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In Situ Photoconversion of Multicolor Luminescence and Pure White Light

Emission Based on Carbon Dot-Supported Supramolecular Assembly

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MAIN TEXT

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ABSTRACT: Constructing multicolor photoluminescence materials that allow for the integration of suitable external stimuli in order to control luminescence color conversions is a challenging objective. Multicolor luminescent output that is regulated in an in situ photo-controlled manner is not a common phenomenon. Herein, a two-stage assembled photoluminescence supramolecular assembly is described, that displays in situ phototuning broad-spectrum output. Benefiting from the reversible photo-switched constitutional interconversion of diarylethenes, the fluorescence of a guest molecule, styrylpyridinium-modified diarylethene, can be switched on/off by alternating ultraviolet and visible light irradiation. Upon complexing with a host, cucurbit[8]uril, the fluorescence intensity of the resulting binary supramolecular nanofiber shows a drastic enhancement when compared with that of the free guest, which can also be quenched and recovered reversibly by light irradiation. Significantly, such cationic supramolecular nanofibers also interact with anionic carbon dots to form broad-spectrum output ternary supramolecular assemblies, the fluorescence of which can be changed efficiently from yellow to blue in an in situ photo-controlled manner. Pure white light emission can be realized expediently in the luminescence color conversion process. The use of light as an external stimulus to regulate fluorescent color conversion provides us with an opportunity to design and construct more advanced anti-counterfeiting materials as well as visual display instruments.

■ INTRODUCTION

The construction of multicolor photoluminescence materials that can be tuned in a simple manner has attracted an increasing amount of attention¹, because of their potential applications as biological image reagents,² stretchable displays,³ multidimensional sensors,⁴ photoelectric devices,⁵ and light-emitting diodes⁶. The conventional methods of acquiring such materials are mainly physical blending or the covalent connections of chromophores with complementary fluorescence in accurate ratios.⁷ Meanwhile, luminescent materials that can be tuned for white-light emission are particularly important, since they can be applied practically in the fields of visual display media and lighting instruments.⁸ Thus far, development of this type of white-light emission materials is limited by the available preparation protocols or device fabrication techniques.⁹ In contrast to the conventional approaches, several noncovalent strategies have been developed in recent years for the synthesis of multicolor emission materials that respond to external stimuli.¹⁰ For example, Yam et al.¹¹ have constructed a new class of amphiphilic anionic platinum(II) assemblies that display tunable fluorescence emission by variation of solvent polarity. Ito et al.¹² have reported a gold(I) isocyanide complex that exhibits different

photoluminescent colors, which are interconvertible by treatment with acetone and mechanical shearing in the solid state. Tian et al.¹³ have utilized two-armed fluorescent molecules as guests and γ -cyclodextrin as the host to construct novel host–guest complexes, which show multicolor luminescence that relies on intramolecular charge transfer. Tao et al.¹⁴ have reported that various fluorescent emissions can be obtained by adjusting the electronic distributions of the *p*-phenylenevinylene units within the hydrophobic cavity of cucurbit[8]uril. The use of external stimuli to regulate multicolor luminescent output provides a convenient way to construct new classes of photoluminescence materials. Among these external stimuli, light is of particular interest, because of its clean, noninvasive, and remote-controlling nature.¹⁵ Reports of such fluorescence color conversion processes that regulated in a reversible photo-controlled manner are still rare.

Among the various building blocks¹⁶ for constructing photo-responsive supramolecular systems, diarylethene derivatives (DAEs), on account of their generally (i) rapid light response, (ii) no spontaneous thermal reversion, and (iii) outstanding photochromic properties,¹⁷ have led to many successful applications as optical switches,¹⁸ molecular modulators,¹⁹ helical transfers controller,²⁰ and information storage media.²¹ Meanwhile, carbon dots (CDs) are a popular class of carbon nanomaterials with the promising advantages of non-toxicity, eco-friendliness, high fluorescence, biocompatibility, good water solubility, excellent photostability, and easy surface functionalization.²² CDs have been, therefore, widely utilized as photosensitizers,²³ imaging-guided nanocarriers,²⁴ multidimensional memory materials,²⁵ and biosensors.²⁶ In this investigation, a photo-responsive multicolor luminescent supramolecular assembly is constructed (Scheme 1) from anionic carbon dots (CDs) and cationic supramolecular nanofibers that are composed of the styrylpyridinium-modified perfluorocyclopentene diarylethene derivative **1**²⁺ and cucurbit[8]uril (CB[8]). On account of the reversible photoisomerization of the diarylethene unit, the fluorescence of styrylpyridinium-modified diarylethene guest molecule **1**²⁺ and the resulting binary supramolecular nanofibers **1**²⁺⊂CB[8] can be switched on/off by alternating ultraviolet and visible light irradiation. Such photo-responsive cationic supramolecular nanofibers further co-assembled with the anionic CDs to form ternary supramolecular nanoparticles, which exhibit colorful emission capacities—i.e., gold, yellow, khaki, and blue—while the luminescent colors of the ternary supramolecular assembly can be interconverted efficiently by irradiated with ultraviolet and visible light. White light emission is also realized in the fluorescent color conversion process.

■ RESULTS AND DISCUSSION

Synthetic routes to the guest molecules **1**•2Cl / **2**•Cl and carbon dots are summarized in Schemes S1 and S2, respectively. The guest molecule **1**•2Cl is obtained in 78% yield by the condensation of the diarylethene derivative **6** with 4-picolinium iodide, followed by counterion exchanges. Employing a similar synthetic method, the reference compound **2**•Cl with better water solubility is prepared in 88% yield. The carbon dots are prepared by means of thermal pyrolysis with citric acid as the carbon source and diethylenetriamine as the surface passivation agent, followed by amide condensation with succinic anhydride. The UV-vis absorption spectrum (Figure 1b) of the open-ring form (**OF-1**²⁺) of **1**²⁺ shows a strong absorption peak at 395 nm. After irradiation of aqueous solutions of **OF-1**²⁺ with UV light (254 nm, 2.0 min), the absorption peak at 395 nm decreases gradually with a blue shift of ~10 nm, and is accompanied by the appearance of a new absorption band centered on 625 nm. Three clear isosbestic points are observed at 311, 360, and 465 nm in the UV-vis absorption spectrum. The photocyclization quantum yield (Φ_{o-c}) for **1**²⁺ is determined (Table S1) to be 0.30. In addition, color changes (Figure 1a) from pale yellow to blue are observed after UV light irradiating an aqueous solution of **1**²⁺. These phenomena are attributed²⁷ to (Figure 1a) the transformation of **1**²⁺ from the open-ring form (**OF-1**²⁺) to the closed-ring form (**CF-1**²⁺). When an aqueous solution of the closed-ring isomer (**CF-1**²⁺) is irradiated with visible light (>600 nm, 1.5 min), the color of the **CF-1**²⁺ solution returns to being pale yellow, and its UV-vis absorption spectrum almost resembles (Figure S8) that of **OF-1**²⁺. When excited at 395 nm, the fluorescence spectrum of **1**²⁺ reveals (Figure 1c) an emission peak centered on 560 nm. Upon irradiation with UV light (254 nm, 2.0 min), the fluorescence intensity of **1**²⁺ is gradually quenched by 82%, and the fluorescence quantum yield of **1**²⁺ decreased (Table S3) from 0.63% to 0.010%, because of the self-absorption effect of **CF-1**²⁺. The fluorescent on/off quenching ratio ($R_{on/off}$) is calculated to be 5 by comparing the intensity at maximum emission wavelength before and after UV light irradiation. When the aqueous solution is irradiated with visible light (>600 nm, 1.0 min), the fluorescence is recovered. These results indicate that the constitution of **1**²⁺ can go back and forth between the open-ring and closed-ring isomers upon alternating irradiation with UV and visible light. The excitation light shows the negligible influence on the photoisomerization reaction of **1**²⁺ under current experimental conditions (Figure S12), but this influence is inevitable when using a high-intensity excitation light source (Figure S13). Moreover, this interconversion could be repeated (Figures 1d and S8) for several cycles without appreciable light fatigue. The ring-closing conversion yield of **1**²⁺ is measured (Figure S11a) to be 79% upon irradiation with UV light. Furthermore, **CF-1**²⁺ can return quantitatively to **OF-1**²⁺ upon

irradiation (Figure S11b) with visible light. All these results demonstrate the excellent photo-switched constitutional interconversion of **1**²⁺.

Cucurbit[8]uril (CB[8])²⁸, which possesses high binding affinity towards a variety of neutral or cationic guest molecules, has been used widely as the assembler of supramolecular architectures in aqueous solutions. It can therefore be anticipated that CB[8] will bind to positively charged **1**²⁺. The supramolecular assembly consisting of **1**²⁺ and CB[8] is studied by UV-vis absorption, NMR, and fluorescence spectroscopies. Upon stepwise addition of CB[8] to an aqueous solution of **OF-1**²⁺, the absorption peak of **1**²⁺ centered on 395 nm shifts gradually (Figure S14) to 420 nm. The binding constant K_a between **OF-1**²⁺ and CB[8] is determined (Figure S14, inset) to be $1.19 \times 10^7 \text{ M}^{-1}$ in aqueous solutions. A Job's plot shows (Figure S15) a maxima at a molar fraction of 0.5, indicating 1:1 stoichiometric ratio between **OF-1**²⁺ and CB[8]. On account of the high affinity between CB[8] and **OF-1**²⁺, supramolecular nanofibers are obtained readily. Transmission electron microscope (TEM, Figure 2b) and atomic force microscope (AFM, Figure S16) images reveal one dimensional (1D) fine nanofibers with lengths of several micrometers. The average hydrodynamic diameter has been determined (Figure S17) to be 470 nm upon equimolar mixing of CB[8] and **1**²⁺, indicating the formation of highly polymerized supramolecular assemblies. In order to verify the binding mode of CB[8] with **1**²⁺, reference compound **2**⁺ with better water-solubility is synthesized (Figures S5–7). In the ¹H NMR spectrum (Figure S18) of a 2:1 mixture of **2**⁺ and CB[8], the chemical shifts of the protons (H₄ and H₅) residing on the C=C double bonds and those (H₆ and H₇) residing on the phenylene groups show upfield shifts ($\Delta\delta = 0.45, 0.42, 0.64$, and 0.30 ppm for H₄, H₅, H₆, and H₇, respectively). The resonances of the protons (H₁, H₂, and H₃) residing on the pyridinium rings and those (H₈ and H₉) on the thiophene rings are shifted downfield ($\Delta\delta = 0.03$ – 0.22 ppm). In addition, correlation peaks are observed (Figure S19) between pyridinium protons (H₁, H₂, and H₃) and thiophene protons (H₈ and H₉) in the two-dimensional (2D) ROESY spectrum. Combining the 1D and 2D NMR spectra, it is clear that CB[8] is bound to the phenylene and C=C double bond units. Considering the intermolecular steric effect between the pyridinium and perfluorocyclopentene units in the binary supramolecular assembly **OF-1**²⁺⊂CB[8], the bonding mode of **OF-1**²⁺⊂CB[8] is confirmed by theoretical calculations using the unilateral styrylpyridinium-modified diarylethene (**3**⁺) as a model compound. The energy optimized structure for the **OF-3**⁺⊂CB[8] is determined to be that (Figure S20 and Table S2) in which the phenylene group and C=C double bond units are included in the cavity of CB[8], which is in agreement with the result from ¹H NMR spectroscopy.

Upon complexation with CB[8], the fluorescence intensity of **OF-1**²⁺ increases (Figure 2a)

16-fold as compared to the inherent fluorescence of **OF-1²⁺**, benefiting from the fact that free rotation of C=C double bond is restricted by CB[8].²⁹ The fluorescence quantum yield of **OF-1²⁺⊂CB[8]** is measured to be 12% (Table S3), a value which is considerably higher than that (0.63%) for free **OF-1²⁺**. The maximum emission wavelength is shifted bathochromically from 560 to 582 nm as a result of the J-stacking of styrylpyridinium units in the cavity of CB[8].^{14,30} Furthermore, upon irradiation with the UV light (254 nm, 2.0 min), the fluorescence intensity of **1²⁺⊂CB[8]** is quenched (Figure 2c) gradually by 97%. The fluorescent on/off quenching ratio ($R_{\text{on/off}}$) is calculated to be 30. The fluorescence quenching can also be detected by the naked eye, wherein the bright yellow fluorescence of the **OF-1²⁺⊂CB[8]** gradually turns dark (Figure 2d, inset) upon irradiation with UV light. When the nonluminous aqueous solution is irradiated with visible light (>600 nm, 40 s), the yellow fluorescence recovers (Figure 2d, inset). These results indicate that the constitutions of **1²⁺** can also be interconverted between **OF-1²⁺** and **CF-1²⁺** in the host-guest complex upon light irradiation. The UV-vis absorption spectra also confirm the outstanding photo-conversation cycles of **1²⁺⊂CB[8]**. Upon irradiating aqueous solutions of **OF-1²⁺⊂CB[8]** with UV light (254 nm, 2.0 min), the strong absorption peak at 420 nm decreases gradually (Figure S21) and a new absorption peak appears at 655 nm with a 30 nm bathochromic shift compared with that (625 nm) of **1²⁺**. The UV-vis absorption spectrum reverts back to its original state upon visible light irradiation (>600 nm, 1.0 min). These processes can be repeated (Figure S22) through several cycles without appreciable light fatigue, an observation which is consistent (Figure 2d) with the fluorescence spectra.

In order to utilize fully the photo-switched fluorescence behavior of the present cationic supramolecular assembly **1²⁺⊂CB[8]**, carbon dots (CDs) are employed as a model substrate in the construction of a broad-spectrum output luminescent supramolecular material. There are several inherent advantages of introducing CDs into the **1²⁺⊂CB[8]** supramolecular assembly. (1) The fact that CDs emit bright blue fluorescence, which is complemented by the yellow fluorescence of the supramolecular nanofiber²⁵ **1²⁺⊂CB[8]**. (2) The spherical distribution of charged functional groups on the surface of CDs provide an ideal adhesion agent to promote the co-assembly with the cationic supramolecular nanofiber, **1²⁺⊂CB[8]**. Experimentally, CDs are prepared by thermal pyrolysis and their morphology is investigated by TEM. A number of black dots with a size distribution of 4–8 nm are observed in the TEM image (Figure 3a). The average hydrated radius is 5.5 nm (Figure S23). The high resolution TEM analysis reveals (Figure S24) the lattice spacings of CDs to be 0.39 nm and 0.22 nm, corresponding to the diffraction peaks (Figure S25) centered on 40.2° and 22.5° in the powder X-ray diffraction

pattern, respectively. This result indicates that the CDs are highly crystalline.³¹ The UV-vis absorption spectrum of the CDs exhibits (Figure S26a) two absorption peaks at 260 and 360 nm, which can be assigned³² to the π - π^* transitions of the aromatic ring components and the n - π^* transition of the C=O group on the surface of CDs, respectively. When excited at 360 nm, strong blue fluorescence is observed (Figure 3c, inset). The fluorescence spectrum reveals (Figure 3c) an emission peak centered on 458 nm, and the corresponding fluorescence quantum yield (Table S3) is 31%. The fluorescence spectra of the CDs are excitation wavelength- and pH-dependent. The emission peaks are shifted (Figure 3d) bathochromically from 458 to 537 nm, when the excitation wavelengths are changed from 360 to 480 nm. The fluorescence intensity increases (Figure S26b) progressively with a change in pH from 2 to 7. There are no obvious changes in fluorescence intensity upon further enhancing the basicity. The functional groups on the surface of the CDs have been studied by Fourier transform infrared (FTIR) spectroscopy and X-ray photoelectron spectroscopy (XPS). The broad vibrational bands in the range of 3000–3500 cm^{-1} can be assigned to O–H and N–H stretching vibrations. The vibration bands at 1633, 1550, 1417, and 1259 cm^{-1} can be ascribed to the stretching vibration of C=O, bending vibration of N–H, and the stretching vibrations of C–N and C–O, respectively (Figure S27). The XPS spectrum shows (Figure S28) strong signal peaks for the elements C, O, and N. These results demonstrate that there are a lot of carboxylic acids and amide groups distributed on the surface of the CDs. In addition, the zeta potentials of the CDs decrease (Figure S29) gradually from 30.6 to –48.1 mV, when the pH is changed from 2 to 9. In a neutral aqueous solution, the zeta potential of the CDs (Figure 3b) is –32.3 mV, indicating that the carboxylic acids are the dominant functional groups on the surface of the CDs.

On account of the good luminescent behavior and the anionic surface functional groups of the CDs, the luminescent supramolecular assembly (CDs@ $\mathbf{1}^{2+}$ CB[8]) can be readily constructed by electrostatic interactions between the cationic supramolecular nanofibers $\mathbf{1}^{2+}$ CB[8] and the anionic CDs. The scanning electron microscopy (SEM) image of the supramolecular assembly CDs@ $\mathbf{1}^{2+}$ CB[8] shows (Figure S30b) numerous spherical nanoparticles with an average diameter of ~700 nm, consistent with the features observed (Figure S30a) in the TEM image. A dynamic light scattering (DLS) experiment reveals (Figure S31) the average hydrodynamic diameter of the CDs@ $\mathbf{1}^{2+}$ CB[8] nanoparticles to be 757 nm in neutral aqueous solution. In addition, the zeta potential of CDs@ $\mathbf{1}^{2+}$ CB[8] supramolecular nanoparticles is determined (Figure S32) to be –16.9 mV, a value which is higher than that (–32.3 mV) of CDs. All of these results demonstrate consistently that the cationic $\mathbf{1}^{2+}$ CB[8] supramolecular nanofibers are able to co-assemble efficiently with the anionic CDs, courtesy

of electrostatic interactions. $\text{CDs@OF-1}^{2+}\text{CB[8]}$ exhibit multicolor fluorescence emission. With the gradual addition of $\text{OF-1}^{2+}\text{CB[8]}$ to the 1 $\mu\text{g/mL}$ aqueous solution of CDs, the fluorescence emission peak, centered on 582 nm, derived from $\text{1}^{2+}\text{CB[8]}$ increases gradually, and the fluorescence emission peak centered on 458 nm originating from the CDs undergoes (Figure 4a) a slight decrease. The fluorescence colors of the resulting ternary supramolecular assembly (Figure 4c) change linearly from bright blue to yellow, as indicated by the CIE 1931 chromaticity diagram. It is noteworthy that white-light emission (0.34, 0.33) is observed when the concentration of the $\text{OF-1}^{2+}\text{CB[8]}$ supramolecular assembly is $\sim 1.17 \mu\text{M}$. The quantum yield of the resulting white-light emission is $\sim 6.8\%$. Benefiting from the reversible photo-controlled luminescent behavior of $\text{1}^{2+}\text{CB[8]}$, the fluorescence of the resultant ternary supramolecular assembly ($\text{CDs@1}^{2+}\text{CB[8]}$) can also be regulated by alternating UV and visible light irradiation. The fluorescence spectra of the $\text{CDs@1}^{2+}\text{CB[8]}$ shows (Figure 4b) two emission bands centered on 458 and 582 nm, which can be ascribed to the fluorescence emissions of CDs and $\text{1}^{2+}\text{CB[8]}$, respectively. Upon irradiation with UV light (254 nm, 1.5 min), the emission peak centered on 582 nm decreases dramatically, accompanied by a slight decrease in the emission peak at 458 nm. All of the luminescent color coordinates have been calculated and plotted (Figure 4d) in the CIE 1931 chromaticity diagram, wherein the fluorescence colors of $\text{CDs@1}^{2+}\text{CB[8]}$ change linearly from yellow to blue. The pure white light emission (0.33, 0.33) is also realized upon UV light irradiation for ~ 24 s, and the corresponding ratio of OF-1^{2+} to CF-1^{2+} is 32:68. The aqueous solutions of $\text{CDs@1}^{2+}\text{CB[8]}$ exhibit (Figure 4e) colorful emissions—i.e., gold, yellow, khaki, and blue—with prolonging the UV light irradiation time. Subsequently, when the solution is irradiated with visible light (>600 nm, 30 s), the fluorescence emission of the solution returns to yellow, indicating that the fluorescence colors of $\text{CDs@1}^{2+}\text{CB[8]}$ can be regulated efficiently by light. In control experiments, the emission peak of the binary supramolecular assembly (CDs@OF-1^{2+}) in the range of 500 to 700 nm is weaker obviously than that of $\text{CDs@OF-1}^{2+}\text{CB[8]}$ (Figure S33), indicating that electrostatic interactions between the CDs and OF-1^{2+} do not effectively restrict the free rotation of C=C double bonds in OF-1^{2+} . There is no appreciable change (Figure S34) of fluorescence color when only the cationic OF-1^{2+} is gradually added to the 1 $\mu\text{g/mL}$ aqueous solution of CDs. Upon irradiation an aqueous solution of CDs@OF-1^{2+} with UV light (254 nm, 1.5 min), there is also no appreciable change in the fluorescence color (Figure S35). These results confirm that CB[8] is an indispensable component in these spectral tunable luminescent supramolecular materials. When considering the rational design of supramolecular assembly

with photo-responsive multicolor luminescence output, there are three main points need to be considered. (1) The introduction of outstanding photo-responsive components, which are indispensable for the desired photo-controlled behavior. (2) The complementary fluorescence color of fluorophores, a factor which is crucial for constructing the broad-spectrum output materials. (3) The appropriate noncovalent bonding interactions to integrate the fluorophores. They are important in constructing the multicolor luminescence supramolecular assembly. It is believed that these design principles, not only provide ideas for designing broad-spectrum outputs photoluminescence materials, but also serve as a reference for constructing more advance anti-counterfeiting materials and photo-controlled molecular devices.

■ CONCLUSION

A photo-switched multicolor luminescent supramolecular assembly has been realized from anionic carbon dots and cationic supramolecular nanofibers that are comprised of a styrylpyridinium-modified diarylethene derivative $\mathbf{1}^{2+}$ and cucurbit[8]uril. On account of the reversible photo-switchable constitutional interconversion properties of dithienylethene, the fluorescence of $\mathbf{1}^{2+}$ and the binary supramolecular nanofibers can be switched on/off by alternating UV and visible light irradiation in aqueous solutions. In addition, a ternary luminescent supramolecular assembly shows multicolor fluorescence emissions, which can be interconverted efficiently in an in situ photo-controlled manner. Pure white light emission can also be realized conveniently in the photo-switched luminescence color conversion process. These smart supramolecular luminescent materials have potential applications in the design and synthesis of anti-counterfeiting materials and photo-controlled molecular devices.

■ ASSOCIATED CONTENT

Supporting Information

Detailed synthetic procedures, microscopic morphology, and spectroscopic (NMR, HRMS, UV-vis absorption, and fluorescence) characterization data for the guest molecules $\mathbf{1} \cdot 2\text{Cl}$ and $\mathbf{2} \cdot \text{Cl}$, carbon dots (CDs), supramolecular fibers ($\mathbf{1}^{2+} \subset \text{CB}[8]$) and the ternary supramolecular assembly ($\text{CDs} @ \mathbf{1}^{2+} \subset \text{CB}[8]$). This information is available free of charge via the internet at <http://pubs.acs.org>.

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Notes

The authors declare no competing financial interest.

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Captions to Scheme and Figures

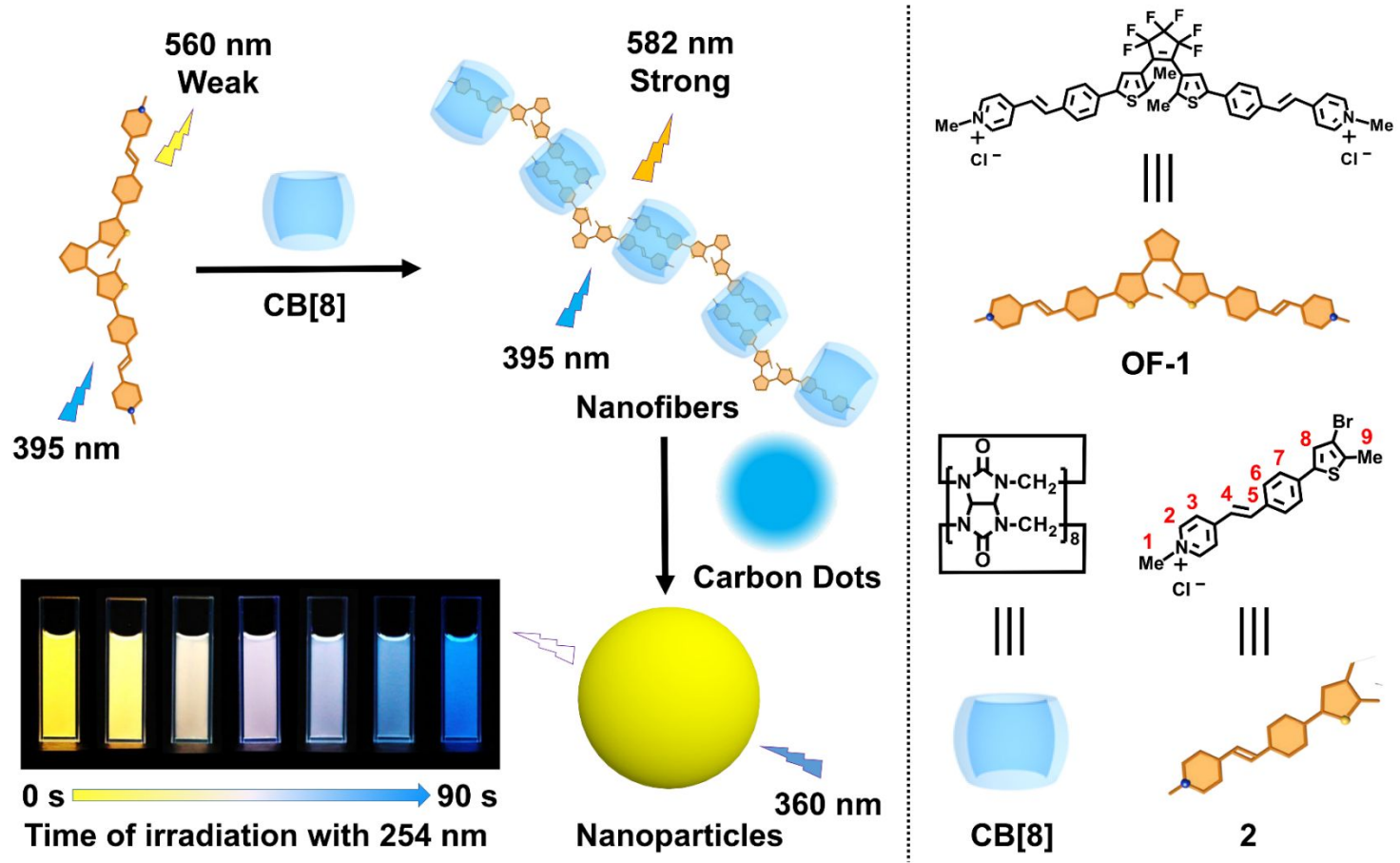
Scheme 1. Schematic illustration of the in situ photo-switched multicolor luminescence supramolecular assembly.

Figure 1. (a) Structural formulas and photographic image of the changes of $\mathbf{1}^{2+}$ upon alternating irradiation with UV and visible light. (b) UV-vis absorption spectra and (inset) absorbance intensity changes of $\mathbf{OF-1}^{2+}$ at 625 nm ($[\mathbf{OF-1}^{2+}] = 2.0 \times 10^{-5}$ M, 298 K) upon irradiation with UV light (254 nm, 2.0 min). (c) Emission spectra and (inset) emission intensity changes of $\mathbf{OF-1}^{2+}$ at 560 nm ($[\mathbf{OF-1}^{2+}] = 5.0 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 395$ nm, 298 K) upon irradiation with UV light (254 nm, 2.0 min). (d) Absorbance changes at 625 nm of $\mathbf{1}^{2+}$ ($[\mathbf{1}^{2+}] = 2.0 \times 10^{-5}$ M, 298 K) upon alternating irradiation with UV (254 nm, 2.0 min) and visible (>600 nm, 1.5 min) light in aqueous solutions.

Figure 2. (a) Emission spectra and (inset) emission intensity changes of $\mathbf{OF-1}^{2+}$ at 582 nm ($[\mathbf{OF-1}^{2+}] = 5.0 \times 10^{-6}$ M, $[\text{CB}[8]] = 0\text{--}10.5 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 395$ nm, 298 K) upon gradual addition of CB[8] in an aqueous solution. (b) TEM image of $\mathbf{OF-1}^{2+}\text{CB}[8]$ supramolecular assembly. (c) Emission spectra and (inset) emission intensity changes of $\mathbf{OF-1}^{2+}\text{CB}[8]$ at 582 nm ($[\mathbf{OF-1}^{2+}\text{CB}[8]] = 5.0 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 395$ nm, 298 K) upon irradiation with UV light (254 nm, 2.0 min). (d) Emission spectra and (inset) emission intensity changes of $\mathbf{OF-1}^{2+}\text{CB}[8]$ at 582 nm ($[\mathbf{OF-1}^{2+}\text{CB}[8]] = 5.0 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 395$ nm, 298 K) upon alternating irradiation with UV (254 nm, 2.0 min) and visible light (>600 nm, 40 s) in an aqueous solution.

Figure 3. (a) TEM image of the anionic CDs. (b) Zeta potential of the anionic CDs in a neutral aqueous solution. (c) Excitation and emission spectra of the anionic CDs ($[\text{CDs}] = 3$ $\mu\text{g/mL}$, 298 K). (d) Emission spectra of the anionic CDs ($[\text{CDs}] = 3$ $\mu\text{g/mL}$, 298 K) at various excitation wavelengths in a neutral aqueous solution.

Figure 4. (a) Emission spectra of CDs ($[\text{CDs}] = 1$ $\mu\text{g/mL}$, $[\mathbf{OF-1}^{2+}\text{CB}[8]] = 0\text{--}3 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 360$ nm, 298 K) with gradual addition of $\mathbf{OF-1}^{2+}\text{CB}[8]$ in a neutral aqueous solution. (b) Emission spectral changes of $\text{CDs@}\mathbf{1}^{2+}\text{CB}[8]$ ($[\text{CDs}] = 1$ $\mu\text{g/mL}$, $[\mathbf{OF-1}^{2+}\text{CB}[8]] = 3 \times 10^{-6}$ M, $\lambda_{\text{ex}} = 360$ nm, 298 K) upon irradiation with UV light (254 nm, 1.5 min) in a neutral aqueous solution. (c) The 1931 CIE chromaticity diagram illustrating the luminescent color changes of CDs with the gradual addition of $\mathbf{OF-1}^{2+}\text{CB}[8]$, corresponding to (a). (d) The 1931 CIE chromaticity diagram depicting the luminescent color changes of $\text{CDs@}\mathbf{1}^{2+}\text{CB}[8]$ upon irradiation with UV light (254 nm, 1.5 min), corresponding to (b). (e) Fluorescence photographs of $\text{CDs@}\mathbf{1}^{2+}\text{CB}[8]$ ($[\text{CDs}] = 3$ $\mu\text{g/mL}$, $[\mathbf{OF-1}^{2+}\text{CB}[8]] = 9 \times 10^{-6}$ M) with increasing the UV light (254 nm, 1.5 min) irradiation time in a neutral aqueous solution.



Scheme 1

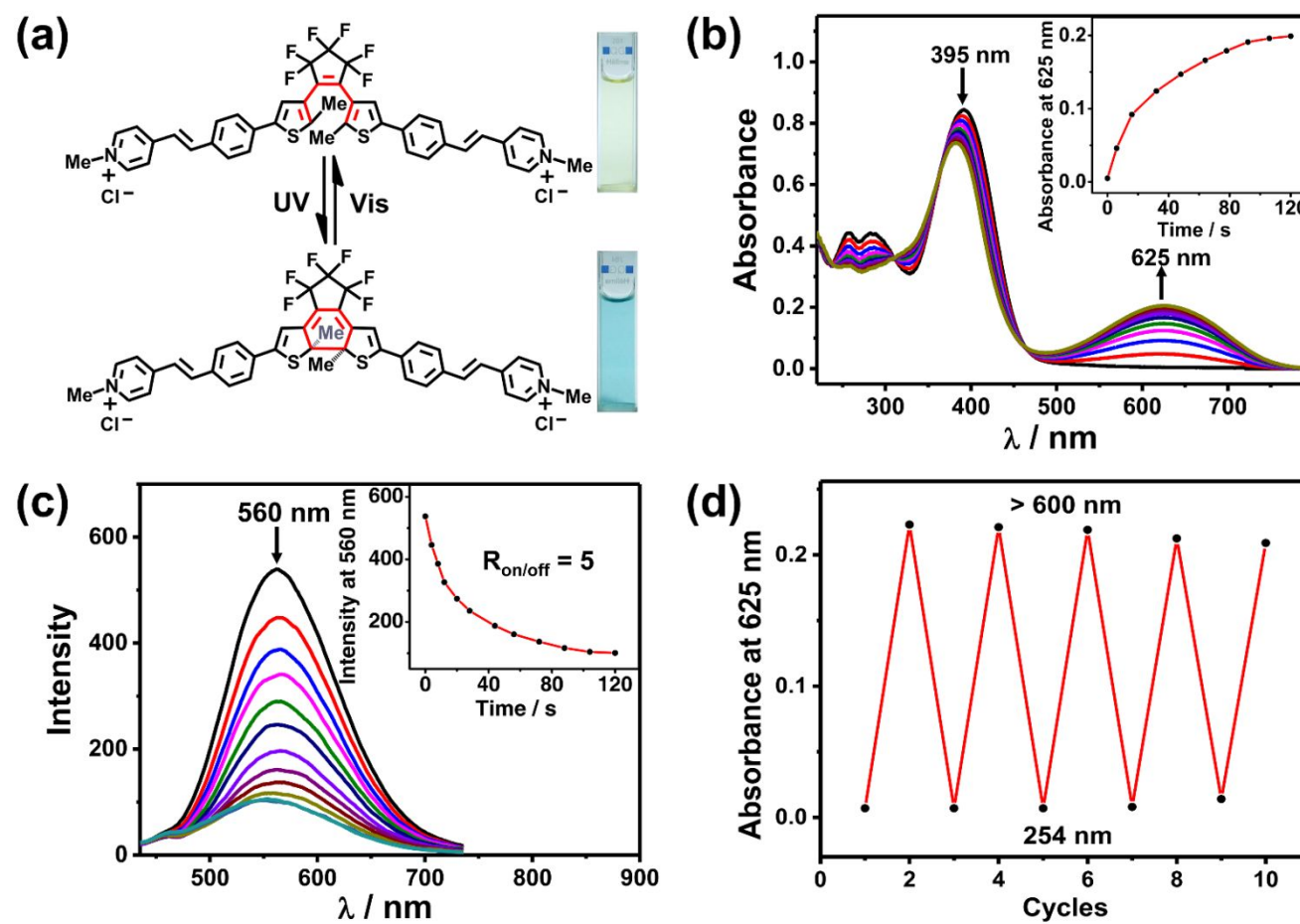


Figure 1

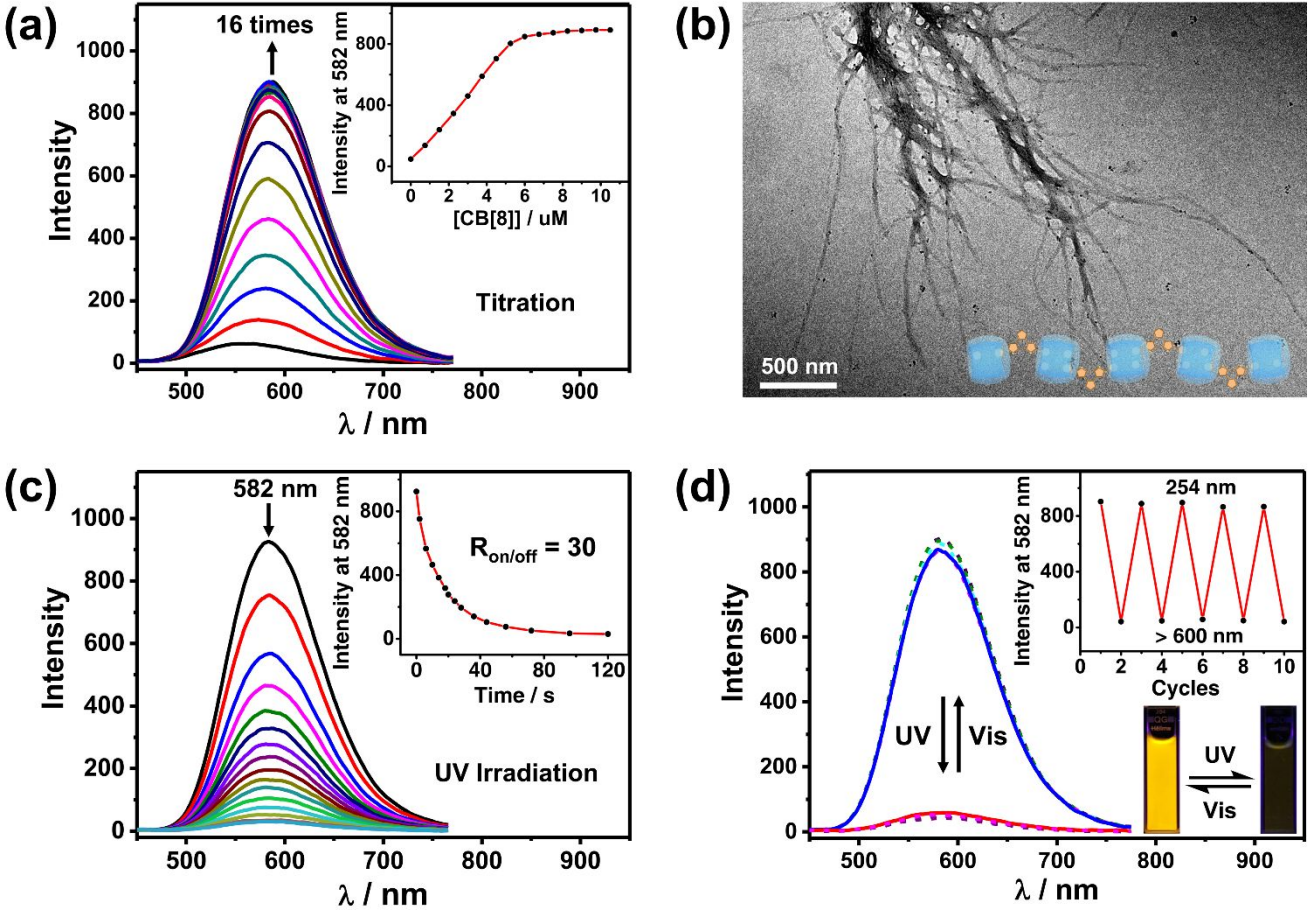


Figure 2

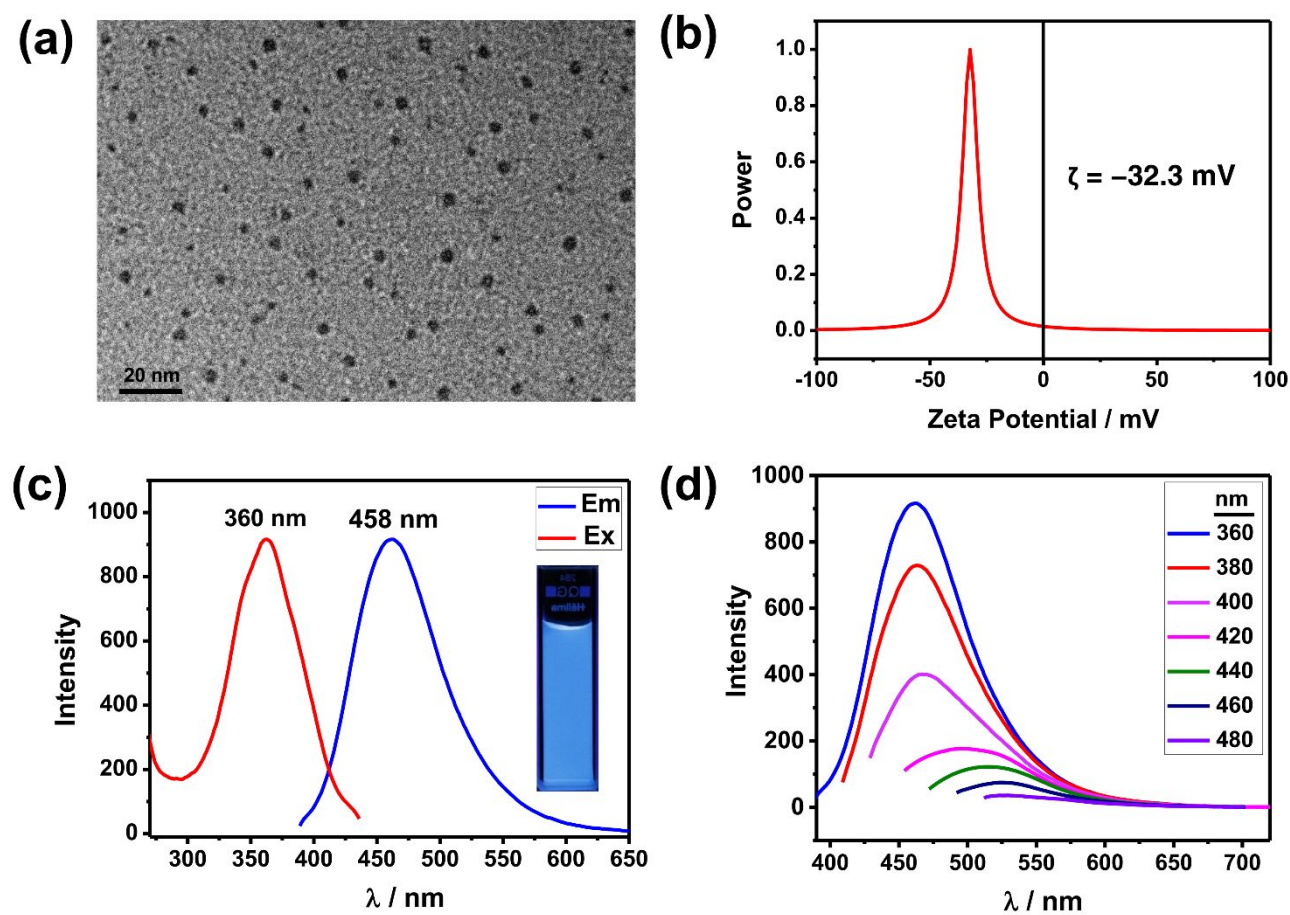


Figure 3

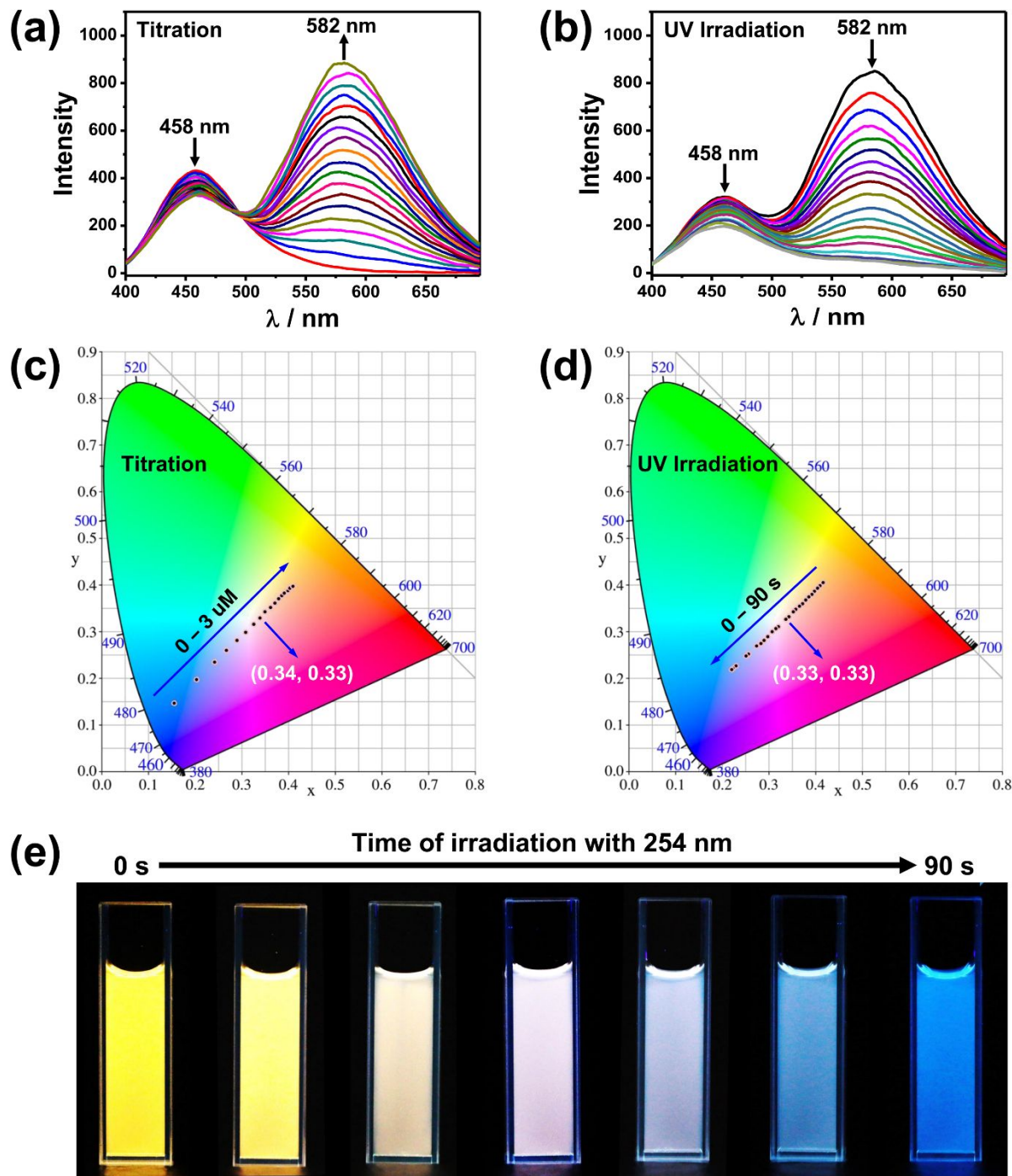


Figure 4

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