

Hydroxynaphththyridine-Derived Group III Metal Chelates: Wide Band Gap and Deep Blue Analogues of Green Alq₃ (Tris(8-hydroxyquinolate)aluminum) and Their Versatile Applications for Organic Light-Emitting Diodes

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Abstract: A series of group III metal chelates have been synthesized and characterized for the versatile application of organic light-emitting diodes (OLEDs). These metal chelates are based on 4-hydroxy-1,5-naphththyridine derivatives as chelating ligands, and they are the blue version analogues of well-known green fluorophore Alq₃ (tris(8-hydroxyquinolinato)aluminum). These chelating ligands and their metal chelates were easily prepared with an improved synthetic method, and they were facially purified by a sublimation process, which enables the materials to be readily available in bulk quantity and facilitates their usage in OLEDs. Unlike most currently known blue analogues of Alq₃ or other deep blue materials, metal chelates of 4-hydroxy-1,5-naphththyridine exhibit very deep blue fluorescence, wide band gap energy, high charge carrier mobility, and superior thermal stability. Using a vacuum-thermal-deposition process in the fabrication of OLEDs, we have successfully demonstrated that the application of these unusual hydroxynaphththyridine metal chelates can be very versatile and effective. First, we have solved or alleviated the problem of exciplex formation that took place between the hole-transporting layer and hydroxynaphththyridine metal chelates, of which OLED application has been prohibited to date. Second, these deep blue materials can play various roles in OLED application. They can be a highly efficient nondopant deep blue emitter: maximum external quantum efficiency η_{ext} of 4.2%; Commission Internationale de L'Eclairage x, y coordinates, $\text{CIE}_{x,y} = 0.15, 0.07$. Compared with Alq₃, Beq₂ (beryllium bis(benzoquinolin-10-olate)), or TPBI (2,2',2''-(1,3,5-phenylene)tris(1-phenyl-1*H*-benzimidazole)), they are a good electron-transporting material: low HOMO energy level of 6.4–6.5 eV and not so high LUMO energy level of 3.0–3.3 eV. They can be ambipolar and possess a high electron mobility of $10^{-4} \text{ cm}^2/\text{V s}$ at an electric field of $6.4 \times 10^5 \text{ V/cm}$. They are a qualified wide band gap host material for efficient blue perylene ($\text{CIE}_{x,y} = 0.14, 0.17$ and maximum η_{ext} 3.8%) or deep blue 9,10-diphenylanthracene ($\text{CIE}_{x,y} = 0.15, 0.06$ and maximum η_{ext} 2.8%). For solid state lighting application, they are desirable as a host material for yellow dopant (rubrene) in achieving high efficiency (η_{ext} 4.3% and η_p 8.7 lm/W at an electroluminescence of 100 cd/m² or η_{ext} 3.9% and η_p 5.1 lm/W at an electroluminescence of 1000 cd/m²) white electroluminescence ($\text{CIE}_{x,y} = 0.30, 0.35$).

1. Introduction

Efficient electroluminescence (EL) was first reported by Tang and Van Slyke using green light-emitting tris(8-hydroxyquinolinato)aluminum (Alq₃).¹ Very few materials attract attention as much as Alq₃ does in organic light-emitting diodes (OLEDs).² Alq₃ has been used as a green emitter, a common electron-transporting material, and a host material for saturated green

and red fluorescent dopants.³ Moreover, due to the rigid ball-like geometry, high glass transition temperature ($T_g \sim 175^\circ\text{C}$),^{3c} and polymorphic nature,⁴ Alq₃ is readily sublimed to form amorphous thin films, which is beneficial to the fabrication and operation lifetime of OLEDs. To tune its fluorescence color to

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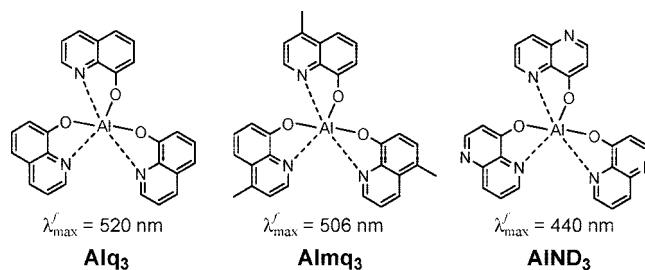
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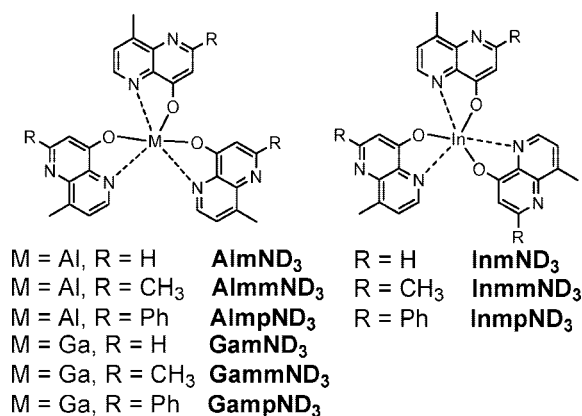
the red or blue region, some structurally modified **Alq₃** derivatives have been developed in past two decades to match the requirement in the RGB full color display or white light illumination application. However, most structurally modified **Alq₃** derivatives become less thermally stable and less volatile under reduced pressure. Very often, they fail in the fabrication of OLEDs by a thermal-vacuum-deposition process.

Photophysical and theoretical studies have already provided insight into the electron distribution of the HOMO/LUMO orbitals of **Alq₃**.^{2,5} Whereas the highest electron density of LUMO is found on pyridine ring, the electrons of HOMO orbitals are located mostly on the phenoxide side of the ligand. An electron-withdrawing substituent on the *para*-position of the phenoxide ring will deplete the HOMO electron density lowering energy level of the filled states. Following the same rationale, the substitution of a electron-donating group on the *para*-position of the pyridine ring will promote the LUMO electron density raising the energy level of the vacant states. Either structural modification results in increasing energy of electrons involved in frontier orbital transition ($\pi-\pi^*$) and an emission that is blue-shifted relative to that of parent **Alq₃** ($\lambda_{\text{max}}^f \sim 514$ nm in toluene and 524 nm in dichloromethane). Although there are several blue or near-blue light-emitting aluminum chelates, **Alq₃** derivatives showing deep blue fluorescence ($\lambda_{\text{max}}^f < 450$ nm) have not been realized yet.^{2,6} Surveying literature, we have found that a methyl substituent on the pyridine moiety or aza (nitrogen) replacement of “CH” of the phenoxide moiety of **Alq₃** is most attractive in blue-shifting fluorescence color, preserving the rigid and globular structure and good volatility of parent **Alq₃**. Examples are tris(4-methyl-8-quinolinato)aluminum (**AlmQ₃**) and tris(4-[1,5]naphthyridinolato)aluminum (**AIND₃**) showing blue-shifted fluorescence at 506 and 440 nm, respectively (Scheme 1). Whereas the greenish **AlmQ₃** was reported with EL performance,^{6c-f} application of blue **AIND₃** for OLEDs is still literature unknown to date.² Herein, in addition to the 4-hydroxy-1,5-naphthyridine (**ND**) aluminum chelate, which was efficiently prepared by our improved synthesis, we report the facile synthesis and full characterization of 4-hydroxy-8-methyl-1,5-naphthyridine (**mND**), 2,8-dimethyl-4-hydroxy-1,5-naphthyridine (**mmND**), and 4-hydroxy-2-phenyl-1,5-naphthyridine (**mpND**) metal chelates (Scheme 2). For **mND**, **mmND**, and **mpND** chelating ligands, a whole series of chelates with group III metals (aluminum, gallium, and indium) were also synthesized and characterized. We employ blue-shifting factors of both

Scheme 1. Chemical Structures of **Alq₃**, **AlmQ₃**, and **AIND₃**



Scheme 2. Chemical Structures of Group III Metal Chelates of 8-Hydroxy-1,5-naphthyridine Derivatives



a methyl substituent and aza (nitrogen element) as the replacement of “CH” in one structure. To our surprise, these metal chelates are in fact new substances and previously unknown. In addition to the thermal stability, they were characterized for deep blue fluorescence and EL in applications of OLEDs. Furthermore, due to their wide band gap and electron-deficient nature, we will demonstrate that these metal chelates are feasible as electron-transporting layer (ETL) material, the host material for highly efficient blue fluorescence dopants, or blue host material for yellow dopants to generate white EL in solid state lighting (SSL) applications.

2. Results and Discussion

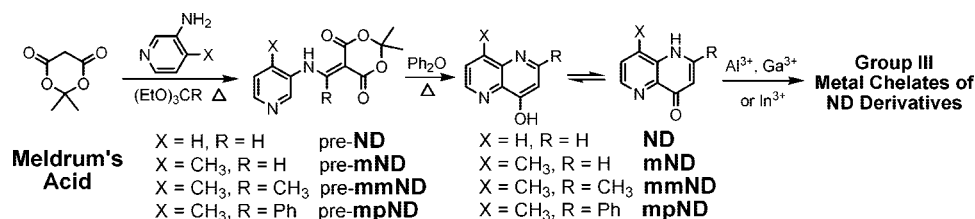
2.1. Synthesis and Structural Characterization. Parent 4-hydroxy-1,5-naphthyridine (or 1,5-naphthyridin-4-ol) is best known to be prepared by so-called “EMME synthesis” from 3-aminopyridine and diethyl ethoxymethylenemalonate.⁷ After the intramolecular cyclization, the hydrolysis of the resulting ester and thermo-decarboxylation of the acid afford 4-hydroxy-1,5-naphthyridine in unsatisfactory overall yields (18–31%).⁸ Alternatively, a shorter and more convenient procedure known as the Cassis method in the synthesis of a wide range of 4-*H*-quinolones (tautomeric forms of hydroxyquinoline) is using

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Scheme 3. Synthetic Routes to a Series of 8-Hydroxy-1,5-naphthyridine Metal Chelates

2,2,6-trimethyl-4*H*-1,3-dioxin-4-one (Meldrum's acid) as an effective diketone equivalent precursor.⁹ Therefore, we adopted the effective Cassis method in the synthesis of the new naphthyridine derivatives, **mND**, **mmND**, and **mpND** (Scheme 3). For comparison purposes in the study herein, parent chelating ligand **ND** was also synthesized by the more convenient Cassis method.

In the first step of synthesis, four pyridylaminomethylene Meldrum's acid derivatives (pre-**ND**, pre-**mND**, pre-**mmND**, pre-**mpND**) were all obtained in reasonably good yields, 87, 78, 71, and 58%, respectively. Among four species, relatively low yields (58%) of pre-**mpND** can be attributed to the bulky phenyl group that causes sterical hindrance in the formation of the methylene-bridge-head between the amino substituent and Meldrum's acid. In the following step of a ring closure reaction, heating Meldrum's acid derivatives in diphenyl ether afforded **ND**, **mND**, **mmND**, and **mpND** in 47–77% yields, respectively. Among them, **ND** (1,5-naphthyridin-4-ol or 8-hydroxy-1,5-naphthyridine) was obtained in lowest yields (47%), and it is due to formation of an undesired structural isomer (1,7-naphthyridin-4-ol) in the ring closure reaction. The same cause is believed to occur in the previous "EMME synthesis" of **ND**, which was prepared in even lower yields.⁸ Nevertheless, the methyl or phenyl substituents of **mND**, **mmND**, or **mpND** are simply from the starting materials of either 3-amino-4-picoline or triethyl orthoacetate/orthobenzoate. A different alkyl group or substituted aromatic ring can be readily incorporated onto **ND** with appropriate starting materials. Through the case of **ND**, it is conceivable that the Cassis method illustrated herein is more convenient and versatile. From the Cassis method, we obtained **ND** in ~40% overall synthetic yields, which is in fact better than 18–31% overall synthetic yields from the literature reported "EMME synthesis".⁸ Moreover, the advantage of the Cassis method is the simpler and more reliable synthetic operation, a two-step procedure instead of a four-step procedure. In "EMME synthesis", a tedious purification process has to be performed four times after each step in the synthetic sequence. The newly synthesized 4-hydroxy-1,5-naphthyridine derivatives were successfully converted into group III metal chelates. However, due to the different reactivity of the metal starting material and the different solubility of the metal chelate product, various reaction conditions and the isolation/purification method were adopted in the final metal chelation reactions, which is somewhat variant from the conventional method preparing **Alq₃** (see Experimental Section for details). Basically, the synthesis and purification of hydroxynaphthyridine group III metal chelates are not much difficult than those for the easily prepared and purified **Alq₃**. They can be readily obtained with volume production in a conventional synthetic laboratory.

These metal chelates were fully characterized by ¹H and ¹³C NMR, mass spectroscopy, and elemental analysis, and they were consistent with proposed structures. Particularly, a single crystal X-ray structure of **AIND₃** was obtained. Its ORTEP drawing is

displayed in Figure 1.¹⁰ Similar to that of **Alq₃**, the meridional (*mer*) configuration of **AIND₃** is clearly evident by the structure diagram and it is consistent with its complicated ¹H NMR spectrum. Interestingly, unlike **AlmND₃** or **GamND₃**, **InmND₃** exhibited a simple ¹H NMR spectrum (four sharp and well-separated proton resonances), an indication of the facial (*fac*) configuration of **InmND₃**. However, recent evidence has demonstrated that the simple ¹H NMR spectrum of **Inq₃** (and hence **InmND₃**) is due to the rapid fluxional transitions between *mer* and *fac* configurations on the NMR time scale.¹¹

The higher level of electron deficiency of **ND** than 8-hydroxyquinoline is evident by the significantly smaller *pK_a* 2.85 of **ND** than *pK_a* 5.13 of 8-hydroxyquinoline.¹² **ND** derivatives have been known for keto–enol tautomerism in polar organic solvents (see Scheme 3).⁷ We found that the extent of naphthyridone tautomeric forms (and hence the *pK_a* value) can be gauged by the ¹H NMR signal (chemical shift) of the proton next to the hydroxy substituent of 8-hydroxyquinoline, **ND**, and **mND**, which locates at 7.09, 6.52, and 6.50 ppm, respectively (see Figure 2). We expect that even smaller *pK_a* values are present for **mmND** and **mpND**. The smaller *pK_a* values (due to the electron deficiency) render **ND** derivatives weaker δ -donors on the phenoxide side of the chelating ligand. As the result of a weaker δ -donor, the average Al–O bond distance of **AIND₃** (1.867 ± 0.03 Å) is longer than that (1.856 ± 0.02 Å) of **Alq₃**.⁴ On the pyridine side of the chelating ligand, the average Al–N bond distance of **AIND₃** (2.025 ± 0.03 Å) is shorter than that (2.051 ± 0.02 Å) of **Alq₃**, and this can be attributed to the "seesaw-like" binding mode of a bidentate chelating ligand. Otherwise, it can be attributed to the stronger N δ -donor in **AIND₃** than in **Alq₃** because of the electron-donating methyl substituent *para* to the N δ -donor. We believe that the strength of the δ -donor (or the electron deficiency) of chelating ligands is one of the reasons why aluminum metal chelates of **mND**, **mmND**, or **mpND** all have a shorter fluorescence peak wavelength and are better in blue color purity than **AIND₃** or **Alq₃** (see Table 1 for fluorescence data).

2.2. Photophysical Properties and Energy Levels. Whereas it is the desired deep blue color ($\lambda_{\text{max}}^{\text{fl}}$ 415–417 nm) of **AlmND₃**, **AlmmND₃**, and **AlmpND₃**, a less satisfactory sky blue fluorescence ($\lambda_{\text{max}}^{\text{fl}}$ 433 nm) of the previously known **AIND₃** is clearly shown in their solution fluorescence images (see Figure 3). In the solid state, all four aluminum chelates display red-shifted fluorescence wavelength $\lambda_{\text{max}}^{\text{fl}}$ 425 nm (**AlmpND₃**), 431

(10) Crystal data for **AIND₃·CH₂Cl₂**: C₂₅H₁₇AlCl₂N₆O₃; *F_w* = 547.33, Triclinic, *P*1, *Z* = 2, *F*(000) = 560. Cell dimensions: *a* = 7.9946(7) Å, *b* = 12.0879(11) Å, *c* = 13.1153(12) Å, α = 70.346(2)°, β = 82.950(2)°, γ = 84.082(2)°, *V* = 1181.98(18) Å³, $2\theta_{\text{max}}$ = 50.0°, ρ_{calcld} = 1.538 mg/m³. Of 8777 reflections, 4157 were independent, 334 parameters, *R*(*F*_o) = 0.0588 (for reflections with *I* > 2 σ (*I*)), *R_w*(*F*_o) = 0.1663 (for reflections with *I* > 2 σ (*I*)). The GoF on *F*² was 0.977.

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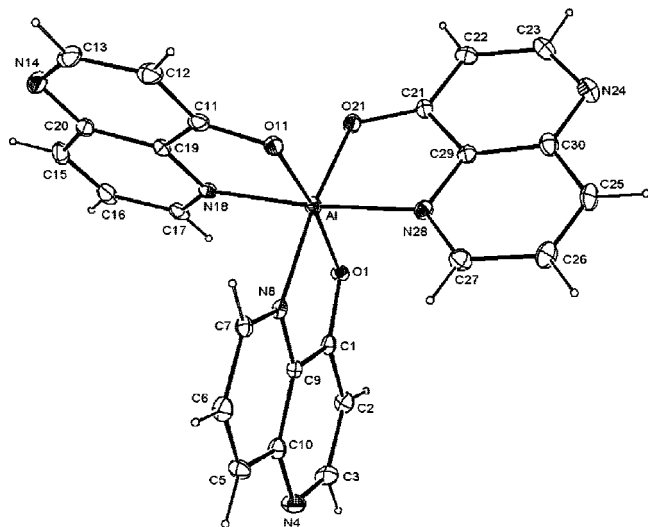


Figure 1. Molecular structure of **AlND₃** determined by X-ray diffraction analysis.

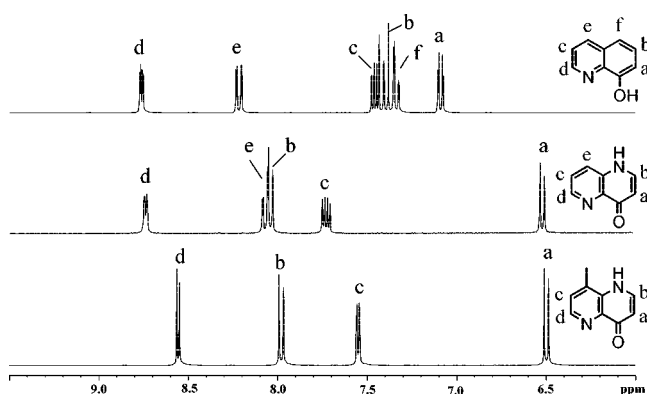


Figure 2. ¹H NMR spectra of 8-hydroxyquinoline, **ND**, and **mND** in CD₃OD from top to bottom, respectively.

Table 1. Optical and Thermal Properties of Metal Chelates of 4-Hydroxy-1,5-naphthyridine Derivatives

metal chelates	solution ^a		solid		Φ_f (%)	T_g (°C)	T_c (°C)	T_m (°C)	LUMO/HOMO (eV) ^b
	$\lambda_{\max}^{ab}$ (nm)	$\lambda_{on\ set}^{ab}$ (nm)	$\lambda_{\max}^{fl}$ (nm)	Φ_f^{fl} (%)					
Alq₃	388, 443	524	20	516	40	174	367	412	3.1/5.9
AlND₃	341, 381	433	47	447	45	122	262	412	-
AlmND₃	338, 370	415	45	431	43	196	277	420	3.0/6.4
						387			
AlmmND₃	335, 372	416	45	419	39	233	^c	436	-
AlmpND₃	326, 376	417	45	425	6	204	278	370	-
GamND₃	341, 374	431	45	439	41	^c	^c	414	3.2/6.5
GammND₃	339, 376	432	42	434	52	212	^c	430	-
GampND₃	327, 383	432	25	439	11	185	275	381	-
InmND₃	344, 382	436	37	445	12	183	^c	361	3.3/6.5
InmmND₃	332, 381	437	36	436	22	221	^c	372	-
InmpND₃	329, 386	437	23	446	8	186	224	370	-

^a In dichloromethane. ^b HOMO energy was determined as the edge of HOMO energy level or the ionization potential of the material; LUMO energy was determined as the lowest photoexcitation state energy from the on-set absorption energy in absorption spectra. ^c Not observed.

nm (**AlmND₃**), and 447 nm (**AlND₃**), except $\lambda_{\max}^{fl}$ 419 nm of **AlmmND₃** (Table 1). In dichloromethane solution, all four aluminum chelates have reasonably good fluorescence quantum yields (Φ_f) around 45–47% but more or less suffer from fluorescence concentration quenching in the solid state (Table 1).

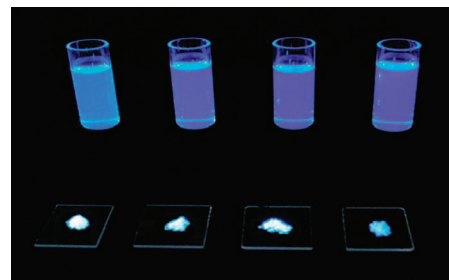


Figure 3. Solution (in dichloromethane) and solid state fluorescence image of **AlND₃**, **AlmND₃**, **AlmmND₃**, and **AlmpND₃** from left to right, respectively.

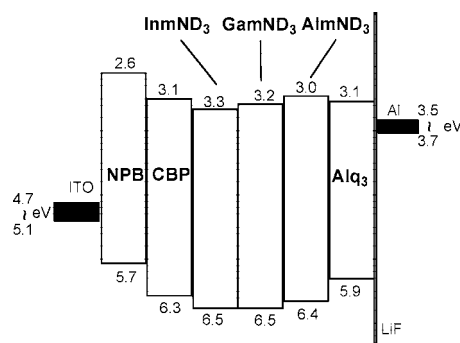


Figure 4. Energy alignment of **AlmND₃**, **GamND₃**, **InmND₃**, **Alq₃**, **NPB**, and **CBP** in OLED devices based on ITO anode and LiF/Al cathode.

Among them, **AlmpND₃** has the most severe fluorescence quenching in the solid state with a Φ_f of only 6%. Such solid-state fluorescence quenching is universal for **mpND**-based chelates (**AlmpND₃**, **GampND₃**, and **InmpND₃**). This can be rationalized by the molecular contact (probably π – π interaction) through protruded phenyl substituents in the solid state. With no exception of the three chelating ligands **mND**, **mmND**, and **mpND**, fluorescence of these metal chelates was observed to be reduced with increasing atomic number of the metal ion from Al to Ga and then In (see Φ_f data in Table 1), known as the heavy atom effect that increases in the rate of intersystem crossing.^{2,13}

In the search for the origin of the wide band gap nature, **AlmND₃**, **GamND₃**, and **InmND₃** were found to have a HOMO energy level around 6.4–6.5 eV, which is significantly lower than 5.9 eV of **Alq₃** (Figure 4). The LUMO energy level of three **mND** metal chelates is around 3.0–3.3 eV, which is similar to or just a bit lower than 3.1 eV of **Alq₃**. This is perfectly logical because the prominent electron deficient feature (N aza substituent) locates on the HOMO of the molecule, which is the pyridin-4-olate ring of the chelating ligand, instead of the LUMO of three **mND** metal chelates, which is the *para*-methylpyridine (or 4-picoline) ring of the chelating ligand. It can be further identified that either HOMO or LUMO is stabilized by the higher atomic number of the central metals, Al, Ga, and In. Such a stabilization effect happens more prominently in the LUMO than in the HOMO. As a result, the energy band gap of three **mND** metal chelates decreases in the order **AlmND₃** > **GamND₃** > **InmND₃**. Relative to that of **Alq₃**, a significantly low HOMO energy level and moderately low LUMO energy level enlarge the energy band gap of these hydroxynaphthyridine-based metal chelates.

(13) Burrow, P. E.; Sapochak, L. S.; McCarty, D. M.; Forrest, S. R.; Thompson, M. E. *Appl. Phys. Lett.* **1994**, *64*, 2718.

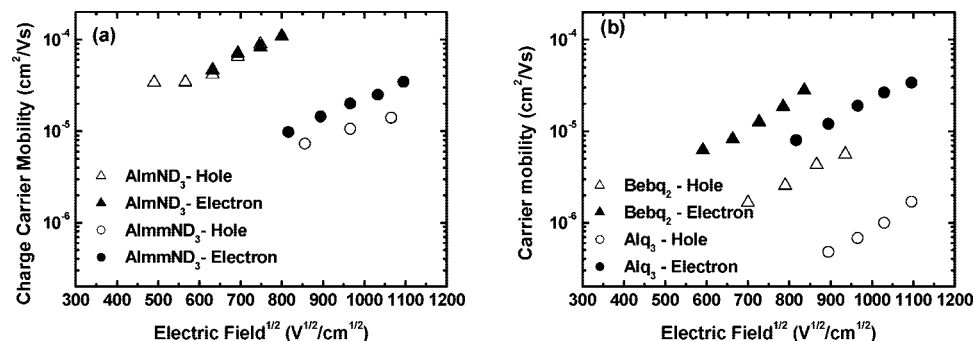


Figure 5. Hole and electron mobility vs the square root of the applied electric field of **AlImND₃** and **AlImND₃** (a) and **Alq₃** and **Bebq₂** (b).

2.3. Charge-Transporting Properties and Charge Carrier Mobilities. In terms of HOMO and LUMO energy levels, hydroxynaphthylidine-based metal chelates are better hole-blocking or electron-transporting materials than **Alq₃** or other deep blue fluorophores. Most known deep blue fluorophores are based on triarylamines, such as NPB, or nonheteroatom-containing polycyclic aromatic hydrocarbons (PAHs), such as anthracene, perylene, pyrene, or spirobifluorene compounds. Compared with those of hydroxynaphthylidine metal chelates, such nonmetal chelate deep blue fluorophores have significantly higher HOMO and LUMO energy levels of 5.1–5.8 eV and 2.0–2.7 eV (assuming band gap energy is ca. 3.1 eV) below vacuum level, respectively.¹⁴

The HOMO–LUMO energy level of the material shown above is one of the determining factors for electron-transporting or hole-transporting properties of OLED materials. The charge carrier (electron or hole) mobility is the other decisive characteristic that influences the charge-transporting nature and hence the efficiency performance of OLEDs. The relative magnitude of the electron and hole mobility of the material indicates the extent of charge balancing, the efficiency of charge recombination, and hence the EL efficiency of OLEDs. We measured the intrinsic charge carrier (hole or electron) mobility in a bulk film (0.6–1 μm) of **AlImND₃** and **AlImND₃** using the optical time-of-flight (TOF) technique that has been described before. In addition, we also took a measurement on **Alq₃** and **Bebq₂** with the same measuring system for comparison and accuracy checking. Figure 5 shows the field dependence of the hole and electron drift mobility of **AlImND₃**, **AlImND₃** (Figure 5a), **Alq₃**, and **Bebq₂** (Figure 5b). First, the electron mobility of **Alq₃** was determined to be $\sim 10^{-5}$ cm²/V s at an electric field of 6.4×10^5 V/cm, which agrees with those obtained previously.¹⁵ Second, the electron mobility of **Bebq₂** was determined to be $\sim 10^{-4}$ cm²/V s at the same electric field, which is 1 order of magnitude higher than that of **Alq₃** and consistent with the report that **Bebq₂** is a better electron-transporting material than **Alq₃**.¹⁶ Devices with an electron-transporting layer (ETL) of **Bebq₂** have been demonstrated with lower driving voltage and longer operation lifetime, when compared with ones with **Alq₃** as an ETL.¹⁶ Third, we found that the charge carrier (either hole or electron) mobility of **AlImND₃** is higher than that of **AlImND₃** by nearly 1 order of magnitude (Figure 5a). In fact, checking

the extrapolated data in the range of electric field of $(3.6\text{--}6.4) \times 10^5$ V/cm, the charge carrier (either hole or electron) mobility of **AlImND₃** is the highest among all (Figure 5a and 5b). Therefore, **AlImND₃** is probably a better electron-transporting material than **AlImND₃** because of its high electron mobility. In addition, **AlImND₃** seems to be the only material that is ambipolar because of its very similar hole and electron mobility determined by TOF technique. On the other hand, **AlImND₃** may be a better nondopant deep blue emitter because of its low hole mobility, which limits the amount of hole carrier on **AlImND₃** and enhances the charge balance in normally hole-dominated devices, one crucial factor for high efficiency OLEDs. As shown in the following sections (2.7 and 2.8), our OLED results are consistent with the forgoing derivation from the TOF data of **AlImND₃** and **AlImND₃**.

2.4. Thermal Properties. Figure 6 shows the differential scanning calorimetry (DSC) thermograms of **Alq₃**, **AlND₃**, **AlImND₃**, **GamND₃**, **AlImND₃**, and **GammND₃**. In DSC measurements, these metal chelates were first taken from the sample that was purified by a sublimation process. With such prethermal-annealed samples (sublimed-scarped samples), we found that DSC thermograms often show nothing but weak endothermic step transitions, indicative of the glass phase transition temperature (T_g). To reveal DSC signals of other phase transition temperatures, such as crystallization temperature (T_c), melting temperature (T_m), or polymorphic phase transition temperature (T_p), the sublimed-scarped samples of metal chelates were redissolved in dichloromethane and then evaporated until dryness under reduced pressure at room temperature. In Figure 6, each metal chelate is displayed with two types of DSC traces. First the heating scan, positioned as the top scan, has a heating temperature beyond the large endothermic signal, T_m . The other scans marked with the sequence number of heating scans or cooling scans in Figure 6 are DSC traces with a measuring temperature less than T_m , except for **AlmpND₃**, **GampND₃**, and **ImnpND₃** (see Figure S1 for their DSC traces). Two types of heating thermograms were displayed because these metal chelates often exhibit thermal decomposition right after their melting transition, which was evident from observations from a polarized optical microscope (POM). With samples prepared under such conditions, **Alq₃** shows T_g (on-set T_g) at 174 °C, a broad and small endotherm peaking at ~ 355 °C (assigned as T_p) immediately followed by an exotherm peaked at 367 °C (T_c), and finally a large endothermic signal at 412 °C (T_m). Whereas a T_g of 174 °C and T_m of 412 °C are rather consistent with literature data for **Alq₃**,^{3c,4,6f,17} the coupled endotherm and

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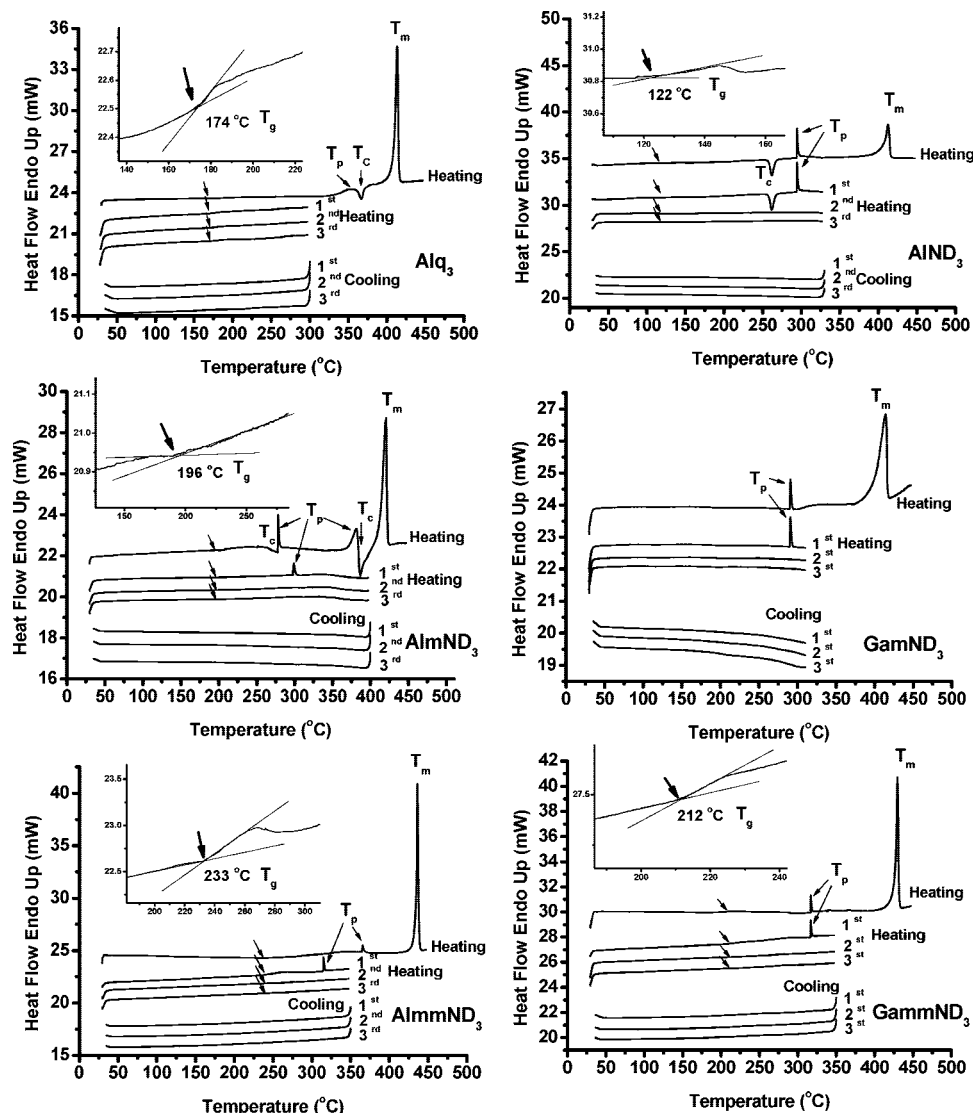


Figure 6. DSC thermograms of Alq_3 , AlND_3 , AlmND_3 , GamND_3 , AlmmND_3 , and GammND_3 .

exotherm have been shown before, although they were reported at higher temperatures of 393 and 396 °C, respectively.^{4,6f,18} It is interesting to know that Alq_3 will be vaporized (sublimed) when heated around melting temperatures at atmospheric pressure.^{6f} We have a similar observation (under POM) for Alq_3 and most metal chelates reported herein as well. Such an observation is indirect evidence that these new metal chelates preserve their superior volatility as well as Alq_3 . Unlike those structurally modified Alq_3 blue derivatives, volatile hydroxynaphthridine metal chelates reported herein have no problem in the fabrication of OLED by thermal-vacuum-deposition processes.

Alq_3 has been well-known for its polymorphic and racemic nature and its multiple phase transitions that have been studied

in great detail recently.^{4,6f,18} It has been suggested that the exotherm observed for Alq_3 in DSC traces is due to crystallization, indicative of the instability of the glassy phase of Alq_3 . Some new hydroxynaphthridine metal chelates, such as AlmND_3 , GamND_3 , GammND_3 , InmND_3 , and InmmND_3 , show no discernible exotherm in DSC traces, a good sign of the morphological stability of their glassy phase. Except for AlND_3 ($T_g \sim 122$ °C) and GamND_3 (no detectable T_g in DSC traces), all metal chelates reported herein show T_g values around 183–233 °C, which is higher than 174 °C for Alq_3 (Table 1). Higher T_g 's of metal chelates often imply a higher morphological stability or a glass phase stability, which is beneficial to the operation lifetime of the multiple-thin-film OLEDs. Furthermore, for all aluminum chelates but AlmpND_3 , higher melting temperatures than Alq_3 were observed by DSC (Table 1). This is a plus because the melting transition of these aluminum chelates was rapidly followed by the thermal decomposition. A higher melting temperature means a higher thermal stability in the case of the aluminum chelates studied herein.

Metal chelates based on mpND are different from the rest. Among all metal chelates, AlmpND_3 , GampND_3 , and InmpND_3

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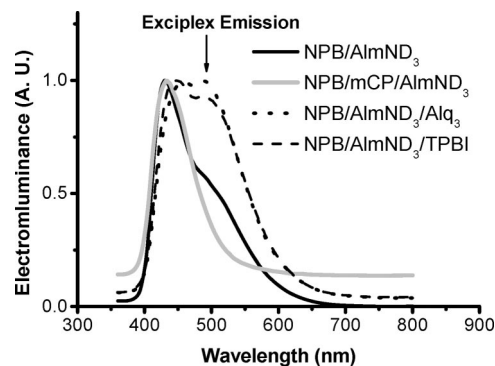


Figure 7. EL spectra of four NPB/AlmND₃-containing OLEDs: ITO/multiple organic layers/LiF/Al.

were found with relatively low T_g 's of 185–204 °C, and they are easily recognizable in repeated heating scans in DSC measurements (Figure S1). Their T_g 's were always observed prior to a cold crystallization peak without discernible polymorphic phase transition signals (as T_p marked for AlND₃, AlmND₃, and AlmmND₃ in Figure 6). AlmpND₃, GampND₃, and InmpND₃ are also a few exceptions that they do not decompose following the melting transition. Such a difference in thermal properties of mpND-based metal chelates can be all attributed to the protruded phenyl substituent on a symmetric and nearly globular structure of naphthyridinolate metal chelates.

2.5. Elimination of Adverse Exciplex Formation in Non-Doped Blue OLEDs. Deep blue fluorescence of AlmND₃ or AlmmND₃ is very attractive for highly demanding blue OLEDs. A simple nondoped bilayer EL device ITO/NPB(30 nm)/AlmND₃(30 nm)/LiF(5 nm)/Al(150 nm) was first fabricated by sequential thermal-vacuum-deposition of NPB, AlmND₃, LiF, and Al onto ITO (indium tin oxide)-coated glass substrate. Unfortunately, in addition to the emission band peaking around 430–450 nm of AlmND₃, a pronounced emission band around 480–540 nm appears in the EL spectrum of the device (Figure 7).

Such a long wavelength EL devastates the EL efficiency and also impairs the blue color purity of OLEDs. A similar observation was also found for AlmmND₃ OLEDs. To clarify the origin of the long wavelength EL, three more control devices were then fabricated: ITO/NPB(30 nm)/mCP(30 nm)/AlmND₃(30 nm)/LiF(5 nm)/Al(150 nm), ITO/NPB(30 nm)/AlmND₃(30 nm)/Alq₃(30 nm)/LiF(5 nm)/Al(150 nm), and ITO/NPB(30 nm)/AlmND₃(30 nm)/TPBI(30 nm)/LiF(5 nm)/Al(150 nm). From their EL spectra (Figure 7), we can firmly conclude that the EL around 480–540 nm is due to the exciplex emission occurring at the interface of NPB and AlmND₃. It is rather common that electron-deficient materials (such as Alq₃ and 1,3,4-oxadiazole compounds) have a propensity to form exciplex with electron-rich hole-transporting materials (such as NPB triarylamine species).¹⁹ By lowering the HOMO energy level of the hole-transporting material (i.e., reducing the HOMO energy level difference between the hole-transporting material and AlmND₃),^{19d} the problem of the exciplex associated with AlmND₃ can be largely alleviated. As shown in Figure 7, having

Table 2. Electroluminescence Characteristics of Nondoped AlmND₃ OLEDs Containing CBP with Different Layer Thickness^a

CBP thickness [nm]	max. luminance, voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	max. efficiency [%; cd/A, lm/W]	$\lambda_{\text{max}}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
3	5555, 15	119, 0.39, 5.68	0.53, 0.79, 0.33	456	0.18, 0.20
5	4971, 15	162, 0.86, 5.63	1.08, 1.01, 0.45	450	0.15, 0.11
10	5070, 15	277, 1.63, 8.37	1.79, 1.51, 0.53	448	0.15, 0.10
15	5606, 15	315, 1.63, 9.55	1.74, 1.68, 0.53	452	0.16, 0.12

^a Devices have the configuration of ITO/NPB(40 nm)/CBP(x nm)/AlmND₃(30 nm)/Alq₃(20 nm)/LiF(0.5 nm)/Al(150 nm), x = 3, 5, 10, and 15, respectively. ^b At current density of 20 mA/cm².

a low HOMO energy level around 6.1 eV,²⁰ arylamine mCP (1,3-di(9H-carbazol-9-yl)benzene) is effective in preventing the exciplex emission as evident in device ITO/NPB(30 nm)/mCP(30 nm)/AlmND₃(30 nm)/LiF(5 nm)/Al(150 nm), although such a device is very poor in EL efficiency and brightness. Also having a low HOMO energy level around 6.3 eV,²¹ another hole-transporting material CBP (4,4'-di(9H-carbazol-9-yl)bi-phenyl) was thus inserted between NPB and AlmND₃. By the variation of CBP layer thickness, we have successfully eliminated the exciplex emission and optimize the performance of the blue devices ITO/NPB(40 nm)/CBP(x nm)/AlmND₃(30 nm)/Alq₃(20 nm)/LiF(5 nm)/Al(150 nm) optimized with x = 10 (Table 2). From such CBP-inserted nondoped AlmND₃ OLEDs, we have successfully achieved deep blue ELs (CIE_{x,y} = 0.15, 0.10) with a reasonably good external quantum efficiency (η_{ext}) reaching 1.79% (or 1.63% at 20 mA/cm²) and a maximum brightness of 5070 cd/m² (or 277 cd m⁻² at 20 mA/cm²) (Figure 8). To the best of our knowledge, we believe that a similar exciplex problem of AlmND₃ has been found for AlND₃ before and the problem has not been solved until this study.²²

2.6. As Electron-Transporting Material in Non-Doped AlmND₃ OLEDs. Similar to green emitter Alq₃, deep blue emitter AlmND₃ potentially can be used as an electron-transporting material in OLEDs. To gauge such viability, four nondoped AlmND₃ OLEDs containing four different ETL materials were fabricated: ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃(30 nm)/ETL(20 nm)/LiF(5 nm)/Al(150 nm), where the ETL material is TPBI (2,2',2''-(1,3,5-phenylene)tris(1-phenyl-1H-benzimidazole), Beq₂ (beryllium bis(benzoquinolin-10-olate), Alq₃, or AlmND₃. As data have shown in Table 3 and Figure 9, in terms of η_{ext} , nondoped AlmND₃ OLEDs are most efficient when AlmND₃ is employed as the ETL material in OLEDs, although it is not as bright as others. However, in terms of power efficiency (η_p), nondoped AlmND₃ OLEDs having TPBI or Beq₂ are more efficient than ones using Alq₃ or AlmND₃ as the ETL materials (Table 3). We notice that AlmND₃ provide OLEDs a relatively low current density and a relatively high turn-on voltage (Figure 9), which are comparable with those of Alq₃ but inferior to those of TPBI or Beq₂. Nevertheless, such results validate the usage of AlmND₃ as an ETL material in OLEDs.

2.7. AlND₃, AlmND₃, AlmmND₃, and AlmpND₃ for High Efficiency Non-Doped Deep Blue OLEDs. Having solved the problem of exciplex EL and demonstrated the electron-

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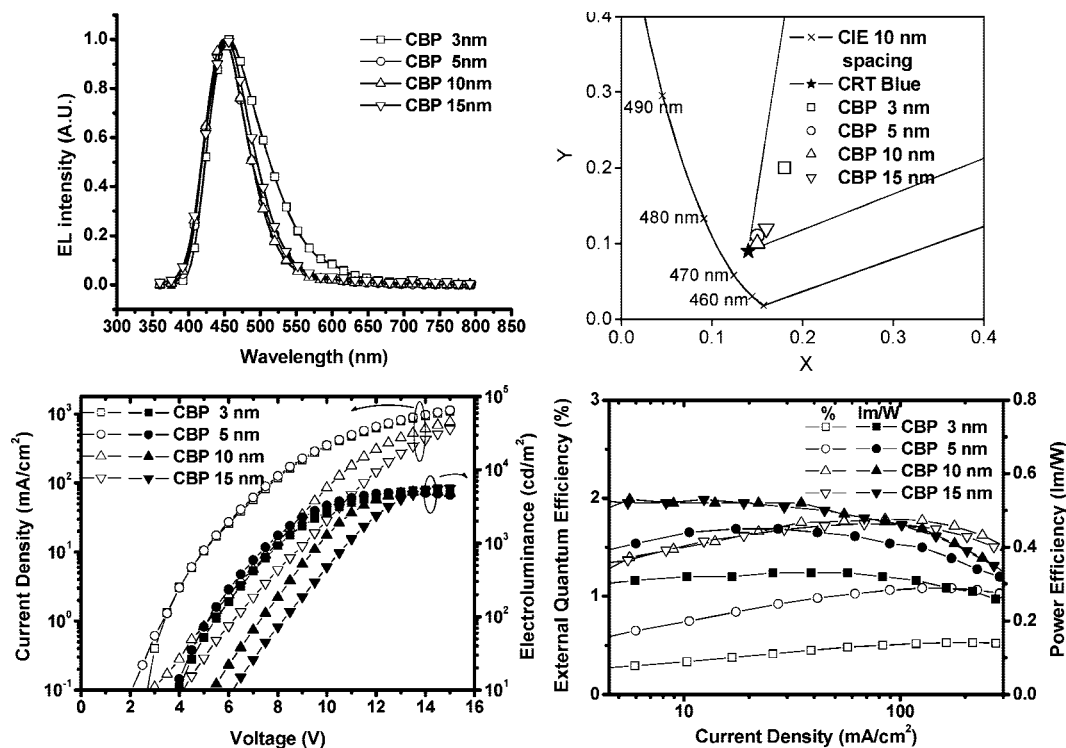


Figure 8. EL characteristics of ITO/NPB(40 nm)/CBP(*x* nm)/AlmND₃(30 nm)/Alq₃(20 nm)/LiF(0.5 nm)/Al(150 nm), *x* = 3, 5, 10, and 15, respectively.

Table 3. Electroluminescence Characteristics of Nondoped AlmND₃ OLEDs with Various ETL Materials^a

ETL	max. luminance and voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	max. efficiency [% , cd/A, lm/W]	$\lambda_{\max}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
TPBI	6000, 15	240, 1.15, 6.2	1.86, 1.79, 0.86	454	0.16, 0.13
Bebq ₂	7140, 15	340, 1.67, 5.6	1.81, 1.84, 0.98	458	0.15, 0.13
Alq ₃	5070, 15	277, 1.63, 8.4	1.79, 1.51, 0.53	448	0.15, 0.10
AlmND ₃	5240, 15	310, 1.85, 6.8	1.96, 1.65, 0.84	452	0.15, 0.10

^a ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃(30 nm)/ETL(20 nm)/LiF(5 nm)/Al(150 nm). ^b At 20 mA/cm².

transporting property of AlmND₃, we are ready to examine and find the best aluminum chelates of ND, mND, mmND, and mpND for nondoped deep blue OLEDs. Having AlmND₃ as the electron-transporting layer, four nondoped OLEDs were based on AIND₃, AlmND₃, AlmmND₃, and AlmpND₃: ITO/NPB(40 nm)/CBP(10 nm)/Blue Al chelates(30 nm)/AlmND₃(20 nm)/LiF(0.5 nm)/Al(150 nm) were fabricated and characterized (Table 4 and Figure 10).

As it can be anticipated from their solution or solid-state fluorescence wavelength, the AlmND₃ OLED showed a much deeper blue EL with 1931 Commission Internationale de L'Eclairage *x, y* coordinates (CIE_{*x, y*} = 0.15, 0.09) than that of AIND₃ OLED (CIE_{*x, y*} = 0.15, 0.19). However, we do not anticipate that AlmND₃ OLED performs better than does AIND₃ OLED considering its fluorescence quantum yield. Since fluorescence quantum yields of AIND₃ are higher than those of AlmND₃ either in solution or in the solid state, we attribute such a result to the charge balancing in the OLED, which has a great influence on the efficiency of the charge-recombination in the OLED. More surprisingly, the AlmmND₃ OLED shows the shortest EL wavelength ($\lambda_{\max}^{\text{EL}}$ ~436 nm), the highest η_{ext} of 4.18% (or 4.11% at 20 mA/cm²), and the brightest EL of 445 cd/m² at 20 mA/cm². Considering the structural difference among AIND₃, AlmND₃, and AlmmND₃, it is hard to conceive

that an insignificant methyl substituent or the number of methyl substituents can exert that much on η_{ext} and EL brightness. Plausibly, among four aluminum chelates, we may recognize AlmmND₃ as an exceptional one based on the red-shifted fluorescence from solution to the solid state. The red-shifting fluorescence is 443→447 nm, 415→431 nm, 416→419 nm, and 417→425 nm for AIND₃, AlmND₃, AlmmND₃, and AlmpND₃, respectively. Energywise, this corresponds to a red-shifting energy of 723, 895, 172, and 451 cm⁻¹, respectively for four aluminum chelates. AlmmND₃ has the smallest red-shifted fluorescence among all. Accordingly, we surmise that, in the solid state, AlmmND₃ aggregates in a quite different fashion from the other aluminum chelates. Compared with other aluminum chelates, such aggregation enables high efficiency of charge recombination (or charge balance) and thus the η_{ext} (or η_{P}) of AlmmND₃ OLEDs. For AlmmND₃, the low hole mobility (lower than its electron mobility) helps in balancing the charge of AlmmND₃ OLEDs.

We have also examined the performance of blue OLEDs with GamND₃, GammND₃, InmND₃, or InmmND₃ as the nondoped light-emitting layer (Table S1). As it can be anticipated from the fact that a heavy atom will quench the fluorescence, indium chelates performed the worst in terms of EL efficiency or brightness. However, this is not exactly the case for gallium chelates. Comparing data in Table 4 and Table S1, the GammND₃ OLED is worse than the AlmmND₃ OLED but the GamND₃ OLED is better than the AlmND₃ OLED. Nevertheless, once again, we verified the same trend of OLED performance; namely, mmND is better than mND for gallium or indium metal chelates as nondoped deep blue light-emitting materials for OLEDs.

2.8. Nondoped AlmmND₃ OLEDs with Various Metal Chelates as Electron-Transporting Layer. Knowing all group III metal chelates are potential ETL material in OLEDs, we fabricated a series of AlmmND₃-based nondoped blue OLEDs

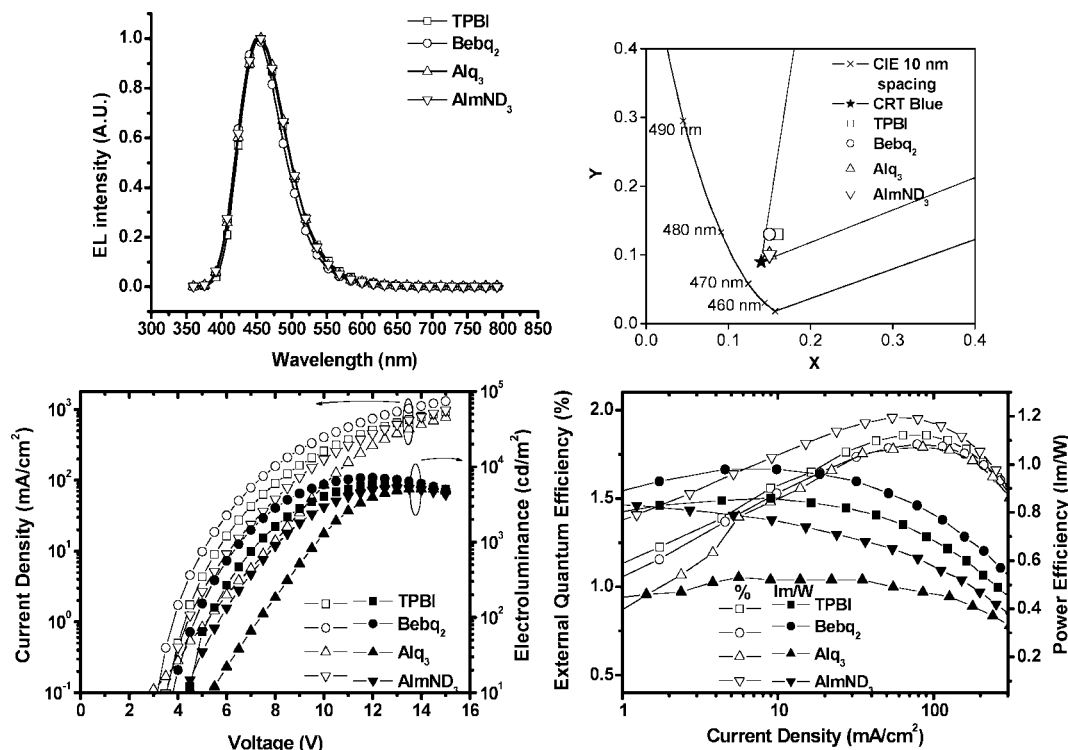


Figure 9. EL characteristics of ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃(30 nm)/ETL(20 nm)/LiF(0.5 nm)/Al(150 nm), where ETL is TPBI, Bebq₂, Alq₃, and AlmND₃, respectively.

Table 4. Electroluminescence Characteristics of Nondoped OLEDs of AlND₃, AlmND₃, AlmmND₃, and AlmpND₃ blue Al chelates^a

Al chelates	max. luminance and voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	max. efficiency [%; cd/A, lm/W]	$\lambda_{\max}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
AlND ₃	2286, 15	216, 0.75, 8.31	0.76, 1.09, 0.45	466	0.15, 0.19
AlmND ₃	3824, 15	279, 1.86, 7.57	1.91, 1.43, 0.77	448	0.15, 0.09
AlmmND ₃	3792, 15	445, 4.11, 9.36	4.18, 2.27, 0.86	436	0.15, 0.07
AlmpND ₃	4078, 15	306, 1.58, 9.63	1.62, 1.58, 0.71	452	0.15, 0.12

^a Devices have the configuration of ITO/NPB(40 nm)/CBP(10 nm)/Blue Al chelates(30 nm)/AlmND₃(20 nm)/LiF(0.5 nm)/Al(150 nm). ^b At current density of 20 mA/cm².

with a variation of ETL material, AlND₃, AlmND₃, AlmmND₃, AlmpND₃, GamND₃, or InnND₃ (Table 5 and Figure S2). Here, for a fair comparison, the AlmmND₃ OLED having AlmND₃ as ETL material is a refabricated one, not the same one in Table 4 and Figure 10. As the data show in Table 5, all devices exhibit virtually the same $\lambda_{\max}^{\text{el}}$ 432 nm and a very similar deep blue color chromaticity CIE_{x,y} = 0.15–0.16, 0.07–0.08, consistent with all OLEDs having the same deep blue nondoped emitter AlmmND₃. Regardless of ETL materials, all deep blue OLEDs have a relatively high EL efficiency with η_{ext} being more than 3.0%. In terms of EL efficiency, the AlmmND₃ OLED having AlmND₃ as the ETL material is the most outstanding. This OLED has η_{ext} reaching 3.77% and η_p over 0.80 lm/W at a current density of 20 mA/cm². From OLED results in section 2.7 and 2.8, it is very clear that AlmmND₃ is the most efficient nondoped deep blue emitter but AlmND₃ is the best ETL material for it.

2.9. AlmND₃ as the Host Material for Blue Perylene and Deep Blue 9,10-Diphenylanthracene Dopants. To demonstrate the wide band gap nature of hydroxynaphthridine metal chelates, we have fabricated a series of AlmND₃ OLEDs containing perylene or 9,10-diphenylanthracene (DPA) as highly efficient blue or deep blue dopant materials. Perylene is a highly

fluorescent blue emitter with $\lambda_{\max}^{\text{fl}}$ 467 nm (in cyclohexane) and Φ_f 94%.²³ However, in the solid state blue perylene becomes a poor fluorophore, a yellow one, due to the severe molecular aggregation that brings about concentration quenching, and a diminished and red-shifting fluorescence takes place. The fluorescence of DPA is even bluer and stronger showing $\lambda_{\max}^{\text{fl}}$ 438 nm (in cyclohexane) and Φ_f 100%.²³ Similar to blue perylene, deep blue DPA has the inherent problem of crystallizing when deposited as a thin film that prohibits its nondopant usage in a device.²⁴ Whereas some reports are available for the dopant usage of perylene,²⁵ there is no known literature case of the dopant usage of DPA in OLEDs to date. This can be simply attributed to the lack of appropriate host material, a wide band gap one, required for such a deep blue DPA dopant.

We took 0.5, 2, and 4 wt % dopant concentration of both perylene and DPA in the fabrication of dopant-based AlmND₃ OLEDs. The EL characteristics of both series of devices are displayed in Figures 11 and 12, and their data are summarized in Table 6.

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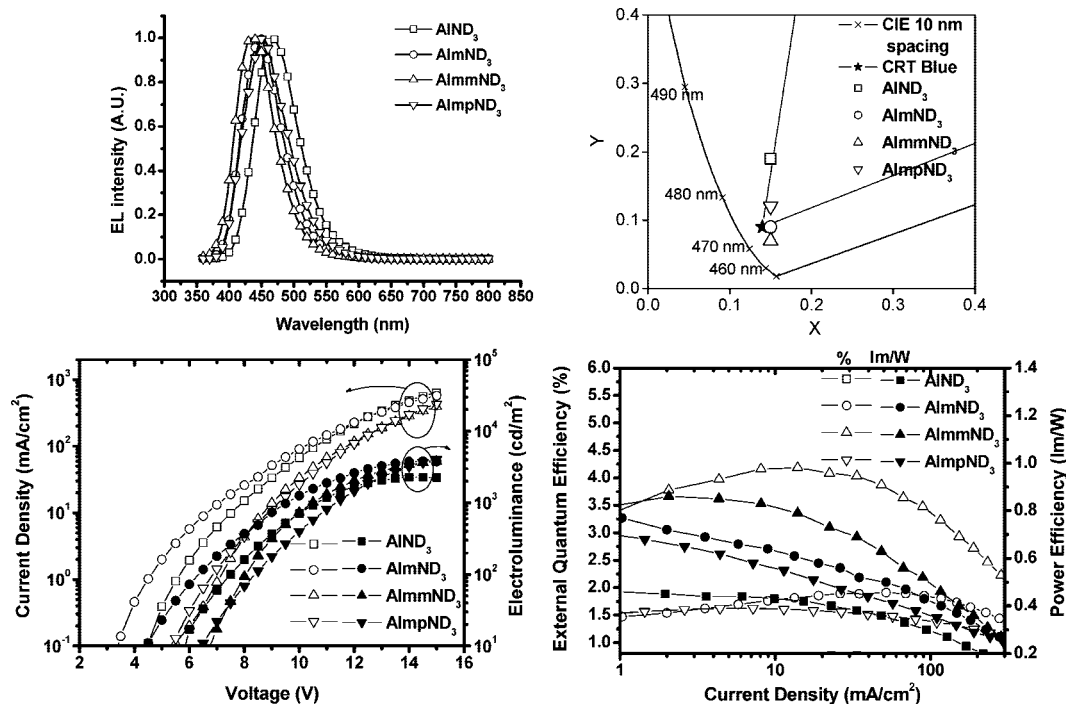


Figure 10. EL characteristics of ITO/NPB(40 nm)/CBP(10 nm)/Blue Al chelate(30 nm)/ETL(20 nm)/LiF(0.5 nm)/Al(150 nm), where Blue Al chelate is AlmND_3 , AlmND_3 , AlmmND_3 , and AlmpND_3 , respectively.

Table 5. Electroluminescence Characteristics of Nondoped OLEDs of AlmmND_3 with Various Metal Chelates as ETL Materials^a

ETL	max. luminance and voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	max. efficiency [%; cd/A, lm/W]	$\lambda_{\text{max}}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
AlND_3	3313, 15	369, 3.40, 8.92	3.67, 2.00, 1.04	432	0.15, 0.07
AlmND_3	4444, 15	401, 3.77, 7.75	3.79, 2.00, 0.94	432	0.15, 0.07
AlmmND_3	4879, 15	329, 3.01, 8.60	3.27, 1.78, 0.84	432	0.16, 0.07
AlmpND_3	5487, 15	376, 3.23, 8.24	3.41, 1.99, 1.13	432	0.16, 0.08
GamND_3	4694, 15	416, 3.67, 8.18	3.68, 2.09, 0.93	432	0.15, 0.07
ImmND_3	3848, 15	376, 3.62, 8.62	3.67, 1.91, 0.91	432	0.15, 0.07

^a Devices have the configuration of ITO/NPB(40 nm)/CBP(10 nm)/ AlmmND_3 (30 nm)/ETL(20 nm)/LiF(0.5 nm)/Al(150 nm). ^b At current density of 20 mA/cm².

Dopant concentration 0.5 wt% was found to be optimum for either perylene or DPA. The vibronic emission observed at the low energy side of the major EL band is the distinct feature of a not so blue perylene EL, indicating a sufficient Förster energy transfer between the AlmND_3 host and perylene dopant. Compared to the undoped device (second entry of Table 4), the maximum EL efficiency of the perylene-doped device was enhanced more than four times to ~ 3.06 lm/W, or more than three times to ~ 4.67 cd/A. At a current density of 20 mA/cm², η_{ext} was 3.13%, which was enhanced 1.7 times compared to an undoped device. The maximum electroluminescence also increased to 12 420 cd/m², a greater than 3-fold enhancement than that of the undoped device, although this is at a price of deep blue color purity. The perylene-doped device had a chromaticity of $\text{CIE}_{x,y} = 0.14, 0.17$, less deep blue than $\text{CIE}_{x,y} = 0.15, 0.09$ of the undoped device (second entry of Table 4). Nonetheless, the performance of the 0.5 wt% perylene-doped AlmND_3 OLED reported herein is one of the best perylene dopant devices known in literature.²⁵ Similar enhancement of the OLED performance was also observed for DPA-doped devices, although the enhancement was to a smaller extent when compared with that of perylene-doped OLEDs. However, the

deep blue color purity was elevated to a higher level, $\text{CIE}_{x,y} = 0.15, 0.06$. After the comparison with emission spectra of solution DPA, solid state AlmND_3 , and solid state DPA (fluorescence spectra shown in Figure 12), the EL spectra of the DPA-doped AlmND_3 OLED can be recognized as a coemission from both DPA and AlmND_3 . Since the main emission wavelength of DPA is shorter than that for AlmND_3 , the coemission EL observed for the DPA-doped AlmND_3 OLED gave rise to the deepest blue color (see inserted 1931 CIE chromaticity diagram in Figure 12). Considering the inadequate overlapping of the absorption spectrum of DAP and the emission spectrum of AlmND_3 (not shown in Figure 12), we can conceive that Förster energy transfer between

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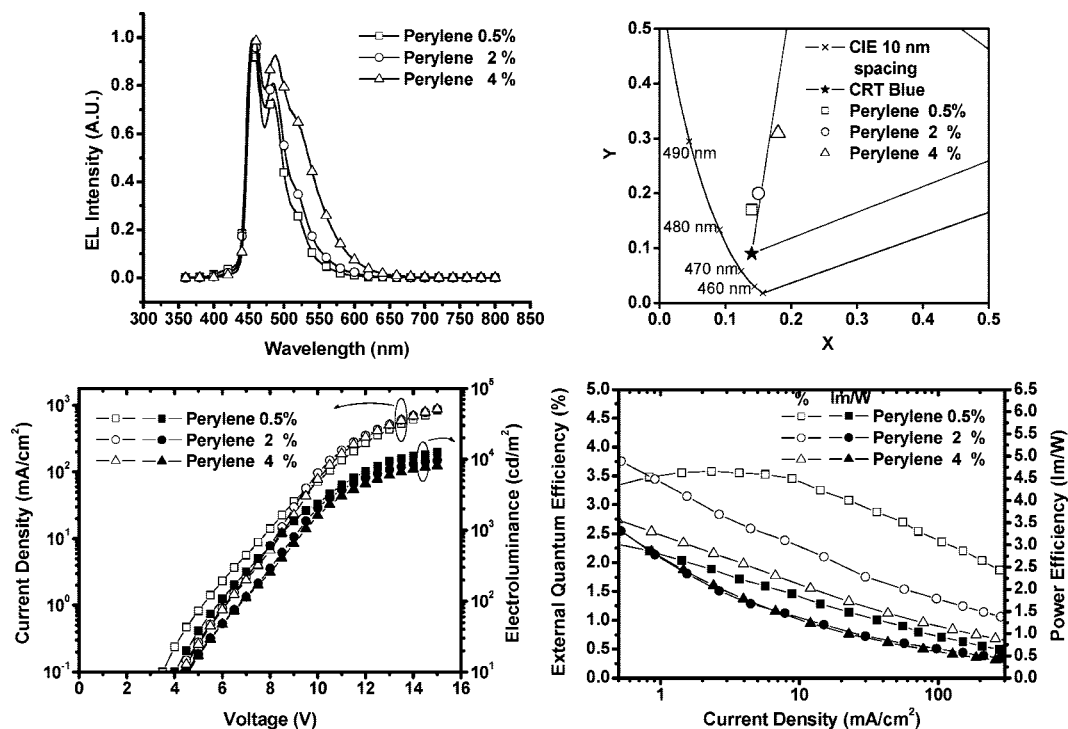


Figure 11. EL characteristics of ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃:perylene(*x* %, 30 nm)/AlmND₃(20 nm)/LiF(0.5 nm)/Al(150 nm), where *x* is 0.5, 2, or 4, the weight percent of perylene dopant.

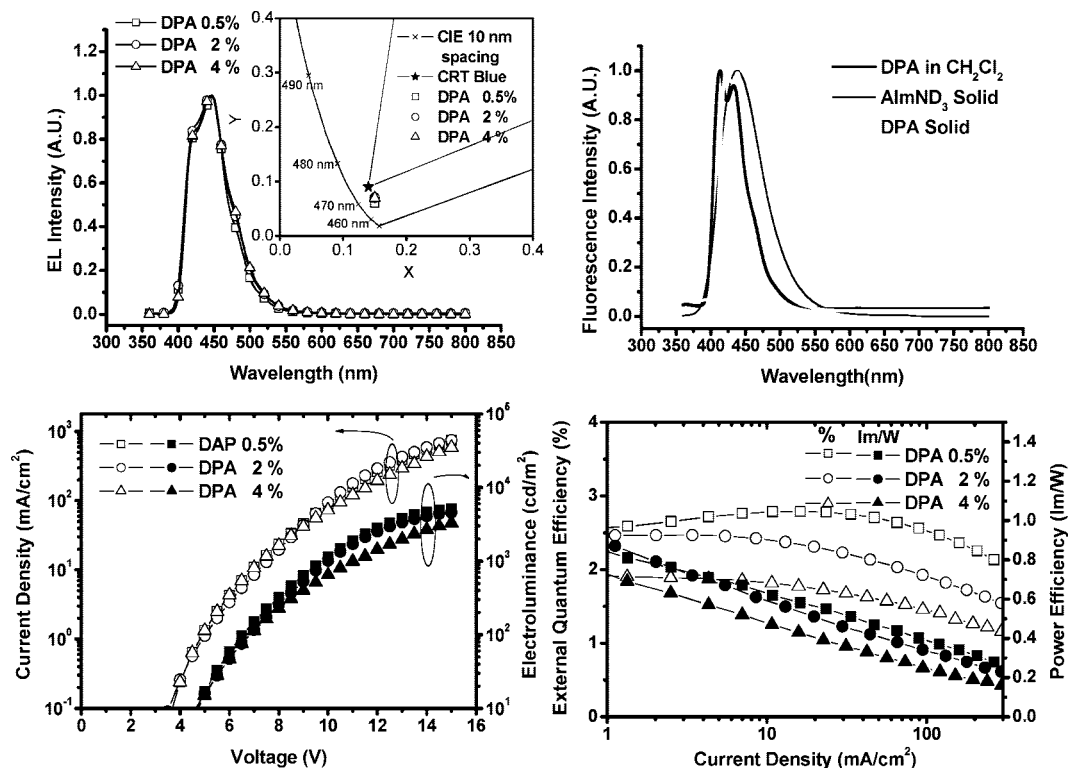


Figure 12. EL characteristics of ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃:DPA(*x* %, 30 nm)/AlmND₃(20 nm)/LiF(0.5 nm)/Al(150 nm), where *x* is 0.5, 2, or 4, the weight percent of DPA dopant

the AlmND₃ host and DPA dopant is somewhat incomplete, even though the HOMO energy level of the AlmND₃ host is low enough for the DPA dopant (but the LUMO energy level of AlmND₃ is not high enough for the DPA dopant). Despite such a coemission, to the best of our knowledge, this is the first observation of a deep blue EL from DPA.

2.10. AlmND₃ as the Host Material for High Efficiency White OLEDs with Rubrene Yellow Dopant. Yellow fluorophore rubrene is probably the most commonly used dopant accompanying blue fluorescent host material in the fabrication of white OLEDs.^{26,27} Fluorescence spectrum of AlmND₃ is partially overlapping with the absorption spectrum of rubrene

Table 6. Electroluminescence Characteristics of **AlmND₃** OLEDs with Various Dopant (Perylene or 9,10-Diphenylanthracene) Concentrations^a

dopant concn [wt %]	max. luminance and voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	max. efficiency [%; cd/A, lm/W]	$\lambda_{\text{max}}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
Perylene as Dopant					
0.5	12420, 15	810, 3.13, 8.3	3.56, 4.67, 3.06	456	0.14, 0.17
2	9610, 15	590, 1.98, 8.7	3.77, 3.51, 3.28	456	0.15, 0.20
4	8060, 15	580, 1.38, 8.9	2.73, 5.45, 3.69	458	0.18, 0.31
9,10-Diphenylanthracene (DPA) as Dopant					
0.5	5120, 15	278, 2.78, 5.8	2.80, 1.39, 0.89	448	0.15, 0.06
2	4595, 15	259, 2.31, 5.4	2.47, 1.38, 0.87	446	0.15, 0.07
4	3295, 15	203, 1.75, 5.6	1.90, 1.91, 0.69	446	0.18, 0.07

^a ITO/NPB(40 nm)/CBP(10 nm)/**AlmND₃**:Dopant (*x* %, 30 nm)/**AlmND₃**(20 nm)/LiF(5 nm)/Al(150 nm), where *x* = 0.5, 2, and 4 (weight percent) for perylene or 9,10-diphenylanthracene (DPA) dopant.

^b At 20 mA/cm².

(Figure S3 left), and white fluorescence is feasible with an appropriate ratio of mixed **AlmND₃** and rubrene (Figure S3 right).

We have demonstrated that the deep blue **AlmND₃** is a good host material for rubrene in generating high EL efficiency white OLEDs. A series of OLEDs ITO/NPB(40 nm)/CBP(10 nm)/**AlmND₃**:rubrene(*x* %, 30 nm)/**AlmND₃**(20 nm)/LiF(0.5 nm)/Al(150 nm) with a variation of rubrene dopant concentration 0.5–4 wt% and EL characteristics of such **AlmND₃**:rubrene OLEDs are shown in Figure 13 and their data are summarized in Table 7.

Similar to most rubrene-based white OLEDs, an authentic white color purity (CIE_{x,y} = 0.32, 0.38) was achieved at a very low dopant concentration, 0.5 wt % of **AlmND₃** (first entry of Table 7). Any higher dopant concentration simply impairs the white color purity of OLEDs. For such white OLEDs at a brightness of 100 cd/m², η_{ext} is 4.25%, which is equivalent to an η_{p} of 8.67 lm/W, one of the highest among literature-known rubrene-based white OLEDs.^{26,27} The color rendering index (CRI) of such white (two-color-component) OLEDs was determined to be in the range 45–50, and it is in general inferior to CRI ~80 of three-color-component white OLEDs.²⁸ At practical lighting conditions, i.e., 1000 cd/m², such an **AlmND₃**:rubrene white OLED has an η_{ext} of ~3.9% or η_{p} of 5.1 lm/W (Figure 13). With few exceptions,²⁷ this EL efficiency at 1000 cd/m² also outperforms most literature-known rubrene-based two-element white OLEDs. Also, a doped device with a 4% rubrene concentration is virtually a yellow OLED, CIE_{x,y} = 0.45, 0.51 (Figure 13 and the fourth entry of Table 7). Its high efficiency, 4.56%, 15.83 cd/A, or 11.94 lm/W, at 100 cd/m² and high brightness (1930 cd/m² at 20 mA/cm²) outperforms most currently known yellow OLEDs.²⁹ The high EL efficiency of such rubrene-based white or yellow OLEDs can be attributed to the high electron mobility of **AlmND₃**. High electron mobility is rarely observed for the host material in the white or yellow

OLEDs, and it facilitates the charge balance and hence the EL efficiency of OLEDs.

3. Conclusion

We have reported the first series of group III metal chelates as the authentic deep blue analogues of green **Alq₃**. A large quantity of these deep blue metal chelates are readily available now due to our improved synthesis and facial purification method applicable for them. Unlike several currently known blue analogues of **Alq₃**, these deep blue group III metal chelates are volatile and thermally stable enough for OLED fabrication by a vacuum-thermal-deposition process. For OLEDs, we have overcome the problem of exciplex formation, which has hampered **AlND₃** from practical usage in OLEDs before. High efficiency (maximum $\eta_{\text{ext}} > 4.0\%$) and deep blue (CIE_{x,y} = 0.15, 0.07) nondoped OLEDs were achieved for **AlmmND₃**. The wide band gap of these deep blue metal chelates, 6.4 and 3.0 eV for **AlmND₃** HOMO and LUMO energy levels, respectively, enable their usage as the host material for perylene blue dopant or deep blue 9,10-diphenylanthracene. These group III metal chelates, particularly **AlmND₃** and **AlmmND₃**, have been characterized for the charge carrier mobility by a time-of-flight technique. Both **AlmND₃** and **AlmmND₃** exhibit high electron mobility, comparable with or even higher than that of BeBq₂ or **Alq₃**. Moreover, **AlmND₃** is ambipolar with a similar mobility of 10^{−4} cm²/Vs for both hole and electron. The success of highly efficient white OLEDs (or yellow OLEDs) based on a rubrene dopant is attributed to the high electron (and hole) mobility of the host material, **AlmND₃**. We have demonstrated the versatile and effective application of hydroxynaphthylidine-based group III metal chelates for OLEDs. More high performance OLEDs can be anticipated now due to the availability of long-thought, wide band gap, deep blue group III metal chelates.

4. Experimental Section

General Information. Both solution and solid-state fluorescence quantum yields (Φ_f 's) of the blue metal chelates were determined by the integrating-sphere method.³⁰ Photoluminescence (PL) spectra were recorded on a Hitachi fluorescence spectrophotometer F-4500, and the same spectrophotometer was used to record the EL spectra of OLEDs. Melting points (*T_m*'s), glass transition temperatures (*T_g*'s), and crystallization temperatures (*T_c*'s) of respective compounds were measured via differential scanning calorimetry (DSC) using a Perkin-Elmer DSC-6 differential scanning calorimeter. The HOMO energy levels of the thin-film samples of metal chelates were studied by ultraviolet photoemission spectroscopy (UPS). The experimental detail of UPS measurement has been described before.³¹ LUMO energy levels were estimated by subtracting the energy gap (ΔE) from HOMO energy levels. ΔE was determined by the on-set absorption energy from the absorption spectra of the materials. UV–visible electronic absorption spectra were recorded on a Hewlett-Packard 8453 Diode Array spectrophotometer. The method of time-of-flight (TOF) in measuring charge carrier mobility has been reported before.³² Data collection of the X-ray crystallography analysis was carried out on a Bruker X8APEX CCD

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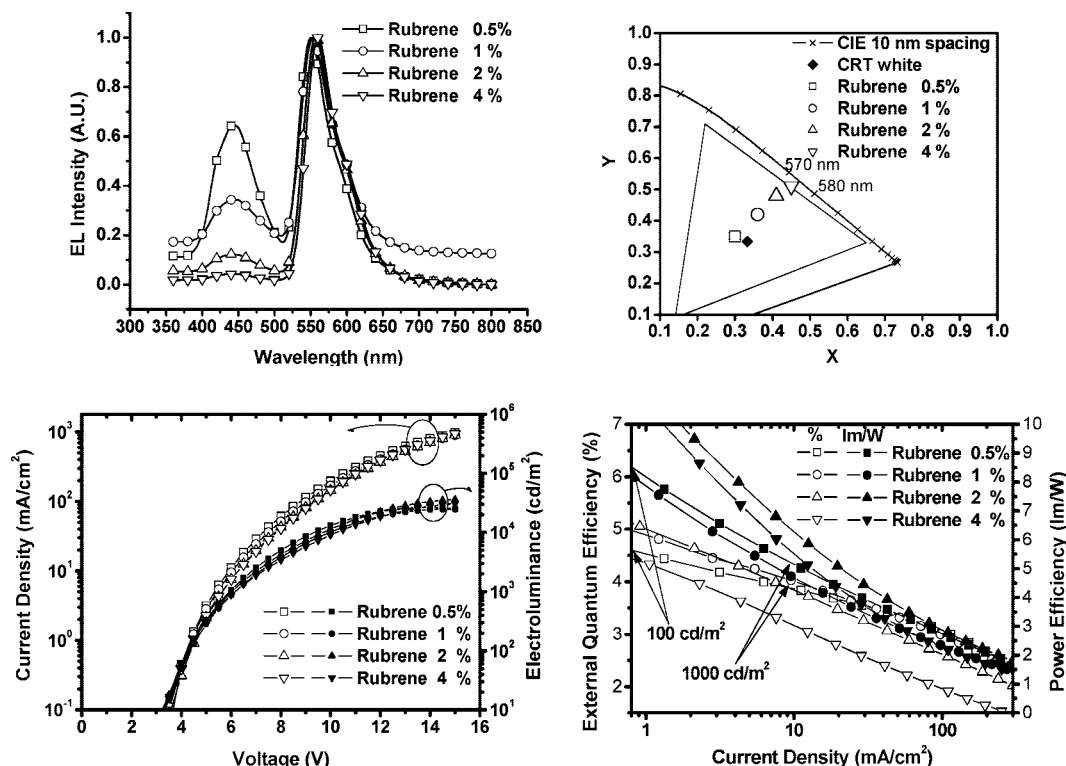


Figure 13. EL characteristics of ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃:rubrene (x %, 30 nm)/AlmND₃ (20 nm)/LiF(0.5 nm)/Al(150 nm), where x is 0.5, 1, 2, or 4, the weight percent of rubrene dopant. EL efficiency (η_{ext} or η_{p}) at 100 or 1000 cd/m² electroluminescence is arrow-marked for the device with 0.5 wt% rubrene dopant.

Table 7. Electroluminescence Characteristics of White OLEDs with AlmND₃ Host and Various Dopant (Rubrene) Concentration^a

rubrene concn [wt %]	max. luminance and voltage [cd/m ² , V]	luminance, efficiency, voltage [cd/m ² , %, V] ^b	efficiency at 100 and 1000 cd/m ² [% , cd/A, lm/W]	$\lambda_{\text{max}}^{\text{el}}$ [nm]	CIE 1931 chromaticity [x, y]
0.5	26 710, 15	1830, 3.67, 6.58	4.25, 11.55, 8.67 3.91, 9.65, 5.13	444, 550	0.30, 0.35
1	24 810, 15	1670, 3.73, 6.75	4.97, 11.17, 8.20 3.97, 8.93, 4.58	444, 554	0.36, 0.42
2	36 290, 15	2260, 3.47, 7.02	5.35, 17.46, 13.03 3.98, 12.97, 6.78	444, 558	0.41, 0.48
4	30 970, 15	1930, 2.79, 7.05	4.56, 15.83, 11.94 3.24, 11.25, 5.77	444, 560	0.45, 0.51

^a Devices have the configuration of ITO/NPB(40 nm)/CBP(10 nm)/AlmND₃:rubrene x %(30 nm)/AlmND(20 nm)/LiF(0.5 nm)/Al(150 nm), x = 0.5, 1, 2, 4, respectively. ^b At current density of 20 mA/cm.

diffractometer at 100 K for AIND₃·CH₂Cl₂ single crystals. The experimental detail of X-ray diffraction and their data process in solving the crystal structure can be found elsewhere.^{30b}

The fabrication of OLEDs and their EL characterization also have been described before.^{19b,33} The device was placed close to the photodiode such that all the forward light entered the photodiode. The effective size of the emitting diodes was 3.14 mm², which is significantly smaller than the active area of the photodiode detector, a condition known as “under-filling”, satisfying the measurement protocol.³⁴ This is one of the most conventional ways in measuring the EL efficiency of OLEDs, although sometimes experimental errors may arise due to the non-Lambertian emission of OLEDs.³⁵ The color rendering index (CRI) of white OLEDs was measured by a spectroradiometer (Specbos 1201, JETI Technishe Instrumente GmbH).

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¹H and ¹³C NMR spectra were recorded on a Bruker AMX-400 MHz or AVA-400 MHz Fourier-transform spectrometer at room temperature. Elemental analyses (on a Perkin-Elmer 2400 CHN Elemental Analyzer) and electron impact (EI), fast atom bombardment (FAB), or matrix-assisted laser desorption/ionization time-of-flight (MALDI-TOF) mass spectra (on a VA Analytical 11-250J or 4800 MALDI TOF/TOF Analyzer) were recorded by the Elemental Analyses and Mass Spectroscopic Laboratory in-house service of the Institute of Chemistry, Academic Sinica.

Materials. For the materials used in device fabrication, Beq₂, DPA (9,10-diphenylanthracene), perylene, and rubrene are commercially available materials and they were used without further purification. NPB (1,4-bis(1-naphthylphenylamino)biphenyl), CBP (4,4'-bis(9-carbazolyl)-2,2'-biphenyl), mCP (1,3-bis(9-carbazolyl)-benzene), TPBI (2,2',2''-(1,3,5-phenylene)tris[1-phenyl-1H-benzimidazole]), and Alq₃ were prepared via published methods and were subjected to gradient sublimation prior to use.

Synthesis of Tris(4-hydroxy-1,5-naphthyridinato)aluminum (AIND₃). To a toluene solution (4.5 mL) was added ND (0.15 g, 1.0 mmol) and aluminum triisopropoxide (0.07 g, 0.34 mmol). The

reaction mixture was heated at reflux under nitrogen for 1 h. During the reaction, a formation of milky white solid emitting blue fluorescence was observed. After cooling, the reaction solution was filtered to isolate a white solid. After washing with *n*-hexanes, the solid was subjected to zone-temperature sublimation. The pure product was isolated as a white solid. Yield: 76% (0.12 g). ^1H NMR (400 MHz, CD_2Cl_2): δ 8.77 (d, 1H, J = 4.75 Hz), 8.73–8.67 (m, 4H), 8.56 (d, 1H, J = 8.54 Hz), 8.51–8.47 (m, 2H), 7.73 (dd, 1H, J = 8.56 Hz, J = 4.72 Hz), 7.56 (dd, 1H, J = 8.60 Hz, J = 4.75 Hz), 7.49 (dd, 1H, J = 8.50 Hz, J = 4.74 Hz), 7.29 (d, 1H, J = 4.59 Hz), 6.93 (d, 1H, J = 5.33 Hz), 6.90 (d, 1H, J = 5.34 Hz), 6.81 (d, 1H, J = 5.32 Hz). ^{13}C NMR (100 MHz, CDCl_3): δ 165.17, 164.98, 164.65, 156.28, 155.73, 155.70, 144.80, 144.43, 144.17, 144.07, 142.39, 141.87, 141.45, 141.37, 138.84, 138.18, 138.01, 125.13, 124.55, 109.80, 109.62, 109.08. FAB-MS: calcd 462.10, m/z = 463.1 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{24}\text{H}_{15}\text{AlN}_6\text{O}_3$: C, 62.31 (62.34); H, 2.93 (3.27); N, 17.88 (18.17).

Synthesis of Tris(4-hydroxy-8-methyl-1,5-naphthyridinato)aluminum (AlmND₃). This compound was synthesized in the same manner as AlND₃, except that **mND** (1.00 g, 6.2 mmol) was used instead of **ND**. The pure product was isolated as a white solid. Yield: 86% (0.90 g). ^1H NMR (400 MHz, CDCl_3): δ 8.73–8.70 (m, 3H), 8.57 (d, 1H, J = 4.90 Hz), 8.52 (d, 1H, J = 4.91 Hz), 7.51 (d, 1H, J = 4.81 Hz), 7.42 (d, 1H, J = 4.90 Hz), 7.27 (d, 1H, J = 4.76 Hz), 7.11 (d, 1H, J = 4.94), 6.94 (d, 1H, J = 5.34 Hz), 6.91 (d, 2H, J = 5.34 Hz), 2.84 (s, 3H), 2.79 (s, 3H), 2.77 (s, 3H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.58, 165.44, 165.12, 154.74, 154.22, 154.20, 153.22, 152.70, 152.56, 143.95, 143.76, 143.70, 143.65, 143.62, 141.75, 137.99, 137.31, 137.18, 125.25, 125.23, 124.69, 109.68, 109.40, 108.91, 17.76, 17.67. FAB-MS: calcd 504.15, m/z = 505.1 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{27}\text{H}_{21}\text{AlN}_6\text{O}_3$: C, 64.20 (64.28); H, 4.33 (4.20); N, 16.36 (16.66).

Synthesis of Tris(4-hydroxy-2,8-dimethyl-1,5-naphthyridinato)aluminum (AlmmND₃). This compound was synthesized in the same manner as AlND₃, except that **mmND** (0.50 g, 2.9 mmol) was used instead of **ND**. The pure product was isolated as a white solid. Yield: 86% (0.45 g). ^1H NMR (400 MHz, CDCl_3): δ 8.46 (d, 1H, J = 4.88 Hz), 8.43 (d, 1H, J = 4.88 Hz), 7.42 (d, 1H, J = 4.88 Hz), 7.34 (d, 1H, J = 4.88 Hz), 7.18 (d, 1H, J = 4.92 Hz), 7.00 (d, 1H, J = 4.96 Hz), 6.81 (s, 2H), 6.80 (s, 1H), 2.80 (s, 3H), 2.76 (s, 3H), 2.74 (s, 3H), 2.61 (s, 3H), 2.60 (s, 6H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.03, 164.76, 164.24, 163.59, 152.01, 151.57, 151.40, 143.09, 143.05, 142.99, 142.74, 142.48, 140.61, 136.89, 136.23, 136.09, 125.08, 124.60, 109.67, 109.26, 108.89, 26.36, 26.29, 17.60, 17.53. FAB-MS: calcd 546.20, m/z = 547.2 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{30}\text{H}_{27}\text{AlN}_6\text{O}_3$: C, 65.54 (65.93); H, 4.94 (4.98); N, 15.37 (15.38).

Synthesis of Tris(4-hydroxy-8-dimethyl-2-phenyl-1,5-naphthyridinato)aluminum (AlmpND₃). This compound was synthesized in the same manner as AlND₃, except that **mpND** (0.20 g, 0.9 mmol) was used instead of **ND**. The pure product was isolated as a white solid. Yield: 77% (0.16 g). ^1H NMR (400 MHz, d_6 -DMSO): δ 8.59 (d, 1H, J = 4.90 Hz), 8.49 (d, 1H, J = 4.92 Hz), 8.23–8.20 (m, 6H), 7.78 (d, 1H, J = 5.00 Hz), 7.69 (d, 1H, J = 4.91 Hz), 7.58–7.42 (m, 14H), 2.85 (s, 3H), 2.83 (s, 3H), 2.80 (s, 3H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.79, 165.75, 165.51, 161.48, 161.04, 153.20, 152.82, 152.69, 143.43, 143.32, 143.12, 141.23, 140.46, 140.25, 139.94, 137.32, 136.68, 136.52, 129.65, 129.46, 129.29, 128.71, 128.67, 128.60, 127.72, 127.65, 127.62, 125.39, 124.88, 17.55, 17.47. FAB-MS: calcd 732.24, m/z = 733.3 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{45}\text{H}_{33}\text{AlN}_6\text{O}_3$: C, 73.84 (73.76); H, 4.37 (4.54); N, 11.37 (11.47).

Synthesis of Tris(4-hydroxy-8-methyl-1,5-naphthyridinato)-gallium (GamND₃). To a water solution (7.2 mL) were added **mND** (0.14 g, 0.87 mmol) and gallium chloride (0.06 g, 0.34 mmol). The reaction mixture was heated and stirred at $\sim 40^\circ\text{C}$. An excess amount of potassium acetate was added to the solution to change the solution acidity from pH 3–4 to pH 7–8. During the reaction (ca. 1 h), the formation of a milky white solid emitting blue

fluorescence was observed. After cooling, the reaction solution was filtered to isolate the white solid. After washing with water, the solid was vacuum-dried and finally subjected to zone-temperature sublimation. The pure product was isolated as a white solid. Yield: 67% (0.11 g). ^1H NMR (400 MHz, CDCl_3): δ 8.70 (d, 3H, J = 5.36 Hz), 8.62 (d, 1H, J = 4.83 Hz), 8.57 (d, 1H, J = 4.90 Hz), 7.54 (d, 1H, J = 4.80 Hz), 7.46 (d, 1H, J = 4.71 Hz), 7.31 (d, 1H, J = 4.78 Hz), 7.27 (d, 1H, J = 4.89), 6.96 (d, 1H, J = 5.37 Hz), 6.93 (d, 1H, J = 5.42 Hz), 6.92 (d, 1H, J = 5.36 Hz), 2.84 (s, 3H), 2.80 (s, 3H), 2.78 (s, 3H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.18, 164.98, 164.74, 154.43, 153.96, 153.51, 153.04, 153.00, 143.89, 143.81, 143.41, 143.15, 141.22, 136.08, 135.45, 135.34, 125.23, 125.14, 124.71, 109.61, 109.51, 108.97, 17.95, 17.85. FAB-MS: calcd 546.09, m/z = 547.0 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{27}\text{H}_{21}\text{GaN}_6\text{O}_3$: C, 59.03 (59.26); H, 4.00 (3.87); N, 15.09 (15.36).

Synthesis of Tris(4-hydroxy-2,8-dimethyl-1,5-naphthyridinato)gallium (GammND₃). This compound was synthesized in the same manner as **GamND₃**, except that **mmND** (0.60 g, 3.4 mmol) was used instead of **mND**. The pure product was isolated as a white solid. Yield: 83% (0.56 g). ^1H NMR (400 MHz, CDCl_3): δ 8.51 (d, 1H, J = 4.83 Hz), 8.48 (d, 1H, J = 4.86 Hz), 7.45 (d, 1H, J = 4.78 Hz), 7.38 (d, 1H, J = 4.90 Hz), 7.23 (d, 1H, J = 4.89 Hz), 7.15 (d, 1H, J = 4.90 Hz), 6.83–6.82 (m, 3H), 2.80 (s, 3H), 2.77 (s, 3H), 2.75 (s, 3H), 2.61 (s, 3H), 2.59 (s, 6H). ^{13}C NMR (100 MHz, CDCl_3): δ 164.63, 164.56, 164.36, 163.94, 163.40, 163.33, 152.32, 151.88, 143.31, 143.25, 142.20, 142.00, 140.07, 135.02, 134.38, 134.26, 125.10, 124.96, 124.62, 109.57, 109.32, 108.88, 26.31, 17.80, 17.71. FAB-MS: calcd 588.14, m/z = 589.0 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{30}\text{H}_{27}\text{GaN}_6\text{O}_3$: C, 61.50 (61.14); H, 4.63 (4.62); N, 14.17 (14.26).

Synthesis of Tris(4-hydroxy-8-dimethyl-2-phenyl-1,5-naphthyridinato)gallium (GampND₃). This compound was synthesized in the same manner as **GamND₃**, except that **mpND** (0.20 g, 8.5 mmol) was used instead of **mND**. The pure product was isolated as a white solid. Yield: 69% (0.15 g). ^1H NMR (400 MHz, CDCl_3): δ 8.62 (d, 1H, J = 4.88 Hz), 8.59 (d, 1H, J = 4.88 Hz), 8.15–8.12 (m, 6H), 7.53–7.40 (m, 14H), 7.33 (d, 1H, J = 5.04 Hz), 7.30 (d, 1H, J = 4.96 Hz), 2.93 (s, 3H), 2.92 (s, 3H), 2.86 (s, 3H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.42, 165.31, 165.13, 161.13, 160.78, 160.70, 153.50, 153.15, 153.11, 143.63, 143.59, 142.77, 142.62, 140.66, 140.39, 140.19, 139.91, 135.45, 134.82, 134.70, 129.64, 129.46, 129.32, 128.69, 128.67, 128.60, 127.69, 128.61, 125.38, 125.24, 124.87, 106.99, 106.61, 106.24, 17.73, 17.65. FAB-MS: calcd 774.19, m/z = 775.2 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{45}\text{H}_{33}\text{GaN}_6\text{O}_3$: C, 69.80 (69.69); H, 4.27 (4.29); N, 10.72 (10.84).

Synthesis of Tris(4-hydroxy-8-methyl-1,5-naphthyridinato)-indium (InmND₃). To a water solution (7.2 mL) were added **mND** (0.15 g, 0.94 mmol) and indium chloride (0.07 g, 0.32 mmol). The reaction mixture was stirred and heated at reflux temperature. An excess amount of potassium acetate was added to the solution to change the solution acidity from pH 3–4 to pH 7–8. During the reaction (ca. 1 h), the formation of a milky white solid was observed. However, the blue fluorescence from such a precipitate was barely discernible. After the removal of water by vacuum distillation, the resulting solid was subjected to Soxhlet extraction by dichloromethane for 24 h. The extracted dichloromethane solution was evaporated until dryness, and the solid residue was further purified by zone-temperature sublimation. The pure product was isolated as an off-white solid. Yield: 76% (0.14 g). ^1H NMR (400 MHz, CDCl_3): δ 8.67 (d, 3H, J = 5.40 Hz), 8.33 (d, 3H, J = 4.80 Hz), 7.51 (dd, 3H, J = 4.79 Hz, J = 0.71 Hz), 6.95 (d, 3H, J = 5.41 Hz), 2.83 (s, 9H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.73, 153.71, 153.68, 144.30, 144.22, 136.46, 125.04, 110.75, 18.23. FAB-MS: calcd 592.07, m/z = 593.0 ($\text{M}+\text{H}^+$). Anal. Found (Calcd) for $\text{C}_{27}\text{H}_{21}\text{InN}_6\text{O}_3$: C, 54.79 (54.75); H, 3.46 (3.57); N, 14.07 (14.19).

Synthesis of Tris(4-hydroxy-2,8-dimethyl-1,5-naphthyridinato)indium (InmmND₃). This compound was synthesized in the same manner as **InmND₃**, except that **mmND** (0.17 g, 0.98 mmol) was used instead of **mND**. The pure product was isolated as an

off-white solid. Yield: 81% (0.17 g). ^1H NMR (400 MHz, CDCl_3): δ 8.22 (d, 3H, $J = 4.80$ Hz), 7.43 (dd, 3H, $J = 4.82$ Hz, $J = 0.68$ Hz), 6.84 (s, 3H), 2.79 (s, 9H), 2.53 (s, 9H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.19, 163.02, 152.57, 143.74, 143.04, 135.36, 124.89, 110.57, 26.15, 18.04. FAB-MS: calcd 634.12, $m/z = 635.1$ ($\text{M} + \text{H}^+$). Anal. Found (Calcd) for $\text{C}_{30}\text{H}_{27}\text{InN}_6\text{O}_3$: C, 56.49 (56.80); H, 4.24 (4.29); N, 12.96 (13.25).

Synthesis of Tris(4-hydroxy-8-dimethyl-2-phenyl-1,5-naphthyridinato)indium (InmpND_3). This compound was synthesized in the same manner as InmND_3 , except that mpND (0.50 g, 2.1 mmol) was used instead of mND . However, the precipitate formed during the reaction was observed for blue fluorescence, and it was isolated by filtration. After washing with water, the solid was vacuum-dried and finally subjected to zone-temperature sublimation. The pure product was isolated as an off-white solid. Yield: 51% (0.30 g). ^1H NMR (400 MHz, CDCl_3): δ 8.34 (d, 3H, $J = 4.80$ Hz), 8.15–8.12 (m, 6H), 7.53–7.51 (m, 6H), 7.49–7.40 (m, 9H), 2.91 (s, 9H). ^{13}C NMR (100 MHz, CDCl_3): δ 165.99, 160.24, 153.78, 144.04, 143.62, 140.05, 135.77, 129.44, 128.63, 127.57, 125.15, 107.86, 17.99. MOLDI-TOF MS: calcd 820.17, $m/z =$

821.15 ($\text{M} + \text{H}^+$). Anal. Found (Calcd) for $\text{C}_{45}\text{H}_{33}\text{InN}_6\text{O}_3$: C, 65.78 (65.86); H, 3.95 (4.05); N, 10.26 (10.24).

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Supporting Information Available: The synthesis and structural characterization data of chelating ligands, **ND**, **mND**, **mmND**, **mpND**, and their precursors are described. DSC thermograms of **AlmpND₃**, **GampND₃**, and **InmpND₃** (Figure S1); EL characteristics (Figure S2) and data (Table S1) of nondoped **GamND₃**, **GammND₃**, **InmND₃**, or **InmmND₃** OLEDs; absorption and photoluminescence spectra of **AlmND₃** and rubrene as well as their solution fluorescence images (Figure S3) are given. The crystallographic information is listed in Table S2 and file (.cif) also given for **AIND₃·CH₂Cl₂**. This material is available free of charges via the Internet at <http://pubs.acs.org>.

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