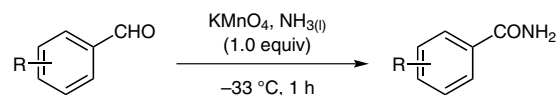


Direct Conversion of Aromatic Aldehydes into Benzamides via Oxidation with Potassium Permanganate in Liquid Ammonia

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Abstract Oxidation of aromatic aldehydes by KMnO₄ in liquid ammonia gives amides directly. The reaction proceeds satisfactorily when the aldehydes are activated by electron-withdrawing substituents on the ring.

Key words aldehydes, amides, aminals, oxidation, liquid ammonia

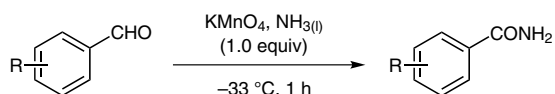
The amide functionality is ubiquitous, and amides are also versatile educts in organic synthesis; thus there is continuing interest in their synthesis.¹ Of particular interest are possibilities of direct conversion of simple substrates into amides which bypass conventional routes, such as reaction of activated carboxylic acids with ammonia or partial hydrolysis of nitriles. In recent years there have been many reports on direct conversion of aromatic aldehydes and benzylic alcohols into amides via oxidative reactions in the presence of an ammonia source. Thus treatment of aromatic aldehydes with aqueous ammonia and iodine and hydrogen peroxide gave amides in excellent yields.² Other ways to afford this transformation consist of conversion of aldehydes and primary alcohols into amides via reaction with ammonium carbonate and *tert*-butyl hydroperoxide with Et₄N⁺I⁻ catalyst.³ Oxidative conversion of aldehydes and alcohols into amides via reaction with oxygen and various ammonia sources catalyzed by transition metals is described in a few papers.⁴ New examples of oxidative amidation of toluenes and styrenes⁵ and oxidative acylation with aldehydes⁶ have been reported recently.

Amongst a plethora of solvents used in organic synthesis and industry, liquid ammonia is of great interest because of its unique properties. Owing to its basicity and hydrogen-bond-accepting ability, liquid ammonia solvates efficiently inorganic cations such as Li⁺, Na⁺, K⁺; whereas, in spite of its formal protic character, it does not efficiently solvate anions.⁷ As a consequence it behaves as a dipolar aprotic solvent, dissolving inorganic salts and sodium or potassium salts of organic anions: alkoxides, enolates, and carbanions.⁷ Moreover, in liquid ammonia solution, anions

of these salts exhibit high nucleophilic activity; thus it is the solvent of choice for many reactions involving carbanions⁸ and for industrial organic synthesis.⁹ Additionally, it is inexpensive, easy to handle under anhydrous conditions, and easy to remove and recover. Although ammonia can be oxidized and its oxidation to nitric acid is a major industrial process, liquid ammonia is resistant to common oxidants and forms stable solutions of potassium permanganate even at room temperature under pressure.¹⁰ A KMnO₄ solution in liquid ammonia has been efficiently applied to oxidation of σ^H adducts of carbanions and other nucleophiles to nitroarenes.¹¹

Ammonia is also a moderately active nucleophile capable of forming σ^H adducts with highly electrophilic arenes, particularly azines, and so a solution of KMnO₄ in liquid ammonia is an efficient agent for oxidative amination of such arenes.¹⁰ We therefore predicted that treatment of an aldehyde with such a solution should result in addition of ammonia to the carbonyl group followed by oxidation of the adduct to form an amide. Indeed, when solid KMnO₄ was added to a solution of benzaldehyde in liquid ammonia the characteristic color of the solution slowly turned dark brown and, after evaporation of ammonia, the residue contained benzamide. After experimentation a convenient procedure was found in which the aldehyde was dissolved in liquid ammonia and, after stirring for one hour, KMnO₄ (1.0 or 1.5 equiv) was added and the mixture stirred for another hour. Excess permanganate was quenched with Na₂SO₃ (3 equiv), the ammonia was evaporated, the residue treated with 6 M hydrochloric acid, and the product extracted with ethyl acetate or filtered, washed with aqueous NaHCO₃, and the product recrystallized. The results of these experiments are presented in Table 1.¹²

Aldehydes containing a halide or electron-withdrawing substituents gave the corresponding amides in good yields (Table 1, entries 2–8). The yields for nitrobenzaldehydes are somewhat lower due to side reactions and partial decomposition. Benzaldehyde and 1-naphthylcarboxyaldehyde gave the corresponding amides in moderate yield, together with small amounts of carboxylic acids. A careful analysis of the mass balance for the reaction of 4-cyanobenzalde-

Table 1 Direct Oxidation of Aromatic Aldehydes to Amides

Entry	Aldehyde (R)	Yield (%)
1	H	39
2	4-Cl	75 ^a
3	4-Br	61
4	3-Br	63
5	4-NC	78 ^b
6	4-O ₂ N	50
7	3-O ₂ N	56 ^c
8	4-F ₃ C	54
9	2,3-(CH=CH) ₂	42 ^c
10	2-F	24
11	3-F	21
12	4-MeO	0 ^d

^a Yield of carboxylic acid: 15%.^b Yield of carboxylic acid: 20%.^c Conditions: 1.5 equiv KMnO₄.^d Yield of carboxylic acid: 85%.

hyde showed that, even in this case, side formation of the carboxylic acid occurred. Its sodium salt was isolated in 20% yield, together with a high yield of the expected 4-cyanobenzamide (78%). Similarly, 4-chlorobenzaldehyde gave a 75% yield of the amide and about 15% of the acid.

Aldehydes which are less prone to undergo nucleophilic addition to the carbonyl group, such as anisaldehyde (Table 1, entry 11), *ortho*-substituted aldehydes (2,3-dichlorobenzaldehyde, 9-anthracenecarboxaldehyde), or aliphatic aldehydes (*n*-octanal), form carboxylic acids exclusively. 2-Fluorobenzaldehyde is a borderline case as it gave a nearly equimolar mixture of 2-fluorobenzamide and 2-fluorobenzoic acid, together with some unreacted aldehyde.

All benzamides obtained in the reactions summarized in Table 1 are known compounds and their identity and purity was confirmed by ¹H NMR, ¹³C NMR, IR spectroscopy and comparison of melting points with literature data.

Oxidation of aldehydes with aqueous solutions of KMnO₄ at various pH values has been the subject of thorough mechanistic studies.¹³ However, the mechanistic conclusions from these studies cannot be extended to the oxidation in liquid ammonia. It appears that conversion of aldehydes into amides in liquid ammonia proceeds via initial addition of ammonia to the carbonyl group. This supposition is supported by the observed effects of substituents in the aromatic rings on the reaction. Oxidation of the hypothetical intermediates – aminals – could proceed via abstraction of hydride anions analogously to hypothetical

pathways of oxidation of σ^H adducts in oxidative nucleophilic substitution of hydrogen in arenes.¹⁴ On the other hand, when the dimethyl acetal of 4-bromobenzaldehyde was subjected to the reaction conditions, after acidic work-up only 4-bromobenzaldehyde was obtained. This result could be explained by much faster oxidation of the aminal than the acetal function by KMnO₄, or by assuming that the intermediate leading to an amide is an imine rather than an aminal.

In conclusion, we have described a simple and practical synthetic procedure by which aromatic aldehydes can be directly converted into benzamides. The scope of the reaction is somewhat limited by the concurrent formation of carboxylic acids, but these side products can be readily removed from the desired amides by simple washing with aqueous base.

Supporting Information

Supporting information for this article is available online at <http://dx.doi.org/10.1055/s-0034-1378920>.

References and Notes

- (a) Allen, C. L.; Wolliams, M. J. *Chem. Soc. Rev.* **2011**, *40*, 3405. (b) Valeur, E.; Bradley, M. *Chem. Soc. Rev.* **2009**, *38*, 606. (c) Humpwey, J. M.; Chamberin, A. R. *Chem. Rev.* **1997**, *97*, 2243.
- Shie, J.-J.; Fang, J.-M. *J. Org. Chem.* **2003**, *68*, 1158.
- Wang, G.; Yu, Q.-Y.; Chen, S.-Y.; Yu, X.-Q. *Org. Biomol. Chem.* **2014**, *12*, 414.
- Nie, R.; Shi, J.; Xia, S.; Shen, L.; Chen, P.; Hou, Z.; Xiao, F.-S. *J. Mater. Chem.* **2014**, *22*, 18115; and references cited therein.
- (a) Wang, Y.; Yamaguchi, K.; Mizuno, N. *Angew. Chem. Int. Ed.* **2012**, *51*, 7250. (b) Vanjari, R.; Guntreddi, T.; Nand Singh, K. *Org. Lett.* **2013**, *15*, 4908. (c) Sharif, M.; Gong, J.-L.; Langer, P.; Beller, M.; Wu, X.-F. *Chem. Commun.* **2014**, *50*, 4747.
- (a) Ekone-Kovi, K.; Wolf, C. *Org. Synth.* **2010**, *87*, 1. (b) Giguère-Bisson, M.; Yoo, W.-J.; Li, C.-J. *Org. Synth.* **2011**, *88*, 14. (c) Möhlmann, L.; Ludwig, S.; Blechert, S. *Beilstein J. Org. Chem.* **2013**, *9*, 602. (d) Wang, W.; Zhao, X.-M.; Wang, J.-L.; Geng, X.; Gong, J.-F.; Hao, X.-Q.; Song, M.-P. *Tetrahedron Lett.* **2014**, *55*, 3192.
- Reichardt, C. *Solvents and Solvent Effects in Organic Chemistry*; Wiley-VCH: Weinheim, **2003**, 3rd ed.
- Mąkosza, M. *Tetrahedron* **1968**, *24*, 175.
- (a) Ji, P.; Atherton, J.; Page, M. I. *Org. Biomol. Chem.* **2012**, *10*, 5732. (b) Ji, P.; Atherton, J.; Page, M. I. *J. Org. Chem.* **2011**, *76*, 1425. (c) Ji, P.; Atherton, J.; Page, M. I. *J. Org. Chem.* **2011**, *76*, 3286.
- (a) van der Plas, H. C.; Woźniak, M. *Croat. Chem. Acta* **1968**, *59*, 33. (b) Szpakiewicz, B.; Grzegorek, M. *Russ. J. Org. Chem.* **2004**, *40*, 829.
- Mąkosza, M.; Staliński, K. *Chem. Eur. J.* **1997**, *3*, 2025.
- General Procedure for the Oxidation of Aldehydes to Amides in Liquid Ammonia**
Under an argon atmosphere, liquid NH₃ (25 mL) was condensed in a two-neck round-bottom flask immersed in a dry ice cooling bath and equipped with a dry ice reflux condenser. Aldehyde

(7.34 mmol) was added, and the resulting solution (or suspension) was stirred for 1 h. KMnO_4 (7.34 mmol, 1.16 g) was added, the cooling bath was removed, and the reaction mixture was stirred for another hour with gentle reflux of NH_3 . Na_2SO_3 (22.0 mmol, 2.78 g) was added, the reflux condenser was removed, and the NH_3 was allowed to evaporate spontaneously. The dark-brown residue was treated with 6 M HCl (30 mL), and the resulting precipitate was filtered, washed with H_2O (100 mL) and sat. aq NaHCO_3 (20 mL). All products were recrystallized from EtOH.

3-Bromobenzamide

Colorless crystals; mp 152–156 °C (lit.^{6d} 157–159 °C). IR (KBr): ν_{max} = 3353, 3175, 1659, 1623, 1564, 1427, 1389, 1123, 1066, 901, 794 cm^{-1} . ^1H NMR (400 MHz): δ = 7.44 (1 H, t, J = 7.8 Hz), 7.56 (1 H, br s), 7.73 (1 H, ddd, J = 7.8, 2.0, 1.0 Hz), 7.90 (1 H, dm, J = 7.8 Hz), 8.08 (1 H, t, J = 1.8 Hz), 8.15 (1 H, br s). ^{13}C NMR (100 MHz): δ = 122.5, 127.5, 131.1, 131.4, 134.9, 137.4, 167.3.

(13) Wiberg, K. B.; Steward, R. J. *Am. Chem. Soc.* **1955**, 77, 1786.

(14) Mąkosza, M.; Staliński, K. *Tetrahedron Lett.* **1998**, 39, 3575.

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