

Lists of structure factors, anisotropic displacement parameters, bond distances and angles involving H atoms and selected torsion angles have been deposited with the IUCr (Reference: AB1116). Copies may be obtained through The Managing Editor, International Union of Crystallography, 5 Abbey Square, Chester CH1 2HU, England.

References

- Frazier, C. C. & Cockerham, M. P. (1987). *J. Opt. Soc. Am.* **B4**, 1899–1903.
 Gilmore, C. J. (1983). *MITHRIL. Computer Program for the Automatic Solution of Crystal Structures from X-ray Data*. Department of Chemistry, Univ. of Glasgow, Scotland.
 Johnson, C. K. (1976). *ORTEPII*. Report ORNL-5138. Oak Ridge National Laboratory, Tennessee, USA.
 Molecular Structure Corporation (1987). *TEXSAN. TEXRAY Structure Analysis Package*. Revised. MSC, 3200 Research Forest Drive, The Woodlands, TX 77381, USA.
 Molecular Structure Corporation (1988). *MSC/AFCS Diffractometer Control Software*. MSC, 3200 Research Forest Drive, The Woodlands, TX 77381, USA.

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N-Phenyl-2-phenyliminotetrahydro-1,3-thiazole-2-carbothioamide, C₁₆H₁₅N₃S₂

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Abstract

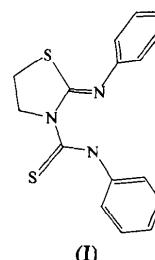
The thiazolidine ring in the title compound has an envelope conformation, *E*₄, with a C atom out of the plane of the ring. The C—S and C—N bond lengths clearly show the typical resonance of the thiourea group. The molecules are linked by van der Waals forces and an intramolecular N···N hydrogen bond is present.

Comment

A large number of thiazole or thiazine derivatives may be obtained by the reaction of haloalkylisothiocyanates with amines. Compounds containing the thiazoline or thiazine structural fragment have biological properties as potential herbicides and anti-fungal agents (Behner & Stendel, 1971; Sumitomo

Chemical Co. Ltd, 1982). Some of them also display medical properties (Johnson, Boukma & Kim, 1970).

The two phenyl groups in the title compound, (I), are planar (the maximum deviation from the least-squares plane is 0.007 Å). Bond lengths and angles are normal. The thiourea group is also planar (the maximum deviation from the least-squares plane is 0.006 Å). The S—C distance of 1.661 (2) Å and the C—N distances of 1.390 (3) and 1.344 (3) Å clearly show the resonance of the thiourea system (Kunchur & Truter, 1958; Valle, Cojazzi, Busetti & Mammi, 1970).



(I)

The thiazolidine ring is not planar, and in terms of the ring-puckering coordinates (Cremer & Pople, 1975) the amplitude and phase magnitudes are $Q = 0.353$ (3) Å and $\varphi = -33.6$ (4)° for the sequence of atoms S1, C2, N3, C4 and C5. The conformation of the ring is an envelope with C5 lying 0.316 (3) Å out of the least-squares plane of the ring. However, Nardelli (1983a) asymmetry parameters (Table 3) show a departure from ideal *E* symmetry. The substituent atoms C31 and N21 are −0.143 (2) and 0.296 (2) Å from the least-squares plane. Bond distances and angles agree well with those found for a

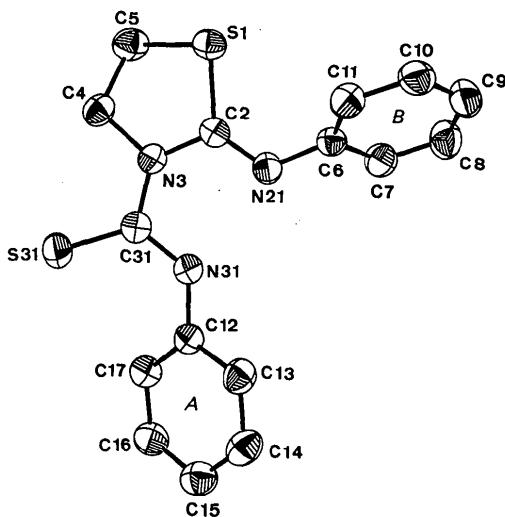


Fig. 1. An ORTEPII (Johnson, 1976) view of the molecule viewed down the *c* axis, showing the atomic numbering scheme. Displacement ellipsoids are shown at the 50% probability level.

planar thiazolidine ring (Rasmussen, Villani & Griffin, 1984).

The dihedral angles between the least-squares planes of the different groups are: phenyl ring *A*–phenyl ring *B* 66.0 (1)°, phenyl ring *A*–thiourea 54.1 (1)°, phenyl ring *A*–thiazolidine ring 56.0 (1)°, phenyl ring *B*–thiourea 38.1 (1)°, phenyl ring *B*–thiazolidine ring 21.5 (1)° and thiourea–thiazolidine ring 16.7 (1)°. The packing of the molecules is governed by van der Waals forces; there is an intramolecular hydrogen bond, N31…N21 = 2.596 (3) Å, N31—H…N21 = 138°.

Experimental

The title compound was obtained by the reaction of 2-phenyl-amino-2-thiazoline with phenyl isothiocyanate. Crystals suitable for data collection were grown from a benzene/methanol solution.

Crystal data

$C_{16}H_{15}N_3S_2$	Mo $K\alpha$ radiation
$M_r = 313.43$	$\lambda = 0.71069 \text{ \AA}$
Orthorhombic	Cell parameters from 25
$P2_12_12_1$	reflections
$a = 11.040 (1) \text{ \AA}$	$\theta = 2-30^\circ$
$b = 18.373 (1) \text{ \AA}$	$\mu = 0.332 \text{ mm}^{-1}$
$c = 7.474 (1) \text{ \AA}$	$T = 293 \text{ K}$
$V = 1516.0 (3) \text{ \AA}^3$	Prism
$Z = 4$	$0.5 \times 0.4 \times 0.3 \text{ mm}$
$D_x = 1.36 \text{ Mg m}^{-3}$	Colourless
$D_m = 1.37 \text{ Mg m}^{-3}$	
D_m measured by flotation in bromobenzene/ethanol	

Data collection

Enraf–Nonius CAD-4 diffractometer	$\theta_{\max} = 20^\circ$
$\omega/2\theta$ scans	$h = 0 \rightarrow 15$
Absorption correction:	$k = 0 \rightarrow 25$
empirical (Walker & Stuart, 1983)	$l = 0 \rightarrow 7$
$T_{\min} = 0.655$, $T_{\max} = 1.360$	3 standard reflections monitored every 100 reflections
2515 measured reflections	intensity variation: insignificant
2515 independent reflections	
2394 observed reflections [$I \geq 3\sigma(I)$]	

Refinement

Refinement on F	$(\Delta/\sigma)_{\max} = 0.004$
$R = 0.032$	$\Delta\rho_{\max} = 0.2 \text{ e \AA}^{-3}$
$wR = 0.033$	$\Delta\rho_{\min} = -0.2 \text{ e \AA}^{-3}$
$S = 1.67$	Extinction correction: none
2394 reflections	Atomic scattering factors from <i>International Tables for X-ray Crystallography</i> (1974, Vol. IV)
190 parameters	
H-atom parameters not refined	
Unit weights applied	

Table 1. Fractional atomic coordinates and equivalent isotropic displacement parameters (\AA^2)

	x	y	z	U_{eq}
S1	0.92022 (7)	0.17050 (3)	0.95164 (9)	0.0413 (2)
S31	0.64341 (6)	-0.05540 (4)	0.91507 (12)	0.0486 (2)
N3	0.8163 (2)	0.0440 (1)	0.9032 (3)	0.0338 (6)
N21	1.0266 (2)	0.0367 (1)	0.9473 (3)	0.0392 (6)
N31	0.8838 (2)	-0.0746 (1)	0.9103 (3)	0.0388 (6)
C2	0.9305 (2)	0.0749 (1)	0.9342 (3)	0.0342 (6)
C4	0.7172 (2)	0.0972 (1)	0.8937 (4)	0.0422 (8)
C5	0.7736 (3)	0.1695 (2)	0.8462 (4)	0.0507 (9)
C6	1.1413 (2)	0.0673 (1)	0.9842 (3)	0.0355 (7)
C7	1.2382 (3)	0.0445 (2)	0.8817 (4)	0.0466 (8)
C8	1.3538 (3)	0.0691 (2)	0.9179 (5)	0.0568 (10)
C9	1.3730 (3)	0.1165 (2)	1.0607 (5)	0.0579 (10)
C10	1.2776 (3)	0.1384 (2)	1.1632 (5)	0.0520 (10)
C11	1.1608 (3)	0.1139 (2)	1.1279 (4)	0.0436 (8)
C12	0.8857 (2)	-0.1514 (1)	0.9199 (3)	0.0341 (6)
C13	0.9780 (3)	-0.1829 (2)	1.0201 (4)	0.0453 (8)
C14	0.9876 (3)	-0.2587 (2)	1.0298 (4)	0.0522 (9)
C15	0.9069 (3)	-0.3019 (1)	0.9396 (4)	0.0492 (9)
C16	0.8166 (3)	-0.2706 (2)	0.8383 (4)	0.0458 (8)
C17	0.8048 (3)	-0.1956 (1)	0.8282 (4)	0.0390 (7)
C31	0.7876 (2)	-0.0297 (1)	0.9084 (3)	0.0344 (6)

Table 2. Selected geometric parameters (\AA , °)

S1—C2	1.764 (2)	S31—C31	1.661 (2)
S1—C5	1.800 (3)	N3—C4	1.468 (3)
N3—C2	1.402 (3)	N21—C2	1.276 (3)
N3—C31	1.390 (3)	N31—C12	1.414 (3)
N21—C6	1.413 (3)	C4—C5	1.509 (4)
N31—C31	1.344 (3)	C6—C11	1.389 (3)
C6—C7	1.381 (3)	C8—C9	1.393 (5)
C7—C8	1.381 (4)	C10—C11	1.391 (4)
C9—C10	1.363 (4)	C12—C17	1.388 (3)
C12—C13	1.389 (3)	C14—C15	1.371 (4)
C13—C14	1.398 (4)	C16—C17	1.387 (3)
C15—C16	1.377 (4)		
C2—S1—C5	90.8 (1)	C4—N3—C31	118.7 (2)
C2—N3—C31	126.4 (2)	C2—N3—C4	114.0 (1)
C2—N21—C6	122.7 (2)	C12—N31—C31	128.6 (2)
N3—C2—N21	122.5 (2)	S1—C2—N21	126.5 (1)
S1—C2—N3	110.9 (1)	N3—C4—C5	106.9 (2)
S1—C5—C4	106.0 (2)	N21—C6—C11	122.3 (2)
N21—C6—C7	117.6 (2)	C7—C6—C11	119.7 (2)
C6—C7—C8	120.4 (2)	C7—C8—C9	119.6 (2)
C8—C9—C10	119.8 (3)	C9—C10—C11	120.9 (3)
C6—C11—C10	119.2 (2)	N31—C12—C17	123.2 (3)
N31—C12—C13	116.9 (2)	C13—C12—C17	119.6 (2)
C12—C13—C14	119.8 (2)	C13—C14—C15	120.1 (2)
C14—C15—C16	119.8 (2)	C15—C16—C17	120.9 (2)
C12—C17—C16	119.5 (2)	N3—C31—N31	114.6 (2)
S31—C31—N31	125.6 (1)	S31—C31—N3	119.6 (1)

Table 3. Asymmetry parameters for the thiazolidine ring

ΔC_s (S1)	0.1511 (15)	ΔC_2 (S1)	0.1583 (10)
ΔC_s (C2)	0.2472 (14)	ΔC_2 (C2)	0.0588 (10)
ΔC_s (N3)	0.2488 (13)	ΔC_2 (N3)	0.0580 (10)
ΔC_s (C4)	0.1553 (12)	ΔC_2 (C4)	0.1522 (11)
ΔC_s (C5)	0.0118 (13)	ΔC_2 (C5)	0.1893 (10)

Preliminary Weissenberg photographs indicated that the crystal belonged to the orthorhombic system with systematic absences $h00$ h odd, $0k0$ k odd, $00l$ l odd, consistent with $P2_12_12_1$. Corrections were made for Lorentz–polarization effects. The structure was solved by direct methods using *MULTAN80* (Main *et al.*, 1980). The H atoms were assigned the same isotropic displacement parameters as the atoms to which they were bonded. All calculations were carried out using the crystallographic programs of the *XRAY70* System (Stewart,

Kundell & Baldwin, 1970). Bond lengths and angles were calculated using PARST (Nardelli, 1983b).

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Lists of structure factors, anisotropic displacement parameters and H-atom coordinates, and a plot of the crystal packing have been deposited with the IUCr (Reference: AB1103). Copies may be obtained through The Managing Editor, International Union of Crystallography, 5 Abbey Square, Chester CH1 2HU, England.

References

- Behner, O. & Stendel, W. (1971). Ger. Patent 2.015.675; *Chem. Abs.* (1972), 76 46195q.
- Cremer, D. & Pople, J. A. (1975). *J. Am. Chem. Soc.* **97**, 1354–1358.
- Johnson, C. K. (1976). ORTEPII. Report ORNL-5138. Oak Ridge National Laboratory, Tennessee, USA.
- Johnson, G. A., Boukma, S. J. & Kim, E. G. (1970). *J. Pharmacol. Exp. Ther.* **171**, 80–87.
- Kunchur, N. R. & Truter, M. R. (1958). *J. Chem. Soc.* **517**, 2551–2557.
- Main, P., Fiske, S. J., Hull, S. E., Lessinger, L., Germain, G., Declercq, J.-P. & Woolfson, M. M. (1980). MULTAN80. *A System of Computer Programs for the Automatic Solution of Crystal Structures from X-ray Diffraction Data*. Univs. of York, England, and Louvain, Belgium.
- Nardelli, M. (1983a). *Acta Cryst. C***39**, 1141–1142.
- Nardelli, M. (1983b). *Comput. Chem.* **7**, 95–98.
- Rasmussen, C. R., Villani, F. J. Jr & Griffin, E. A. (1984). *Acta Cryst. C***40**, 2120–2122.
- Stewart, J. M., Kundell, F. A. & Baldwin, J. C. (1970). *The XRAY70 System*. Computer Science Center, Univ. of Maryland, College Park, Maryland, USA.
- Sumitomo Chemical Co. Ltd. (1982). Jap. Patent 82/144 205 (Kokai Tokkyo Koho); *Chem. Abs.* (1983), 98 12930q.
- Valle, G., Cojazzi, G., Busetti, V. & Mammi, M. (1970). *Acta Cryst. B***24**, 468–477.
- Walker, N. & Stuart, D. (1983). *Acta Cryst. A***39**, 158–166.