

NITROSYLATION OF 3a-SUBSTITUTED

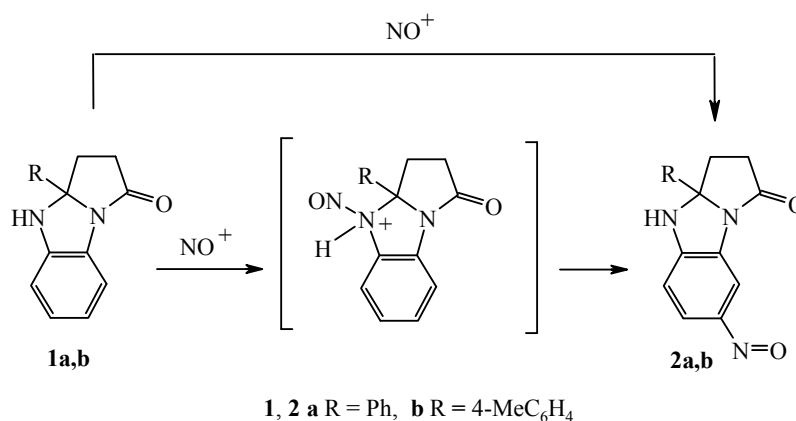
2,3,3a,4-TETRAHYDRO-1H-BENZO[d]-

PYRROLO[1,2-a]IMIDAZOL-1-ONES

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Keywords: 2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-ones, intramolecular NO⁺ cation migration, nitrosylation, Fischer-Hepp rearrangement.

We report the nitrosylation of previously synthesized 3a-substituted 2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-ones **1a,b** in the reaction of 5R-3H-furan-2-ones with 1,2-diaminobenzene [1]. The nitrosylation reaction was carried out at 0–5°C in aqueous acetonitrile. Nitrous acid was generated *in situ* and introduced directly to react with the substrate. The crystalline products have a color characteristic for nitroso compounds.



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Since imidazolones **1a,b** possess two reaction sites capable of reacting with electrophilic reagents, we could expect the formation of products of N-nitrosylation and/or C-nitrosylation.

The nitrosonium cation is a weak electrophile and, thus, attack at the secondary amino group in imidazolones **1a,b** is more likely and would obey kinetic control. However, the NMR spectral data indicate the formation of compounds **2a,b**, which were identified as 7-nitroso-3a-R-2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-ones. The formation of these nitroso products is possible through a Fischer-Hepp rearrangement, which proceeds under acid catalysis conditions as an intramolecular migration of the NO⁺ cation, leading to the thermodynamically more stable C-nitrosylation products **2a,b**.

The ¹H NMR spectra were taken on a Varian-400 spectrometer at 400 MHz at 20–25°C in CDCl₃ with TMS as internal standard.

7-Nitroso-3a-phenyl-2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-one (2a). Excess concentrated sulfuric acid was added to a solution of NaNO₂ (0.07 g, 1 mmol) in cold water (3–5 ml), followed by a solution of compound **1a** (0.25 g, 1 mmol) in 4:1 acetonitrile-water (10 ml) and then water (50 ml). The mixture was stirred for 2 h at room temperature. The precipitate was filtered off and washed with water to give 0.21 g (75%) pale-yellow crystalline nitroso derivative **2a**; mp 96–98°C. ¹H NMR spectrum, δ, ppm (*J*, Hz): 1.62 (1H, br.s, NH); 3.61 (2H, t, *J* = 6.2, H-2); 3.88 (2H, t, *J* = 6.2, H-3); 7.47 (1H, d, *J* = 2.0, H-8); 7.49–7.65 (5H, m, C₆H₅); 8.12 (1H, d, *J* = 8.0, H-6); 8.25 (1H, d, *J* = 8.0, H-5). Found, %: C 69.00; H 5.15; N 15.23. C₁₆H₁₃N₃O₂. Calculated, %: C 68.81; H 4.69; N 15.05.

3a-(4-Methylphenyl)-7-nitroso-2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-one (2b) was obtained in 73% yield analogously from compound **1b** (0.26 g, 1 mmol) and NaNO₂ (0.07 g, 1 mmol); mp 120–122°C. ¹H NMR spectrum, δ, ppm (*J*, Hz): 1.68 (1H, br.s, NH); 2.26 (3H, s, CH₃); 3.24 (2H, t, *J* = 6.0, H-2); 3.44 (2H, t, *J* = 6.2, H-3); 7.21, 7.39 (4H, d, *J* = 8.0, *p*-Tol); 7.47 (1H, d, *J* = 2.0, H-8); 8.12 (1H, d, *J* = 8.0, H-6); 8.25 (1H, d, *J* = 8.0, H-5). Found, %: C 69.49; H 5.06; N 14.69. C₁₇H₁₅N₃O₂. Calculated, %: C 69.61; H 5.15; N 14.33.

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REFERENCES

1. V. S. Grinev, O. A. Amal'chieva, A. Yu. Yegorova, and E. V. Lyubun', *Zh. Org. Khim.*, **46**, 1376 (2010).