

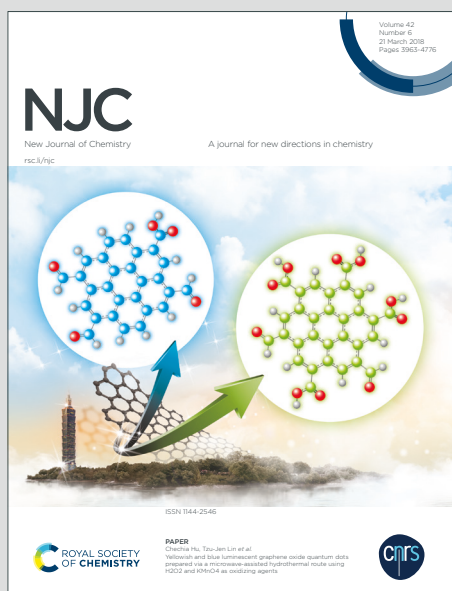
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Very thermostable energetic materials based on a fused-triazole: 3,6-diamino-1H-[1,2,4]triazolo[4,3-b][1,2,4]triazole

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Abstract: 3,6-Diamino-1H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (**1**) and its energetic salts (**2-9**) were designed and synthesized based on a fused-triazole backbone with two C-amino groups as substituents. Their physicochemical and energetic properties were measured or calculated. Among them, compound **1** exhibits superior thermostability ($T_{d(\text{onset})}$: 261 °C), surpassing its analogues 3,7-diamino-7H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (DATT, 219 °C) and 3,6,7-triamino-7H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (TATOT, 245 °C). The differences in thermal stabilities were further investigated by determining the lowest bond dissociation energies (BDE) where a positive correlation between the stability of the molecules and the lowest BDE values is observed. The results show that **1** with the highest value for the lowest BDE has a superior thermostability in comparison to DATT and TATOT. The energetic salts (**2-9**) also exhibit remarkable thermal stabilities as well as low impact and friction sensitivities. The fused-triazole backbone 1H-[1,2,4]triazolo[4,3-b][1,2,4]triazole with two C-amino groups as substituents is shown to be a promising building block for construction of very thermally stable energetic materials.

Introduction

Nitrogen rich five-membered heterocycles have become an important option for designing and synthesizing energetic materials owing to their high positive heats of formation and high densities.¹⁻⁴ Heterocyclic rings including furazan, pyrazole, triazole, tetrazole and their derivatives provide varied backbones for the construction of energetic molecules.^{5,6} Among them, 1H-1,2,4-triazole has been shown to be a promising building block in the design of high performance energetic materials. The reported energetic molecules based on 1H-1,2,4-triazole exhibit high densities and remarkable thermal stabilities.⁷⁻¹¹ Nineteen different salts based on 3,5-dinitro-1,2,4-triazolates that show low sensitivities and high thermal stabilities were synthesized.¹² 3-Nitro-1,2,4-triazole-5-one (NTO) is another powerful insensitive triazole-based explosive, whose detonation performance is comparable to RDX and HMX.¹³⁻¹⁵

Relative to a single triazole ring, fused heterocyclic rings have higher nitrogen content and better thermostability. The numerous C-N bonds and N-N bonds in fused heterocycles contribute markedly to the powerful storage capacity of energetic materials.¹⁶ 3,6,7-Triamino-7H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (TATOT) is representative of a compound based on a fused triazole backbone.^{17,18} TATOT and its ionic derivatives exhibit good thermal stability and high detonation

performance. In 2017, a combination of the TATOT cation with 5-dinitromethyl-3-trinitromethyl-1H-1,2,4-triazolate resulted in a high-energy salt with a density of 1.90 g cm⁻³ and an outstanding detonation performance.¹⁹ 3,7-Diamino-7H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (DATT) and its energetic salts have also been reported as fused heterocyclic compounds with high energy densities.⁴

The amino group plays an important role in enhancing sensitivity and improving thermal stability via formation of inter-/intra-molecular hydrogen bonds.²⁰ Therefore, 3,6-diamino-1H-[1,2,4]triazolo[4,3-b][1,2,4]triazole (**1**) was designed with a fused triazole as the backbone and two C-amino groups as substituents for study in this work (Fig. 1). Compound **1** has a backbone similar to TATOT and DATT, but with a different number (TATOT) and positions (DATT) of the amino groups. It is interesting to note that the latter two characteristics make a major difference in the thermostability of energetic molecules. The energetic properties of the monocationic and dicationic salts (**2-9**) based on **1** have also been investigated.

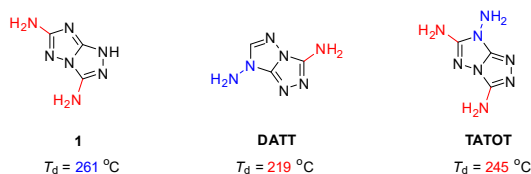
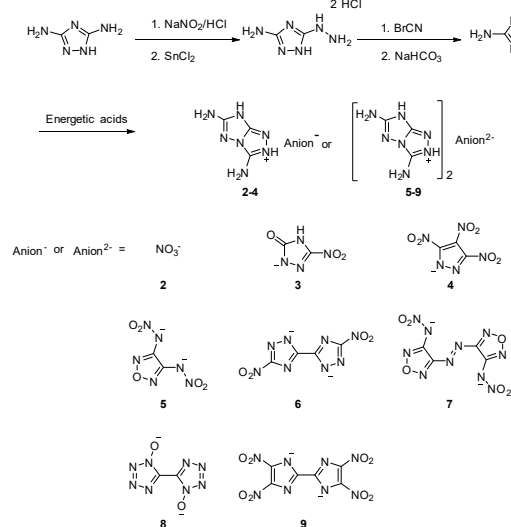


Fig.1 The structures and decomposition temperatures of energetic compounds (**1**, DATT and TATOT) based on [1,2,4]triazolo[4,3-b][1,2,4]triazole.

Results and discussion

Synthesis

The synthetic route is shown in Scheme 1. 3-Amino-5-hydrazino-1,2,4-triazole dihydrochloride was prepared by reacting 1,3-diamino-1,2,4-triazole with sodium nitrite in dilute HCl and followed by adding SnCl₂ in concentrated HCl solution. Then 3-amino-5-hydrazino-1,2,4-triazole dihydrochloride was dissolved in dilute hydrochloric acid (2 M) and cyanogen bromide was added. The reaction mixture was heated to 60 °C and stirred at reflux for 2 h. The reaction mixture was concentrated with air and treated with a saturated NaHCO₃ solution to give crude product **1**, which was purified further by recrystallization in hot water. The monocationic salts (**2-4**) and dicationic salts (**5-9**) were synthesized by reacting **1** with various energetic acids at 60 °C for 2 h in water.



Scheme 1. Synthesis of **1** and its energetic salts (**2-9**).

¹⁵N NMR spectroscopy

The ¹⁵N NMR spectra (Fig. 2) of **1**, **2**, and **5** were determined in *d*₆-DMSO. In the spectrum of **1**, there are seven signals, the nitrogen atom signals of the two amino groups are observed at the highest field ($\delta = -337.9$ ppm (N6) and $\delta = -328.0$ ppm (N7)). The fused triazole backbone has signals at $\delta = -211.4$ (N1), $\delta = -170.3$ (N2), $\delta = -180.8$ (N3), $\delta = -156.2$ (N4) and $\delta = -246.2$ ppm (N5). In the spectra of **2** and **5**, due to the interactions with different anions, the nitrogen signals of the amino groups on the cation are shifted to lower field ($\delta = -326.9$ ppm (N6); $\delta = -324.0$ ppm (N7)). There are only six signals for **2**, the signals of N4 and N1 could not be detected due to rapid proton exchange. The chemical shift of the nitrogen atom in the nitrate ion (N8) is

found at $\delta = -3.8$ ppm. Due to the symmetric structure of nitramino-furazanate in **5**, three signals were found at $\delta = 12.1$ ppm (N8); $\delta = -21.4$ ppm (N10); $\delta = -176.1$ ppm (N9) for the anion. The two signals for nitrogen atoms in amino groups were observed at high field. However, only one signal was detected for the fused-triazole backbone. When two drops of D₂O were added (bottom of Fig. 2), one more signal ($\delta = -189.6$ ppm (N3)) was observed although the others in the fused-triazole backbone are still absent.

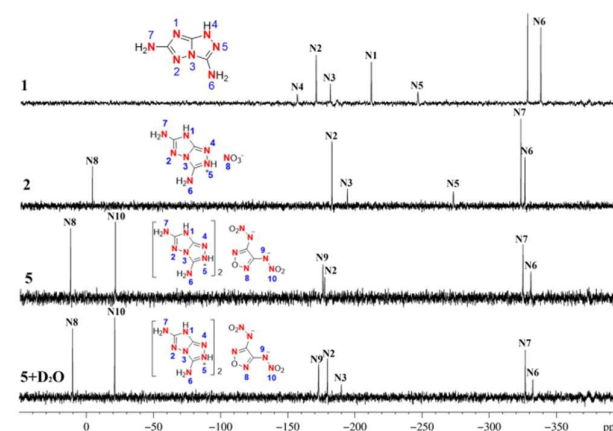


Fig. 2 ¹⁵N NMR spectra of **1**, **2**, **5** and **5+D₂O**.

Crystal structures

Compound **1** crystallizes in the monoclinic space group P2₁/n with the density of 1.766 g cm⁻³ at 99.9 K. The crystal structure of **1** is shown in Fig. 3a. All the atoms are essentially in the same plane. The torsion angles are N6-C2-N2-N1 = 176.55°; N2-N1-C1-N5 = -178.44°; N1-C1-N5-C3 = 179.87°; C3-N4-N3-C2 = -177.63°; N1-C1-N3-N4 = -179.62°. The C-N bond lengths of two C-amino groups are 1.348 Å (C2-N6) and 1.373 Å (C3-N7), which are shorter than that (C3-N5 = 1.377 Å) in the fused triazole backbone. Additionally, there are several hydrogen bonds between the molecules in the crystal structure (Fig. 3b). Both C-amino groups participate in hydrogen bonding (N6-H6A...N2 2.198 Å, N6-H6B...N4 2.040 Å).

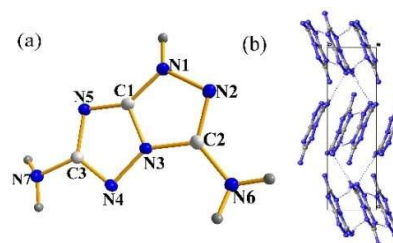


Fig. 3 (a) Molecular structure of **1**. (b) Packing diagram of **1** along the *a* axis.

Compound **4H₂O** (Fig. 4a) crystallizes as a monoclinic in the monoclinic space group P2₁/n at 100 K, with a calculated density of 1.794 g cm⁻³. The bond lengths of the two C-amino groups are N6-C1 = 1.328 Å and C2-N7 = 1.338 Å, respectively. They are shorter

than those in **1** owing to the introduction of an anion. The packing diagram of **4**·H₂O is shown in Fig. 4b viewed along the *a* axis, which is strongly influenced by strong hydrogen bonds involving intermolecular and intramolecular interactions (N6–H6B···O6, N4–H4···N8 and O1W–H1WB···N2). The others are given in the Supporting Information.

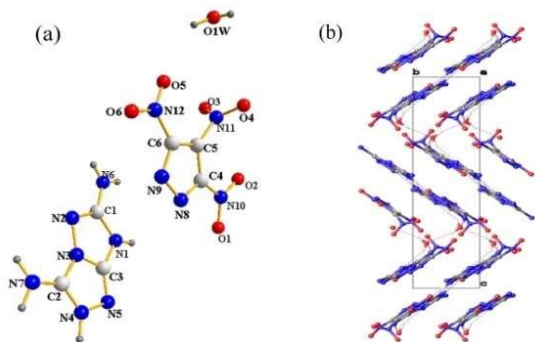


Fig.4 (a) Molecular structure of **4**·H₂O. (b) Packing diagram of **4**·H₂O along the *a* axis.

Physicochemical and energetic properties

The physicochemical properties of **1-9** are summarized in Table

1. All the compounds show good densities between 1.68 to 1.78 g cm⁻³; among them, **2** and **9** have a density of 1.78 g cm⁻³, which is comparable to RDX (1.80 g cm⁻³). All compounds have densities greater than DATT (1.71 g cm⁻³), except **7**. The values of the heats of formation were calculated by using the Gaussian 09 suite of programs.²¹ Compounds **5-8** have high positive heats of formation. Especially for **8**, it has the highest positive heat of formation at 1246.6 kJ mol⁻¹. The heats of formation of all the new molecules exceed DATT (480.3 kJ mol⁻¹) and RDX (70.3 kJ mol⁻¹).

Thermal stability is one of the key properties for energetic materials, considering their practical applications.^{22,23} The thermal behavior of these compounds was determined by differential scanning calorimetry (DSC) at a heating rate of 5 °C min⁻¹. Most of the compounds have very high thermal decomposition temperatures (onset), ranging from 202 (**4**) to 293 °C (**8**). Generally, the decomposition temperatures of these salts are higher than those of RDX (204 °C) and DATT (219 °C). The combination of 3,6-diamino-fused triazole cations and anions increases a large number of intermolecular and intramolecular interactions improving the thermal stabilities of these energetic salts.

Table 1. Physicochemical and energetic properties of **1-9** compared with RDX, TATOT and DATT.

Comp	$\rho^{[a]}$ (g cm ⁻³)	$D_v^{[b]}$ (m s ⁻¹)	$P^{[c]}$ (GPa)	$\Delta H_f^{[d]}$ (kJ mol ⁻¹ /kJ g ⁻¹)	$T_{dec}^{[e]}$ (°C)	$IS^{[f]}$ (J)	$FS^{[g]}$ (N)
1	1.71	7891	21.0	302.6/2.18	261	>40	>360
2	1.78	8531	28.1	167.3/0.83	219	15	240
3	1.76	8101	24.5	291.0/1.08	232	>40	>360
4	1.74	8695	32.1	810.4/2.25	202	>40	>360
5	1.72	8421	27.1	1128.7/2.41	238	20	>360
6	1.72	8050	23.9	1088.9/2.16	277	36	>360
7	1.68	8215	25.6	1124.8/1.87	163	40	>360
8	1.75	8591	27.7	1426.6/3.18	293	20	>360
9	1.78	8237	26.6	915.1/1.54	272	>40	>360
RDX	1.80	8795	34.9	70.3/0.32	204	7.5	120
DATT ^[h]	1.71	8461	25.4	480.3/3.45	219	40	360
TATOT ^[i]	1.72	9385	29.7	446.7/2.90	245	40	360

[a] Density measured by a gas pycnometer at 25 °C; [b] Calculated detonation velocity; [c] Calculated detonation pressure; [d] Calculated molar enthalpy of formation in solid state; [e] Temperature of decomposition (onset); [f] Impact sensitivity; [g] Friction sensitivity; [h] Ref 5; [i] Ref 16.

Based on the similar fused triazole as the skeleton, **1**, DATT and TATOT have different decomposition temperatures. Compound **1** begins to decompose at 261 °C, which is higher than DATT (219 °C) and TATOT (245 °C). These differences encouraged us to explore the structure further. The thermal stabilities of the energetic materials are closely related to the initial stages of the first decomposition, which can be deduced on the basis of bond dissociation energies (BDE).²⁴ The BDE levels were obtained by using the Gaussian 09 program. Possible initial pyrolysis initiation bonds were considered by breaking the following bonds: two C-amino groups in **1**, C-amino and N-amino groups in DATT and two C-amino and an N-amino groups in TATOT. The dissociation energies of the corresponding bonds are shown in Fig.5. The lowest BDE of **1** (Fig. 5a) is 116.5 kJ mol⁻¹ (C-NH₂), while for DATT (Fig. 5b) and TATOT (Fig. 5c), the values are 23.5 kJ mol⁻¹ (N-NH₂) and 70.8 kJ mol⁻¹ (N-NH₂), respectively. The BDE values of N-amino groups are lower than C-amino groups in this system. Considering their structural similarities, the differences in the thermal stabilities should originate from the presence of N-amino groups on DATT and TATOT. The molecule having the higher value of the lowest BDE, tends to be more stable thermally. Compound **1** with the largest value of the lowest BDE (116.5 kJ mol⁻¹) has the best thermal stability (*T*_d: 261 °C), and DATT, with the smallest value of lowest BDE (23.5 kJ mol⁻¹) exhibits the worst thermal stability (*T*_d: 219 °C).

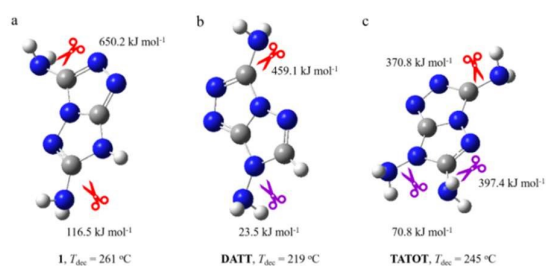


Fig. 5 The bond dissociation energies (BDE) of (a) **1**, (b) DATT⁵ and (c) TATOT¹⁸.

The detonation pressures (*P*), and velocities (*D*_v) were calculated by using the EXPLO5 (v6.01) program, based on the calculated heats of formation and measured densities.²⁵ The detonation performances are shown in Table 1. The calculated detonation velocities range from 7891 m s⁻¹ to 8695 m s⁻¹ which are superior to TNT (6881 m s⁻¹) and are comparable to RDX (8795 m s⁻¹). The calculated detonation pressures lie in the range from 21.0 GPa to 32.1 GPa, which are higher than TNT (19.5 GPa). Among these molecules, **4** demonstrates outstanding

detonation velocity (8695 m s⁻¹) and superior detonation pressure (32.1 Pa).

A fine balance between low sensitivity and high detonation performance is highly desirable for energetic materials.²⁶ The mechanical stimuli were measured by using friction tester and standard BAM drop hammer techniques.²⁷ The impact sensitivities (*IS*) and friction sensitivities (*FS*) of all compounds are listed in Table 1. All of the energetic salts are insensitive to both impact and friction stimuli, with impact sensitivity over 15 J and friction sensitivity up to 360 N (except **2**). The insensitive properties of **1-9**, together with their good thermal stabilities and comparable detonation properties, highlight their application potential as energetic materials.

Conclusions

A series of energetic molecules were designed and synthesized based on heterocyclic triazole-fused skeletons. The physical and chemical properties, including density, decomposition temperature, friction sensitivity, impact sensitivity as well as detonation performance were investigated. Except for **7**, the densities of compounds **1-9** are above 1.7 g·cm⁻³; the densities of **2** and **9** are comparable to that of RDX. Compound **4** has the best detonation performance with a detonation velocity of 8695 m s⁻¹ and a detonation pressure of 32.1 GPa. All compounds are highly insensitive to both friction and impact stimuli. The series of compounds have relatively high thermal stabilities with the highest decomposition temperature at 293 °C (**8**). BDE investigation shows that **1** has the higher value of the lowest BDE value and has a better thermal stability than its analogues DATT and TATOT. These molecules exhibit promising properties and may find applications as energetic materials especially as heat-resistant explosives due to their superior thermal stabilities.

Experimental Section

Caution! The compounds in this work are potentially energetic materials that could explode under certain conditions (e.g., impact, friction, or electric discharge). Appropriate safety precautions, such as the use of safety shields in a fume hood and personal protection equipment (safety glasses, face shields, earplugs, as well as gloves) should always be observed when handling these materials.

General methods

All reagents were purchased from AKSci or Alfa Aesar in analytical grade and were used as supplied. The ¹H and ¹³C spectra were recorded on a 300 MHz (Bruker AVANCE 300) nuclear magnetic resonance spectrometer. Chemical shifts for ¹H and ¹³C NMR spectra are given with respect to external (CH₃)₄Si. [D₆] DMSO was used as a locking solvent unless otherwise stated. The ¹⁵N spectra

were recorded on a 500 MHz (Bruker AVANCE 500) nuclear magnetic resonance spectrometer. IR spectra were recorded using KBr pellets with a FT-IR spectrometer (Thermo Nicolet AVATAR 370). Density was determined at room temperature by employing a Micromeritics AccuPyc II 1340 gas pycnometer. Thermal decomposition (onset) points were recorded using a dry nitrogen gas purge and a heating rate of 5 °C min⁻¹ on a differential scanning calorimeter (DSC, TA Instruments Q2000). Elemental analyses (C, H, N) were performed with a Vario Micro cube Elementar Analyser. Impact and friction sensitivity measurements were made using a standard BAM Fallhammer and a BAM friction tester.

Computational Methods

The gas phase enthalpies of formation for **1** was calculated based on G2 method. The solid-state heat of formation was calculated with Trouton's rule according to equation (1) (*T* represents either the melting point or the decomposition temperature when no melting occurs prior to decomposition).²⁸

$$\Delta H_{\text{sub}} = 188 / \text{Jmol}^{-1} \text{K}^{-1} \times T \quad (1)$$

For energetic salts, the solid phase enthalpy of formation is obtained using a Born-Haber energy cycle.²⁹ For compounds which are hydrates (**4**·H₂O and **7**·2H₂O), the solid-phase enthalpy of formation is obtained by adding the gas phase heat of formation of the anhydrous compound to that of water (-241.8 kJ mol⁻¹).³⁰

Crystal Structure Analysis

Single colorless needle crystals (**1**) of dimensions 0.13 × 0.13 × 0.10 mm³, and single yellow needle crystals (**4**·H₂O) of dimensions 0.33 × 0.05 × 0.02 mm³ were selected and mounted on a nylon loop with paratone oil on a XtaLAB Synergy, Dualflex, HyPix diffractometer. The crystals were kept at a steady *T* = 99.9(3) K and *T* = 100.00(10) K during data collection, respectively. The structures were solved with the ShelXT³¹ solution program using dual methods and by using Olex2³² as the graphical interface. The model were refined with ShelXL 2018/3³³ using full matrix least squares minimisation on *F*².

Synthesis

Synthesis of **1**

3-Hydrazinium-5-amino-1*H*-1,2,4-triazolium dichloride³⁴⁻³⁶ (1.87 g, 10.0 mmol) was dissolved in dilute hydrochloric acid (25 mL, 2 M), cyanogen bromide (1.50 g, 14.0 mmol) was added, and the solution was heated with stirring at 60 °C for 2 h. The solution was held at reflux for an additional 2 h. After cooling to 25 °C, the solution was concentrated to give a white precipitate

which was collected by filtration. The white solid was dissolved in distilled water (15 mL). Sodium bicarbonate (0.84 g, 10.0 mmol) was added, and after stirring 15 min at 25 °C, the white solid (1.08 g, 78% yield) was filtered, washed with water (15 mL × 2) and dried in air.

1: White solid. *T*_d (onset): 247 °C. ¹H NMR (*d*₆-DMSO): δ 11.50 (br), 6.02 (s, 2H), 5.45 (s, 2H) ppm. ¹³C NMR (*d*₆-DMSO): δ 169.7, 155.3, 141.3 ppm. ¹⁵N NMR (*d*₆-DMSO): δ -156.2, -170.3, -180.8, -211.4, -246.2, -328.0, -337.9 ppm. IR (KBr): $\tilde{\nu}$ = 3414, 3333, 3150, 1685, 1647, 1617, 1587, 1540, 1491, 1437, 1357, 1255, 1184, 1133, 1088, 1012, 847, 751, 678, 626 cm⁻¹. C₃H₅N₇ (139.12): Calcd: C 25.90, H 3.62, N 70.48 %. Found: C 25.57, H 3.58, N 69.28 %.

General procedure for synthesis of **2-4**

The energetic acid (2.0 mmol) was dissolved in water (10 mL), then **1** (0.28 g, 2.0 mmol) was added while stirring. The reaction mixture was heated to 60 °C and stirred for 2 h. The precipitate formed, was collected by filtration, and washed with water (5 mL × 2), and ethanol (5 mL × 2) and dried in vacuum to give the products **2-4**.

2: Yellow solid. *T*_d (onset): 219 °C. ¹H NMR (*d*₆-DMSO): δ 12.40 (br), 8.02 (br, 2H), 7.05 (br, 2H) ppm. ¹³C NMR (*d*₆-DMSO): δ 159.7, 148.2, 140.6 ppm. IR (KBr): $\tilde{\nu}$ = 3238, 3157, 1690, 1637, 1569, 1527, 1460, 1360, 1177, 1047, 988, 841, 818, 680 cm⁻¹. C₃H₆N₈O₃ (202.13): Calcd: C 17.83, H 2.99, N 55.44 %. Found: C 17.86, H 3.13, N 55.52 %.

3: White solid. *T*_m: 185 °C. *T*_d (onset): 232 °C. ¹H NMR (*d*₆-DMSO): δ 6.22 (s, 2H), 5.92 (br, 2H) ppm. ¹³C NMR (*d*₆-DMSO): δ 168.9, 155.2, 154.7, 148.9, 141.2 ppm. IR (KBr): $\tilde{\nu}$ = 3563, 3477, 1674, 1507, 1391, 1311, 1175, 1120, 1055, 1022, 976, 839, 782, 711, 701, 682, 632, 606 cm⁻¹. C₅H₇N₁₁O₃ (269.19): Calcd: C 22.51, H 2.62, N 57.24 %. Found: C 22.41, H 2.71, N 57.51 %.

4·H₂O: White solid. *T*_d (onset): 222 °C. δ ¹H NMR (*d*₆-DMSO): δ 13.10 (br), 8.03 (br, 2H), 7.05 (br, 2H) ppm. ¹³C NMR (*d*₆-DMSO): δ 159.5, 148.2, 146.9, 140.5, 122.0 ppm. IR (KBr): $\tilde{\nu}$ = 3577, 3443, 3331, 1676, 1641, 1577, 1528, 1460, 1356, 1318, 1299, 1158, 1085, 847, 807, 767, 633 cm⁻¹. C₆H₈N₁₂O₇ (360.20): Calcd: C 20.01, H 2.24, N 46.66 %. Found: C 20.02, H 2.26, N 46.92 %.

General procedure for synthesis of **5-9**

The energetic acid (1.0 mmol) was dissolved in water (10 mL), then **1** (0.28 g, 2.0 mmol) was added with stirring. The reaction mixture was heated at 60 °C and stirred for 2 h. The precipitate formed, was collected by filtration, and washed with water (5 mL ×

2), and ethanol (5 mL $\times$ 2) and dried in vacuum to give the products **5-9**.

5: White solid. $T_d(\text{onset})$: 238 °C. ^1H NMR (d_6 -DMSO): δ 8.92 (br, 2H), 7.18 (br, 2H) ppm. ^{13}C NMR (d_6 -DMSO): δ 163.7, 151.1, 150.5, 140.9 ppm. IR (KBr): $\tilde{\nu}$ = 3433, 3367, 1711, 1690, 1649, 1517, 1487, 1452, 1413, 1358, 1022, 991, 889, 864, 842, 814, 662 cm^{-1} . $\text{C}_8\text{H}_{12}\text{N}_2\text{O}_5$ (468.31): Calcd: C 20.62, H 2.63, N 59.82 %. Found: C 20.51, H 2.63, N 59.91 %.

6: White solid. $T_d(\text{onset})$: 277 °C. ^1H NMR (d_6 -DMSO): δ 11.80 (br), 6.18 (s, 2H), 5.62 (br, 1H) ppm. ^{13}C NMR (d_6 -DMSO): δ 169.2, 164.6, 154.9, 151.2, 141.2 ppm. IR (KBr): $\tilde{\nu}$ = 3414, 3331, 3154, 1688, 1617, 1542, 1490, 1441, 1397, 1358, 1256, 1185, 1135, 1088, 1011, 848, 750, 679, 626 cm^{-1} . $\text{C}_{10}\text{H}_{12}\text{N}_{22}\text{O}_4$ (504.36): Calcd: C 23.81, H 2.42, N 61.10 %. Found: C 23.84, H 2.43, N 61.70 %.

7·2H₂O: Orange solid. $T_d(\text{onset})$: 163 °C. ^1H NMR (d_6 -DMSO): δ 13.00 (br), 7.98 (s, 2H), 6.96 (br, 1H) ppm. ^{13}C NMR (d_6 -DMSO): δ 160.0, 159.2, 152.2, 148.4, 140.6 ppm. IR (KBr): $\tilde{\nu}$ = 3428, 3365, 3184, 1680, 1606, 1588, 1522, 1457, 1388, 1335, 1272, 1173, 1079, 1050, 1018, 991, 946, 882, 838, 820, 771, 750, 689, 654, 604 cm^{-1} . $\text{C}_{10}\text{H}_{16}\text{N}_{24}\text{O}_8$ (600.39): Calcd: C 20.00, H 2.69, N 55.99 %. Found: C 20.28, H 2.65, N 56.61 %.

8: White solid. $T_d(\text{onset})$: 292 °C. ^1H NMR (d_6 -DMSO): δ 7.07 (br) ppm. ^{13}C NMR (d_6 -DMSO): δ 164.8, 151.6, 140.8, 135.1 ppm. IR (KBr): $\tilde{\nu}$ = 3368, 3194, 1700, 1669, 1517, 1421, 1384, 1362, 1343, 1235, 1168, 1114, 1061, 996, 869, 822, 783, 743, 701, 679 cm^{-1} . $\text{C}_8\text{H}_{12}\text{N}_{22}\text{O}_2$ (448.33): Calcd: C 21.43, H 2.70, N 68.73 %. Found: C 21.43, H 2.86, N 67.51 %.

9: Yellow solid. $T_d(\text{onset})$: 272 °C. ^1H NMR (d_6 -DMSO): δ 7.74 (br) ppm. ^{13}C NMR (d_6 -DMSO): δ 161.2, 149.2, 142.3, 140.7, 139.7 ppm. IR (KBr): $\tilde{\nu}$ = 3427, 3340, 3151, 1697, 1678, 1651, 1626, 1543, 1502, 1482, 137, 1302, 1253, 1207, 1115, 1024, 949, 856, 835, 813, 752 cm^{-1} . $\text{C}_{12}\text{H}_{12}\text{N}_{22}\text{O}_8$ (592.37): Calcd: C 24.33, H 2.04, N 52.02 %. Found: C 24.23, H 2.10, N 51.87 %.

Conflict of interest

The authors declare no conflict of interest.

Acknowledgments

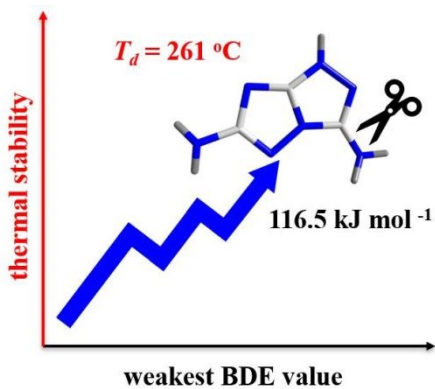
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The fused-triazole backbone 1*H*-[1,2,4]triazolo[4,3-*b*][1,2,4]triazole with two *C*-amino groups gave a highly thermally energetic compound.