



Full paper/Mémoire

One-pot synthesis of highly substituted pyrroles using nano copper oxide as an effective heterogeneous nanocatalyst

Hamid Saeidian*, Morteza Abdoli, Robab Salimi

Department of Science, Payame Noor University (PNU), PO Box 19395-4697, Tehran, Iran

ARTICLE INFO

Article history:

Received 1 December 2012

Accepted after revision 7 February 2013

Available online 15 March 2013

Keywords:

Nano copper oxide

Highly substituted pyrroles

Multicomponent reactions

Nanocatalyst

DFT calculations

ABSTRACT

A convenient one-pot four-component reaction of aromatic aldehydes, β -keto esters and nitromethane in the presence of an amine and 10 mol% CuO nanoparticles for the synthesis of highly substituted pyrroles is described. Excellent yields, easily available and less expensive catalyst and easy work-up are the key features of the present method. The NMR spectrum was coupled with quantum chemical calculations in DFT approach using the hybrid B3LYP exchange–correlation functional to confirm the structures of (1-benzyl-4-(4-chlorophenyl)-2-methyl-1*H*-pyrrol-3-yl)(phenyl)methanone **14** and 1-(1-benzyl-4-(4-chlorophenyl)-2-phenyl-1*H*-pyrrol-3-yl)ethanone **14'**.

© 2013 Académie des sciences. Published by Elsevier Masson SAS. All rights reserved.

1. Introduction

Multicomponent reactions (MCRs) have emerged as a vital field of chemistry and a powerful tool in modern synthetic organic chemistry, allowing the facile synthesis of complicated compounds in a one-pot reaction without the isolation of any intermediates. Clearly, for multi-step synthetic procedures the number of reaction and purification steps is among the most important criteria for the efficiency and practicability of the process and should be as low as possible. Therefore, it has drawn the attraction of organic chemists to develop novel MCRs for a broad range of products [1].

On the other hand, pyrrole nucleus is a versatile and valuable building block for a number of biologically and pharmaceutically active compounds (Fig. 1). A range of pharmacological properties, such as antibacterial, antitumor, anti-inflammatory, antioxidant, anti-anginal and antifungal activities of this important class of heterocycles are found in the literature [2].

The known used routes for synthesizing these compounds are Hantzsch [3], Knorr [4], and Paal–Knorr [5] reactions. However, most of these methods employ expensive catalysts, multi-steps synthetic operations, long reaction times or harsh reaction conditions. Thus, it is clearly evident that the need for the development of new and flexible methods is required. Recently, literature has highlighted the importance of nanosized materials in several scientific and technological areas, and many research councils have intensified investments in nanotechnology for the coming years [6]. The surface of metal oxides, such as TiO₂, Al₂O₃, ZnO, CuO and MgO, exhibits both Lewis acid and Lewis base character [7]. They are excellent adsorbents for a wide variety of organic compounds and increase the reactivity of the reactants. In any metal oxide, surface atoms make a distinct contribution to its catalyst activity. In powder particles, the number of surface atoms is a large fraction of the total. The high surface area to volume ratio of metal oxide nanoparticles is mainly responsible for their catalytic properties [8]. High yield, selectivity and recyclability have been reported for a variety of nanocatalyst-based organic reactions [9]. Nano copper oxide is certainly one of the most interesting metal oxides, because it has surface properties, which suggest that a very rich organic chemistry may occur there [10]. Although a few number of new pyrrole

* Corresponding author.

E-mail address: h_porkaleh@yahoo.com (H. Saeidian).

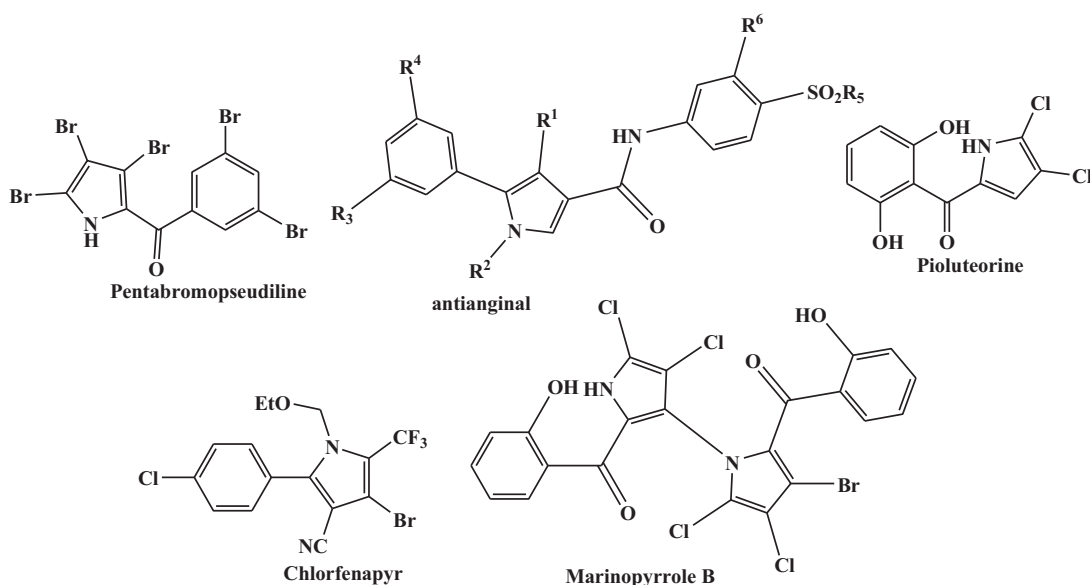


Fig. 1. Some pyrrole-containing bioactive compounds.

syntheses, including MCRs [2c], have been reported in recent years, no example is known for the synthesis of highly substituted pyrroles, using CuO nanoparticles.

In the context of our general interest in the synthesis of heterocycles [11] and following our research on nanoparticles applications in organic synthesis as nanocatalysts [12], herein, we propose a facile synthesis of highly substituted pyrroles using CuO nanoparticles as an efficient, commercially available and non-toxic nanocatalyst via a four-component reaction (Scheme 1). A joint of experimental and DFT calculations was used to confirm the structure of (1-benzyl-4-(4-chlorophenyl)-2-methyl-1H-pyrrol-3-yl)(phenyl)methanone **14** and 1-(1-benzyl-4-(4-chlorophenyl)-2-phenyl-1H-pyrrol-3-yl)ethanone **14'**. It was shown that DFT methods perform NMR spectra calculations very well and give accurate results [13].

2. Experimental

2.1. General information

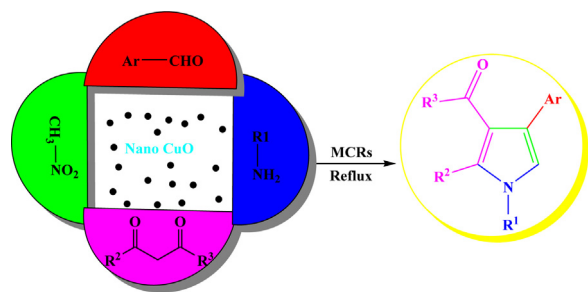
All the chemicals required for the synthesis of highly substituted pyrroles were purchased from Sigma-Aldrich

(St. Louis, MO, USA), Fluka (Neu-Ulm, Germany) and Merck (Darmstadt, Germany) companies and were used as received. Nano copper oxide, with the mean particle size 40 nm and purity 99.9%, was purchased from Nano Pioneers (Mashhad, Iran) company. The synthesized compounds gave all satisfactory spectroscopic data. A Bruker (DRX-500 Avanes) NMR was used to record the ^1H -NMR and ^{13}C -NMR spectra. All NMR spectra were determined in CDCl_3 at ambient temperature.

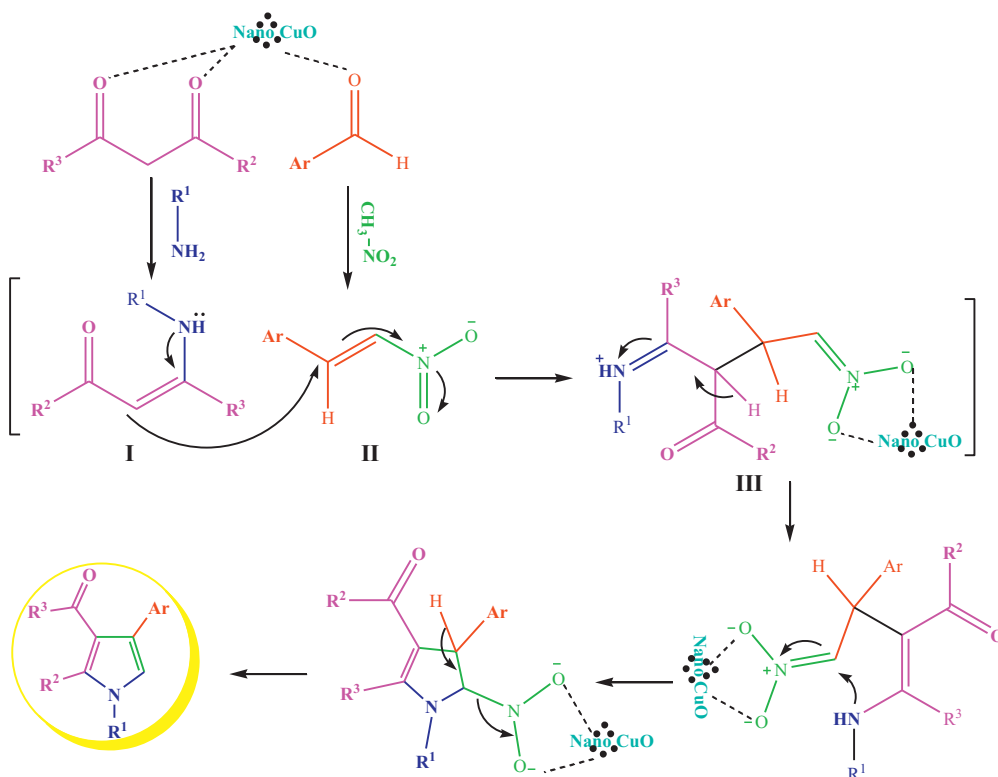
All geometry optimizations and frequency calculations of all species were carried out using the Gaussian 03 program [14]. Density Functional Theory with the Becke three parameters hybrid functional (DFT-B3LYP) calculations were performed with a 6-311++G (2d, 2p) basis set for all atoms. Vibrational frequencies were calculated at the same level to ensure that each stationary point is a real minimum. Harmonic-oscillator approximation was also used for the thermodynamic partition functions. After geometry optimization and frequency calculations, zero-point energies (ZPEs) and thermal corrections are obtained at 298 K. The NMR computations of absolute shieldings were performed using the Gauge-Independent Atomic Orbital (GIAO) method [15].

2.2. General procedure for synthesis of highly substituted pyrroles

To a stirred suspension of CuO nanoparticles (10 mol%) in nitromethane (5 mL) were added an aryl aldehyde (1 mmol), a β -keto ester (1 mmol) and an amine (1.2 mol). The reaction mixture was stirred at 100–105 °C for 12 h. The progress of the reaction was monitored by TLC. After completion of the reaction, the catalyst was recovered by filtration. The nano CuO was reused in subsequent reaction without losing any significant activity. The mixture was concentrated under vacuum. The residue was purified by preparative TLC (eluent: petroleum ether/ethyl acetate 5:1)



Scheme 1. Synthesis route of highly substituted pyrroles via a four-component reaction using nano CuO.



Scheme 2. Suggested reaction mechanism for the formation of desired highly substituted pyrroles.

to afford the desired compound. The structure of the products was confirmed by $^1\text{H-NMR}$, $^{13}\text{C-NMR}$ and comparison with authentic samples prepared by reported methods [16,17].

¹ Representative spectroscopic data: Table 1, Entry 1: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 2.51 (s, 3H), 3.71 (s, 3H), 5.10 (s, 2H), 6.63 (s, 1H), 7.11 (d, $J = 6.8$ Hz, 2H), 7.30–7.43 (m, 8H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.56, 50.57 (2C), 110.76, 120.59, 126.18, 126.21, 126.57, 127.65, 127.84, 128.96, 129.11, 135.80, 136.60, 136.73, 166.33. Table 1, Entry 2: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 2.52 (s, 3H), 3.73 (s, 3H), 5.09 (s, 2H), 6.62 (s, 1H), 7.12 (d, $J = 8.4$ Hz, 2H), 7.32–7.43 (m, 7H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.64, 50.61, 50.63, 110.68, 120.67, 125.09, 126.57, 127.78, 127.92, 128.98, 130.47, 132.02, 134.34, 136.61, 136.90, 166.10. Table 1, Entry 5: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 2.52 (s, 3H), 3.64 (s, 3H), 5.12 (s, 2H), 6.60 (s, 1H), 7.11 (d, $J = 8.4$ Hz, 2H), 7.23–7.40 (m, 7H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.39, 50.57, 50.65, 112.13, 120.57, 120.93, 122.91, 126.03, 126.57, 127.64, 127.82, 128.98, 129.11, 131.61, 134.35, 136.01, 136.76, 166.01. Table 1, Entry 7: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 1.16 (t, $J = 7.2$ Hz, 3H), 2.51 (s, 3H), 4.20 (q, $J = 7.2$ Hz, 2H), 5.10 (s, 2H), 6.63 (s, 1H), 7.11 (d, $J = 6.8$ Hz, 2H), 7.30–7.44 (m, 8H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.51, 14.08, 50.55, 59.41, 111.07, 120.44, 126.14, 126.22, 126.58, 127.53, 127.81, 128.95, 129.30, 135.88, 136.45, 136.78, 165.91. Table 1, Entry 8: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 1.21 (t, $J = 7.2$ Hz, 3H), 2.53 (s, 3H), 4.22 (q, $J = 7.2$ Hz, 2H), 5.10 (s, 2H), 6.63 (s, 1H), 7.12 (d, $J = 7.2$ Hz, 2H), 7.32–7.46 (m, 7H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.61, 14.22, 50.58, 59.52, 110.96, 120.57, 125.08, 127.56, 127.66, 127.90, 128.97, 129.01, 129.34, 130.65, 131.98, 134.44, 165.68. Table 1, Entry 12: $^1\text{H-NMR}$ (400 MHz, CDCl_3): δ 1.18 (t, $J = 7.2$ Hz, 3H), 1.86 (d, $J = 6.8$ Hz, 3H), 2.48 (s, 3H), 4.19 (q, $J = 7.2$ Hz, 2H), 5.40 (q, $J = 6.8$ Hz, 1H), 6.74 (s, 1H), 7.12 (d, $J = 8.4$ Hz, 2H), 7.3–7.35 (m, 7H); $^{13}\text{C-NMR}$ (100 MHz, CDCl_3): δ 11.49, 14.16, 22.18, 55.17, 59.48, 110.83, 116.92, 124.79, 125.83, 127.63, 127.69, 128.92, 130.58, 131.91, 134.69, 136.70, 141.98, 165.81.

3. Results and discussion

To find the optimal conditions, the synthesis of ethyl 1-benzyl-4-(4-chlorophenyl)-2-methyl-1H-pyrrole-3-carboxylate (Table 1, entry 8) was used as a model reaction. A mixture of 4-chlorobenzaldehyde (1 mmol), methyl acetoacetate (1 mmol), benzyl amine (1.3 mol) and nitromethane (5 mL) was stirred under various reaction conditions at 100°C . The reaction was checked without nano CuO as a catalyst, and we did not get any product, while good results were obtained in the presence of CuO nanoparticles after 12 h. On using the optimized amount of catalyst, we found that 10 mol% of CuO nanoparticles could effectively catalyze the reaction for the synthesis of the desired product. Commercially available CuO bulk also was evaluated for the synthesis of the desired title compounds. Clearly, the reaction time, using CuO nanoparticles, has been reduced by three times with higher yield than CuO bulk (81% versus 45%). The optimum reaction conditions used was 10% nano CuO and boiling nitromethane as the reaction medium. The $^1\text{H-NMR}$ spectrum of ethyl 1-benzyl-4-(4-chlorophenyl)-2-methyl-1H-pyrrole-3-carboxylate (Table 1, entry 8) showed two signals (δ 1.21 and 4.22) for the protons of ethoxy group, a singlet (δ 2.52) for the methyl protons of the pyrrole ring, a resonance for the deshielded benzylic proton (δ 5.09), a singlet (δ 6.62) for the aromatic proton of the pyrrole ring. The presence of this signal in $^1\text{H-NMR}$ spectra of synthesized compounds is a good indication of the formation of the desired chemicals. The ^1H -decoupled

Table 1
CuO nanoparticles catalyzed formation of highly substituted pyrroles.

Entry	β -Keto esters	Amines	Aldehydes	Products	Yield (%)
1					75
2					78
3					69
4					65
5					73

Table 1 (Continued)

Entry	β -Keto esters	Amines	Aldehydes	Products	Yield (%)
6					76
7					77
8					81
9					70
10					74

Table 1 (Continued)

Entry	β -Keto esters	Amines	Aldehydes	Products	Yield (%)
11					75
12					68
13					66

^{13}C -NMR spectrum of ethyl 1-benzyl-4-(4-chlorophenyl)-2-methyl-1H-pyrrole-3-carboxylate showed 17 distinct resonances in agreement with the proposed structure.

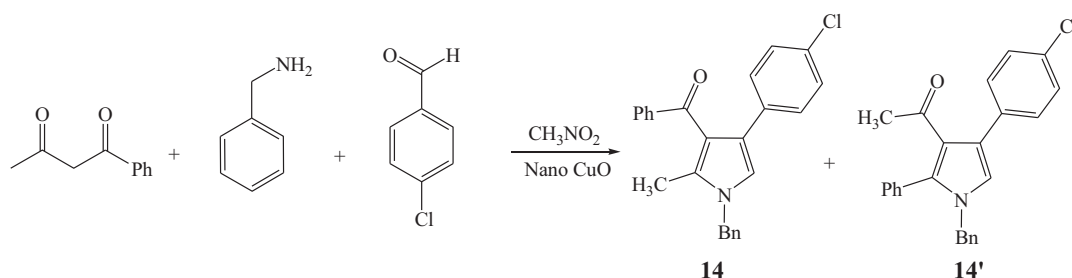
After finding the suitable conditions, we subjected a series of other aromatic aldehydes, primary amines and β -keto ester to obtain the corresponding poly-substituted pyrroles under the optimized reaction conditions. The results are summarized in Table 1. Generally, the reactions were performed well and the reaction rates as well as the yields of products are satisfactory.

The suggested mechanism for this reaction is patterned in Scheme 2. We believe that CuO nanoparticles are coordinated to the oxygen of the aldehyde and activate it for nucleophilic attack [18]. On the other hand, CuO nanoparticles facilitate the enaminezation of the β -keto

esters. Subsequently, the enamine **I** reacts with nitrostyrene **II** to provide the Michael adduct **III**. The intermediate **III** undergoes cyclization with a loss of HNO_2 to yield the desired products. The high surface area to volume ratio of CuO nanoparticles is mainly responsible for their catalytic properties [8].

The applicability of the present methodology is further extended by performing the reaction with 4-chlorobenzaldehyde, benzoylacetone, benzyl amine and nitromethane in the presence of CuO nanoparticles (Scheme 3).

This reaction can lead to two structural isomeric products; hence, their NMR spectra are very similar together. In order to provide an assignment of NMR spectra of the studied compounds, the NMR spectra of isomers were coupled with quantum chemical calculations



Scheme 3. CuO nanoparticles catalyzed reaction of 4-chlorobenzaldehyde with benzoylacetone.

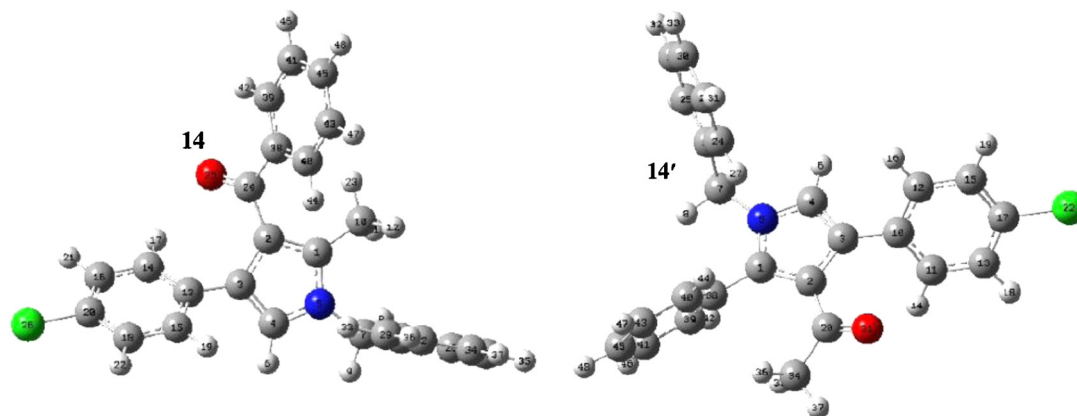


Fig. 2. Numbering optimized structures of **14** (**14'**).

in the DFT approach of hybrid B3LYP exchange-correlation functional to confirm the structures of the synthesized compounds. The optimized structures of **14**(**14'**), are given in Fig. 2 with atom numbers. In order to express the chemical shifts in ppm, the geometry of the TMS molecule had been optimized and then, its NMR spectrum was calculated by using the same method and basis set as in the case of the calculations on the compounds **14**(**14'**).

The calculated isotropic shielding constants were then transformed to chemical shifts relative to TMS. The calculated spectra are compared with the experimental data. Theoretical and experimental chemical shifts of NMR of **14**(**14'**) are given in Table 2.

As it can be seen from Table 2, the NMR shifts obtained by the DFT method at the B3LYP/6-311++G (2d,2p) level are in reasonable agreement with the experimental values of the ^1H -NMR and ^{13}C -NMR spectra.

Table 2
Experimental and calculated ^1H -NMR and ^{13}C -NMR chemical shifts (ppm) for **14**(**14'**).

Atom	14		14'	
	$\delta_{\text{Calcd.}}$	$\delta_{\text{Exp.}}$	$\delta_{\text{Calcd.}}$	$\delta_{\text{Exp.}}$
H_{CH_3}	1.95	2.35	1.84	2.33
H_{Bn}	5.27	5.15	5.07	5.14
$\text{H}_{\text{pyrrole}}$	7.14	6.76	6.73	6.74
C_{CH_3}	11.94	11.36	31.11	11.41
C_{Bn}	54.03	50.60	52.97	50.56
$\text{C}_{\text{C=O}}$	198.81	194.16	198.15	193.95

4. Conclusions

In conclusion, we have reported an efficient procedure for the synthesis of highly substituted pyrroles using CuO nanoparticles as a non-toxic and inexpensive heterogeneous nanocatalyst. The method also offers some other advantages, such as clean reaction, low loading of catalyst and use of various substrates, which make it a useful and attractive strategy for the synthesis of highly substituted pyrroles. A joint of experimental and DFT calculations was also used to confirm the structure of (1-benzyl-4-(4-chlorophenyl)-2-methyl-1H-pyrrol-3-yl)(phenyl)methanone **14** and 1-(1-benzyl-4-(4-chlorophenyl)-2-phenyl-1H-pyrrol-3-yl)ethanone **14'**.

Acknowledgements

We would like to acknowledge the financial support from Payame Noor University (PNU). We also are thankful to Mr. Biglari for NMR spectral support.

References

- [1] (a) A. Dömling, W. Wang, K. Wang, Chem. Rev. 112 (2012) 3135; (b) P.J. Boissarie, Z.E. Hamilton, S. Lang, J.A. Murphy, C. Suckling, Org. Lett. 13 (2011) 6184; (c) V. Rama, K. Kanagaraj, T. Subramanian, P. Suresh, K. Pitchumani, Catal. Commun. 26 (2012) 39; (d) B. Das, N. Bhunia, M. Lingaiah, Synthesis (2011) 3471; (e) M. Adib, E. Sheikhi, H.R. Bijanzadeh, Synlett (2012) 85; (f) M.M. Heravi, L. Ranjbar, F. Derikvand, F.F. Bamoharram, Catal. Commun. 8 (2007) 289.

- [2] (a) H. Fan, J. Peng, M.T. Hamann, J.F. Hu, *Chem. Rev.* 108 (2008) 264;
(b) J. Nuss, M. Williams, R. Mohan, R. Martin, T.L. Wang, K. Aoki, H. Tsuruoka, N. Hayashi, T.U.S. Homma, Patent 20110263673A1, 2011.
(c) V. Estevez, M. Villacampa, J.C. Menendez, *Chem. Soc. Rev.* 39 (2010) 4402;
(d) G.W. Gribble, in: A.R. Katritzky, C.W. Rees, E.F.V. Scriven (Eds.), *Comprehensive Heterocyclic Chemistry II*, 2, Pergamon Press, Oxford, 1996, p. 207;
(e) M. Reisser, G. Maas, *J. Org. Chem.* 69 (2008) 4913.
- [3] T.A. Moss, T. Nowak, *Tetrahedron Lett.* 53 (2012) 3056.
- [4] A. Alberola, A.G. Ortega, M.L. Sádaba, C. Sañudo, *Tetrahedron* 55 (1999) 6555.
- [5] K. Aghapoor, L. Ebadi-Nia, F. Mohsenzadeh, M.M. Morad, Y. Balavar, H.R. Darabi, *J. Organomet. Chem.* 708–709 (2012) 25.
- [6] (a) C. Beaudry, S. Allaoui, *Res. Policy* 41 (2012) 1589;
(b) K. Radad, M. Al-Shraim, R. Moldzio, W.D. Rausch, *Environ. Toxicol. Phar.* 34 (2012) 661.
- [7] K. Tanabe, *Solid Acids and Bases*, Academic Press, New York, 1970.
- [8] A.T. Bell, *Science* 299 (2003) 1688.
- [9] (a) M.Z. Kassae, F. Movahedi, H. Masrouri, *Synlett* (2009) 1326;
(b) K.V.V. Satyanarayana, P.A. Ramaiah, Y.L.N. Murty, M.R. Chandra, S.V.N. Pammi, *Catal. Commun.* 25 (2012) 50;
(c) M. Stratakis, H. Garcia, *Chem. Rev.* 112 (2012) 4469;
(d) V. Polshettiwar, R. Luque, A. Fihri, H. Zhu, M. Bouhrara, J.M. Basset, *Chem. Rev.* 111 (2011) 3036.
- [10] (a) H.R. Shaterian, F. Moradi, M. Mohammadnia, *C.R. Chimie* 15 (2012) 1055;
(b) S.G. Babu, R. Karvembu, *Ind. Eng. Chem. Res.* 50 (2011) 9594;
(c) V.K. Akkilagunta, R.R. Kakulapati, *J. Org. Chem.* 76 (2011) 6819;
(d) A. Shamsian, *Synlett* (2011) 1640.
- [11] (a) F.M. Moghaddam, B. Koushki, H. Saeidian, S. Nourian, Z. Mirjafary, M. Kiamehr, *Pure Appl. Chem.* 83 (2011) 709, and references cited therein;
(b) F.M. Moghaddam, H. Saeidian, Z. Mirjafary, S. Taheri, S. Kheirjou, *Synlett* (2009) 1047, and references cited therein.
- [12] (a) F.M. Matloubi, H. Saeidian, *Mater. Sci. Eng. B* 139 (2007) 265;
(b) Z. Mirjafary, H. Saeidian, A. Sadeghi, F.M. Moghaddam, *Catal. Commun.* 9 (2008) 299;
(c) H. Saeidian, F.M. Moghaddam, A. Pourjavadi, S. Barzegar, R. Soleyman, A. Sohrabi, *J. Braz. Chem. Soc.* 20 (2009) 466.
- [13] (a) B. Brycki, I. Kowalczyk, J. Werner, T. Borowiak, I. Wolska, *J. Mol. Struct.* 791 (2006) 137;
(b) M. Ghiasi, A.A. Oskouie, H. Saeidian, *Carbohydrate Res.* 348 (2012) 47;
(c) C.F. Tormena, G.V. Da Silva, *J. Chem. Phys. Lett.* 398 (2004) 466.
- [14] M.J. Frisch, G.W. Trucks, H.B. Schlegel, G.E. Scuseria, M.A. Robb, J.R. Cheeseman, J.A. Montgomery Jr., T. Vreven, K.N. Kudin, J.C. Burant, J.M. Millam, S.S. Iyengar, J. Tomasi, V. Barone, B. Mennucci, M. Cossi, G. Scalmani, N. Rega, G.A. Petersson, H. Nakatsuji, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, M. Klene, X. Li, J.E. Knox, H.P. Hratchian, J.B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R.E. Stratmann, O. Yazyev, A.J. Austin, R. Cammi, C. Pomelli, J.W. Ochterski, P.Y. Ayala, K. Morokuma, G.A. Voth, P. Salvador, J.J. Dannenberg, V.G. Zakrzewski, S. Dapprich, A.D. Daniels, M.C. Strain, O. Farkas, D.K. Malick, A.D. Rabuck, K. Raghavachari, J.B. Foresman, J.V. Ortiz, Q. Cui, A.G. Baboul, S. Clifford, J. Cioslowski, B.B. Stefanov, G. Liu, A. Liashenko, P. Piskorz, I. Komaromi, R.L. Martin, D.J. Fox, T. Keith, M.A. Al-Laham, C.Y. Peng, A. Nanayakkara, M. Challacombe, P.M.W. Gill, B. Johnson, W. Chen, M.W. Wong, C. Gonzalez, J.A. Pople, *Gaussian 03, Revision C.02*, Gaussian Inc, Wallingford CT, 2004 (*Gaussian 03, Revision E.01*, Gaussian, Inc., Pittsburgh PA, 2003).
- [15] K. Wolinski, J.F. Hilton, P.J. Pulay, *Am. Chem. Soc.* 112 (1990) 8251.
- [16] S. Maiti, S. Biswas, U. Jana, *J. Org. Chem.* 75 (2010) 1674.
- [17] G.R. Reddy, T.R. Reddy, S.C. Joseph, K.S. Reddy, M. Pal, *RSC Adv.* 2 (2012) 3387.
- [18] G. Pandey, R.P. Singh, A. Garg, V.K. Singh, *Tetrahedron Lett.* 46 (2005) 2137.