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Iron-Catalyzed Primary C–H Amination of Sulfamate Esters and Its Application in the Synthesis of Azetidines

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ummary of main observation and conclusion The direct amination of unactivated primary C–H bonds is extremely challenging due to their inert nature. Here we report an intramolecular primary C–H amination of sulfamate esters using an iron catalyst derived from iron(II) triflate and bipyridine. An array of cathiazinanes were synthesized in moderate to good yields, which were further converted into biologically important azetidines by a one-pot procedure. I his research demonstrates the potential of applying simple nitrogen ligands in iron-catalyzed C–H functionalization and offers an accessible alternative to state-of-the-art iron-nitrene chemistry.

Background and Originality Content

Nitrogen-containing heterocycles are omnipresent in many synthetic intermediates, natural products, and pharmaceuticals.^[1] n last several decades, considerable endeavors have been committed to develop efficient methods toward their synthesis.^[2] As one of the most straightforward and economical methods, direct C–H amination of ubiquitous hydrocarbons by means of tansition metal mediated nitrene insertion has drawn numerous attentions of the community.^[3,4] Significant advances have been achieved albeit primarily employing source-limited metals (i.e., nodium, iridium, and ruthenium).^[3g–3n] Recently, remarkable examples of non-precious metal-catalyzed nitrene insertion bactions^[4] have been emerged as a powerful C–N bond formation method by the contributions of Che,^[5] White,^[6] Zhang,^[7] Betley,^[8] and other.^[4,9]

To date, the nitrene insertion reactions of aliphatic C(sp³)–H onds are mainly limited to activated substrates such as allylic and benzylic C–H bonds.^[3,4] Amination of unactivated aliphatic C–H honds is of substantial challenge owing to their thermodynamic tability, in particular for primary C–H bonds with a bonddissociation energy (BDE) of 100.5 kcal/mol.^[10] Only a handful of examples capable of primary C–H amination have been known. river and co-workers realized an intramolecular reaction of aryl azides catalyzed by a thermally robust Rh(II) complex for the ynthesis of indolines (Scheme 1A).^[3i] In 2010, the Zhang group reported an excellent cobalt-catalyzed amination of the primary C– H bond of phosphoryl azides (Scheme 1B).^[7a] In 2013, the Betley eroup employed an iron(II)-dipyrrinato catalyst [(^{Ad}DP^{CIAr})FeCI(OEt₂)] to achieve the amination of a variety of C–H bonds, but exhibited poor reactivity toward primary C–H bonds (Scheme 1C).^[8b] Remarkably, a bis-NHC-stabilized iron(III) porphyrin catalyst, [Fe(TDCPP)(IMe)₂]I, was developed by Che et al. enabling the amination of inert aliphatic C–H bonds in good yields.^[5c] In 2017, amination of aryl azides was disclosed by the Plietker group using a nucleophilic iron catalyst Bu₄N[Fe(CO)₃(NO)] (Scheme 1D).^[11] Scheme 1 Transition-metal-catalyzed nitrene insertion of primary C–H

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bonds



Additionally, a notable manganese-catalyzed amination of lfamate esters under mild conditions was accomplished by the White group in 2015, and a wide range of substrates were tolerated ir excellent reactivity and selectivity (Scheme 1E).^[6b]

Utilizing abundant iron salts and readily available nitrogen ligands to generate the active iron catalysts in situ has advantages in practical synthesis due to their easy accessibility.^[12] Seminal is using the "iron salt + ligand" protocol were independently disclosed by the Che^[12a] and Chan^[12b] groups. In 2019, our group developed an iron/aminopyridine-catalyzed amination of aliphatic condary and tertiary C–H bonds.^[13] Encouraged by these results, we sought to further expand this simple iron system to realize more challenging primary C–H amination (Scheme 1F). Herein, we report is successful development of an intramolecular primary C–H bond amination of sulfamate esters using an iron catalyst derived from Fe(OTf)₂ and bipyridine. Furthermore, a variety of azetidines, conformationally rigid scaffold widely existed in potential drug candidates and other bio-important compounds,^[14] are accessed om the amination products.

Results and Discussion

We commenced our studies with 2-methyl-2-phenylpropyl sulfamate ester **1a** as a model substrate to optimize reaction

conditions (see Tables S1–S6 in the SI for details). The representative results were outlined in Table 1. Among the examined ligands (entries 1-7, and Table S1), tridentate ligands (L1-L3) generally showed inferior reactivity to bidentate ligands (L4-L7). Aminopyridine-type ligands, previously used in ironcatalyzed amination of sulfamate esters and sulfonamides,[12] resulted in poor reactivity. We were delighted to find that the use of bipyridine L4 offered 60% NMR yield of the desired oxathiazinane 2a (entry 4). Examination of 1,10-phenanthrolinetype ligands (L5–L7) revealed that dichloro-substituted L7 is more efficient (entry 7). Although slightly lower yield was obtained with L4 in comparison with L7 (60% vs 63%, entries 4 vs 7), we decided to use L4 as the ligand for further optimization due to its readily availability and low cost. A number of oxidants were tested, and PhI(OCOCF₃)₂ was found superior. Further examination of iron sources, solvents, and temperature (entries 11-15) revealed that the use of Fe(OTf)₂ at 100 °C in acetonitrile improved the isolation yield to 81%, which was identified as the optimal reaction conditions (entry 15). Control experiment without the use of ligand resulted in no product formation (entry 16).

Table 1 Condition optimizations



entry ^a	iron salt	ligand (x mol%)	oxidant	2a (%) ^b
1	Fe(ClO ₄) ₂	L1 (20)	PhI(OCOCF ₃) ₂	20
2	Fe(ClO ₄) ₂	L2 (20)	PhI(OCOCF₃)₂	11
3	Fe(ClO ₄) ₂	L3 (20)	PhI(OCOCF₃)₂	28
4	Fe(ClO ₄) ₂	L4 (30)	PhI(OCOCF ₃) ₂	60 (55) ^c
5	Fe(ClO ₄) ₂	L5 (30)	PhI(OCOCF ₃) ₂	52
6	Fe(ClO ₄) ₂	L6 (30)	PhI(OCOCF ₃) ₂	44
7	Fe(ClO ₄) ₂	L7 (30)	PhI(OCOCF ₃) ₂	63
8	Fe(ClO ₄) ₂	L4 (30)	PhI(OAc) ₂	22
9	Fe(ClO4) ₂	L4 (30)	PhI(DMM)	33
10	Fe(ClO ₄) ₂	L4 (30)	PhI(OPiv) ₂	21
11	FeCl ₂	L4 (30)	PhI(OCOCF ₃) ₂	33 ^c
12	Fe(acac) ₂	L4 (30)	PhI(OCOCF ₃) ₂	48 ^c
13	Fe(OAc) ₂	L4 (30)	PhI(OCOCF ₃) ₂	70 ^c
14	Fe(OTf) ₂	L4 (30)	PhI(OCOCF ₃) ₂	73 ^c
15 ^d	Fe(OTf)₂	L4 (30)	PhI(OCOCF₃)₂	81 ^c
16 ^e	Fe(OTf) ₂	_	PhI(OCOCF ₃) ₂	N.D.

^oReaction conditions: **1a** (0.2 mmol), iron salt (0.02mmol, 10 mol%), oxidant (0.4 mmol), and 4 Å MS (50 mg) in MeCN (2 mL) at 80 °C for 3 h. ^bNMR yield

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Chin. J. Chem. 2019, 37, XXX-XXX

using 1,3,5-trimethoxybenzene as an internal standard. Isolated yield. ^{*d*}100 °C. PhI(DMM), phenyliodonium dimethylmalonate.^[15] ^{*e*}Without ligand. N.D., not detected.

Table 2 Substrate scope^a



Reaction conditions: **1** (0.4 mmol), Fe(OTf)₂ (0.04 mmol, 10 mol%), **L4** (0.12 mmol, 30 mol%), PhI(OCOCF₃)₂ (0.8 mmol), and 4 Å MS (100 mg) in MeCN (4 mL) at 100 °C for 3 h. b 0.2 mmol scale.

Next, the substrate scope was explored (Table 2). Substrates (1b-1f) bearing electron-deficient groups, such as F, Cl, and Br, on the aryl ring generated the corresponding oxathiazinanes 2b-2f in J5–82% yields. Substrate **1g** containing an α -naphthyl group, provided the corresponding product 2g in 59% yield. para-Isobutylnd para-methyl-substituted 2-aryl-2-methylpropyl sulfamates (2h, 2i) were also investigated and resulted in moderate yields (44-Down, Previously, synthesis of oxathiazinane **2j** from the corresponding starting material 1j catalyzed by an iron hthalocyanine catalyst ([FePc]·SbF₆) resulted in only 3% yield.^[6b] An significant improvement (64% yield) was achieved by the use of manganese *tert*-butylphthalocyanine catalvst [Mn(^tBuPc)] SbF₆).^[6b] By contrast, our chemistry afforded **2j** in 82% yield under the standard reaction conditions. Additionally, an ester-containing substrate 1k was also tolerated delivering the xathiazinane 2k in 53% yield. Secondary alcohol-derived sulfamate substrate 1I was compatible to form the corresponding roduct 21 in 44% yield. A sulfonamide substrate was also subjected to the standard reaction conditions, but failed to afford any desired product (Scheme S1 in the SI).

Scheme 2 Investigation of site selectivity



Systematic studies by White and Che disclosed that the reactivity trends of amination using both heme-like^[6a] and nonheme^[5b] iron catalysts are in accordance to the C-H bond dissociation energies. The selectivity of this method toward sulfamate esters with multiple C-H bond types were also studied (Scheme 2). The amination of substrate 1m, bearing both propargylic and primary C-H bonds, occurred at the propargylic site exclusively to provide 2m in moderate yield (Scheme 2A). Treatment of substrate 1n under the standard conditions delivered a 5:2 mixture of secondary C-H versus primary C-H amination products (Scheme 2B). Although sterically hindered, tertiary C-H amination is much more preferred over primary C-H amination, as showcasing by the reactions of substrates 10 and 1p (Scheme 2C). Moreover, with substrate 1q, six-membered ring oxathiazinae 2q is formed preferentially over the five-membered ring product 2q' (Scheme 2D). Those trends of selectivity, intrinsically driven by the substrates themselves, are in line with previous observations.[5b,6a]

According to the literature precedent^[5b,c,8b], a possible reaction pathway was proposed (Scheme 3). Treatment of substrate **1** with PhI(OCOCF3)₂ generates compound **II**, which upon reacts with the iron catalyst I leading to the formation of an imidoiron species **III** together with iodobenzene. Then direct C–H insertion or H-atom abstraction/radical recombination^[5c,8b] of **III** yields the desired product **2** and regenerates the catalyst.

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Scheme 3 Proposed reaction pathway



As showcased in literature, oxathiazinanes are valuable intermediates for heterocycles, [16a] 1,3-amino ethers, [16b] and other oactive targets.^[16] To further demonstrate the utility of this ironcatalyzed amination reaction, a one-pot procedure was used to nvert oxathiazinanes synthesized above to azetidines, a structurally and biologically important class of N-heterocyclic mpounds.^[14] Protection of the amination products **2** with CbzCl afforded N-Cbz oxathiazinanes 3. Without chromatographic purification of the intermediates 3, NaI was added followed by NaH generate azetidines 4 through sequential ring-opening/ringclosing steps.^[17] As shown in Table 3, the azetidines 4a-4m were nthesized in 50-79% yields by this one-pot procedure. Delightedly, 3,3-dimethylazetidine 4j was also generated in 51% vield. Notably, 3,3-dimethyl-azetidine is an important pattern in a number of biologically active molecules, such as an dopamine receptor antagonist zetidoline,^[14c] a PDE4 inhibitor,^[14d] and an acetylcholine receptor agonist^[14e] as shown in Figure 1.

Table 3 One-pot synthesis of azetidines from oxathiazinanes^a



^{α}Reaction conditions: **2** (0.1 mmol), CbzCl (0.2 mmol), NaH (0.2 mmol in THF (1 mL) at room temperature for 30 min; NaI (0.15 mmol) in DMF (1 mL) at 60 °C for 1 h, then NaH (0.3 mmol) at 40 °C for 3 h. ^bO.2 mmol.



Figure 1 Selected examples of biologically relevant azetidines.

Conclusions

In summary, we have developed an iron-catalyzed intramolecular primary C–H amination of sulfamate esters. A range of oxathiazinanes were produced in moderate to good yields using an iron catalyst derived from $Fe(OTf)_2$ and bipyridine. The synthetic utility of the amination reaction to construct azetidines is demonstrated by a one-pot procedure. Given the readily availability and high reactivity of the catalyst toward aliphatic substrates, this method offers a practical alternative to heterocycle synthesis.

Experimental

General procedure for Fe-catalyzed C–H amination.

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Chin. J. Chem. 2019, 37, XXX-XXX

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To a 10 mL vial equipped with a magnetic stirring bar were added Fe(OTf)₂ (14.2 mg, 0.04 mmol, 10 mol%), **L4** (18.7 mg, 0.12 mmol, 30 mol%), and 2 mL of MeCN. After the mixture was stirred at 30 °C for 30 min, substrate **1** (0.4 mmol), PhI(OCOCF₃)₂ (344.0 ng, 0.8 mmol, 2 equiv), 4 Å MS (100 mg), and another 2 mL of MeCN were added. Then the vial was sealed and the mixture was stirred at 100 °C for 3 h. The reaction was cooled to room temperature, filtered through a pad of celite, and washed with CH₂Cl₂ (3 × 5 mL). The filtrate was concentrated under reduce ressure, and the residue was purified by flash chromatography on silica gel (200 ~ 300 mesh) to give the desired products **2**.

Supporting Information

The supporting information for this article is available on the WWW under https://doi.org/10.1002/cjoc.2018xxxxx.

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