Number 22, 1965 567

A Stable Methylzirconium Compound

By J. R. Surtees

(Division of Organic Chemistry, C.S.I.R.O., Chemical Research Laboratories, Melbourne, Australia)

We report the preparation of the first stable alkylzirconium compound, namely chlorobis(cyclopentadienyl)(methyl)zirconium (II) by the route shown in equation (1). A mixture of dichloro- μ -oxotetra(cyclopentadienyl)dizirconium (I)^{1,2} and trimethylaluminium (2 mols.) in benzene was heated at 65-70° for 3.5 hr. After evaporation of benzene, the residue was dissolved in tetrahydrofuran; subsequent treatment with light petroleum gave a solid whose proton magnetic resonance (p.m.r.) spectrum showed no significant Crystallization aluminium-methyl resonance. from benzene-petroleum afforded the methylzirconium compound (53%), decomp. 191—193°. Elemental analyses, molecular-weight determination and spectroscopic measurements were consistent with structure (II).

The p.m.r. spectrum (in benzene, with tetramethylsilane as internal reference) shows two sharp peaks at δ 0.42 (presumed to be due to ZrCH₃)³ and 5.78 p.p.m. (π -C₅H₅), of which the intensities

are in the ratio 3:10. Structure (III; R=Me) is excluded by the infrared spectrum, which lacks the intense absorption at 750—775 cm.⁻¹ associated with the Zr–O–Zr system.^{1,2,4}

The methylzirconium compound (II) is rapidly hydrolysed by water to give methane in quantitative yield, together with the oxide (I).

$$\begin{array}{ccc} \text{(I)} & (\text{C}_5\text{H}_5)_2\text{Zr}(\text{Cl}) \cdot \text{O} \cdot \text{Zr}(\text{Cl}) (\text{C}_5\text{H}_5)_2 \, + \, 2\text{Me}_3\text{Al} \\ \downarrow & \end{array}$$

(II)
$$2(C_5H_5)_2Zr(Cl)Me + Me_2Al \cdot O \cdot AlMe_2$$

(III)
$$(C_5H_5)_2ZrR\cdot O\cdot ZrR(C_5H_5)_2$$

While this work was in progress Brainina, Dvoryantseva, and Freidlina⁴ announced the preparation of two arylzirconium compounds (III; R = Ph, $p\text{-MeC}_6H_4$), by the reaction of aryl-lithium reagents with either (I) or dichlorobis(cyclopentadienyl)zirconium.

(Received, May 3rd, 1965; Com. 292.)

¹ E. M. Brainina, R. Kh. Freidlina, and A. N. Nesmeyanov, Doklady Akad. Nauk S.S.S.R., 1964, 154, 1113.

² A. F. Reid, J. S. Shannon, J. M. Swan, and P. C. Wailes, Austral. J. Chem., 1965, 18, 173.

² T. S. Piper and G. Wilkinson, J. Inorg. Nuclear Chem., 1956, 3, 104; similar data are reported for methyl derivatives of other transition metals.

⁴ E. M. Brainina, G. G. Dvoryantseva, and R. Kh. Freidlina, Doklady Akad. Nauk S.S.S.R., 1964, 156, 1375