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# Facile benzofuran synthesis: Palladium-catalyzed carbonylative Suzuki coupling of methyl 2-(2-iodophenoxy)acetates under CO gas-free conditions



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# ABSTRACT

A palladium-catalyzed carbonylative Suzuki coupling of methyl 2-(2-iodophenoxy)acetates with arylboronic acids has been developed. The reactions were performed under CO gas-free conditions and the obtained products act as a direct precursor for the synthesis of highly functionalized benzofuran derivatives.

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Since the pioneering work by Heck and co-workers in the 1970s,<sup>1</sup> palladium-catalyzed carbonylation reaction has drawn increasing attention due to wide applications in both academic and industrial fields.<sup>2</sup> Carbonylation reactions represent a powerful method for the preparation of carbonyl-containing compounds, including ketones, aldehydes, carboxylic acids and their derivatives. Although significant progress has been made, some limitations still exist; for example reaction outcomes are typically worse with *ortho*-substituted substrates due to steric effects and carbon monoxide gas is usually needed which limits the application of the corresponding methods in laboratories.

On the other hand, diarylketones frequently appear in natural products, pharmaceuticals, and biologically active compounds.<sup>3</sup> Many synthetic methods have been developed for diarylketone synthesis, including the Friedel-Crafts acylation of arenes,<sup>4</sup> Fries-rearrangement,<sup>5</sup> and the acylation of benzoic acid derivatives with organometallic reagents.<sup>6</sup> Alternatively, the palladium-catalyzed carbonylative Suzuki coupling reaction offers another approach to construct diarylketones. Recently, we developed a practical and convenient synthetic method for diarylketone synthesis *via* the palladium-catalyzed carbonylative Suzuki coupling reactions

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of aryl halides with arylboronic acids, with formic acid as the CO source.<sup>7</sup> However, under these reaction conditions, dramatically decreased yields were obtained when a functional group was located at the *ortho* position of the aryl halides. Hence, this remains a challenge which needs to be solved.

Additionally, benzofuran derivatives have been applied in many areas and the development of new synthetic procedures is an attractive area of research.<sup>8</sup>

Herein, we report the continuation of our work on the preparation of diarylketones *via* palladium-catalyzed carbonylative Suzuki coupling reaction with formic acid as the CO source and the application of the obtained products as a direct precursor for the synthesis of highly functionalized benzofuran derivatives.

Initially, methyl 2-(2-iodophenoxy)acetate and phenyl boronic acid were used as model substrates, with formic acid as the CO source,  $Pd(OAc)_2$  as the catalyst,  $PPh_3$  as the ligand and  $K_2CO_3$  as the base, in toluene at 100 °C for 16 h. Gratifyingly, the desired product was obtained in 43% yield (Table 1, entry 1). We then examined the efficiency of various ligands. Monodentate ligands  $PCy_3$ and  $BuPAd_2$  gave higher yields, while Xphos resulted in a lower yield (Entries 2–4). Bidentate ligands DPPF and DPPE also provided the desired product in decreased yields (Entries 5–6). Next, various bases including  $Cs_2CO_3$ ,  $Ag_2CO_3$ ,  $Et_3N$ , DIPEA, and DBU were investigated (Entries 7–11), where  $Cs_2CO_3$  provided the best results. Solvent screening showed that toluene was the optimal solvent for



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Table 1

Screening of the reaction conditions.<sup>a</sup>



_	Entry	Catalyst	Ligand	Base	Solvent	Yield (%) <sup>b</sup>
	1	$Pd(OAc)_2$	$PPh_3$	K <sub>2</sub> CO <sub>3</sub>	Toluene	43
	2	$Pd(OAc)_2$	PCy <sub>3</sub>	K <sub>2</sub> CO <sub>3</sub>	Toluene	47
	3	$Pd(OAc)_2$	$BuPAd_2$	K <sub>2</sub> CO <sub>3</sub>	Toluene	55
	4	$Pd(OAc)_2$	XPhos	K <sub>2</sub> CO <sub>3</sub>	Toluene	13
	5	$Pd(OAc)_2$	DPPF	K <sub>2</sub> CO <sub>3</sub>	Toluene	37
	6	$Pd(OAc)_2$	DPPE	K <sub>2</sub> CO <sub>3</sub>	Toluene	34
	7	$Pd(OAc)_2$	$BuPAd_2$	$Cs_2CO_3$	Toluene	67
	8	$Pd(OAc)_2$	$BuPAd_2$	$Ag_2CO_3$	Toluene	23
	9	$Pd(OAc)_2$	$BuPAd_2$	Et <sub>3</sub> N	Toluene	61
	10	$Pd(OAc)_2$	$BuPAd_2$	DIPEA	Toluene	52
	11	$Pd(OAc)_2$	$BuPAd_2$	DBU	Toluene	47
	12	$Pd(OAc)_2$	$BuPAd_2$	$Cs_2CO_3$	DMSO	0
	13	$Pd(OAc)_2$	$BuPAd_2$	$Cs_2CO_3$	$CH_3CN$	Trace
	14	$Pd(OAc)_2$	$BuPAd_2$	$Cs_2CO_3$	DCE	63
	15	$Pd(OAc)_2$	$BuPAd_2$	$Cs_2CO_3$	1,4-Dioxane	55
	16	$Pd_2(dba)_2$	$BuPAd_2$	$Cs_2CO_3$	Toluene	53
	17	$PdCl_2(PPh_3)_4$	$BuPAd_2$	$Cs_2CO_3$	Toluene	0
	18	PdCl <sub>2</sub>	BuPAd <sub>2</sub>	$Cs_2CO_3$	Toluene	51
	19	$Pd(TFA)_2$	BuPAd <sub>2</sub>	$Cs_2CO_3$	Toluene	62

<sup>a</sup> Reagents and conditions: methyl 2-(2-iodophenoxy)acetate (0.5 mmol), phenyl boronic acid (1.0 mmol), catalyst (5.0 mol%), ligand (10.0 mol% for monodentate ligand; 5.0 mol% for bidentate ligand), base (1.5 mmol), formic acid (2.0 mmol), acetic anhydride (2.0 mmol), solvent (4.0 mL), 100 °C, 16 h.

<sup>b</sup> GC yield, with dodecane as an internal standard.

this transformation (Entries 12-15). Other palladium pre-catalysts, including Pd<sub>2</sub>(dba)<sub>3</sub>, PdCl<sub>2</sub>(PPh<sub>3</sub>)<sub>4</sub>, PdCl<sub>2</sub>, and Pd(TFA)<sub>2</sub> were also studied, however, no further improvement of the reaction outcome could be obtained (Entries 16-19).

With the best reaction condition in hand, we then turned our attention to examining the substrate scope.<sup>9</sup> As shown in Scheme 1, aryl boronic acids with a single electron-donating group gave the corresponding products in good yields (3ab-ae), while the 3,5dimethyl groups resulted in a lower yield (3af). Trifluoromethoxy substituted iodides gave the corresponding product in moderate yield (3ag). A moderate yield was also obtained with para-chloro substitution (3ah). Furthermore, 1-naphthyl and 2-naphthyl boronic acids were also studied; the corresponding products were isolated in 62% and 72% yield, respectively (3ai-aj). 3-Thiophenyl boronic acid was also tolerated and afforded the corresponding product in 50% yield (3ak). For aryl iodides, substitution at the meta-position to the iodo group with chloro and ester groups provided the corresponding products in 45% and 42% yields, respectively (3ba-ca).

Under our standard reaction conditions, we then explored a one-pot benzofuran synthesis reaction. After completion of the carbonylative Suzuki coupling reaction, sodium methoxide and methanol were added to the reaction mixture, affording benzofuran (4) in 40% yield after heating at 130 °C for 6 h.







Scheme 1. Substrate scope for the carbonylative Suzuki coupling reaction. Reagents and conditions: aryl iodides (0.5 mmol), aryl boronic acid (1.0 mmol),  $Pd(OAc)_2 \ \ (5.0\ mol\%), \ \ BuPAd_2 \ \ (10.0\ mol\%), \ \ Cs_2CO_3 \ \ (1.5\ mmol), \ \ formic \ \ acid$ (2.0 mmol), acetic anhydride (2.0 mmol), toluene (4.0 mL), 100 °C, 16 h, isolated vields.

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# A. Supplementary data

Supplementary data associated with this article can be found, in the online version, at http://dx.doi.org/10.1016/j.tetlet.2017.09. 063.

### References

- 1. (a) Heck RF. J Am Chem Soc. 1968;90:5546-5548;
- (b) Schoenberg A, Bartoletti I, Heck RF. J Org Chem. 1974;39:3318-3326; (c) Schoenberg A, Heck RF. J Org Chem. 1974;39:3327–3331.
- (a) For selected reviews, see: Sumino S, Fusano A, Fukuyama T, Ryu I. Acc Chem Res. 2014:47:1563-1574:
- (b) Friis SD, Lindhardt AT, Skrydstrup T. Acc Chem Res. 2016;49:594-605; (c) Wu X-F. RSC Adv. 2016;6:83831-83837;
- (d) Peng J-B, Qi X, Wu X-F. ChemSusChem. 2016;9:2279-2283;
- (e) Peng J-B, Qi X, Wu X-F. Synlett. 2017;28:175-194;
- (f) Gabriele B, Mancuso R, Salerno G. Eur J Org Chem. 2012;6825-6839. 3. (a) Jiang Y, Tu P. Chem Pharm Bull. 2005;53:1164-1166;
- (b) Nilar L-HD, Nguyen G, Venkatraman K-Y, Sim L, Harrison J. Phytochemistry.

2005;66:1718-1723;

- (c) Rancon S, Chaboud A, Darbour N, et al. *Phytochemistry*. 2001;57:553–557;
  (d) Ito H, Nishitani E, Konoshima T, Takasaki M, Kozuka M, Yoshida T. *Phytochemistry*. 2000:54:695–700:
- (e) Li J-C, Nohara T. Chem Pharm Bull. 2000;48:1354–1355.
- 4. (a) Song CE, Shim WH, Roh EJ, Choi JH. Chem Commun. 2000;1695–1696;
  - (b) Fürstner A, Voigtländer D, Schrader W, Giebel D, Reetz MT. Org Lett. 2001;3:417-420;
  - (c) Nara SJ, Harjani JR, Salunkhe MM. J Org Chem. 2001;66:8616-8620;
  - (d) Ross J, Xiao JL. Green Chem. 2002;4:129-133;
  - (e) Gmouh S, Yang H, Vaultier M. Org Lett. 2003;5:2219–2222;
  - (f) Fillion E, Fishlock D, Wilsily A, Goll JM. J Org Chem. 2005;70:1316–1327;
  - (g) Jang DO, Moon KS, Cho DH, Kim JG. Tetrahedron Lett. 2006;47:6063-6066;
  - (h) Sartori G, Maggi R. Chem Rev. 2006;106:1077-1104;
  - (i) Dzudza A, Marks TJ. J Org Chem. 2008;73:4004-4016.
- 5. (a) Bellus D, Hrdlovic P. Chem Rev. 1967;67:599–609;
- (b) Sibi MP, Snieckus V. J Org Chem. 1983;48:1935–1937.
- 6. (a) Labadie JW, Stille JK. J Am Chem Soc. 1983;105:6129-6137; (b) Hatanaka Y, Fukushima S, Hiyama T. Tetrahedron. 1992;48:2113-2126;
  - (c) Goossen LJ, Ghosh K. Angew Chem Int Ed. 2001;40:3458–3460;
  - (d) Huang YC, Majumdar KK, Cheng C-H. J Org Chem. 2002;67:1682-1684;
  - (e) Zapf A. Angew Chem Int Ed. 2003;42:5394-5399;
  - (f) Wang D, Zhang Z. Org Lett. 2003;5:4645–4648;
  - (g) Tatamidani H, Yokota K, Kakiuchi F, Chatani N. J Org Chem. 2004;69:5615–5621;

(h) Zhang Y, Rovis T. J Am Chem Soc. 2004;126:15964-15965;

(i) Silbestri GF, Bogel-Masson R, Lockhart MT, Chopa AB. J Organomet Chem. 2006;691:1520–1524;
(j) Jafarpour F, Rashidi-Ranjbar P, Kashani AO. Eur J Org Chem. 2011;2128–2132;

- (k) Sharma P, Rohilla S, Jain N. J Org Chem. 2017;82:1105–1113.
- 7. Qi X, Jiang L-B, Li H-P, Wu X-F. Chem Eur J. 2015;21:17650–17656.
- 8. (a) For selected examples, see: Mandali PK, Chand DK. Synthesis. 2015;1661-1668:
- (b) Varela-Fernández A, González-Rodríguez C, Varela JA, Castedo L, Saá C. Org Lett. 2009;11:5350–5353;
- (c) Wang S, Li P, Yu L, Wang L. Org Lett. 2011;13:5968-5971;
- (d) Wang J-R, Manabe K. J Org Chem. 2010;75:5340-5342;
- (e) Eidamshaus C, Burch JD. Org Lett. 2008;10:4211-4214;
- (f) Carril M, SanMartin R, Tellitu I, Dominguez E. Org Lett. 2006;8:1467–1470; (g) Anxionnat B, Pardo DG, Ricci G, Rossen K, Cossy J. Org Lett. 2013;15:3876–3879.
- 9. General procedure: Pd(OAc)<sub>2</sub> (5 mol%), BuPAd<sub>2</sub> (10 mol%), phenylboronic acid (1.0 mmol, 2 equiv) and Cs<sub>2</sub>CO<sub>3</sub> (1.5 mmol) were transferred into an oven-dried tube filled with nitrogen. Toluene (4.0 mL), acetic acid-2 (2-iodophenoxy)-methyl ester (0.5 mmol, 1 equiv) and Et<sub>3</sub>N (4 mmol) were added into the reaction tube. Then a cooled mixture of formic acid (2.0 mmol) and acetic anhydride (2.0 mmol), which was stirred for 1.0 h at 60 °C, was rapidly added to the reaction tube and the reaction mixture stirred for 14–16 h at 100 °C. Upon completion, the reaction mixture was concentrated and purified by silica gel column chromatography.