

Nickel/N-Heterocyclic Carbene Complex-Catalyzed Enantioselective Redox-Neutral Coupling of Benzyl Alcohols and Alkynes to Allylic **Alcohols**

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Supporting Information

ABSTRACT: The nickel-catalyzed enantioselective redox-neutral coupling of alcohols and alkynes to access chiral allylic alcohols is reported. The reaction proceeds via a hydrogen transfer process under ambient temperature, converting abundant feedstock alcohols and alkynes to chiral allylic alcohols with high stereoselectivities in one chemical step. Key to the success of this process was the development of a bulky chiral N-heterocyclic carbene, (R,R,R,R)-SIPE, a chiral version of SIPr, as the ligand for nickel. Notably, we found that the utilization of



P(OPh)₃ as secondary ligand for nickel was crucial to inhibit the isomerization of products.

KEYWORDS: asymmetric catalysis, nickel, ligand design, N-heterocyclic carbenes, transfer hydrogenation, redox-neutral coupling, alkynes, alcohols

nantioenriched allylic alcohols represent important chiral building blocks because of their synthetic versatility and the common occurrence of this substructure in a variety of pharmaceuticals and natural products. Consequently, the general asymmetric construction of chiral allylic alcohols is an important objective in organic synthesis.¹ Among the reported methods, the classical asymmetric addition of vinyl organometallics to carbonyl compounds² is a general method to synthesize chiral allylic alcohols, although vinyl organometallics need to be preformed, often through multistep procedures (Scheme 1A). Alternatively, the nickel-catalyzed asymmetric reductive coupling of aldehydes and alkynes³ has been advanced as a more efficient approach for the preparation of chiral allylic alcohols (Scheme 1B).⁴ While the direct use of widely available alkynes as coupling partner in lieu of a vinyl organometallic in this process is advantageous, stoichiometric organometallic reductants (e.g., ZnEt₂ or Et₃B) are still required.

Recently, we questioned whether an alcohol could serve as both a "greener" hydride source and also as a precursor for a transient aldehyde for the overall redox-neutral asymmetric coupling of alkynes and alcohols to form chiral allylic alcohols (Scheme 1C).⁵⁻⁷ If successful, such a direct asymmetric C-H alkenylation of alcohols via hydrohydroxyalkylation of alkynes represents a nearly ideal method to access chiral allylic alcohols given the abundance of both starting materials and excellent atom-, step-, and redox economy of the process.⁸ We noted that the Krische group,⁹ and the Matsubara group,¹⁰ has developed elegant alcohol-alkyne coupling reactions for the synthesis of achiral allylic alcohols using ruthenium and nickelbased catalysts, respectively. However, the methods for the

Scheme 1. Synthesis of Chiral Allylic Alcohols

(A) Metal-catalyzed asymmetric addition of vinyl organometallics to aldehydes

$$\begin{array}{c} O \\ R^{1} \underbrace{\downarrow}_{H} & R^{2} \underbrace{\swarrow}_{MX} & \underbrace{L^{*}M (cat.)}_{R^{1}} & O^{H} \\ M = Zn, Al \text{ etc.} & R^{1} \underbrace{\swarrow}_{R^{2}} \end{array}$$

(B) Ni-catalyzed asymmetric reductive coupling of aldehydes and alkynes

$$\begin{array}{c} O \\ R^{1} \\ H \\ H \\ H \\ H \\ H \\ R^{2} \\ \hline R^{3} \\ \hline R^{3} \\ \hline R^{3} \\ \hline R^{4} \\ \hline R^{2} \\ \hline R^{3} \\ \hline R^{2} \\ \hline R^{3} \\ \hline R^{2} \\ \hline R^{2}$$

(C) Ni-catalyzed asymmetric redox-neutral coupling of alcohols and alkynes



enantioselective C-H alkenylation of alcohols with alkynes to form chiral allylic alcohols, have previously not been reported.

Despite recent advances made in the field of asymmetric redox-neutral coupling of alcohols and unsaturated hydrocarbons,^{5–7} base metal-catalyzed enantioselective examples

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remain elusive.¹¹ One factor that has impeded the development of asymmetric base metal-catalyzed reactions is the lack of suitable chiral ligands with efficient control of both reactivity and selectivity.

Chiral N-heterocyclic carbenes (NHCs) have become increasingly common in asymmetric catalysis,¹² especially for nickel(0)-catalysis,^{3b,c,13} as a result of their modular preparation, steric and electronic tunability, and robustness as monodentate ligands for transition-metals. We recently developed a bulky chiral NHC ligand, namely, ANIPE (Figure 1, L1),¹⁴ which could be considered a chiral version of IPr,



Figure 1. NHCs designed and applied in this study.

whose design draws inspiration from Gawley's carbene (L2).¹⁵ The use of L1 enabled a highly enantioselective copper catalyzed hydroboration of α -olefins.¹⁴ Subsequently, Cramer reported a nickel-catalyzed enantioselective intramolecular alkylation of pyridone substrates based on a similar ligand system.^{13g} Herein, we describe the development of SIPE ligands (Figure 1), a new class of bulky chiral analogues of saturated NHCs, like SIPr. It is practically important that the precursor of SIPE ligands, air- and moisture-stable salts, could be easily prepared on multigram scale.¹⁶ The SIPE ligands were successfully applied to the nickel-catalyzed enantioselective redox-neutral coupling of alcohols and alkynes to form chiral allylic alcohols. To the best of our knowledge, this is the first example of base metal-catalyzed enantioselective transfer hydrogenative C–C bond forming reactions.

We began our study by examining the coupling of benzyl alcohol (1a) and 4-octyne (2a) to form chiral allylic alcohol 3a in the presence of 2 mol % Ni(cod)₂ and chiral ligand. Initially, a range of commonly used chiral phosphine and N-heterocyclic carbene ligands were tested, none of which provided the desired product 3a (see Supporting Information (SI)). The use of our previously disclosed ligand L1, however, gave encouraging results, providing product 3a in 68% yield and 92.5:7.5 e.r. with an E/Z selectivity of 88:12 (Table 1, entry 1). Although the use of L2 resulted in lower yield (entry 2), the use of L3, a newly design saturated NHC led to an improvement in the yield and selectivity (entry 3, E/Z 94:6, 95.5:4.5 e.r.). During the course of optimization, we found that partial isomerization¹⁷ of 3a occurred under the reaction conditions, which we ascribed to a sequential hydronickelation and β -H elimination process initiated by a nickel hydride species.¹⁸ Various Lewis basic additives, such as PPh₃, ethyl acrylate, Et₃N, and phosphites, were investigated as secondary ligands, whose coordination might stabilize the nickel catalyst and thus suppress isomerization events (entries 4-9). We were pleased to find that the electron-deficient phosphite ligand $P(OPh)_{3}^{19}$ in combination with L3 at room temperature delivered 3a in 76% isolated yield and 95.5:4.5 e.r., with a

Table 1. Ligand Screening and Reaction Optimization

OF Ph 1a 0.2 mn	ł + r nol	n-Pr 1 2 a 1.5 equiv	Ni(cod) ₂ (2.0 mol%) NHC/HCI (2.0 mol%) NaHMDS (4.0 mol%) additive (2.0 mol%) cyclohexane (0.2 M) rt, 24 h	Ph	H n-Pr 3a
entry	NHC	additive	yield (%) ^a	E/Z^{a}	e.r. ^b
1	L1		68	88:12	92.5:7.5
2	L2		42	88:12	91.5:8.5
3	L3		84	94:6	95.5:4.5
4	L3	PPh ₃	<2	nd	nd
5	L3	ethyl acrylate	68	93:7	nd
6	L3	dimethyl fumar	ate <2	nd	nd
7	L3	NEt ₃	60	90:10	nd
8	L3	$P(OEt)_3$	69	93:7	nd
9	L3	$P(OPh)_3$	82(76)	99:1	95.5:4.5
10	L4	$P(OPh)_3$	84	97:3	94.5:5.5
11	L5	$P(OPh)_3$	72	96:4	94.5:5.5
12	L6	$P(OPh)_3$	70	96:4	95:5
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^{*a*}Determined by NMR analysis with 1,1,2,2-tetrachloroethane as internal standard using crude sample; isolated yield shown in parentheses. ^{*b*}Determined by HPLC analysis with a chiral stationary phase.

dramatically improved E/Z ratio of 99:1 (entry 9). Interestingly, other members of the SIPE family of ligands exhibited slightly lower levels of selectivity (L4–L6, entries 10–12).

With our optimized conditions, we next set out to investigate the scope of this novel enantioselective alcoholalkyne coupling protocol (Table 2). First, various benzylic alcohols were coupled with 4-octyne. Both electron-rich and electron-poor benzylic alcohols were suitable, and electronic properties had little influence on the enantioselectivity, although electron-poor benzylic alcohols required a higher catalyst loading to achieve good conversion. Generally, chiral allylic alcohol products were obtained in moderate to high yields (52–86%) and high stereoselectivities (\geq 96:4 *E*/*Z*) and enantioselectivities (94:6-95.5:4.5 e.r. in most cases). Importantly, this mild protocol tolerated many functional groups, such as ethers (3e, 3i, 3j, and 3y) and silvl ethers (3z), a thioether (3q), a trifluoromethyl ether (3f), aryl halides (F, Cl) (3g and 3k), a benzofuran (3h), a trifluoromethyl (3m), an ester (3n), and a ketone (3o). It is noteworthy that an aryl chloride was compatible with reaction conditions (3k), considering well-established, mild protocols for nickelcatalyzed cross coupling of aryl chlorides.⁴ Remarkably, simple aliphatic alcohol (phenylpropanol) could also be employed in this transformation, providing the coupling product 3r with moder ate enantioselectivity, though with excellent control of double bond geometry.

Next, we examined the scope of alkynes coupling partner. In addition to 4-octyne, we found that symmetric internal alkynes, including 3-hexyne, 5-decyne, 6-dodecyne, and 7-tetradecyne, were all competent substrates, providing products in 94:6–96:4 e.r. and 74–92% yields (3s-3v). In the case of 2-butyne (the smallest internal alkyne), bulkier ligand L6 was employed, affording product 3w in 86.5:13.5 e.r.²⁰ Given that efficient enantiocontrol of a nickel-catalyzed reductive coupling of dialkyl internal alkynes and aldehydes is challenging, the enantioselectivities obtained in this novel protocol are considerably high.³ Moreover, the use of unsymmetric internal



^aYields of isolated products on 0.2 mmol scale. ^bUsing 5 mol % catalyst (x = 5). ^cUsing 10 mol % catalyst (x = 10). ^dtoluene as the solvent. ^eUsing 20 mol % catalyst (x = 20, without P(OPh)₃). ^fUsing L6/HCl as the ligand. ^g20 mol % benzaldehyde was added.

alkynes, 3-octyne and an aryl-substituted alkyne, gave products as regioisomeric mixtures with high stereoselectivities for both regioisomers (3x and 3x', 4a and 4a'). In the case of an alkynylarene, 20 mol % benzaldehyde was added to promote the initial oxa-nickelacycle formation. Finally, regioisomeric products (3y and 3y', 3z and 3z') from alkynes containing ethers could also be obtained in good yields and stereoselectivities. The absolute configuration of the product 3a was determined as the R-configuration by Mosher ester analysis (see SI).

A series of experiments were performed to get insight into the mechanism. First, a deuterium-labeling experiment using d_2 -1a in the model reaction gave 87% deuterium incorporation at the olefinic position of 3a (Figure 2A). On the basis of this result and previous observations,¹⁰ we proposed a catalytic cycle as outlined in Figure 2B: (1) Alcohol dehydrogenation initially generates the corresponding aldehyde. (2) A subsequent oxidative cyclization of aldehyde, alkyne and Ni(0) catalyst furnishes oxanickelacyle A,²¹ which is protonated by alcohol to afford acyclic vinylnickel intermediate **B**. (3) A subsequent β -H elimination of **B** provides nickel hydride complex **C** and regenerates the aldehyde. (4) Finally, reductive elimination affords the allylic alcohol product and regenerates the nickel catalyst. In support of this proposal, when d2-1**a** and nondeuterated 1**e** was used in a crossover reaction, partial incorporation at olefinic position was observed for both products, indicative of scrambling (Figure 2C). In contrast, no deuterium scrambling on benzylic position of d2-**3a** and d1-**3e** was observed, which suggest the dehydrogenation step of benzylic alcohols is probably irreversible and





Figure 2. Mechanistic studies and proposed catalytic circle

oxidative cyclization is fast. In addition, both intermolecular and intramolecular KIEs were observed (KIE = 2.1 or 2.0, Figure 2D). Moreover, the X-ray crystal structure of Ni(0)complex 5 was obtained. The use of 5 as catalyst gave similar results to the reaction using nickel catalyst generated in situ, consistent with monomeric nickel-L3 complex being the active catalyst (Figure 2E). Finally, an erosion in E/Z ratio of 3a was observed in the control experiment in the absence of P(OPh)₃ using 5 as catalyst, further confirm the important role of the additive (Figure 2E).

In summary, we have developed the first nickel-catalyzed enantioselective redox-neutral coupling of alcohols and alkynes, which also represents the first enantioselective base metal-catalyzed transfer hydrogenative C-C bond forming reaction. This mild protocol utilizes two feedstock chemicals and earth-abundant nickel catalysts to form chiral allylic alcohols with high stereoselectivities in one chemical step. Essential for the success of the process was the development of SIPE-type ligands and the utilization of a basic additive to suppress the product isomerization. Efforts to expand the scope of the reaction by further ligand design and exploration of wider application of these ligands are ongoing in our laboratory.

ASSOCIATED CONTENT

S Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acscatal.8b04198.

Experimental procedures, spectroscopic data, and NMR spectra of all products (PDF)

Crystallographic data for 5 (CCDC 1843290) (CIF)

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Notes

The authors declare the following competing financial interest(s): A patent has been filed.

Ph

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3a. 73%

E/Z = 98.2 95.5 e r

without P(OPh)₃: 92%, E/Z = 92:8, 95:5 e.r.

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(C) Deuterium scrambling experiment

-MeO)Ph

1e

(0.1 mmol)

2a

(0.3 mmol)

2a (0.3 mmol)

2a

(0.3 mmol)

99%D

(D) ¹H/²H KIE measurements

, 99%D

(E) Complex 5 as the catalyst

d2**-1a**

(0.1 mmol)

1a or d2-1a

(0.2 mmol)

d1-1a

(0.2 mmol)

1a

(0.2 mmol)

ОН

'standard

conditions

22

(0.3 mmol)

'standard conditions

 $k_H / k_D = 2.1 \pm 0.6$

'standard conditions

d1-3a : d1-3a' = 2:1

70%, 94.5:5.5 e.r.

5 (2.0 mol%)

P(OPh)₃ (2.0 mol%)

cyclohexane (0.2 M)

rt. 24 h

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DEDICATION

Dedicated to Prof. Xiyan Lu on the occasion of his 90th birthday.

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ОН

(p-MeO)Ph

0%D

d1-3e, 41%, 92:8 e.r.

3a or d2-3a

d1-3a

Ph Ph

5

42%C

d1-3a

он

d2-3a, 37%, 95:5 e.r

99%D

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