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Requirements for Terpene Cyclizations inside the Supramolecular Resorcinarene Capsule: Bound Water and its Protonation Determine the Catalytic Activity

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Supporting Information Placeholder

ABSTRACT: The elucidation of the requirements for efficient catalysis within supramolecular host systems is an important prerequisite for developing novel supramolecular catalysts. The resorcinarene hexamer has recently been shown to be the first supramolecular catalyst to promote the tail-to-head terpene cyclization in a biomimetic fashion. We herein present the synthesis of a number of resorcinarene-based macrocycles composed of different ratios of resorcinol and pyrogallol units capable of self-assembly, and compare the corresponding assemblies regarding their catalytic activity in the cyclization of monoterpenes. The assemblies were investigated in detail with respect to a number of properties including the encapsulation of substrate, and ion pairs, the structural incorporation of water and the response to externally added acid (HCl). The results obtained strongly indicate that water incorporated into the hydrogen bond network of the self-assembled structure plays an integral role for catalysis, effectively acting as a *proton shuttle* to activate the encapsulated substrate. These findings are also supported by molecular dynamics simulations, providing further insight into the protonation pathway and the relative energies of the intermediates involved.

Introduction

The catalytic power and selectivity displayed by natural enzymes still serves as inspiration and role model for many organic chemists working in the broad field of catalysis.¹ Chemists successfully mimicked some aspects of enzyme catalysis utilizing self-assembled supramolecular host structures.^{2,3}

Numerous reactions within different host structures have been reported, however, their catalytic efficiency and selectivity usually do not rival their natural counterparts. In order to close this gap, and to design new, more efficient catalysts, it is essential to understand the fundamental requirements for catalytic activity in these artificial systems.

One of the most frequently applied supramolecular catalyst is the hexameric capsule **I** (Figure 1a). It self-assembles from six units of resorcinarene **1** and eight molecules of water in apolar media such as chloroform.⁴ A number of reactions involving mainly cationic transition states have been reported using structure **I** as a catalyst.^{2k, 2n, 2q-s} A prime example is the successful catalysis of the tail-to-head terpene (THT) cyclization inside **I** developed by our group,⁵ which utilizes the supramolecular cavity to enable a reaction that is very difficult to perform in bulk solution.⁶ Terpenes form one of the largest classes of natural products with remarkable structural

diversity. Many members exhibit interesting biological activity, making them suitable lead compounds for drug development.⁷ Since most compounds in this class require considerable synthetic effort and are often not available in significant quantity, an efficient method to access these compounds from rather simple precursors would provide a powerful tool to the organic synthetic community. Utilizing capsule **I** as an aromatic cavity with some similarities to natural cyclase enzymes, we were able to showcase some first examples: A four-step total synthesis of isolongifolene,^{5c} the first total synthesis of δ -selinene,^{5d} as well as a four step synthesis of the complex tricyclic presilphiperfolan-1 β -ol,⁸ which is difficult to access *via* other means. Interestingly, the closely related hexamer **II**,⁹ which is formed from six units of pyrogallolarene **2** (Figure 1b) has been found to be catalytically inactive in THT cyclizations.¹⁰ The reason for its inactivity, however, remained unknown. We previously speculated that either its low intrinsic acidity,¹⁰ or its inability to bind ion pairs are the cause for this observation. To clarify this issue, we decided to closely investigate assemblies **I** and **II**, as well as the related macrocycles **3**–**5** (Figure 1d), featuring different ratios of resorcinol and pyrogallol units. Additionally, the electron-deficient tetrafluorinated resorcinarene derivative **6** was selected. As the main test reaction, we chose the THT cyclization of geranyl acetate (**7**) to α -terpinene (**8**, Figure 1c, Scheme S1), a reaction which was shown to undergo a “non-stop” cyclization inside **I**.^{5a}

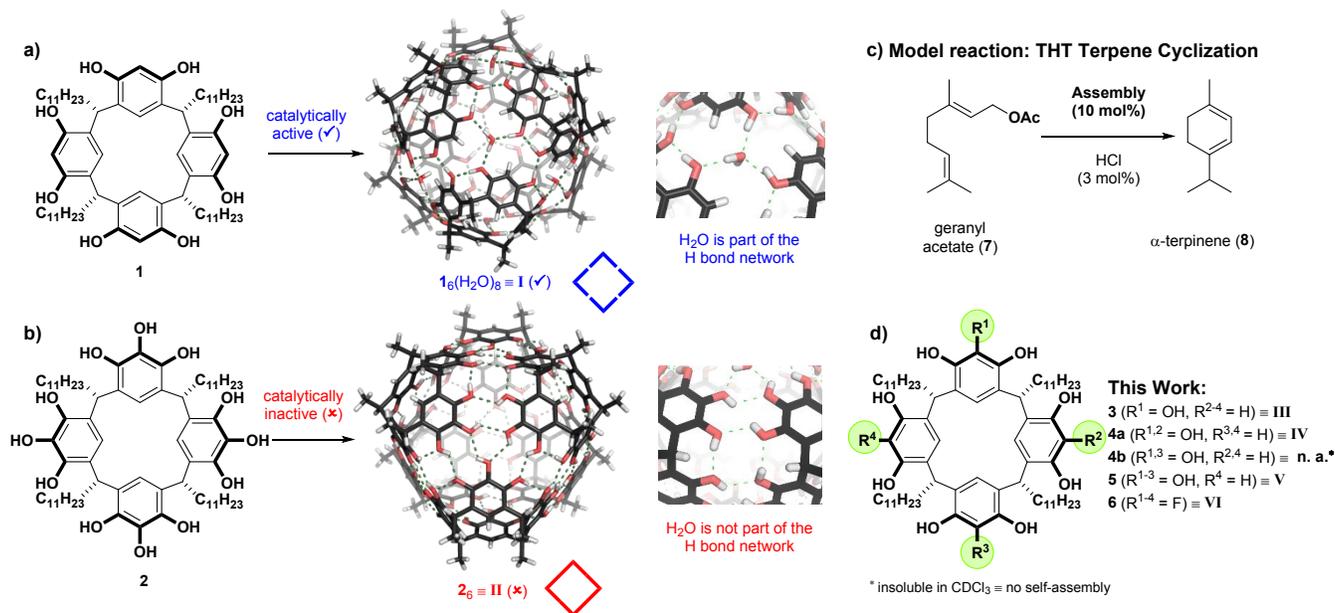
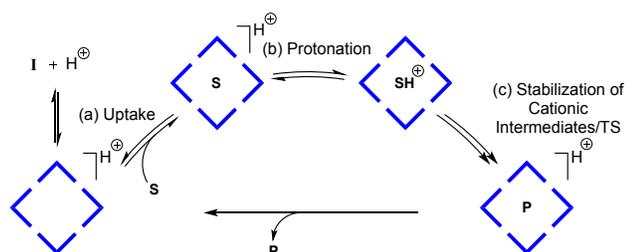


Figure 1: a) Self-assembly of **I** from six units of **1** and eight water molecules, four of which are not fully saturated within the hydrogen bond network and can function as hydrogen bond donor at the inside of the cavity. b) Self-assembly of **II** from six units of **2** forming a fully saturated hydrogen bond network without water. c) Tail-to-head terpene (THT) cyclization of geranyl acetate (**7**) forming α -terpinene (**8**). d) Macrocycles **3 – 6** corresponding to assemblies **III – VI**.

Concerning the prerequisites for catalytic activity, we considered the following steps in the catalytic cycle as potentially decisive (Scheme 1): (a) For a successful conversion inside the molecular capsule, the substrate has to be encapsulated. Therefore, substrate uptake in the different assemblies was explored first. (b) The substrate has to be activated by protonation.^{5b} Protonation in an apolar solvent like chloroform leads to the formation of ion pairs. Consequently, we next investigated the ability of the assemblies to encapsulate ion pairs. (c) Cationic intermediates and transition states have very likely to be stabilized *via* cation- π interactions. Accordingly, we included an electron-deficient derivative (i.e. **6**) featuring four additional fluorine substituents in the study as it should display reduced cation- π stabilization.



Scheme 1: Potentially decisive steps of the catalytic cycle; **S** = Substrate, **TS** = Transition States, **P** = Products.

Here, we report efficient synthetic routes to macrocycles **3 – 6** (Figure 1d), their ability to self-assemble to the hexameric capsules **III – VI**, and provide evidence that water being incorporated into the hydrogen bond network of the corresponding assemblies plays a crucial role for the efficient catalytic cyclization of monoterpenes.

Results and Discussion

Synthesis of Macrocycles 3 – 6 and Investigation of their Self-assembling Properties

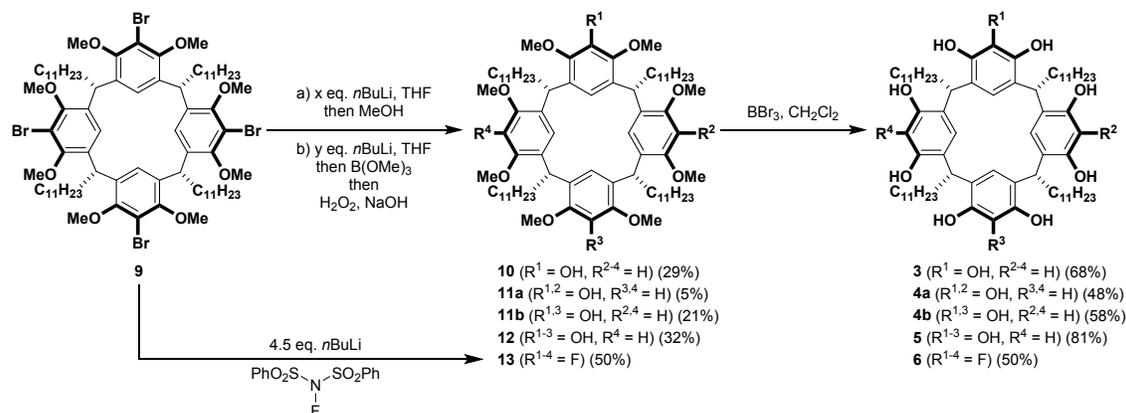
First, it was necessary to develop a reliable route to access the macrocycles **3 – 6**, in pure form on a preparative scale (Scheme 2). While there has been a report about directly cyclizing different ratios of resorcinol and pyrogallol with the corresponding aldehyde to obtain macrocycles **3 – 5** by Atwood *et al.*, these reactions yielded very complex mixtures that were not separable in our hands.^{9d} To circumvent these separation issues, a more controlled route utilizing tetrabromo-derivative **9** was developed (Scheme 2). Compound **9** was accessible *via* literature procedures in two steps from resorcinarene **1** on decagram scale.¹¹ It was partially debrominated by treatment with specific amounts of *n*-butyllithium, and subsequently methanol. The remaining aryl bromides were then converted *in situ* into phenolic moieties using *n*-butyllithium and trimethylborate followed by the addition of NaOH/H₂O₂.¹² This yielded biased mixtures of the octamethylated compounds **10 – 12** which were separable by column chromatography. The tetrafluorinated compound **13** was accessed in a similar way from **9** using halogen-lithium exchange followed by the addition of the electrophilic fluorine reagent *N*-fluorobenzenesulfonimide (NFSI).¹³ The removal of the methyl protecting groups was achieved by stirring in the presence of boron tribromide at room temperature for several days,¹⁴ yielding compounds **3 – 6** in moderate to good yields and in analytically pure form (see SI Chapter 12). The products were recrystallized and subsequently washed with a H₂O/MeOH mixture in order to remove any acid traces (see SI for details). Aside from compound **4b**, featuring two distal pyrogallol units, all macrocycles were soluble in chloroform, and self-assembled to hexameric capsules related to **I** and **II**.

DOSY-NMR experiments¹⁵ provided diffusion values for the new assemblies very close to those of assemblies **I** and **II** (see Table S2). Self-assembly was further corroborated by the fact that assemblies **III** – **VI** show uptake of ammonium salts (i.e. NBu_4Br (**14**), Figure S18) indicated by the strong upfield shift of the guest protons due to the aromatic shielding effect. We, therefore, concluded that the obtained macrocycles form assemblies resembling the structures of **I** and **II**. However, it has to be noted that in the case of the assemblies **III** – **V**, the structures are not as symmetric and clearly defined as in **I** and **II**. Due to the lower degree of symmetry of the macrocycles, different isomers of hexameric capsules are formed. The broad signals in the ^1H -NMR spectra recorded in CDCl_3 suggest that there is more than one assembly present and that these are interconverting. Solid state structures could not be obtained since crystal formation is hindered by the dynamic nature of the assemblies, and the long flexible alkyl chains. Besides, crystal structures do not always correlate with structures present in solution as illustrated by macrocycle **5**.^{9d} While we cannot exclude the presence of the previously reported, X-ray-based donut shaped structure,^{9d} none of the experiments we carried out hinted at such an assembly in solution. In the case of assembly **VI**, the NMR spectrum closely resembles that of the resorcinarene-based capsule **I**, indicating a similar structure. (see SI for the energy optimized structure)

Evaluation of the Catalytic Activity in the Cyclization of Monoterpenes

With assemblies **I** – **VI** in hand, the cyclization of geranyl acetate (**7**, Figure 1c) as the model reaction was investigated employing the optimized conditions reported previously by our group for assembly **I**.^{5b} HCl is added as a co-catalyst, which, unable to catalyze the reaction by itself, works in a synergistic fashion with the assembly. As reported by us,¹⁰ no reaction was observed with assembly **II** (Figure 2e), while assembly **I** (Figure 2a) catalyzed the formation of α -terpinene (**8**) from acetate **7** in 30 – 35% yield. The reaction proceeds in good selectivity with other common cyclization products (Figure 2g) such as eucalyptol (**15**), terpinolene (**16**), limonene (**17**), γ -terpinene (**18**) and isoterpinolene (**19**) being only formed in minor quantities (< 10% yield). Assembly **III** (Figure 2b), formed from monomer **3**, featuring one additional hydroxyl group compared to resorcinarene **1**, also displayed catalytic activity and likewise yielded α -terpinene (**8**) as the main

product, albeit in a lower yield. When employing assembly **IV** (Figure 2c), the conversion of substrate remained significant, but the overall yield of cyclization products dropped significantly. We attribute this fact to a side reaction in which phenolic groups of **4a** react with the highly reactive cationic intermediates of the cyclization reaction. This was also observed for **I** in an earlier report and confirmed by isolation.^{5a} In the current study, these alkylation products were detected by ESI-MS analysis of the reaction mixtures (see Figures S9 – S14) in the cases of the assemblies **I**, **III** and **IV**. DOSY-NMR measurements during the reaction indicate that the assemblies stay largely intact (SI Chapter 7.4). Interestingly, assembly **V** (Figure 2d) with three pyrogallol units showed no catalytic activity in this cyclization reaction. The reaction in the presence of assembly **VI** (Figure 2f) featuring subunits with fluorine substituents proceeded very slowly. Similar results were obtained when the reaction was conducted with the intrinsically more reactive nerol (**20**) as the substrate (see Figure S3 for the reaction profiles). Assemblies **I**, **III** and **IV** showed comparable behavior with only small differences in product selectivity. Initially, mainly α -terpineol (**25**) is formed, which is then further converted to eucalyptol (**15**), the major product after 72 hours (30 – 38% yield). In the case of the fluorinated assembly **VI**, the formation of α -terpineol (**25**) in combination with traces of eucalyptol (**15**) (4% after 72 hours) indicates a reaction profile similar to the one of assembly **I**, with the conversion being, however, again significantly slower. Assembly **II** shows no activity, while the closely related assembly **V** leads to considerable conversion (60% conversion after 72h, Table 1, entry 1). At first glance, this seemed surprising since assembly **V** showed no conversion with geranyl acetate (**7**). In order to understand this observation, control experiments with assembly **V** were conducted (Table 1). No conversion was observed in the absence of **V** with HCl present (entry 2). Blocking assembly **V** with tetrabutyl ammonium bromide (**14**) did not halt the reaction completely (13%, entry 3), indicating that the reaction is not only taking place within the cavity in that case. This was further corroborated by substituting assembly **V** with 4-hexyl resorcinol (**21**) which also led to some conversion of nerol (**20**) (10%, entry 5) after three days. Moreover, the product distribution of entry 1 indicated that the reaction likely takes place outside/on the outer surface of the capsule. The main products are limonene (**17**) and α -terpineol (**25**), which are also formed in regular solution experiments under acidic conditions,¹⁶ and in the presence of 4-hexyl resorcinol (**21**).



Scheme 2: Synthesis of the macrocycles **3** – **6** starting from the literature known compound **9**.¹¹

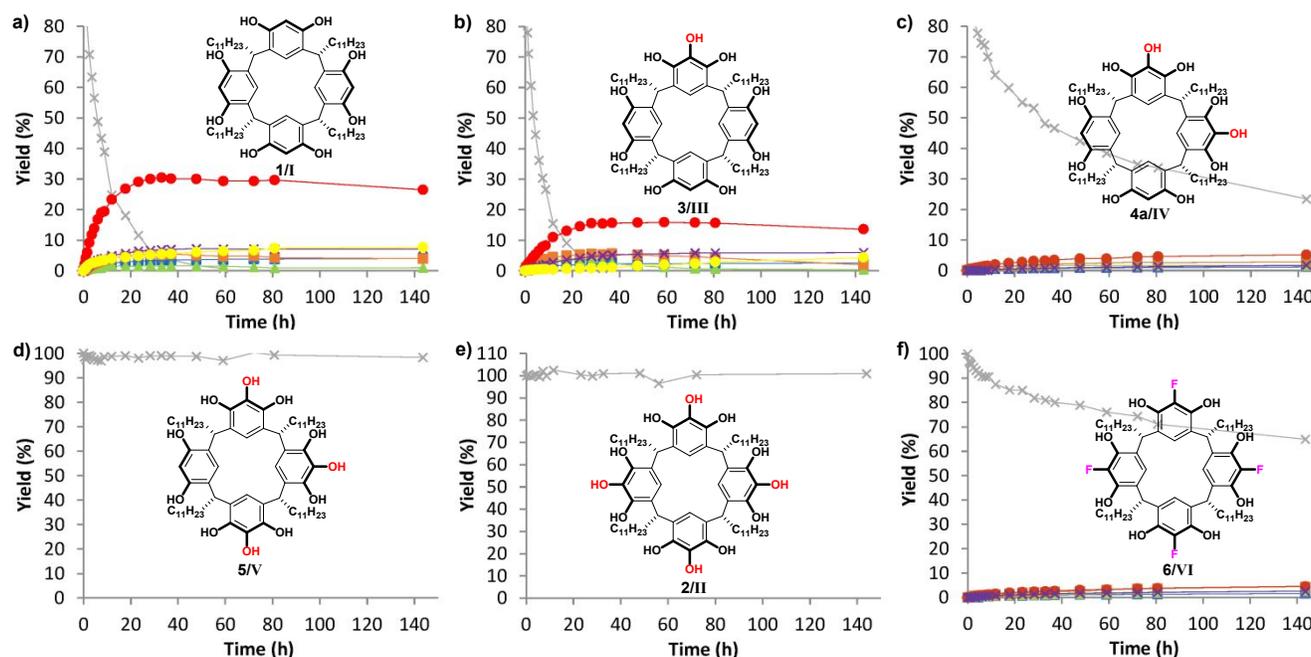
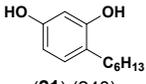


Figure 2 a – f): GC-based reaction profiles of assemblies **I**, **III**, **IV**, **V**, **II** and **VI** with geranyl acetate (**7**) as substrate (reaction conditions: 33.3 mmol/L of **7**, 10 mol% assembly, 3 mol% HCl, in CDCl_3 at 30 °C). g) Substrate **7** and reaction products. h) Additional substrates and guests.

Table 1: Control experiments for the nerol (**20**) cyclization with assembly **V** under standard reaction conditions. (33.3 mmol/L of **20**, 3 mol% HCl, in CDCl_3 at 30 °C)

Entry	Catalyst (mol%)	Additive (mol%)	Conversion (%) ^[a] (72 h)
1	V (10)	-	60
2	-	-	-
3	V (10)	<i>n</i> -Bu ₄ NBr (14) (15)	13
4	-	<i>n</i> -Bu ₄ NBr (14) (15)	-
5	-	 (21) (240)	10

^[a] Determined by GC

Additionally, the absence of eucalyptol (**15**) which is usually formed to a significant degree when the reaction takes place within the assembly (i.e. in assembly **I**) indicates a conversion outside the assembly. Taken together, these results indicate that in this case the reaction is taking place mainly outside/on the outer surface of assembly **V** presumably via hydrogen bonding (Table 1, entry 5), due to the higher reactivity of nerol (**20**). What causes the difference in capsules **II** and **V**? It is likely that the perfect hydrogen bond network^{9b} of **II** prevents activation of nerol *via* hydrogen bonds. It should also be noted that assembly **V** proved to be an inefficient catalyst in the cyclization of other substrates (Figure 2h) such as geraniol (**22**), linalool (**23**), neryl chloroacetate (**24**), all of which showed less than 15% conversion after 72 hours. In the cases where traces of product were detected, these correspond again to limonene (**17**) and α -terpineol (**25**), indicating that the conversion is likely taking place in solution rather than inside **V** (Figures S5–S7). In summary, the cyclization studies revealed that assemblies **I**, **III**, **IV** and **VI** are active supramolecular catalysts for the THT cyclization, while **II** and **V** are inactive.

Uptake Studies

Substrate uptake is a prerequisite for catalysis inside the capsule. In order to elucidate a potential correlation between catalytic inactivity and inability of substrate uptake, a set of NMR experiments was performed. For this purpose, the saturated substrate analogue **26** (Figure 2h), which is not converted in the presence of assemblies **I** – **VI** and HCl, was employed. Host and guest solutions in chloroform were mixed in the ratios used for cyclization (1:10) in the presence of HCl at 0 °C and immediately submitted to NMR spectroscopy. The samples were then stored at 30 °C and monitored by ¹H-NMR. Guest uptake over time was determined *via* ¹H-NMR analysis (Table 2), by comparison of the respective assembly signals (internal standard) and the signal of 'free' **26**. The values given are calculated assuming that directly after addition no encapsulation is observed. Table 2 shows that the amount of non-encapsulated substrate **26** is decreasing with time in all cases, indicating a gradual encapsulation of **26** by all assemblies to varying degrees (for full details see Tables S4 – S9). This is further supported by the appearance of small signals below 0.5 ppm in the respective ¹H-NMR spectra, which continued to increase when the samples were monitored over a longer period of time (see Figures S15, S16). Since all assemblies investigated, including the catalytically inactive assemblies **II** and **V**, showed some uptake of guest **26**, a direct correlation between catalytic activity and encapsulation of substrate could not be established.

Table 2: Encapsulation of substrate analogue **26** by assemblies **I** – **VI** under reaction conditions. (33.3 mmol/L of **26**, 10 mol% assembly, 3 mol% HCl, in CDCl₃ at 30 °C)

Entry	Assembly	Encapsulation after 6 h (%) ^[a]	Encapsulation after 48 h (%) ^[a]
1	I	1.3	3.1
2	III	1.0	1.4
3	IV	0.3	0.7
4	V	0.5	1.5
5	II	0.6	1.0
6	VI	0.3	2.3

^[a] Determined by ¹H-NMR Integration

Encapsulation of Ion Pairs

After encapsulation, protonation of the substrate with HCl in the apolar solvent chloroform will form an ion pair. Therefore, we investigated the ability of the different assemblies to encapsulate ion pairs. While **I** readily encapsulates ammonium salts (e. g. tetrabutylammonium bromide **14**) as ion pairs, the encapsulation of anions in assembly **II** has been shown to be energetically unfavorable.¹⁰ The binding of anions can be easily investigated with ¹⁹F-NMR when utilizing fluorine containing anions (e.g. triflate). Figure 3 shows the ¹⁹F-NMR spectra of the assemblies **I** – **VI** in the presence of 1.0 equivalent of

tetrabutylammonium triflate (**27**). In all cases, the presence of the assembly causes a shift and broadening of the signal corresponding to the triflate anion, with the sole exception being assembly **II** where no effect on the triflate signal was observed. When employing tetrabutylammonium tetrafluoroborate (**28**), similar results were obtained (Figures S19, S20). This is in agreement with the ability of all assemblies beside **II** to encapsulate ammonium salts such as TBAB (see also SI Chapter 8.3). Importantly, the catalytically inactive assembly **V** is able to bind ion pairs in contrast to the inactive assembly **II**. In conclusion, the ability to encapsulate ion pairs does not correlate with the catalytic activity of assemblies **I** – **VI**.

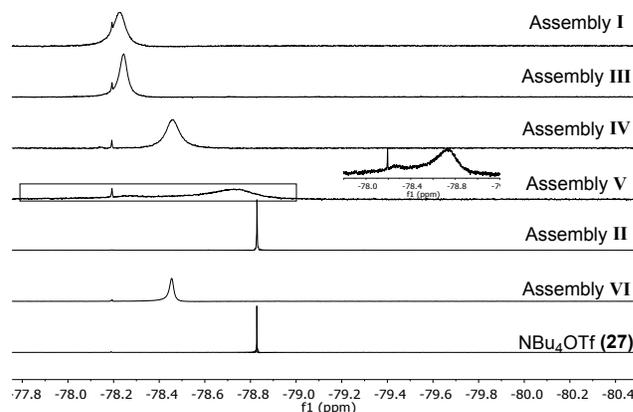


Figure 3: ¹⁹F-NMR spectra of assemblies **I** – **VI** (3.33 mmol/L) in the presence of 1.0 eq. of NBu₄OTf (**27**) in CDCl₃.

Protonation Studies

During the cyclization studies, we recognized that some ¹H-NMR signals of several assemblies changed upon the addition of HCl as co-catalyst. Therefore, this phenomenon was investigated in more detail. For this purpose, chloroform solutions of the assemblies **I** – **VI** were titrated with HCl (3.33 mmol/L of assemblies **I** – **VI**, 0.0 – 1.0 eq. HCl added to the same sample). The results (Figure 4) show that in four cases (**I**, **III**, **IV** and **VI**, Figure 4a – c and f) the signals corresponding to the phenol groups as well as the water signal broaden with increasing amounts of HCl. We attributed these changes to the protonation of the assemblies by the external acid. In contrast, the signals corresponding to the phenol groups of the assemblies **II** and **V** (Figure 4d, e) remain unaffected by the added acid and also the water signal only broadens to a very small extent when compared to the other assemblies (Figures 4a – c and f). This indicates that assemblies **II** and **V** are not protonated by the external acid, at least on the ¹H-NMR (500 MHz) time scale.¹⁷ A protonation of the phenolic groups of **II** and **V** in the presence of unbound water seems unlikely considering the pK_a values of H₃O⁺ (pK_a = 0)¹⁸ and protonated phenols (pK_a ≈ -6).¹⁹ The integrity of the assemblies **I** – **VI** in presence of 1.0 eq. HCl was confirmed by DOSY-NMR experiments. (see SI Chapter 10 for details). Importantly, the observed protonation of the host correlates well with the observed catalytic activity of the respective assemblies in the cyclization of monoterpenes.

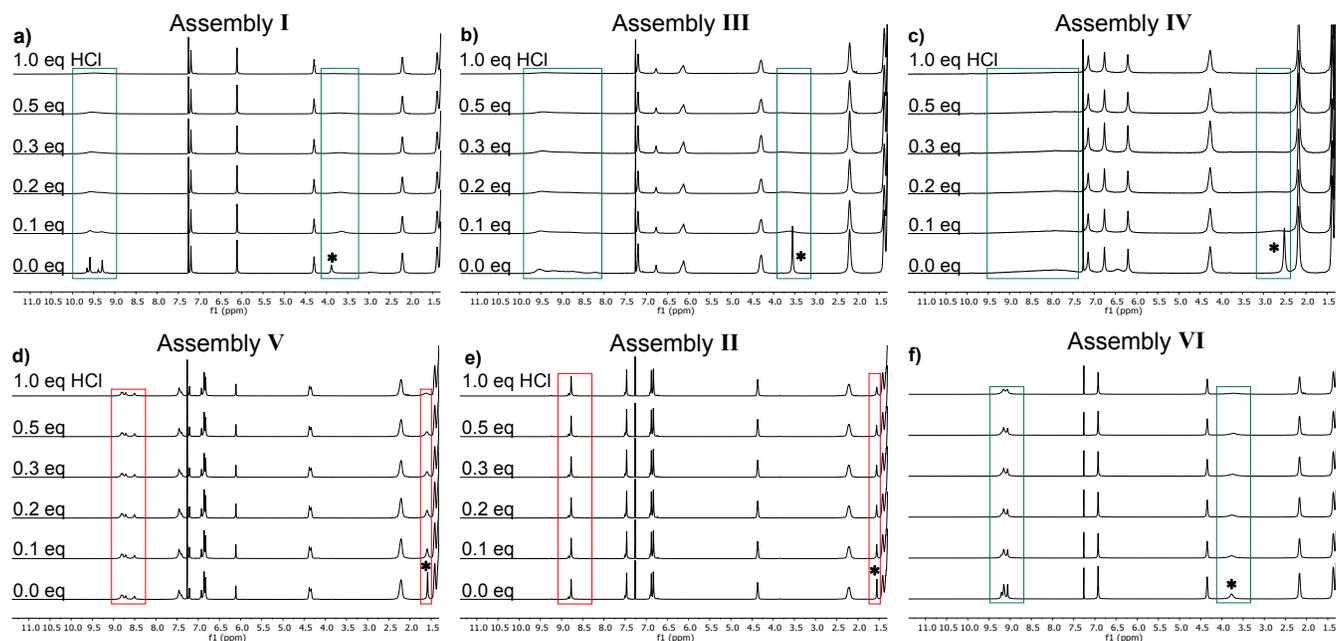


Figure 4: a – f) NMR titration of the respective assembly (3.33 mmol/L) with HCl (0.0 – 1.0 eq.). Representative signals corresponding to phenol moieties and water are highlighted (green for catalytically active, red for catalytically inactive). The water signals are marked with an asterisk (*).

Water as Integral Part of the Assembly

As mentioned before, the main structural difference between assemblies **I** and **II** is the incorporation of water into the hydrogen bond network. This can be directly observed in the $^1\text{H-NMR}$ spectra in CDCl_3 , where the water signal in the presence of **I** is significantly downfield shifted, while it remains largely unaffected with **II** (Figure 4a and e, 0.0 eq. HCl). The

shift is associated with the incorporation of water into the capsular hydrogen bond network.²⁰ Water being incorporated into the assembly and 'free' water in solution are in fast exchange on the $^1\text{H-NMR}$ time scale, therefore, only one averaged water signal is observed. The magnitude of the downfield shift depends on the total amount of water present in the solution; with a low water content leading to a stronger shift. Cohen *et al.* reported a method based on DOSY-NMR experiments to determine the water content in **I** and **II**.²⁰

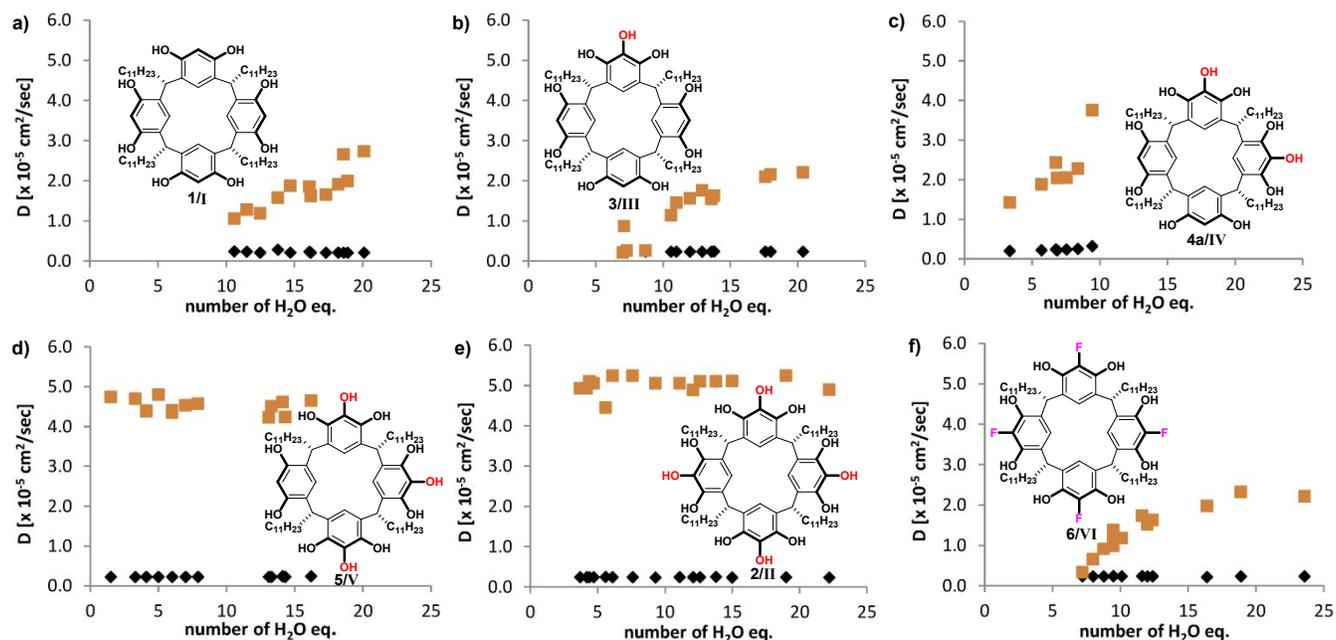


Figure 5: a – f) Influence of the water content on the diffusion coefficients D of the assemblies (\blacklozenge) and water (\blacksquare). Experiments were performed in CDCl_3 solutions of the assemblies (3.33 – 1.67 mmol/L).

If water is not part of the assembly, like in **II**, its diffusion value is independent of the total water amount in the sample (e.g. assembly **II**, Figure 5e). However, if water is an integral part of the capsule, its diffusion value converges to the value of the assembly when the total amount of water decreases (assembly **I**, Figure 5a). Varying the water content of chloroform solutions of the novel catalytically active assemblies **III**, **IV** and **VI** (Figure 5 b, c, f) indicated that in these cases water is part of the hydrogen bond network, similar to structure **I** (Figure 5a). In contrast, the data obtained for the inactive assembly **V** (Figure 5d) implies that water does not take part in the formation of assembly **V**; the same observation as was made for structure **II** (see SI for a molecular model of assembly **V**). It should be noted that in the case of assembly **IV**, due to an overlap between the water signal and signals of the assembly, only a few measurements resulted in reliable data points concerning the diffusion value of water. However, considering that the water signal is shifted significantly and that the diffusion values obtained are significantly lower than the values observed for the assemblies containing no water, we concluded that water is part of the hydrogen bond network of assembly **IV**.

In conclusion, the incorporation of water into the hydrogen bond network of the assembly correlates well with the catalytic activity observed. Interestingly, the assemblies containing water in the hydrogen bond network also displayed enhanced protonation in the presence of HCl (see previous section). These observations indicate a central role of the bound water molecules for the catalytic activity.

Molecular Dynamics Simulations

The results presented in the previous two sections, indicate that water being part of the hydrogen bond network is essential in the protonation of the assembly by an external acid; thereby enabling catalysis within the host structures **I**, **III**, **IV** and **VI**. In order to learn more details about the protonation event inside the capsule, we turned to molecular dynamics (MD) simulations. Specifically, capsule **I** containing encapsulated geranyl acetate (**7**), and external HCl were submitted to QM/MM (DFT, PBE,²¹ +D level²²) MD simulations at 300 K using the CP2K code.²³ The QM region consisted of six resorcinol units belonging to three adjacent resorcinarene molecules, one water molecule, one HCl molecule and geranyl acetate (**7**), whereas the rest of assembly **I** and the chloroform solvent molecules were treated at the MM level (see SI for computational details).

As the proton transfer from HCl to the carbonyl-oxygen of substrate **7** involves separate, relevant metastable states separated by high free energy barriers along the reaction path, standard MD is not suitable. Therefore, the sampling of the process was enhanced using Metadynamics²⁴ (MetaD) implemented in the PLUMED2 code.²⁵ To characterize the complex free energy landscape describing the proton migration from HCl throughout the complex hydrogen bond network, we used the method recently developed by Grifoni *et al.*²⁶

The simulation can be summarized as follows: (i) Energetically, the most favored pathway involves a direct protonation of the water at the capsule surface by HCl (Figure 6a, state **A** → **B**). The alternative pathway involving protonation of a phenol group, followed by rapid proton migration to a water molecule, is less likely (Figure S23). (ii) The formed chloride of the $\text{H}_3\text{O}^+/\text{Cl}^-$ ion pair is stabilized by coordination to the

surrounding phenol groups of the resorcinarene molecules. It replaces the hydronium ion in the hydrogen bond network of capsule **I**

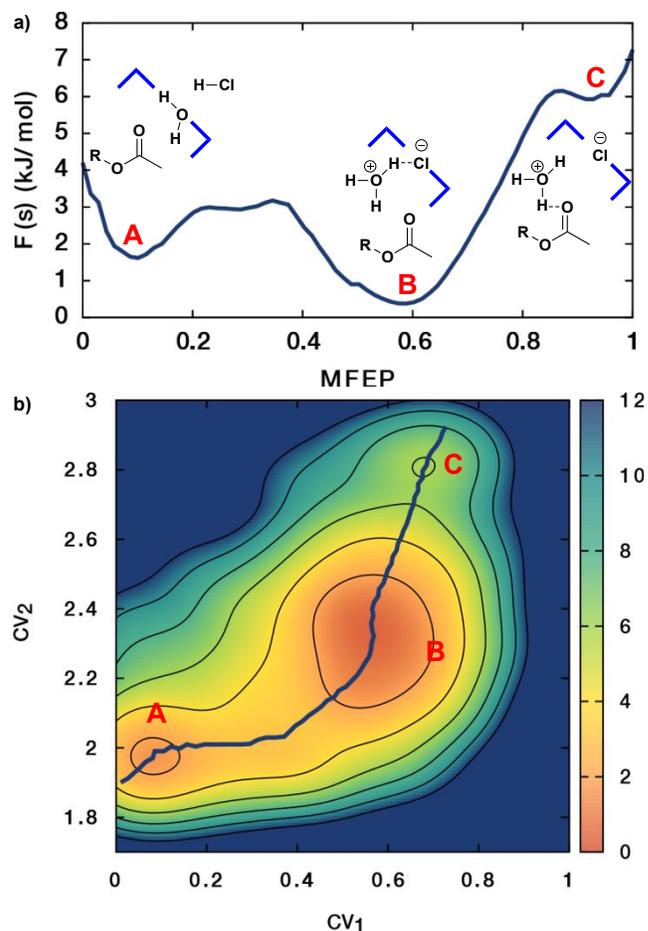


Figure 6: a) Free energy along the minimum free energy path (MFEP) showing the three states' (A, B and C) relative stabilities and free energy barriers separating them. b) Free energy surface for the geranyl acetate (**7**) protonation by means of HCl as co-catalyst for assembly **I** along the coordination of the carbonyl-oxygen with the water oxygen atom (CV_1) and the coordination of the water oxygen atoms with the acidic hydrogen atoms (CV_2). Letters A, B, and C label the minima corresponding to the metastable states referring to the initial neutral state, the intermediate formation of the $\text{H}_3\text{O}^+/\text{Cl}^-$ ion pair, and the final protonation of the carbonyl-oxygen by H_3O^+ . The blue line passing through the three states represents the MFEP connecting them. The color bar represents the energy reported in $\text{kJ}\cdot\text{mol}^{-1}$.

and shifts the hydronium ion below the inner surface of the capsule (state **B**). (iii) Finally, the hydronium ion and the carbonyl of the geranyl acetate substrate form a very compact complex, sharing the proton (state **C**). Similar structures have been observed in other computational studies focusing on the acid catalyzed ester hydrolysis.²⁷

To further analyze the results, we calculated the free energy surface following the reweighting procedure previously described.²⁸ In order to extract useful and chemically meaningful information on the protonation process, two

collective variables (CV_1 and CV_2) describing the progressive protonation reaction of the carbonyl-oxygen have been selected. To be general in describing the fundamental chemical features, we used coordination numbers. These functions count how many atoms of a specific species are found within a cutoff sphere from the central atom. In the present case, CV_1 accounts for the proximity of the carbonyl-oxygen and the water oxygen atoms. At the starting point of the simulation (Figure 6b, state **A**), CV_1 is close to zero; after the formation of the final hydronium-carbonyl complex (state **C**), it will be around one, as the water molecule gets close to the carbonyl-oxygen. The second variable (CV_2) accounts for the number of hydrogen atoms in close proximity to the water-oxygen. At the starting point of the simulation (state **A**), CV_2 is two, corresponding to the number of covalent bonds in the neutral water molecule. As the proton transfer proceeds, CV_2 increases.

The minimum free energy path²⁹ (MFEP) along the free energy surface is depicted in Figure 6a. This curve represents the minimum energy required during the process. It is worth noticing that state **B** in which the H_3O^+/Cl^- ion pair is formed is the most stable state. The chloride anion is well stabilized by the surrounding phenol groups at the surface of the capsule (see Figure S24). The estimated energy barrier for the dissociation process is only approx. $1.5 \text{ kJ}\cdot\text{mol}^{-1}$, much lower than the thermal barrier ($k_B T \sim 2.5 \text{ kJ}\cdot\text{mol}^{-1}$). The picture is different when we consider the formation of the carbonyl/hydronium complex of state **C**. Its formation is slowed down by a larger barrier of approx. $6 \text{ kJ}\cdot\text{mol}^{-1}$ and state **C** is less stable than state **B**. Again, this result is in line with previous computational findings.²⁷

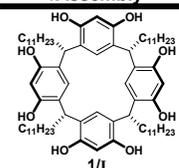
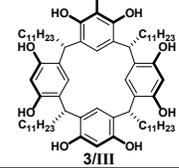
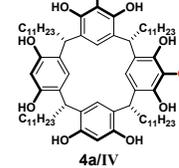
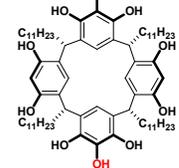
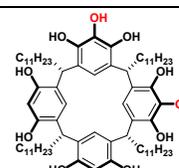
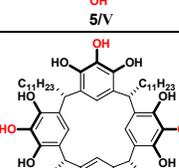
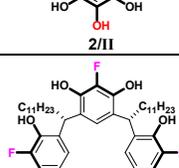
The calculations, therefore, provide further evidence that water is indeed essential for the proton transfer process from the co-catalyst HCl onto the encapsulated substrate.

Conclusion

The experimental results obtained in this study, are summarized in Table 3. Taken together, these results strongly indicate that the decisive requirement for catalytic activity in the context of terpene cyclization is the incorporation of water molecules into the hydrogen bond network of the assembly. The crucial water molecule functions as a proton shuttle, delivering the proton onto the encapsulated substrate. If water is not incorporated into the assembly, this protonation mode is prevented, rendering these hosts catalytically inactive. This holds true even in the case where the host is able to encapsulate the substrate and stabilize ion pairs (assembly **V**, Table 3, entry 5). This finding has furthermore been validated by calculations, which confirm that the proposed intermediary states constitute the minimum free energy path.

Interestingly, the proposed protonation pathway may also provide insights into the puzzling observation that the chloride counter anion of the HCl co-catalyst does not interfere with the cationic cyclization cascade. Control experiments with HCl, as well as other Bronsted or Lewis acids^{5a, 6, 30} usually lead to counter anion quenching of cationic intermediates. Since the chloride anion is bound at the capsule surface (Figure 6a), it is prevented from quenching cationic intermediates inside the cavity of the capsule.

Table 3: Comparison of properties of macrocycles **1–6** and the corresponding assemblies **I–VI** indicating a correlation between water being part of the hydrogen bond network and catalytic activity in the THT cyclization.

Macrocycle / Assembly	Catalysis	Substrate Binding	Anion Binding	Water	Protonation
	✓	✓	✓	✓	✓
	✓	✓	✓	✓	✓
	✓	✓	✓	✓	✓
	n. d. ^[a]				
	✗	✓	✓	✗	✗
	✗	✓	✗	✗	✗
	✓	✓	✓	✓	✓

^[a] Not determined due to the insolubility of **4b** in $CDCl_3$.

Summary

In summary, we presented the synthesis and characterization of four new macrocycles **3** – **5** composed of different ratios of resorcinol and pyrogallol units as well as a new tetrafluorinated resorcinarene derivative **6**. With the exception of compound **4b**, all derivatives self-assemble to dynamic, hexameric hydrogen-bonded assemblies in chloroform solution. The new assemblies **III** – **VI** together with assemblies **I** and **II**, were studied in detail and compared with respect to their catalytic activity in the monoterpene cyclization, their ability to encapsulate terpene substrates and ion pairs, their response to acid additive, and the amount of water incorporated into the hydrogen bond network. The experimental results strongly indicate a correlation between the catalytic activity and water being part of the assembly. QM/MM molecular dynamics simulations provided insights into the specific role of the water molecule in the protonation process of the encapsulated substrate. The incorporated water molecules likely act as a proton shuttle by transferring the proton from the acid co-catalyst HCl to the encapsulated substrate, which initiates the cyclization. Since the chloride counter anion is replacing one water molecule in the hydrogen bond network, it is immobilized and does not interfere with the cationic cyclization cascade reaction. These findings finally reveal the activation mechanism inside capsule **I**, and decipher the prerequisites for catalytic activity. It, therefore, represents an important step towards the rational design of new supramolecular catalyst systems. We expect this finding to make a significant impact on future developments in the field of supramolecular catalysis, since the proposed model is likely transferable to other types of host systems.

Associated content

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI:

Experimental details and NMR spectra of new compounds (PDF)

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Notes

The authors declare no competing financial interest.

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