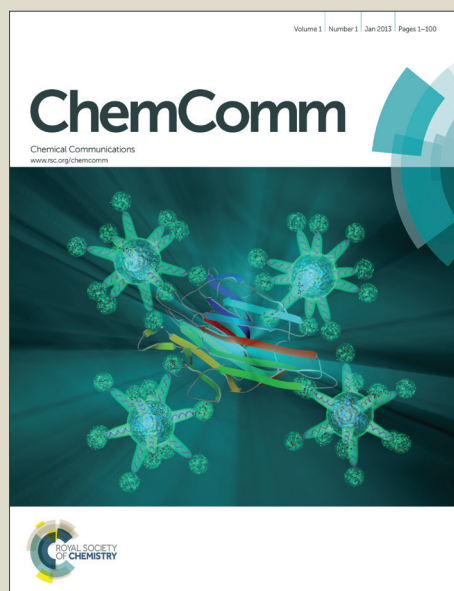


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COMMUNICATION

Al(OH)₃ Facilitated Synthesis of Water-soluble, Magnetic, Radiolabelled and Fluorescent Hydroxyapatite Nanoparticles

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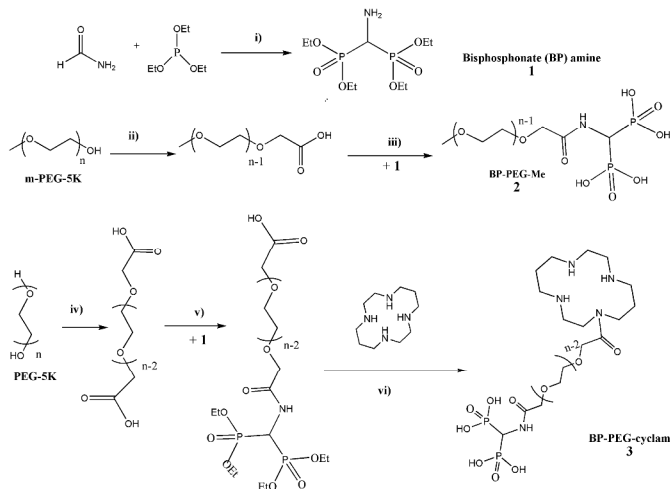
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Magnetic and fluorescent hydroxyapatite nanoparticles were synthesised using Al(OH)₃-stabilised MnFe₂O₄ or Fe₃O₄ nanoparticles as precursors. They were readily and efficiently radiolabelled with ¹⁸F. Bisphosphonate polyethylene glycol polymers were utilised to endow the nanoparticles with excellent colloidal stability in water and to incorporate cyclam for high affinity labelling with ⁶⁴Cu.

Molecular imaging techniques, including magnetic resonance imaging (MRI), positron emission tomography (PET), single photon emission computed tomography (SPECT) and fluorescence optical imaging, play an increasingly important role in clinical diagnosis and management of disease, as well as medical and biological research. Multimodal imaging recently has gained attention because of its potential to overcome the limitations of individual imaging modalities and to provide more accurate and complete physiological information at sites of disease.¹⁻³ Numerous nanoparticles (NPs) have been studied as multimodal imaging contrast agents due to their multi-functionality and potential for surface modification.⁴⁻⁷ An adequate multimodal nanoparticulate contrast agent must be multifunctional, biocompatible and colloidally stable. The NPs should be uniform in morphology and size, so that they share similar *in vivo* behaviour, and chemically stable to ensure that the signal of each modality reflects the same anatomic position.

Hydroxyapatite (HA) has attracted much interest as the basis of multifunctional probes very recently,⁸⁻¹¹ because of its biocompatibility and high affinity for fluoride which allows facile labelling with the positron emitter ¹⁸F. Fluorescent HA can be obtained by either doping with rare earth cations^{8,11} or by conjugation with organic dyes¹⁰. HA is not an ideal fluorescent host matrix, so luminescent rare earth doped HA nanocrystal requires up

to 20% replacement of OH⁻ by F⁻ (maximum theoretical value for fluoride substitution), to minimise the quenching of the excited state of rare earth cations.^{8,12,13} As a result, such HA is no longer suitable for ¹⁸F radiolabelling. It has been reported that magnetic iron oxide NPs can be deposited on the surface of HA aggregates or NPs *via* thermolysis¹¹ or a wet chemistry approach.⁹ One problem that remained unsolved for both synthetic approaches is how to effectively isolate the desired Fe₃O₄/HA composites from the unwanted iron oxides and HA nanoparticles. Moreover, all these multifunctional HA NPs suffer from the problem of aggregation or large size to some extent, which is an obstacle for their biological or medical applications. In this work, we present a novel synthesis of magnetic and fluorescent HA nanocomposites with uniform size and morphology, and excellent colloidal stability in water by using Fe₃O₄ nanoparticles stabilised with Al(OH)₃ as a template. The radiolabelling, magnetic and optical properties were investigated, to demonstrate potential for application as tri-modal probes for MR, PET and optical imaging.



Scheme 1: synthesis of bisphosphonate PEG, **i**) dropwise addition of POCl₃ at -7.5 °C under N₂, extract with CH₂Cl₂/H₂O/CH₂Cl₂ under basic/acidic/basic conditions; **ii**) CrO₃/H₂SO₄, 24 hours; **iii**) react with **1** and *N,N'*-dicyclohexylcarbodiimide (DCC), CH₂Cl₂, 18 h then with TMBS, remove solvent, stir in methanol; **iv**) CrO₃/H₂SO₄, 24 hours; **v**) react with **1** and DCC, CH₂Cl₂, 18 hours; **vi**) react with 1,4,8,11-tetraazacyclotetradecane and DCC, CH₂Cl₂, 18 hours, then with TMBS, remove solvent, stir in methanol for 2 h.

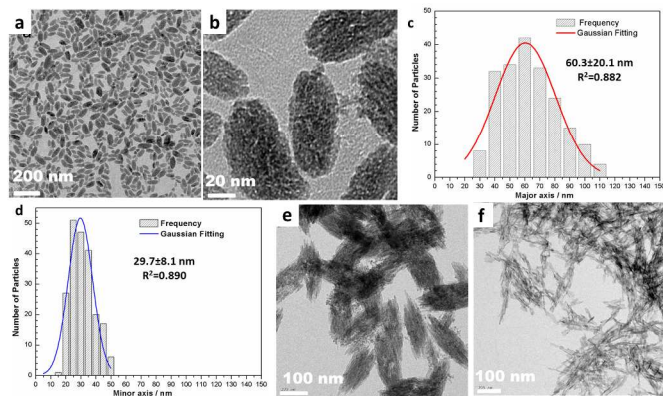


Fig. 1 TEM images **a**) and **b**) and size distribution **c**) and **d**) (major and minor axis, respectively) of Fe₃O₄@HA NPs synthesised from Fe₃O₄@Al(OH)₃ in the presence of BP-PEG-Me; **e**) TEM image of MnFe₂O₄@HA NPs synthesised in absence of BP-PEG-Me **2**; **f**) TEM image of pure HA NPs.

Our strategy is to synthesise HA using water-soluble magnetic Fe₃O₄@Al(OH)₃ or MnFe₂O₄@Al(OH)₃ NPs as templates. The advantages of this approach is the small hydrodynamic size of the template particles and their excellent colloidal stability, provided by the Al(OH)₃ layer as reported previously.¹⁴ More importantly, the layer of Al(OH)₃ can be readily removed as it is soluble under basic pH conditions. The design incorporated bisphosphonate polyethylene glycol (BP-PEG) polymers (Scheme 1) to stabilise NPs after the formation of HA on the surface, to take advantage of the outstanding binding affinity of bisphosphonates to HA.

Bisphosphonate amine **1** was obtained *via* a slightly modified version of the previously reported protocol.¹⁵ PEG carboxylic acids were obtained by oxidation of corresponding PEG polymers with CrO₃/H₂SO₄ *via* the reported protocol.¹⁶ The bisphosphonate (BP) or 1,4,8,11-tetraazacyclotetradecane (cyclam) were grafted to PEG

polymeric chain *via* amide formation mediated by *N,N'*-dicyclohexylcarbodiimide (DCC), followed by deprotection with bromotrimethylsilane (TMBS)¹⁷. BP-PEG-Me **2** and BP-PEG-cyclam **3** were purified by dialysis for over 24 hours using a membrane with a cut-off size of 3500 Da to remove unconjugated small molecules such as bisphosphonate amine **1** and 1,4,8,11-tetraazacyclotetradecane. The conjugation of bisphosphonate and PEG was confirmed by the change in chemical shift in the ³¹P NMR spectrum (from 20 ppm for free bisphosphonate to 12.8 ppm for BP-PEG, see Supporting Information).

Nanoparticulate precursors MnFe₂O₄@Al(OH)₃ and Fe₃O₄@Al(OH)₃ were obtained *via* a method reported by our group previously.¹⁴ Typically, 4 ml Fe₃O₄@Al(OH)₃ colloids (concentration of Fe₃O₄, *ca.* 8 mg/ml) and 200 mg BP-PEG-Me polymers **2** were placed in a 500 ml flask containing 300 ml water. Under stirring, 4 ml 0.1 mol/l NaH₂PO₄ and 4 ml 0.2 mol/l CaCl₂ aqueous solutions were added sequentially. This light brown solution was refluxed overnight after the addition of 30 ml 28% ammonia water. The NPs were collected by centrifugation at 2000 g for 30 minutes, re-dissolved in 10 ml water and freeze dried.

The X-ray powder diffraction (XRD) pattern of NPs indicates co-existence of HA and Fe₃O₄ (or MnFe₂O₄). TEM images show an olive-like morphology for Fe₃O₄@HA and MnFe₂O₄@HA NPs, which is significantly different both in size and in morphology from pure HA NPs synthesised under the same conditions (Fig. 1). Particle analysis by TEM gave a mean size of 60.3 nm (major axis) x 29.7 nm (minor axis) for Fe₃O₄@HA NPs (Fig. 1c&d). MnFe₂O₄@HA NPs displayed a similar aspect ratio to Fe₃O₄@HA NPs, but the size was almost doubled. These results indicate an important role of BP-PEG-Me in reducing particle size.

The hydrodynamic size of NPs during the synthesis was monitored by dynamic light scattering (DLS) experiments. No obvious change was observed after adding the solutions of NaH₂PO₄ and CaCl₂ into the solution of Fe₃O₄@Al(OH)₃ NPs, and it remained around 80 nm. This leads to a hypothesis that phosphate anions adsorb on the highly positive charged surface of Fe₃O₄@Al(OH)₃ NP and then react with the subsequently added Ca²⁺ to form calcium phosphate around the NPs. At the elevated temperature and basic solution environment, the outer layer of calcium phosphate was converted to crystallised HA, meanwhile the Al(OH)₃ was dissolved, resulting in the formation of Fe₃O₄@HA NPs. The diminished peak due to -OH around 3300-3500 cm⁻¹ in the IR spectrum, together with XRD patterns, confirm the replacement of Al(OH)₃ layer by HA. The changes in surface potential of NPs could not be monitored by measuring the zeta potential during the process, since the polymeric PEG imposes a thick hydration layer on NP surface and zeta potential no longer correlates to the surface potential. Therefore the surface potential was monitored in the absence of BP-PEG-Me, and a significant decrease in zeta potential was observed after the addition of NaH₂PO₄ solution, from 42.5 mV to 24.9 mV. This is presumed to be due to the adsorption of phosphate anions on the surface, since the changes in pH should be negligible in this case. The zeta potential slightly increased back to 27.1 mV after adding the CaCl₂ solution, indicating a reaction of calcium cations and phosphate anions. Similar results were also observed for MnFe₂O₄@HA NPs. In this synthesis approach, the positively charged Al(OH)₃ layer is essential for the formation of MnFe₂O₄@HA or Fe₃O₄@HA. Using the MnFe₂O₄ NPs colloids instead of MnFe₂O₄@Al(OH)₃ as precursors, a simple mixture of magnetic MnFe₂O₄ and non-magnetic HA NPs was obtained, identifiable as two kinds of NPs with apparently different morphology and size on TEM images.

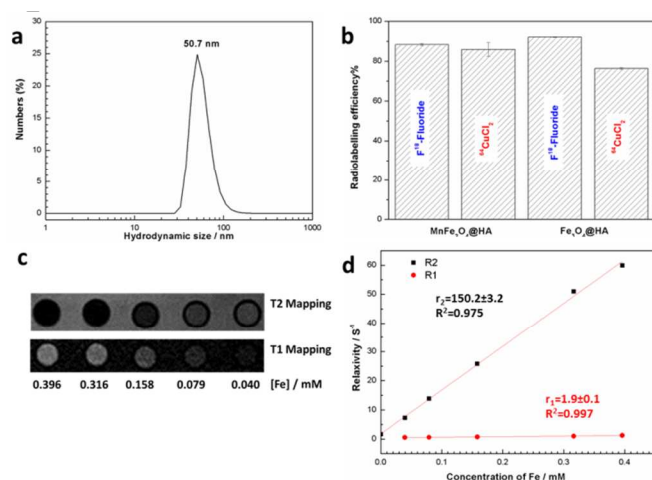


Fig. 2 a) DLS size distribution of 1 mg/ml Fe₃O₄@HA NPs in aqueous solution; b) Radiolabelling of MnFe₂O₄@HA and Fe₃O₄@HA NPs with ^{18}F -fluoride and $^{64}\text{CuCl}_2$ (the latter after sonicating the particles with BP-PEG-cyclam); c) T₁ and T₂ weighted MR images of the solution containing Fe₃O₄@HA NPs, and d) relaxivities of Fe₃O₄@HA NPs. Concentration of iron in the solution was measured by ICP-MS.

Both MnFe₂O₄@HA and Fe₃O₄@HA NPs were coated by BP-PEG-Me polymers **2** during the synthesis, as confirmed by *ca.* 18% mass loss on thermogravimetric analysis (TGA). Due to the strong interactions between the bisphosphonate group of **2** and MnFe₂O₄@HA and Fe₃O₄@HA NPs,^{17, 18} both NPs exhibit long-term colloidal stability in aqueous solution, even in high ionic strength environment such as PBS. The hydrodynamic size of Fe₃O₄@HA and MnFe₂O₄@HA NPs remained at 50.7 nm and 60.3 nm, respectively, for over two months (Fig. 2a). The excellent colloidal stability and small hydrodynamic size of MnFe₂O₄@HA and Fe₃O₄@HA make them potentially suitable for biological or medical applications.

Unsurprisingly, because of the high affinity of fluoride for HA,¹⁹ both NPs exhibit a high radiolabelling efficiency with ^{18}F -fluoride, up to 88.3±0.5% for 0.3 mg MnFe₂O₄@HA NPs and 92.1±0.1% for 0.3 mg Fe₃O₄@NaYF₄ NPs (Fig. 2b). Labelling and purification was readily achieved in less than 23 min. To provide a means of incorporating the positron emitter ^{64}Cu , the NPs were sonicated in 1 mg/ml BP-PEG-cyclam solution for 30 minutes to allow replacement of a fraction of BP-PEG-Me by BP-PEG-cyclam, and free BP-PEG-polymers were removed by centrifugation before mixing with radioactivity. The resulting particles showed a high ^{64}Cu radiolabelling efficiency in a short time (< 5 minutes) (Fig. 2b). Both NPs display essentially the magnetic properties of Fe₃O₄ or MnFe₂O₄ NPs and are active on MR images (Fig. 2c&d). The transverse (r_2) and longitudinal (r_1) relaxivities of Fe₃O₄@HA NPs were measured to be 150.2±3.2 mM⁻¹s⁻¹ and 1.9±0.1 mM⁻¹s⁻¹, respectively, at 3T magnetic field. As expected, the relaxivities of NPs could be improved by altering the ratio of magnetic component and non-magnetic HA, since r_2 is proportional to the volume fraction of magnetic component²⁰. For example, the r_2 of MnFe₂O₄@HA NPs could be dramatically improved from 105.7±3.5 mM⁻¹s⁻¹ to 246.5±15.9 mM⁻¹s⁻¹ by doubling the amount of MnFe₂O₄@Al(OH)₃ while keeping the amount of NaH₂PO₄ and CaCl₂ solutions the same during the synthesis. High transverse relaxivity of these magnetic

hydroxyapatite NPs as well as a high ratio of r_2/r_1 demonstrate their potential application as T₂ contrast agents on MR imaging.

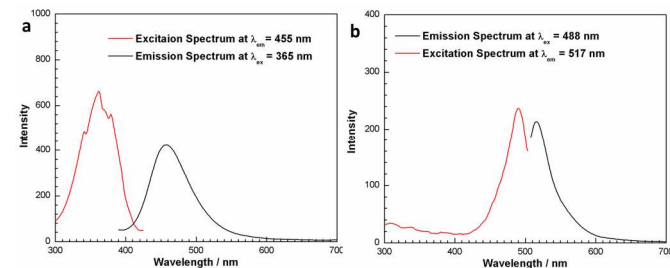


Fig. 3 Emission and excitation spectra of aqueous solutions of Fe₃O₄@HA conjugates with a) Maria blue and b) Alexa Fluor[®] 488.

Fluorescent HA is normally produced either by doping with rare earth cations (Eu or Tb),^{8, 11, 21} or by conjugation with organic dyes.^{10, 22} Here we conjugated the fluorescent dyes Maria blue and Alexa Fluor[®] covalently to the NPs surface using sodium pamidronate as an aminobisphosphonate linker (see Supplementary Information). The amine group of pamidronate is reactive for NHS ester dyes to form stable amide bonds while its bisphosphonate group interacts strongly with the Ca or Fe at the surface of NPs; dyes are thus linked to NPs without the risk of leakage. Fluorescent spectra in Fig. 3 show an emission at 455 nm for the conjugates of Fe₃O₄@HA and Maria blue under excitation at 365 nm, and an emission at 517 nm for the Alexa Fluor[®] 488 conjugates under excitation at 488 nm. More importantly, the fluorescence of these conjugated NPs is stable and strong even after being stored at room temperature for over one month, implying the potential applications as optical contrast.

In summary, we have presented a facile approach to synthesise magnetic and fluorescent hydroxyapatite nanoparticles with a well-defined morphology and uniform size, using Al(OH)₃-stabilised Fe₃O₄ or MnFe₂O₄ NPs as templates. The change from the highly positively charged surface and base-solubility of the Al(OH)₃ layer to the neutral or slightly negative zeta potential and acid solubility of the Fe₃O₄@HA particles is likely to offer alternative biological properties. These NPs are promising candidates for development as tri-modal probes for MR, PET and optical imaging, since they display excellent colloidal stability and high radiolabelling efficiency both for ^{64}Cu and for ^{18}F , as well as fluorescent and magnetic properties. Radiolabelling with other metallic radioisotopes will also be achievable by replacing cyclam with corresponding chelators. This synthesis approach allows us to tune the magnetic properties of particles by altering the ratio of precursors, without decreasing the radiolabelling efficiency or fluorescent property. The flexible conjugation method ensures that dyes with different wavelengths could be selected for different applications. The synthesis strategy for conjugation of BP-PEG-cyclam can also be applied to the conjugation of BP-PEG with peptides or antibodies, leading to the application in targeted imaging.

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Electronic Supplementary Information (ESI) available: conjugation of NPs with dyes, radiolabelling for NPs, NMR spectra, XRD, IR, zeta potential, DLS size distribution, TEM images and TGA data of NPs, fluorescent images of NPs. See DOI: 10.1039/c000000x/

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