table 1), which was used with commercial ZnPh₂ and 5 mol % (–)-MIB.^{69,70} The phenyl addition proceeded with 94% enantioselectivity in toluene and 60% enantioselectivity in diethyl ether (entries 1 and 2). Diethyl ether most likely binds to the MIB-based zinc catalyst, reducing

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its activity and, therefore, enantioselectivity. When less coordinating *tert*-butyl methyl ether (TBME) was used, the diaryl-methanol was obtained with 88% enantioselectivity (entry 3). In the mixed solvent composed of 1:3 TBME:hexanes the enantioselectivity was 89% at rt (entry 4) and 92% at 0 °C (entry 5). Of the solvents examined, only TBME was suitable for the salt metathesis of PhLi with ZnCl₂.

To evaluate the possibility of beginning with aryl bromides, we next generated ZnPh₂ by metalation of 4.5 equiv PhBr with 4 equiv *n*-BuLi in TBME followed by transmetalation with 2 equiv ZnCl₂. After addition of hexanes to precipitate additional LiCl, the in situ generated Ph₂Zn solution was used in place of the commercial Ph₂Zn under otherwise identical conditions (Table 1, entry 6). The expected alcohol product was isolated, but the ee was only 2%. We hypothesized that the Lewis acidic LiCl, generated en route to ZnPh₂, promoted the addition to form the racemate faster than the amino alcohol-based Lewis acid catalyst promoted the asymmetric addition. Other researchers have had varying degrees of success employing either filtration or centrifugation of LiCl and MgX₂ byproducts.^{20–22,33,67}

These salt byproducts are often produced as a fine particulate and are difficult to remove. Although these procedures are quite useful on laboratory scale, filtration or centrifugation of highly air-sensitive materials is less practical on large scale. To overcome this problem, our strategy was to inhibit the LiCl byproduct rather than remove it. A similar approach was devised by Bolm and co-workers involving the addition of Ph₂Zn to aldehydes. These researchers observed a beneficial effect of dimethoxypoly(ethylene glycol) (DiMPEG) on the catalyst enantioselectivity^{43,49} and proposed that DiMPEG suppressed reactions catalyzed by trace achiral Lewis acids, including ZnBr₂ and LiBr, allowing the arylation reaction to proceed via the ligand-accelerated⁷¹ pathway.¹⁹ Although enantioselectivities reached 93%, yields ranged from 8–31% when Ph₂Zn was generated from PhLi and ZnBr₂.⁴³ Furthermore, we had difficulties with reproducibility using DiMPEG in combination with the MIB-based catalyst for enantioselective vinylation of aldehydes.^{72,73}

2.2. Development of a Lithium Chloride Selective Inhibitor.

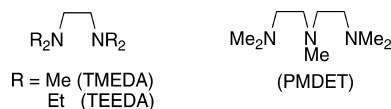
The lack of enantioselectivity with Ph₂Zn generated from PhBr in Table 1 (entry 6) suggested that the achiral LiCl is a more active Lewis acid than the (–)-MIB-based zinc catalyst. There are three important differences between the lithium and zinc Lewis acids: (1) the lithium is more electropositive and probably the stronger Lewis acid, (2) the lithium center is less sterically saturated than the zinc center in the MIB-based catalyst and (3) the lithium chloride has at least two available coordination sites while the MIB-based zinc catalyst has only one accessible site. Based on this analysis, our strategy was to employ bidentate inhibitors that would chelate lithium and bind tightly, but coordinate in a monodentate fashion to the chiral zinc catalyst. Support for this approach was gained through structures of [TMEDA·LiCl]_n, which contain four-coordinate lithium centers with bridging chlorides.^{74,75}

On the basis of this proposal, multidentate amines were screened as LiCl inhibitors in the catalytic enantioselective

Table 2. Examination of Possible LiCl Inhibitors

$\text{PhBr} \xrightarrow[\text{4.5 equiv}]{\begin{array}{l} \text{i) } n\text{-BuLi (4.0 equiv)} \\ \text{ii) } t\text{-BuOMe} \\ \text{iii) ZnCl}_2 \text{ (2.1 equiv)} \end{array}} 2 \text{ Ph}_2\text{Zn} + 4 \text{ LiCl} \xrightarrow[\begin{array}{l} \text{v) (–)-MIB (5 mol\%)} \\ \text{vi) ArCHO, rt} \end{array}]{\begin{array}{l} \text{iv) Inhibitor Hex/Tol} \end{array}} \text{Ar-CH(OH)-Ph} \quad \mathbf{5}$			
Ar=2-naphthyl			
entry	inhibitor	equiv.	ee (%)
1	none	—	2
2	TEEDA	0.2	55
3	TEEDA	0.4	76
4	TEEDA	0.8	83
5	TEEDA	1.0	83
6	TEEDA	1.2	77
7	TEEDA	0.8	89
8	TEEDA	0.8	92 ^b
9	TMEDA	0.8	81
10	TMEDA	1.0	81
11	PMDET	0.1	51
12	PMDET	0.2	80
13	PMDET	0.4	71

^a Solvent = *t*-BuOMe/Hex = 1:3 unless noted. ^b Addition conducted at 0 °C.



phenylation of 2-naphthaldehyde with in situ prepared ZnPh₂ (Table 2). *N,N,N',N'*-Tetraethylethylenediamine (TEEDA) was first examined, because the amino groups are slightly larger than those of the tetramethyl analog TMEDA. Use of 0.2 equiv TEEDA resulted in an increase in the product ee from 2% in the absence of diamine to 55% ee (Table 2, entries 1 and 2). Increasing the amount of TEEDA from 0.2–0.8 and 1.0 equiv resulted in product ee's of up to 83%. A further increase in TEEDA to 1.2 equiv, however, resulted in a slower reaction and a decrease in the product ee to 77%, probably due to inhibition of the MIB-based zinc catalyst by the diamine. It was found that addition of 5 equiv toluene (or hexanes) relative to TBME after transmetalation led to higher enantioselectivity (up to 89%, entry 7). When the temperature of the addition was lowered from rt (entry 7) to 0 °C (entry 8) the enantioselectivity increased to 92%. The same enantioselectivity was obtained with commercial Ph₂Zn in Table 1 (entry 5), indicating that TEEDA is an excellent inhibitor of LiCl. For comparison, TMEDA was examined in Table 2 (entries 9 and 10) and found to be nearly as effective as TEEDA. Pentamethyldiethylenetriamine inhibited LiCl at lower concentrations (0.2 equiv, 80% product ee, entry 12).

2.3. Generation and Application of Mixed Aryl Alkyl Zinc Reagents. As outlined in the Introduction, the background reaction of diarylzinc reagents with aldehydes is often competitive with, or faster than the ligand accelerated pathway⁷¹ with amino alcohol-based catalysts. On the basis of the successful application of mixed aryl alkyl zinc reagents by Bolm and co-workers,³³ we desired to develop an in situ route to these species to increase enantioselectivities in the

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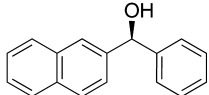
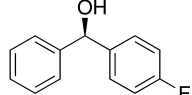
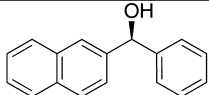
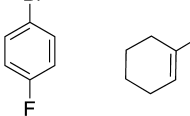
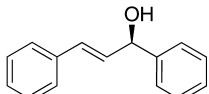
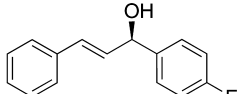
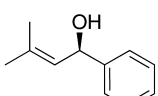
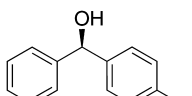
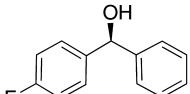
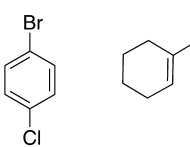
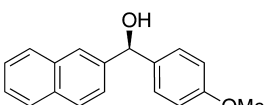
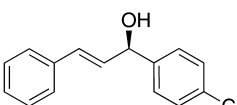
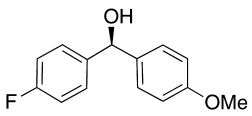
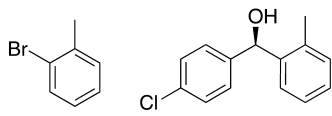
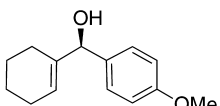
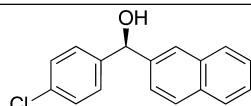
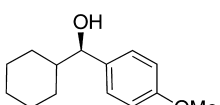
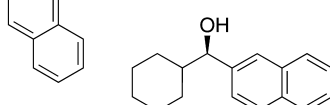
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Table 3. Catalytic Asymmetric Aryl Additions to Aldehydes with ArZnBu Generated from Aryl Bromides

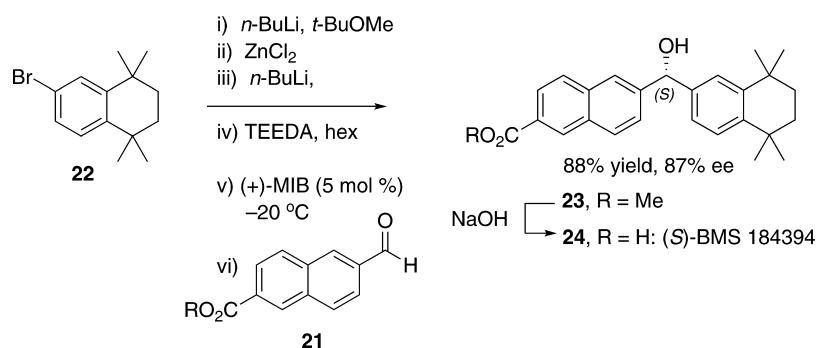
i) <i>n</i> -BuLi (2.0 equiv) ii) ZnCl ₂ (2.1 equiv) iii) <i>t</i> -BuOMe iv) <i>n</i> -BuLi (2.0 equiv)					iv) TEEDA (0.8 equiv) hex or tol v) (–)-MIB (5 mol%) vi) RCHO, 0 °C				
ArBr 2.0 equiv					ArZnBu				
entry	ArBr with Ar=	product	#	yield ee (%)	entry	ArBr with Ar=	product	#	yield ee (%)
1	Ph ₂ Zn/Et ₂ Zn		5	-- 97	10		13	75 97	
2			5	90 97	11		14	64 88	
3			6	95 88	12		15	74 84	
4			7	75 90	13		4	78 95	
5			8	80 96	14		16	55 87	
6			9	96 93	15		17	68 81	
7			10	84 93	16		18	73 95	
8			11	82 83	17		19	79 96	
9			12	84 78	18		20	76 82	

aldehyde arylations. To determine the benefit of the mixed organozinc reagents with MIB, our initial experiments involved conproportionation of a 1:1 ratio of commercial Ph₂Zn and Et₂Zn to generate PhZnEt (eq 1) followed by addition of (–)-MIB and 2-naphthaldehyde at 0 °C. Under these conditions, the enantioselectivity increased from 92% with Ph₂Zn (entry 5, Table 1 and entry 8, Table 2) to 97% with the mixed PhZnEt (Table 3, entry 1).

To prepare the mixed aryl alkyl zinc reagents in situ we chose to avoid the use of dialkylzinc reagents, focusing on the more readily available alkylolithiums. Thus, metalation of PhBr with *n*-BuLi (2 equiv each) and addition of 2.1 equiv ZnCl₂ resulted in the generation of PhZnCl. A second dose of *n*-BuLi (2 equiv) was then added to produce PhZnBu, which was used in combination with 0.8 equiv TEEDA in the asymmetric addition reaction (Table 3, entry 2). Gratify-

ingly, the enantioselectivity with the in situ generated PhZnBu (97%) was equal to the salt-free PhZnEt, despite the 4 equiv of LiCl in the reaction vessel.

To determine the generality of this method, a series of aryl bromides and aldehydes were employed (Table 3). Bromobenzene and 4-substituted aryl bromides bearing -OMe, -F, or -Cl substituents were used in the arylation of benzaldehyde derivatives with 93–97% enantioselectivity. 2-Bromotoluene and 2-bromonaphthalene were added to benzaldehydes with ≥95% enantioselectivity. Aryl additions to α,β-unsaturated aldehydes occurred with 81–90% enantioselectivity (entries 3, 4, 8, 11, 12, 14, and 15). The aliphatic substrate, cyclohexanecarboxaldehyde, underwent aryl addition with 78–82% enantioselectivity (entries 9 and 18). The examples in Table 3 are the first examples of aldehyde arylation beginning with aryl bromides.⁶⁸

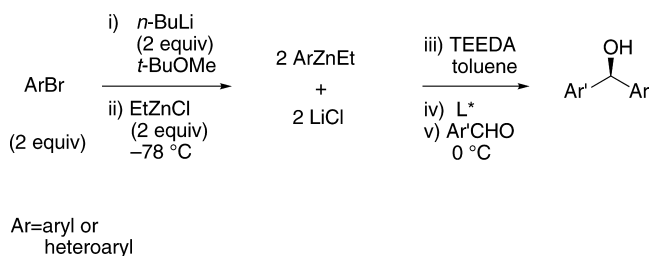
Scheme 2. Synthesis of Enantioenriched **23**, the Key Intermediate in the Synthesis of (*S*)-BMS 184394

Subsequent to our initial communication,⁶⁸ a related report appeared by Pu⁷⁶ employing aryl iodides, and two examples were reported by Harada.^{63,77} Woodward also developed a method using aryl zinc halides in combination with trimethyl aluminum based on the Schlenk equilibrium.⁷⁸

Formal Synthesis of (*S*)-BMS 184394. One example of a biologically active diarylmethanol is BMS 184394 (**24**, Scheme 2), a RAR γ selective retinoid with activity against skin diseases and cancers, in particular breast cancer and acute promyelocytic leukemia.^{79–81} Using conditions outlined in Table 3, 3.0 equiv aryl bromide **22** (Scheme 2) was employed to generate the mixed aryl butyl zinc reagent. TEEDA (1.5 equiv) and hexanes were added followed by (+)-MIB (5 mol %) and aldehyde **21**. The addition product **23** was produced with 87% enantioselectivity in 88% yield (Scheme 2). Conversion to (*S*)-BMS 184394 can be accomplished by saponification of the ester.⁸⁰

2.4. Attempted Heteroaryl Additions to Aldehydes. General, highly enantioselective additions of heteroaryl groups to aldehydes have not been developed. To our knowledge, the only examples of highly enantioselective heteroaryl additions were published in 2008 by Gau and involved the addition of 2-furyl aluminum reagents to ketones.⁸² Heteroaryl groups are among the most important pharmacophors in medicinal chemistry and diheteroarylmethanols have been identified as biologically active structural motifs.¹⁵ Thus, not only would methods for heteroaryl additions to aldehydes increase the classes of enantioenriched diarylmethanols accessible, it would enable the catalytic asymmetric synthesis of diheteroarylmethanols that are currently not directly accessible.

With the goal of introducing asymmetric heteroaryl additions to aldehydes, we applied our arylation procedure to metalation of 3-bromothiophene followed by addition to benzaldehyde. The only modification was to maintain the temperature of the heteroaryl lithium at $-78\text{ }^{\circ}\text{C}$. Unfortunately, no addition product was observed. When the aryl bromides were used under the

Scheme 3. Metathesis with EtZnCl for the Aryl and Heteroaryl Additions to Aldehydes

Ar = aryl or heteroaryl

conditions outlined in Table 3, the salt metathesis was conducted at room temperature for 4.5 h.⁶⁸ At this temperature (3-thienyl)Li readily decomposes.

We hypothesized that the absence of product was due to decomposition of the heteroaryl lithium in the transmetalation step, which was complicated by the limited solubility of ZnCl₂ in TBME at low temperature. To address this problem, we envisaged a more soluble zinc source might undergo transmetalation at lower temperature. Our choice of EtZnCl was based on the large reactivity difference of sp² hybridized Zn–C bonds over their sp³ counterparts. Another advantage of EtZnCl is that only a single equivalent of LiCl forms during the metathesis, whereas ZnCl₂ produces two equivalents (Scheme 3). Lower levels of salt byproduct facilitate inhibition of the LiCl-promoted background reaction.

The synthesis of EtZnCl was initially performed following the method of Woodward and co-workers by combination of ZnCl₂ and ZnEt₂ in THF followed by removal of the solvent under reduced pressure.⁸³ Using EtZnCl prepared in this manner, the transmetalation proceeded at $-78\text{ }^{\circ}\text{C}$ and the desired heteroaryl addition products were obtained. Unfortunately, product yields and ee's varied greatly from run to run. Our unsuccessful attempts to develop asymmetric heteroaryl additions convinced us to first focus on development of a low temperature transmetalation and then revisit enantioselective heteroaryl additions.

2.5. Development of Low Temperature Transmetalation Conditions and Synthesis of Biologically Active 2a. Momentarily stepping away from the more challenging enantioenriched diheteroarylmethanols, we concentrated on developing low temperature conditions for lithium to zinc transmetalations. We attributed the inconsistencies in the previously described heteroaryl additions to the presence of residual zinc-bound THF in the EtZnCl, which was observed by ¹H NMR spectroscopy.⁸³ THF is known to inhibit the MIB-based zinc Lewis acid catalyst.

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Table 4. Synthesis of Aryl Heteroaryl- and Diheteroarylmethanols using 5 mol % (–)-MIB (Unless Noted below)

entry	ArheteroBr	product	#	yield	ee (%)	entry	ArheteroBr	product	#	yield	ee (%)
1			25	68	90	8 ^a			32	57	90
2 ^a			26	72	92	9			33	86	93
3			27	83	93	10 ^a			34	67	80
4			28	60	94	11			35	60	99
5 ^a			29	65	88	12			36	79	94
6 ^a			30	70	81	13			37	61	89
7 ^a			31	70	82						

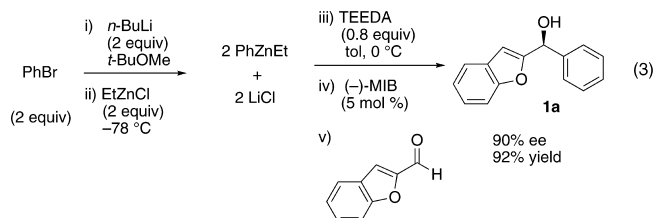
^a MIB (10 mol %). See Supporting Information for details.

On the basis of this hypothesis, an alternative synthesis of EtZnCl was pursued.^{84–86} Using toluene in place of THF required heating ZnEt₂ and sparingly soluble ZnCl₂ at 60 °C for 72 h, after which the solution was filtered to remove any unreacted ZnCl₂. The volatiles were then removed under reduced pressure to afford EtZnCl as a white solid that could be stored under nitrogen for months.

The THF-free EtZnCl was first employed with bromobenzene (eq 3). The transmetalation was conducted at –78 °C to generate PhZnEt, and the addition to 2-benzofurancarbaldehyde was performed at 0 °C. After 12 h, the reaction mixture was quenched with water, worked up, and purified on deactivated silica. We were pleased to isolate the desired addition product **1a** in 92% yield with 90% ee (eq 3). Compound **1a** was converted to the promising breast cancer treatment candidate **2a** (Figure 1), without loss of ee in 41% unoptimized yield (84% based on recovered **1a** of 90% ee) via a Mitsunobu reaction with imidazole.⁸⁷ S_N2 substitutions of this type are known to be very difficult.¹ Alternative methods for this transformation also appear potentially useful.^{89–92}

Recently promising diarylmethanes have been examined as possible inhibitors and receptor agonist candidates, but due to

limited methods to synthesize the diarylmethanols enantioselectively, most of the studies employed racemic material.^{11,93,94} Synthesis of **1a** with 90% ee suggests that this procedure can be used to prepare diarylmethanols and their derivatives with high ee.



2.6. Enantioselective Addition of Heteroaryl Groups to Aldehydes. The heteroaryl addition was attempted with 3-bromothiophene under the conditions employed with bromobenzene to generate **1a** in eq 3. Thus, after metalation of 3-bromothiophene with *n*-BuLi, transmetalation was performed at –78 °C with THF-free EtZnCl. The resulting solution was then warmed to 0 °C and TEEDA, (–)-MIB, and benzaldehyde were added (Table 4). After stirring 12 h, workup, and purification we were pleased to isolate the desired heteroaryl addition

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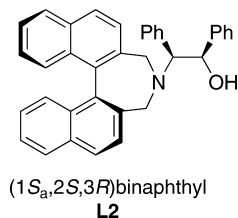


Figure 2. Structure of Chan's ligand (**L2**).

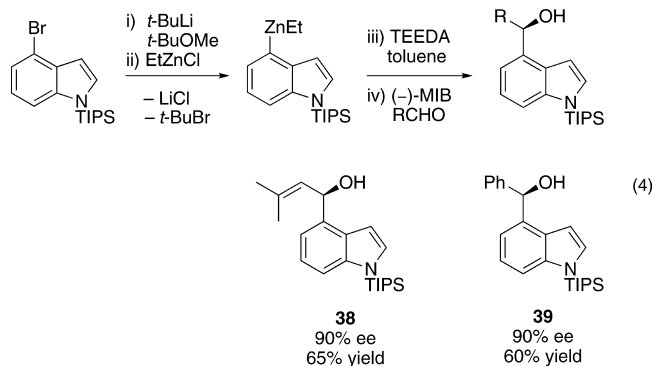
product in 68% yield with 90% ee (Table 4, entry 1). These revised conditions led to reproducible product ee's and yields.

To explore the enantioselective synthesis of diheteroaryl-methanols, the optimized conditions for addition of (3-thienyl)ZnEt to benzaldehyde were employed with heteroaromatic aldehydes. Thus, addition to 5-methyl-2-furan carboxaldehyde, 2-thiophenecarboxaldehyde and 3-benzofurancarboxaldehyde occurred with 92–94% enantioselectivity in 60–83% yield (Table 4, entries 2–4). The differences in yields in entries 1–4 probably arise from a combination of the instability of the heteroaryl organometallic reagents and diminished electrophilicity of the heteroaromatic aldehydes. To develop practical and useful methods, scalability must be demonstrated. Thus, for the synthesis of **27**, precursor to potential drug candidate **3** (Figure 1), the asymmetric addition was scaled to produce 820 mg (83% yield and 93% ee, entry 3).

Other heterocycles such as 3-bromobenzothiophene can also be used in the addition with very good enantioselectivities (81–88% ee, entries 5–7). Employing 2-bromothiophene and benzaldehyde afforded diarylmethanol in 90% ee and 57% yield (entry 8). In a similar fashion 3-furanyl ethyl zinc can be added to benzaldehyde (93% ee, 86% yield, entry 9) and heteroaromatic aldehydes with excellent enantioselectivities (89–99% ee, entries 11–13). 5-Methyl-2-furan carboxaldehyde gave addition product of 80% ee (entry 10). Attempts to add 2-furanylzinc reagents to aldehydes, however, resulted in poor enantioselectivities, probably due to the presence of the coordinating oxygen in close proximity to zinc.

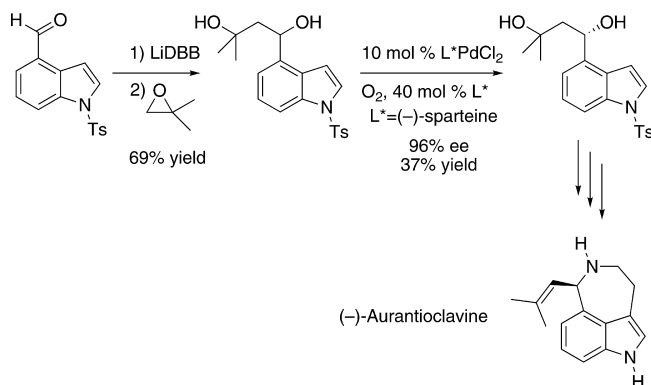
To determine if our method for addition of heteroaromatic groups to aldehydes could be extended to other catalysts, we examined the use of Chan's ligand (**L2**, Figure 2)^{48,95} with 3-bromothiophene and benzaldehyde under the conditions listed in Table 4, which led to product of 90% ee and 70% yield. These results are virtually identical to those in entry 1 (Table 4) with MIB, indicating that our strategy employing TEEDA to inhibit LiCl is applicable to other amino alcohol-based catalysts.

Indoles are regarded as privileged structures in medicinal chemistry and are substructures of an enormous variety of natural products.^{96,97} We therefore turned our attention toward the synthesis of enantioenriched diarylmethanols containing the indole motif. Metalation of *N*-silyl-protected 4-bromoindole with *n*-BuLi was unsuccessful under a variety of conditions, including those in Table 4, most likely due to the electron rich nature of the heterocyclic π -system. More challenging metal–halogen exchange reactions are generally performed with two equiv *t*-BuLi.⁹⁸ In these reactions, the first equiv undergoes the



metal–halogen exchange with the aryl bromide generating *t*-BuBr and the second drives the equilibrium by promoting elimination of the liberated *t*-BuBr to produce isobutylene and LiBr. Unfortunately, diamines that inhibit LiCl had little impact when LiBr was formed. Although we do not understand the intimate differences between LiCl and LiBr at this time, we speculate that weaker bridging Li–Br interactions in [(diamine)LiBr]_n facilitate dissociation of the oligomers, opening a coordination site on lithium. To avoid production of LiBr, a 1:1 ratio of 4-bromoindole to *t*-BuLi was employed, furnishing indole-based diarylmethanols with 90% ee and 60–65% yield (eq 4).

Scheme 4. Structure of (–)-Aurantioclavine and Intermediates in its Synthesis by Stoltz and Coworkers⁹⁹



Enantioenriched indole **38** is a potential intermediate for the synthesis of (–)-aurantioclavine, which is illustrated in Scheme 4 along with related intermediates in the elegant synthesis of this alkaloid by Stoltz and co-workers.⁹⁹

2.7. Tandem Asymmetric Aryl Addition/Diastereoselective Epoxidation. We recently developed a series of tandem reactions involving the asymmetric addition of alkyl,^{100–102} vinyl,^{73,100,103,104}

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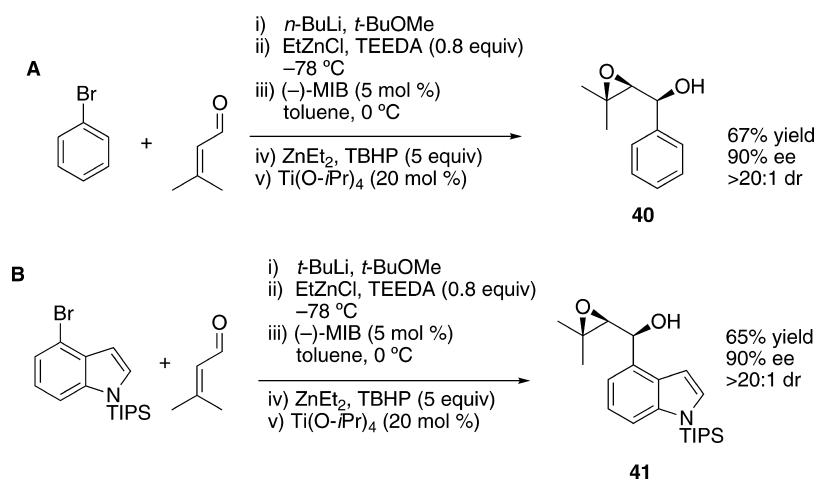
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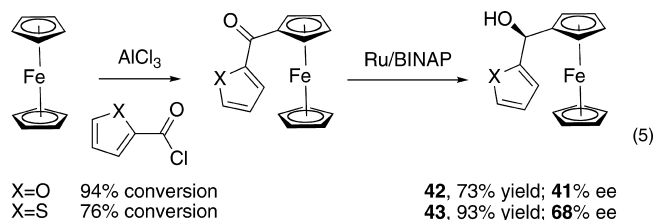
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Scheme 5. Tandem Asymmetric Arylation of Aldehydes/Diastereoselective Epoxidation

or allyl¹⁰⁵ groups to aldehydes and ketones followed by diastereoselective epoxidation to provide epoxy alcohols with three contiguous stereogenic centers.¹⁰⁶ These one-pot procedures rapidly increase molecular complexity in a synthetically efficient fashion. To explore the possibility of performing arylation and heteroarylation of enals followed by diastereoselective epoxidation, we examined asymmetric phenyl addition/oxidation with 3-methyl-2-butanone. As shown in Scheme 5A, using the conditions outlined in Table 4 the catalytic asymmetric phenyl addition was performed. The resulting enantioenriched zinc allylic alkoxide was then treated with Et₂Zn (1 equiv), TBHP (*tert*-butylhydroperoxide, 5 equiv), and 20 mol % titanium tetraisopropoxide at 0 °C. The epoxidation reached completion in 3 h, after which the reaction mixture was quenched, worked up, and the product purified by chromatography to afford the epoxy alcohol in 67% yield with 90% ee and >20:1 dr (as determined by ¹H NMR). The heteroarylation/epoxidation was examined with the TIPS protected 4-bromindole (Scheme 5B). Metalation with *t*-BuLi, transmetalation with EtZnCl, and asymmetric addition as performed in eq 4 was followed by addition of Et₂Zn, TBHP and 20 mol % titanium tetraisopropoxide at 0 °C. Following workup and purification, the enantioenriched indole epoxy alcohol was isolated in 65% yield with 90% ee and >20:1 dr. Interestingly, attempted epoxidation of the isolated indole allylic alcohol product in Equation 4 with *m*-CPBA resulted in formation of the epoxy alcohol in low yield accompanied by several side products. The examples in Scheme 5 indicate that the asymmetric arylation and heteroarylation are compatible with our tandem diastereoselective epoxidation conditions and could be used to prepare an array of functionalized epoxy alcohols.

2.8. Synthesis of Ferrocenylzinc and Applications to Asymmetric Additions. Having developed successful methods for the enantioselective addition of aryl and heteroaryl groups to aldehydes, we focused on the generation of the ferrocenylzinc reagent, (Fc)ZnEt. Highly enantioselective additions of ferrocenylzinc reagents to aromatic and heteroaromatic aldehydes would provide rapid access to heteroaryl ferrocenyl methanols.

Related motifs^{107,108} are precursors to important enantioenriched ferrocene-based ligands such as BoPhoz,¹⁰⁹ Josiphos,¹¹⁰ FERRIPHOS,^{111,112} Pigiphos,^{113,114} PPFA,¹¹⁵ Walphos,¹¹⁶ Taniaphos¹¹⁷ and Trap.^{109,118,119} The ferrocenyl methanol scaffold is often synthesized by CBS^{120–122} or Ru/BINAP reduction of ferrocenyl ketones.^{4,119,123} Asymmetric reduction of heteroaromatic ketone derivatives, however, resulted in only moderate enantioselectivity (X = O, 41% ee; X = S, 68% ee, eq 5).¹²⁴



Beginning with ferrocenyl bromide (FcBr) and applying the conditions used in Table 4 to the generation and addition of

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Table 5. Asymmetric Addition of (Fc)ZnEt to Aldehydes

entry	ligand	ArCHO	product	#	yield ee (%)
1	(-)-MIB			44	75 50
2	L2				86 98
3	L2			43	95 96
4	L2			42	95 96

(Fc)ZnEt to benzaldehyde with (-)-MIB provided product with a disappointing 50% ee (Table 5, entry 1). Inspired by the importance of functionalized ferrocenyl methanols, we screened other amino alcohol ligands. Fortunately, use of Chan's^{48,95} amino alcohol **L2** (Figure 2) with benzaldehyde provided the desired product in 86% yield with 98% ee (Table 5, entry 2). Use of 2-thiophenecarboxaldehyde and 2-furfural in combination with **L2** provided the ferrocene-based ligand precursors with enantioselectivities of 96% and yields of 95% (entries 3 and 4). It is known that substitution of furyl groups for phenyl can lead to an increase in catalyst enantioselectivity.¹²⁵

The high yields and stereochemical purity of functionalized ferrocenes make them attractive building blocks for the construction of new ferrocene-based ligands for asymmetric catalysis.

3. Summary and Outlook

Herein we described versatile methods for the generation of diaryl-, aryl heteroaryl-, and diheteroarylmethanols with high levels of enantioselectivity. The significance of these methods is that asymmetric arylation of aldehydes can now be initiated with aryl bromides, many of which are readily available. Key to the success of our procedures was introduction of a diamine, such as TEEDA. In the absence of TEEDA the addition reaction was promoted by LiCl, generating racemic products. The TEEDA inhibited the LiCl byproduct, allowing the asymmetric addition to proceed via the ligand accelerated pathway.^{71,126} Importantly, in the presence of the diamine it was not necessary to filter,⁴³ centrifuge,⁶⁷ or isolate the pyrophoric arylzinc reagents required with previous procedures, making our method desirable for large scale applications.

We also developed the first method for the synthesis of highly enantioenriched diheteroarylmethanols from readily available heteroaryl bromides. A crucial feature of this approach was the

use of EtZnCl in the transmetalation step with the heteroaryl-lithium at -78°C , at which temperature decomposition of the heteroaryl organometallic species was minimized. Use of EtZnCl in place of ZnCl₂ also halves the amount of LiCl byproduct, which was detrimental to the enantioselectivity in the asymmetric addition and must be inhibited by diamine. This method was also shown to be applicable to the tandem asymmetric addition/diastereoselective epoxidation to generate epoxy alcohols with two stereogenic centers in high enantio- and diastereoselectivity.

Finally, we also described the first examples of generation and highly enantioselective addition of ferrocenyl zinc reagents to aldehydes, opening the door to new enantioenriched ferrocene-based ligands. The straightforward methods introduced herein make possible the synthesis of functionalized, previously inaccessible enantioenriched diheteroarylmethanols. We anticipate that these methods will be useful in medicinal chemistry and asymmetric catalysis.

4. Experimental Section

General Considerations. All reactions were performed under a nitrogen atmosphere with oven-dried glassware using standard Schlenk or vacuum line techniques. The progress of reactions was monitored by thin-layer chromatography (TLC) performed on Whatman precoated silica gel 60 Å K6F plates and visualized by ultraviolet light or by staining with cerium-ammonium-molybdate. *t*-BuOMe was distilled from Na/benzophenone and toluene was dried through alumina columns. TEEDA was distilled and stored under nitrogen. The ¹H NMR and ¹³C{¹H} NMR spectra were obtained on a Bruker Fourier transform NMR spectrometer at either 300 or 500 or 75 or 125 MHz, respectively. ¹H NMR spectra were referenced to tetramethylsilane in CDCl₃ or residual protonated solvent; ¹³C{¹H} NMR spectra were referenced to residual solvent. Analysis of enantiomeric excess was performed using a Hewlett-Packard 1100 Series HPLC and a chiral column. Alternatively, a Berger SFC PionTo was employed when the compounds could not be resolved by HPLC. The optical rotations were recorded using a JASCO DIP-370. Infrared spectra were obtained using a Perkin-Elmer Spectrum 100 Series spectrometer. All reagents were purchased from Aldrich or Acros unless otherwise described. 3-Benzofurancarboxaldehyde was synthesized according to known procedure starting from commercially available 3-methylbenzofuran.¹²⁷ Binaphthyl amino alcohol ligand was synthesized according to Chan's procedure.^{48,95} EtZnCl was synthesized following Guerrero's method.^{84,128} All aldehyde substrates were distilled prior to use. Silica gel (Silicaflash P60 40–63 μm, Silicycle) was used for air-flashed chromatography.

Caution. Dialkylzinc and alkyl lithium reagents are pyrophoric! Care and appropriate laboratory equipment and attire must be used when handling these reagents.

Arylation of Aldehydes. Preparation of (4-Fluoro-phenyl)-(4-methoxy-phenyl)-methanol (10). A nitrogen purged Schlenk flask was charged with 4-bromoanisole (100.1 μL, 0.8 mmol) and *t*-BuOMe (1 mL) and cooled to -78°C . *n*-BuLi (0.32 mL, 2.5 M in hexanes, 0.8 mmol) was added dropwise and the solution was stirred for 1 h and the temperature raised to 0°C . ZnCl₂ (114.5 mg, 0.84 mmol) was added to the reaction mixture and it was stirred for 30 min. Additional *n*-BuLi (0.32 mL, 2.5 M in hexanes, 0.8 mmol) was added to the reaction mixture and the resulting solution was allowed to warm to rt and stirred 4.5 h. Toluene (5 mL) and TEEDA (68 μL, 0.32 mmol) were added to the reaction vessel and the solution was stirred. After 1 h (-)-MIB (4.8 mg, 0.02 mmol) was added and the reaction vessel was cooled to 0°C for 30 min. Finally, 4-fluorobenzaldehyde (43 μL, 0.4 mmol) was added and reaction mixture was stirred at 0°C and monitored by TLC. After completion (12 h), the reaction mixture was quenched with H₂O

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(20 mL) and extracted with ethyl acetate (3 × 20 mL). The organic layer was dried over MgSO₄, filtered, and the solvent was removed under reduced pressure. The crude product was purified by column chromatography on silica gel (hexanes/EtOAc, 95:5) to give **10** (77.7 mg, 84% yield) as a white crystalline solid (mp = 52 °C). The enantiomeric excess was determined by HPLC with a Chiralcel AS-H column (hexanes/2-propanol = 95:5, flow rate = 0.5 mL/min), *t_r* (1) = 20.0 min, *t_r* (2) = 22.1 min, $[\alpha]_D^{20} = +13.8$ (*c* = 0.195, THF, 93% ee); ¹H NMR (C₆D₆, 300 MHz): δ 2.17 (s, 1H), 3.38 (s, 3H), 5.50 (s, 1H), 6.81–6.95 (m, 4H), 7.17–7.29 (m, 4H); ¹³C{¹H} NMR (C₆D₆, 75 MHz): δ 55.0, 75.3, 114.3, 115.4 (d, *J* = 21.2 Hz), 128.4, 128.7 (d, *J* = 8.0 Hz), 136.9, 141.0 (d, *J* = 3.0 Hz), 159.7, 162.6 (d, *J* = 243 Hz); IR (neat): 831, 1033, 1248, 1504, 1609, 2837, 2957, 3422 cm⁻¹; HRMS calcd for C₁₄H₁₃FO₂ (M)⁺: 232.0900, found: 232.0900.

Furan-3-yl(thiophen-2-yl)methanol (35). A nitrogen purged Schlenk flask was charged with 3-bromofuran (67.0 μL, 0.75 mmol) and *t*-BuOMe (1 mL) and cooled to -78 °C. *n*-BuLi (0.3 mL, 2.5 M in hexanes, 0.75 mmol) was then added dropwise and the solution was stirred for 1 h at this temperature. During this time a white precipitate formed. Solid EtZnCl (97.0 mg, 0.75 mmol) was added to the reaction mixture at -78 °C followed by toluene (3 mL). The heterogeneous solution was stirred at -78 °C for 30 min and then warmed at 0 °C. TEEDA (64 μL, 0.30 mmol) was added and the solution stirred for an additional 30 min. (-)-MIB (190 μL, 0.1 M solution in hexanes, 0.019 mmol) was added to the reaction flask and the solution was stirred for 5 min before 2-thiophenecarboxaldehyde (35 μL, 0.37 mmol, dissolved in 1.5 mL of toluene) was added over 1.5 h by syringe pump. The reaction mixture was stirred at 0 °C and monitored by TLC until completion (approximately 10 h). The reaction mixture was diluted with 3 mL EtOAc and quenched with water (5 mL). The organic layer was separated and the aqueous solution extracted with EtOAc (3 × 10 mL). The combined organic layer was washed with brine (5 mL), dried over MgSO₄, filtered, and the solvent was removed under reduced pressure. The crude product was purified by column chromatography on deactivated silica gel (hexanes/EtOAc, 95:5) to give **35** (44.4 mg, 60% yield) as a colorless oil. The enantiomeric excess was determined by HPLC with a Chiralcel OD-H column (hexanes/2-propanol = 97:3, flow rate = 0.5 mL/min), *t_r* (1) = 40.4 min, *t_r* (2) = 47.1 min, $[\alpha]_D^{20} = +18.2$ (*c* = 0.032, CHCl₃, 99% ee); ¹H NMR (CDCl₃, 500 MHz): δ 2.27 (d, *J* = 5.0 Hz, 1H), 6.04 (d, *J* = 5.0 Hz, 1H), 6.44–6.45 (m, 1H), 6.97–6.99 (m, 1H), 7.01–7.02 (m, 1H), 7.30 (dd, *J* = 1.5, 5.0 Hz, 1H), 7.41 (t, *J* = 1.5 Hz, 1H), 7.45–7.46 (m, 1H); ¹³C{¹H} NMR (CDCl₃, 125 MHz): δ 65.8, 109.3, 125.0, 125.6, 126.9, 128.6, 140.1, 143.7, 147.4; IR (neat): 3410, 3108, 2924, 2855, 1759, 1672, 1614, 1507, 1416, 1264, 1230, 1156, 1022 cm⁻¹; HRMS calcd for C₉H₇O₂S (M-H)⁺: 179.0167, found 179.0169.

Asymmetric Addition/Diastereoselective Epoxidation Reactions. Preparation of (3,3-Dimethyloxiran-2-yl)(1-(triisopropylsilyl)-1H-indol-4-yl)methanol (41). A nitrogen purged Schlenk flask was charged with 4-bromo-1-TIPS indole (106.5 mg, 0.3 mmol) and

t-BuOMe (1 mL) and cooled to -78 °C. *t*-BuLi (0.18 mL, 1.7 M in pentane, 0.3 mmol) was added dropwise and the solution was stirred for 1 h. Solid EtZnCl (39.6 mg, 0.3 mmol) was added to the reaction solution at -78 °C. Toluene (3 mL) was next added giving a heterogeneous mixture. The solution was warmed at -10 °C and stirred for 3 h. TEEDA (26 μL, 0.12 mmol) was then added and the solution stirred for an additional 30 min. (-)-MIB (150 μL, 0.1 M solution in hexanes, 0.015 mmol) was added to the reaction flask, the solution was stirred for 5 min, and 3-methyl-2-butenal (14.7 μL, 0.15 mmol, dissolved in 1.5 mL of toluene) was added by syringe pump over 1.5 h. The reaction mixture was stirred at 0 °C and monitored by TLC until the starting aldehyde had been consumed. Upon completion of the asymmetric addition, ZnEt₂ (0.15 mL, 1 M in hexanes, 0.15 mmol) was added followed by TBHP (0.14 mL, 5.5 M in decane, 0.77 mmol). After stirring for 5 min Ti(O-*i*Pr)₄ (30 μL, 1 M in hexanes, 0.03 mmol) was added and the reaction mixture stirred until complete by TLC analysis (approximately 3 h). The reaction mixture was diluted with 3 mL EtOAc and quenched with water (5 mL). The organic layer was next separated and the aqueous solution extracted with EtOAc (3 × 10 mL). The combined organic layer was washed with brine (5 mL), dried over MgSO₄, filtered, and the volatile materials were removed under reduced pressure. The crude product was purified by column chromatography on deactivated silica gel (hexanes/EtOAc, 90:10) to give **41** (36.6 mg, 64.9% yield) as a clear oil. The diastereomeric ratio was determined by ¹H NMR analysis of the crude product (dr > 20:1); $[\alpha]_D^{20} = -6.1$ (*c* = 0.041, CHCl₃, 90% ee); ¹H NMR (CDCl₃, 500 MHz): δ 1.13 (d, *J* = 7.6 Hz, 18H), 1.29 (s, 3H), 1.43 (s, 3H), 1.69 (sept, *J* = 7.6 Hz, 3H), 2.48 (d, *J* = 3.0 Hz, 1H), 3.31 (d, *J* = 8.1 Hz, 1H), 4.89 (dd, *J* = 3.0, 8.1 Hz, 1H), 6.80 (d, *J* = 3.3 Hz, 1H), 7.10–7.14 (m, 1H), 7.30 (d, *J* = 3.3 Hz, 1H), 7.47 (dd, *J* = 2.8, 6.9 Hz, 1H); ¹³C{¹H} NMR (CDCl₃, 125 MHz): δ 12.8, 18.0, 19.5, 24.8, 60.1, 67.3, 72.2, 103.3, 113.9, 117.6, 121.1, 129.2, 131.5, 141.3; IR (neat): 3445, 3135, 3081, 3048, 2948, 2892, 2868, 2760, 2729, 2625, 2559, 2361, 2343, 2246, 2150, 2074, 1892, 1824, 1740, 1675, 1599, 1514, 1463, 1428, 1378, 1345, 1323, 1280, 1248, 1209, 1150, 1124, 1096, 1073 cm⁻¹; HRMS calcd for C₂₂H₃₅NO₂NaSi (M+Na)⁺: 396.2335, found 396.2321.

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Supporting Information Available: Complete experimental procedures, product characterization, and full author list for ref 93. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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