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ARTICLE TYPE

Fluorocarbene, Fluoroolefin, and Fluorocarbyne Complexes of Rh[†]Christopher J. Pell^{†a}, Yanjun Zhu^{†a}, Rafael Huacuja^{†a}, David E. Herbert^{a‡}, Russell P. Hughes^{b*}, and Oleg V. Ozerov^{a*}

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The manuscript reports the synthesis, characterization, and analysis of electronic structure in a series of complexes of small perfluorocarbon ligands with the (PNP)Rh fragment (where PNP is a diarylamido/bis(phosphine) pincer ligand). Reactions of (PNP)Rh(TBE) as the source of (PNP)Rh with CHF₃ and C₂H₅F produced perfluoroalkylidene complexes (PNP)Rh=CF₂ and (PNP)Rh=C(F)(CF₃). (PNP)Rh=CF₂ could also be obtained via the reaction of (PNP)Rh(TBE) with Me₃SiCF₃/CsF, with an admixture of (PNP)Rh(C₂F₄), where TBE = *tert*-butylethylene. Abstraction of fluoride from these neutral (PNP)RhC_xF_y complexes was successful, although only abstraction from (PNP)Rh=CF₂ allowed unambiguous identification of the Rh product, [(PNP)Rh≡CF]⁺. DFT computational studies allowed comparison of relative energies of (PNP)Rh(C₂F₄) and [(PNP)Rh(C₂F₃)]⁺ isomers as well as comparisons between the electronic structure of the =CF₂, C₂F₄, and ≡CF⁺ complexes and their hydrocarbon analogues.

Introduction

Organofluorine chemistry's major impact on the world of industrial chemistry has inspired many investigations into the unique properties that are inherent to molecules and materials containing C-F bonds. Transition metal complexes containing perfluorocarbon ligands are an important subset of these studies since they exhibit distinctive bonding properties¹ and can mediate perfluoroalkyl-carbon bond forming processes.² Group 9 perfluoroalkylidenes have garnered interest in the past decade after Hughes developed a simple reductive method for making Ir=CFR complexes (Figure 1, top) from iridium-fluoroalkyl precursors.³ These complexes have been analyzed in the context of their potential intermediacy in perfluoroolefin metathesis,⁴ and more recently the Baker group has shown that analogous cobalt perfluorocarbenes (Figure 1, top)⁵ are capable of undergoing a [2+1] cycloaddition with :CF₂⁶ and [2+2] cycloaddition with C₂F₄.⁷ Analogous chemistry was also reported for a difluorocarbene complex of Ni(0).⁸ Baker has also shown that cationic cobalt (III) difluorocarbenes could undergo migratory insertion into perfluoroalkyl ligands, possibly providing a blueprint for transition metal catalyzed perfluoroolefin polymerization.⁹

The only family of isolable terminal fluoromethylidyne complexes known to date are the Cp^{*}M(CO)₂(CF) compounds (M = Cr, Mo, W) reported by Hughes and co-workers (Figure 1, bottom).¹⁰ The Andrews group has reported a number of fluoromethylidyne complexes of the general formula X₃M(CF) (Figure 1, bottom; X = halogen) via trapping laser ablated metal atoms in argon/halocarbon matrices at ca. 10 K.¹¹ Most of the isolable terminal carbyne complexes are complexes of metals of groups 6,^{12,13} 7,^{13,14} and 8.^{15,16} A few examples are known for group 5.¹⁷ In group 9, one 18-electron complex has been fully

characterized for Ir by Bergman et al. (Figure 1, bottom),¹⁸ and one square planar 16-electron complex was mentioned in passing for Rh by Werner et al.,¹⁹ as a component of a reaction mixture. The "concentration" of metal carbyne complexes in the middle of the transition metal series can be compared with similar trends for other metal-element multiple bonds.^{20, 21} In this report, we describe the synthesis, characterization, and analysis of electronic structure of a rare cationic fluoromethylidyne complex of Rh, as well as related Rh perfluoroalkylidene complexes.

Hughes' Reductive Route to Perfluoroalkylidenes

Baker's Co Analogue

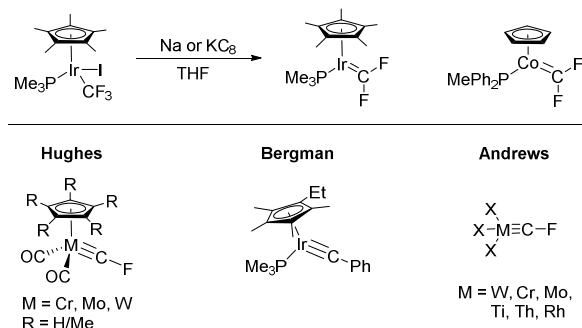


Figure 1. Perfluoroalkylidenes from Hughes and Baker. Isolated fluorocarbynes by Hughes, the Ir carbyne by Bergman, and matrix-trapped fluorocarbynes by Andrews.

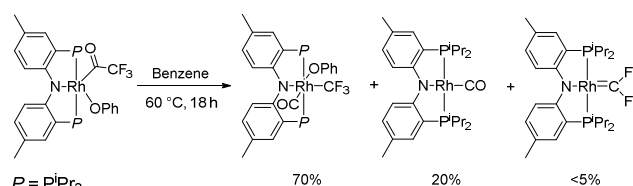
Results and Discussion

Synthesis of CF₂, C₂F₄, and CFCF₃ complexes.

We recently reported reactions of the (PNP)Rh fragment with aryl carboxylates, including aryl-oxygen oxidative addition.²² The (PNP)Rh acyl-oxygen oxidative addition product of phenyl trifluoroacetate, (PNP)Rh(C=OCF₃)(OPh), could be thermalized



to produce (PNP)Rh(CO) and (PNP)Rh(CF₃)(CO)(OPh) as major products. In that report, we noted that some other unidentified products were evident in trace amounts. We continued to be intrigued by one apparent trace product in particular that was consistently observed in 2-5% yield. For it, we observed a doublet of triplets both in the ³¹P{¹H} NMR and ¹⁹F NMR spectra (coupling constants: ¹J_{Rh-P} = 146 Hz, ²J_{Rh-F} = 49 Hz, ³J_{P-F} = 30 Hz). These multiplicities implied a P₂RhF₂ NMR spin system – rather unexpected given the *three* fluorines in the CF₃ group of the starting material. We noted that the ¹⁹F NMR chemical shift was itself uncommon (95.6 ppm) and in the range reported for various difluorocarbene complexes (i.e., M=CF₂).^{5,10a, 23,24} The observed ²J_{Rh-F} = 49 Hz was also similar to that of Grushin's *trans*-(Ph₃P)₂(F)Rh=CF₂, which possessed a ²J_{Rh-F} of 33 Hz.²³



Scheme 1. Initial observation of (PNP)Rh=CF₂.

We hypothesized that this minor side product might be (PNP)Rh=CF₂ and attempted an independent synthesis of it based on the procedure of Grushin et al. that yielded *trans*-(F)(PPh₃)₂Rh=CF₂.²³ Indeed, treatment of (PNP)Rh(TBE) (TBE = *tert*-butylethylene) with CsF/Me₃SiCF₃ (Rupert's reagent) resulted in complete consumption of (PNP)Rh(TBE) and the formation of (PNP)Rh=CF₂ and (PNP)Rh(C₂F₄) in about 85:15 ratio (NMR evidence).²⁵ We were able to isolate (PNP)Rh=CF₂ in 52% yield and of >98% purity by recrystallization. The ³¹P{¹H} NMR and ¹⁹F NMR spectra of (PNP)Rh=CF₂ so obtained were identical to that of the impurity we observed in the reaction in Scheme 1.

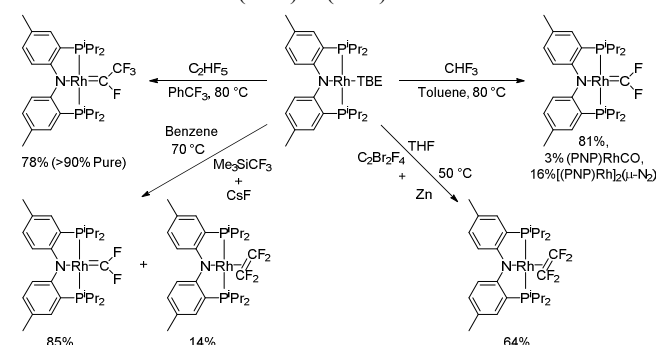
(PNP)Rh(TBE) showed no reaction with Me₃SiCF₃ alone. Similar to the other cases of use of CsF/Me₃SiCF₃, we propose that these reagents generate a CF₃ anion equivalent that displaces TBE and then loses fluoride, resulting in the formal transfer of CF₂ to Rh. Alternatively, CsF/Me₃SiCF₃ could be generating free CF₂ which then binds to Rh. Using the CsF/Me₃SiCF₃ protocol, we could not avoid the formation of (PNP)Rh(C₂F₄) due to the generation of free C₂F₄ from the CsF/Me₃SiCF₃ mixture. C₂F₄ was observed by ¹⁹F NMR spectroscopy in control experiments where CsF and Me₃SiCF₃ were mixed in C₆D₆ and heated at 80 °C. No reaction was observed when (PNP)Rh=CF₂ was treated with another equivalent of CsF/Me₃SiCF₃. This contrasts the reactivity of Baker's difluorocarbene cobalt (I) complexes⁶ which undergo a [2+1] cycloaddition with free :CF₂ to form cobalt tetrafluoroethylene complexes.

To date, we have not been able to formulate a reasonable proposal for how (PNP)Rh=CF₂ could be formed from (PNP)Rh(CO(CF₃)(OPh) (Scheme 1). The formation of M=CF₂ by fluoride migration from M-CF₃ is well precedented and is likely the key step in forming (PNP)Rh=CF₂; the difficulty is with conceiving of a plausible fate of the other atoms of the original phenyl trifluoroacetate molecule.

Goldman et al. documented formation of (PCP)Ir=CF₂ in a reaction of a (PCP)Ir source with HCF₃.²⁴ This reaction

proceeded via C-H oxidative addition of HCF₃ to Ir followed by loss of HF. In a similar vein, we found that (PNP)Rh(TBE) reacted with HCF₃ at 80 °C to provide a mixture of compounds containing (PNP)Rh=CF₂ as a major product (>80%) with (PNP)Rh(CO) and [(PNP)Rh]₂(μ-N₂) as minor products. Commercial HCF₃ contains dinitrogen as an impurity. Hydrolysis of a difluorocarbene complex to a carbonyl complex has precedent,²⁶ but attempts to purposefully hydrolyze (PNP)Rh=CF₂ proved to be unsuccessful, reminiscent of Baker's cobalt fluorocarbene complexes.⁵ It is possible that hydrolysis of (PNP)Rh=CF₂ only takes place in the presence of HF (a by-product of (PNP)Rh=CF₂ generation). We observed no intermediates²⁷ in the reaction of (PNP)Rh(TBE) with HCF₃, which may indicate that dissociation of TBE²⁸ is the rate-limiting step.

An analogous reaction of (PNP)Rh(TBE) with C₂HF₅ was attempted as a potential means to access (PNP)Rh(C₂F₄). However, the major product of this reaction turned out to be a tetrafluoroethylidene complex (PNP)Rh=C(F)(CF₃). Dinitrogen impurity in C₂HF₅ led to the known²⁹ [(PNP)Rh]₂(μ-N₂) as a major side product, whose content could be reduced by degassing C₂HF₅ using the "freeze-pump-thaw" technique. (PNP)Rh=CF₂ was also observed as a side product composing 13% of the reaction mixture when (PNP)Rh(TBE) was treated with 2 atm of



Scheme 2. Synthesis of rhodium fluorocarbenes and tetrafluoroethylene complexes.

C₂HF₅ and heated overnight at 80 °C. (PNP)Rh=C(F)(CF₃) could be isolated in >90% purity with (PNP)Rh(CO) composing the rest of the mixture. The synthesis of (PNP)Rh(C₂F₄) was instead accomplished by thermolysis of (PNP)Rh(TBE) in a solution containing C₂F₄ which was generated in situ by reducing C₂F₄Br₂ with 1.5 eq. of Zn powder at 50 °C in THF. (PNP)Rh(C₂F₄) was isolated in 64% yield as a pure solid.

The presence of multiple NMR-active nuclei provided for information-rich NMR spectra of (PNP)Rh=CF₂, (PNP)Rh=C(F)(CF₃), and (PNP)Rh(C₂F₄). All three complexes displayed C_{2v}-symmetric NMR spectra in solution at ambient temperature. The carbene complexes (PNP)Rh=CF₂ and (PNP)Rh=C(F)(CF₃) displayed characteristic ¹³C NMR resonances at 206.3 and 225.0 ppm. In the CF₂ complex the observation that the two fluorines couple identically to both phosphorus nuclei, and vice versa, is consistent with rapid rotation about the Rh=CF₂ bond on the NMR timescale at room temperature. Likewise the observation of identical coupling of both P-nuclei to all four fluorines in the C₂F₄ complex is consistent with rapid rotation about the Rh-alkene bond axis.



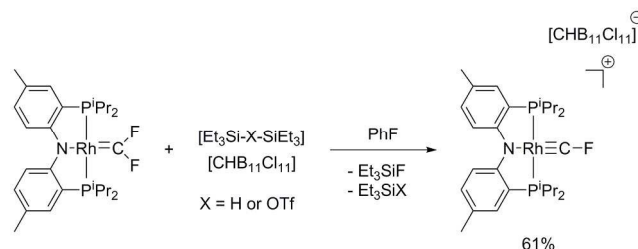
Small energy barriers to these rotations are calculated by DFT (see below).

The identity of (PNP)Rh=CF₂ and (PNP)Rh(C₂F₄) was confirmed by X-ray diffraction studies on suitable single crystals.

Treating the CF₂ or C₂F₄ ligands as occupying a sole coordination site, the coordination environment about Rh is approximately square planar in both molecules. The CF₂ unit in (PNP)Rh=CF₂ lies approximately in that plane, while the C-C vector of the C₂F₄ ligand in (PNP)Rh(C₂F₄) is approximately perpendicular to it. The CF₂ and C₂F₄ ligands evidently exert similar *trans*-influence as the Rh-N distances in (PNP)Rh=CF₂ and (PNP)Rh(C₂F₄) are only different by ca. 0.01 Å. In general, the metrics of the Rh=CF₂ unit in (PNP)Rh=CF₂ are very similar to the Rh=CF₂ unit in *trans*-(Ph₃P)₂(F)Rh=CF₂. The structures of (PNP)Rh=CF₂ and (PNP)Rh(C₂F₄) contain some close C-F...H contacts (C-F...H distances of 2.33-2.45) (F...C distances of 3.1-3.3 Å). While they are probably unavoidable in these molecules, these distances are short enough to be considered weak F...H interactions.³⁰ C-F...H interactions have been observed in pincer-ligated zirconium complexes bearing a trifluoromethyl as a pendant group, which have also exhibited through-space H-F coupling visible in their ¹H NMR spectra.³¹ However (PNP)Rh(C₂F₄) and (PNP)Rh=CF₂ showed no through-space ¹⁹F-¹H coupling to the isopropyl arms.

25 Synthesis of cationic fluoromethylidyne.

With compounds (PNP)Rh=CF₂, (PNP)Rh(=CFCF₃), and (PNP)Rh(C₂F₄) in hand, we contemplated whether one of the fluorides could be removed to yield cationic C_xF_y complexes. Hughes et al. previously demonstrated proton-induced loss of fluoride from α-positions of Ir perfluoroalkyls,^{3a} and Baker recently demonstrated a Lewis-acid abstraction of a fluoride from N-heterocyclic fluoroalkenes to yield polyfluoroalkenyl imidazolium salts.³² There is significant precedent for electrophilic abstraction of an anionic heteroatom substituent from late-metal carbene complexes by a Lewis acid.^{18, 33} Trialkylsilylium cations, in the form of their salts with halogenated carborane anions, are powerful Lewis acids with high affinity for fluoride.³⁴ We and others have exploited them in catalytic C-F activation reactions³⁵ and thus a [R₃Si]⁺ reagent appeared perfect for fluoride abstraction.



Scheme 3. Synthesis of [(PNP)Rh ≡ CF][CHB₁₁Cl₁₁] via fluoride abstraction from (PNP)Rh=CF₂.

Reactions of (PNP)Rh=CF₂, (PNP)Rh(C₂F₄), and (PNP)Rh=C(F)(CF₃) with [Et₃Si-H-SiEt₃][HCB₁₁Cl₁₁]³⁶ or [(Et₃Si)₂OTf][HCB₁₁Cl₁₁]³⁷ all generated the Et₃SiF by-product, indicating that fluoride abstraction took place in all three cases. However, reactions of (PNP)Rh(C₂F₄) and (PNP)Rh=C(F)(CF₃) resulted in mixtures of several products as seen by ¹⁹F NMR

spectroscopy and typically broad or no signals were observed by ³¹P{¹H} NMR spectroscopy. The reaction mixtures produced from the reaction of (PNP)Rh(C₂F₄) or of (PNP)Rh=C(F)(CF₃) with [(Et₃Si)₂OTf][HCB₁₁Cl₁₁] did regenerate the corresponding starting material when treated with CsF. This indicates that fluoride abstraction from these two isomeric complexes generates isomers of [(PNP)Rh(C₂F₃)]⁺ that do not interconvert on the experimental time scale. Although we were not able to identify these compounds experimentally, DFT computational studies were used to investigate possible structures of the [(PNP)Rh(C₂F₃)]⁺ isomers (*vide infra*).

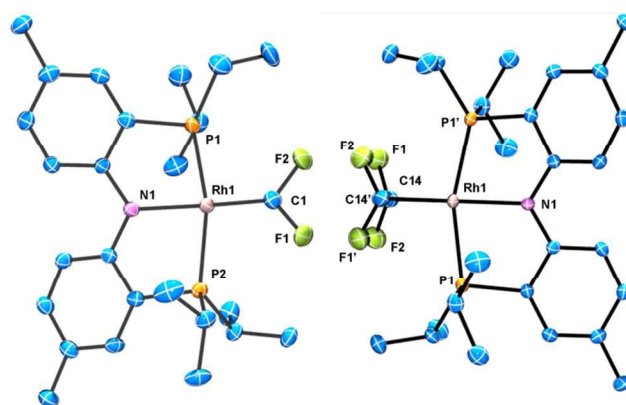


Figure 2. ORTEPs of (PNP)Rh=CF₂ (left) and (PNP)Rh(C₂F₄) (right). The ellipsoids are set at the 50% probability level, and hydrogen atoms are omitted for clarity. Selected bond distances (Å) and angles (°) for (PNP)Rh=CF₂: Rh1-C1, 1.821(4); Rh1-N1, 2.043(3); C1-F1, 1.335(4); C1-F2, 1.348(5); N1-Rh-C1, 171.39(15); F2-C1-F1, 100.8(3); Rh1-C1-F2, 130.1(3); Rh1-C1-F1, 128.6. (PNP)Rh(C₂F₄): Rh1-C14, 2.006(3); Rh1-N1, 2.054(3); C14-F1, 1.378(3); C14-F2, 1.361(3); C14-C14', 1.354(7); C14-Rh-C14', 39.4(2); C14-Rh-N1, 160.28(10).

On the other hand, reaction of (PNP)Rh=CF₂ with [Et₃Si-H-SiEt₃][HCB₁₁Cl₁₁] cleanly and reproducibly generated a new Rh complex that displayed a P₂RhF NMR spin system. The key NMR spectroscopic features of this compound were the unusual ¹⁹F NMR chemical shift (66.2 ppm), the very high ¹J_{C-F} coupling constant of 470 Hz,³⁸ and the rather substantial ²J_{Rh-F} = 136 Hz.

These spectroscopic data are similar to those exhibited by Cp*(CO)₂Mo≡CF, whose ¹⁹F NMR spectrum contained a singlet at 78.15 ppm, with a large ¹J_{C-F} coupling constant of 556 Hz evident by ¹³C NMR spectroscopy.^{10a} Hughes's other Cp(CO)₂M≡CF (M = Cr, W) complexes also exhibited ¹⁹F NMR chemical shifts in this region with high J_{C-F} coupling constants.^{10b} Due to limited solubility in non-interactive solvents and the extensive coupling inherent to the fluoromethylidyne ¹³C NMR resonance in [(PNP)Rh≡CF]⁺, it was not observed by ¹³C{³¹P}, ¹³C{¹H}, nor ¹³C{¹⁹F} NMR spectroscopy.

X-ray quality crystals of [(PNP)Rh≡CF][CHB₁₁Cl₁₁] were studied using X-ray diffraction to yield a structure fully supportive of a fluorocarbene formulation (Figure 3). The structural and NMR spectroscopic features of [(PNP)Rh≡CF]⁺ are best reviewed in comparison with Cl₃Rh≡CF and a few other relevant compounds. Andrews et al. observed IR spectroscopic evidence for Cl₃Rh≡CF in reactions of laser-ablated rhodium atoms with CFCl₃. A DFT calculation of this product predicted a

Rh-C bond length of 1.740 Å and a Rh-C-F bond angle of 143.4°. ^{11d} This compares with our observed Rh-C bond length of 1.702(7) Å and a Rh-C-F bond angle of 173.46°. Although both [(PNP)Rh≡CF]⁺ and Cl₃Rh≡CF are four-coordinate, they contain 5 different numbers of valence electrons: from a hypothetical point of view of a [CF]⁺ ligand, it is attached to a d⁸ Rh center in [(PNP)Rh≡CF]⁺, but to a d⁷ [Cl₃Rh]⁺ fragment in Cl₃Rh≡CF. The geometry of the RhCF unit in [(PNP)Rh≡CF]⁺ is similar to 35 Bergman's iridium carbyne complex (Scheme 1), which possesses an Ir-C bond length of 1.734(6) Å and an Ir-C-C bond angle of 175.7(4). ¹⁸ The Rh-C distance in [(PNP)Rh≡CF]⁺ is ca. 0.12 Å shorter than that in (PNP)Rh=CF₂, consistent with the increase in the Rh-C bond order. The Rh-C bond distance in [(PNP)Rh≡CF]⁺ is also ca. 0.07 Å shorter than that in Werner's 15 *trans*-(PⁱPr₃)₂ClRh=C=C(Me)(H) square planar vinylidene complex. ³⁹

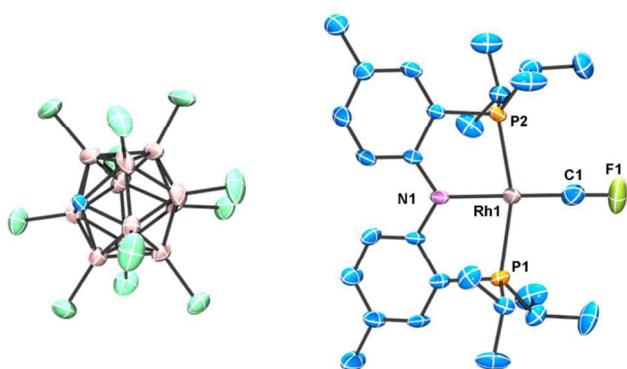


Figure 3. ORTEP of [(PNP)Rh≡CF][HCB₁₁Cl₁₁]. The ellipsoids are set at the 50% probability level, and hydrogen atoms are omitted for clarity. Selected bond distances (Å) and angles (°): Rh1-C1, 1.702(7); Rh1-N1, 2.019(4); C1-F1, 1.257(8); C1-Rh1-N1, 174.1(3); F1-C1-Rh1, 173.4(7).

Computational Studies and Discussion

DFT Structural Studies

Modern Density Functional Theory (DFT) is a powerful tool 20 with which to examine electronic structures and bonding trends in organometallic compounds. ^{40, 41} In addition, application of Natural Bond Orbital (NBO) ⁴²⁻⁴⁷ methods allows insight into some of the subtleties of metal-ligand bonding. ^{42,44,46} The NBO analysis also generates Wiberg Bond Indices (WBI), ⁴⁸ determined 25 within the natural atomic orbital basis, providing one means of estimating bond orders between atoms. Trends in WBI values are also useful in tracking variations in bond multiplicities. The

newly synthesized family of fluorocarbon complexes (PNP)Rh(C₂F₄), (PNP)Rh=CF₂, and [(PNP)Rh≡CF]⁺ prompted a 30 computational comparison with their (hypothetical) hydrocarbon analogues (PNP)Rh(C₂H₄), (PNP)Rh=CH₂ and [(PNP)Rh≡CH]⁺ in order to assess the effects of fluorination on the metal-carbon bonding, and, for the carbene and carbyne complexes, to probe the nature and extent of the multiple bonding between the metal and carbon. In addition, since the CF⁺ ligand is isoelectronic with the well-known ligands NO⁺ and CO, it was of interest to establish trends in metal ligand and intraligand bonding between [(PNP)Rh≡CF]⁺, [(PNP)Rh(NO)]⁺, and (PNP)Rh(CO). Full molecule DFT studies were performed using the M06 35 functional ^{49,50} and the triple-ζ LACV3P**++ basis set, which uses extended core potentials ⁵¹⁻⁵⁴ on heavy atoms and a 6-311G**++ basis ⁵⁵⁻⁵⁸ for other atoms, as implemented in the Jaguar ^{59, 60} suite of programs. Full details are available as Supplementary Information.

45 Selected bond lengths and computed WBI values for the calculated complexes are provided in Table 1, with metric comparisons to the available crystallographic structures reported here. The DFT calculated metrics are in good agreement with crystallographic numbers, giving confidence in the DFT metrics 50 for the unknown complexes. One exception appears in (PNP)Rh(C₂F₄) in which the C-C distance for the coordinated alkene (1.354 Å) is only slightly longer than that in C₂F₄ itself (1.318 Å) ⁶¹ and is much shorter than all other transition metal complexes of this perfluoroalkene in the Cambridge Structure 55 Database. ⁶² In contrast, the DFT calculated value (1.416 Å) is in good agreement with other crystallographically determined values ⁶² and is much more sensible with respect to calculated WBI values (see below).

All the complexes examined here can be formally viewed as 60 square planar d⁸ compounds, i.e., as complexes of a d⁸, three-coordinate fragment (PNP)Rh with neutral or cationic ligands. The NBO perspective of bonding interactions in such compounds ^{42,46} requires 4-electron/3-center bonds between the pair of *trans* ligands such that the alkene, carbene, and carbyne 65 ligands of interest are always involved in a shared bonding interaction with the N of the PNP pincer. Clearly contributions to this shared interaction may be weighted differently in each case, and the WBI values should reflect this.

Comparison of C₂H₄ and C₂F₄ ligands.

70 The bonding between alkenes and transition metal fragments is well understood, ⁶³ but a comparison between C₂H₄ and C₂F₄ coordinated to identical metal-ligand fragments is rare. A classic intramolecular example involves CpRh(C₂H₄)(C₂F₄) ⁶⁴ but



Table 1. Calculated and Crystallographic Bond lengths (Å)^a and Wiberg Bond Indices^b (WBI).

Compound	Rh-P _{ave}	Rh-N	Rh-C _{ave}	C-C	C-X _{ave} (ligand)
(PNP)Rh(C ₂ H ₄)	2.332 <i>0.462</i>	2.098 <i>0.378</i>	2.165 <i>0.453</i>	1.394 <i>1.474</i>	1.088 <i>0.930</i>
(PNP)Rh(C ₂ F ₄)	2.3309(11) <i>0.436</i>	2.054(3) <i>0.352</i>	2.006(3) <i>0.594</i>	1.354(7)^c <i>1.186</i>	1.369(3) <i>0.858</i>
(PNP)Rh=CH ₂	2.338 <i>0.476</i>	2.216 <i>0.225</i>	1.850 <i>1.250</i>	-	1.104 <i>0.955</i>
(PNP)Rh=CF ₂	2.331 2.302(12) <i>0.462</i>	2.156 2.043(3) <i>0.297</i>	1.864 1.821(4) <i>1.168</i>	-	1.325 1.341(5) <i>0.911</i>
[(PNP)Rh≡CH] ⁺	2.378 <i>0.440</i>	2.075 <i>0.332</i>	1.728 <i>1.714</i>	-	1.110 <i>0.942</i>
[(PNP)Rh≡CF] ⁺	2.384 2.337(16) <i>0.432</i>	2.061 2.019(4) <i>0.341</i>	1.740 1.702(7) <i>1.587</i>	-	1.247 1.257(8) <i>1.065</i>
[(PNP)Rh≡C-CF ₃] ⁺	2.384 <i>0.453</i>	2.054 <i>0.330</i>	1.734 <i>1.770</i>	-	1.499 <i>0.960</i>
(PNP)Rh(CO)	2.336 <i>0.462</i>	2.117 <i>0.291</i>	1.855 <i>1.041</i>	-	1.152 <i>2.031</i>
[(PNP)Rh(NO)] ⁺	2.393 <i>0.439</i>	2.019 <i>0.445</i>	1.778 (Rh-N) <i>1.151</i>	-	1.141 (NO) <i>1.917</i>

^a DFT calculated (M06/LACV3P**++) values are in plain text; X ray crystallographic values are in bold. ^b WBI values are in italics. ^c This crystallographic value is questionable. See discussion in the text.

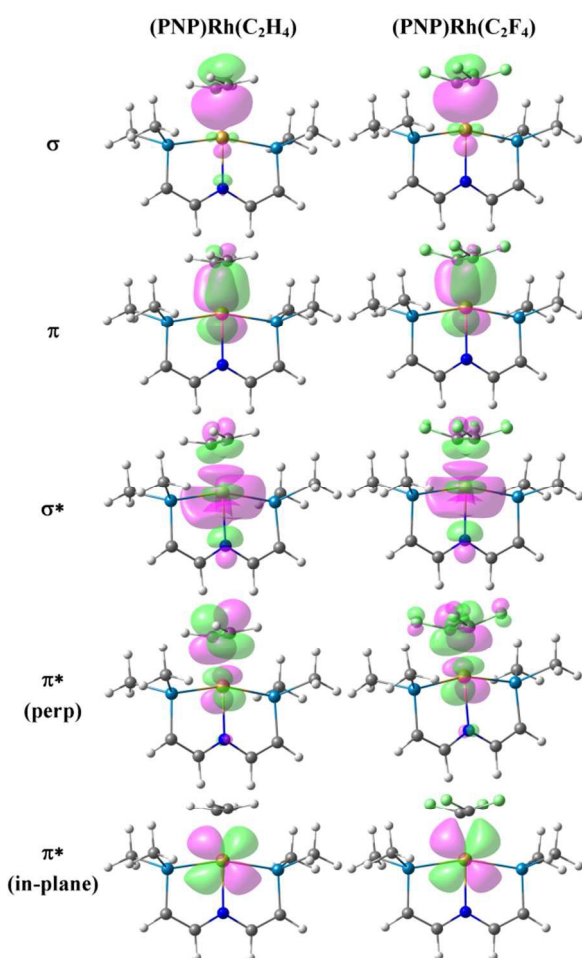


Figure 4. NLMOs for the bonding interactions between C₂H₄ (column 1) and C₂F₄ (column 2) and the (PNP)Rh fragment. For clarity the P'Pr₂ groups have been replaced by PMe₂ groups and the aryl part of the pincer truncated to P-CH=CH-N linkers.

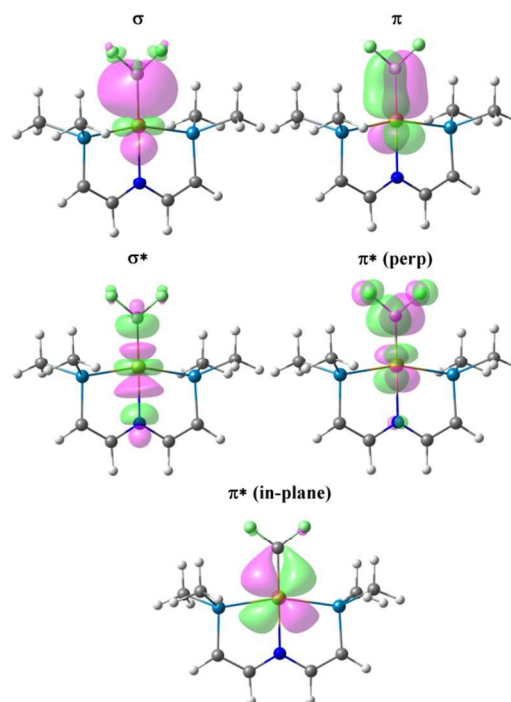


Figure 5. NLMOs for the bonding interactions between the CF₂ ligand and the (PNP)Rh fragment. For clarity the P'Pr₂ groups have been replaced by PMe₂ groups and the aryl part of the pincer truncated to P-CH=CH-N linkers.

the hydrocarbon and fluorocarbon alkenes are necessarily bound to different fragments in this molecule. Figure 4 illustrates the key Natural Localized Molecular Orbitals (NLMOs)^{42,45,46} arising from NBO calculations of interactions between C₂H₄ and C₂F₄ and the truncated⁶⁵ (PNP)Rh fragment. The bonding orbitals (σ and π) are essentially localized on Rh and the alkene ligand, while the corresponding antibonding NLMOs show significant



“tailing” involving the σ and π orbitals on the *trans*-N of the pincer ligand. This “tailing” is indicative of delocalization of these N electrons into the corresponding σ^* and π^* components of the Rh-alkene interaction; it is significantly greater for the σ^* component and corresponds to the 3-center/4-electron bonding expected between *trans*-ligands in a d^8 Rh(I) complex. The WBI values indicate significantly greater reduction in C-C bonding and increase in Rh-C bonding in coordinated C_2F_4 than in C_2H_4 . This is consistent with the shorter Rh-C distances and with the idea of a more metallacyclopropane structure and stronger Rh-C bonding for the fluorinated alkene complex. Not surprisingly, stronger bonding to the fluorinated alkene results in weaker bonding to the *trans* ligand, with correspondingly lower WBI values for the Rh-N bonds (Table 1). In the π^* -perp NMO for the C_2F_4 there is also evidence for delocalization from F-lone pairs (see figure 4).

There are two Rh \rightarrow alkene backbonding options involving π^* -perp or π^* -in-plane. Clearly the ground state conformation of the C_2F_4 complex utilizes the former, but the latter is available for an in-plane C_2F_4 conformation, leading to a low barrier for C_2F_4 rotation. Similar arguments for facile rotation of C_2F_4 ligands in Ru(II) complexes have been put forth elsewhere.⁶⁶ The free energy profile for C_2F_4 rotation was calculated using a truncated version of the PNP ligand (identical to that shown in Figure 4), and is unusual. Relative to the perpendicular conformation observed in the ground state there two transition states were located. The first, lying 7.1 kcal/mol above the ground state

corresponds to a 45 degree rotation about the Rh-alkene bond axis, and the second, lying 8.7 kcal/mol above the ground state is the conformation in which the fluoroalkene lies in the coordination plane. These barriers are low enough in energy that rotation should be fast on the NMR timescale, consistent with the observed NMR data. The barriers contrast with those for the corresponding C_2H_4 analogue, for which the in-plane conformation is a minimum, lying only 0.4 kcal/mol above the ground state, and the 45 degree conformation is a transition state lying 2.5 kcal/mol above the ground state.

Comparison of CH_2 and CF_2 ligands.

Figure 5 presents the corresponding NLMOs for the CF_2 complex with a truncated⁶⁵ PNP ligand. Those for the CH_2 analogue are very similar, except for the fluorine delocalizations into the π^* -NMO, and are not illustrated here but can be found in the ESI (Figure S17). The σ and π NLMOs are consistent with a formal double bond between Rh and the CF_2 (or CH_2) ligands, with the p-orbital on C and d-orbital on Rh providing the π -component. These NLMOs look essentially identical to those in Figure 4, except for a more significant delocalization of the fluorine lone pairs in the π^* -perp NMO. But now there is competition between the rhodium d-orbital and the fluorine lone pair p-orbitals for π -bonding with the carbene carbon, as expected; in (PNP)Rh= CH_2 only the metal can provide this π -bonding. Consequently π -bonding with fluorines diminishes π -bonding with Rh and, relative to the CH_2 complex, the Rh-C WBI decreases significantly and the Rh-C distance increases; notably the C-F WBI is greater than in the previously discussed C_2F_4 complex. In contrast to the alkene ligands discussed above there is overall weaker metal-carbon bonding to CF_2 than CH_2 and the corresponding *trans* Rh-N WBI value is larger for (PNP)Rh= CF_2 .

Facile CF_2 ligand rotation is expected due to the availability of the π^* -perp and π^* -in-plane interactions. Calculations on the truncated⁶⁵ ligand analogue of the CF_2 complex reveal the same trend in conformational energetics observed for the C_2F_4 complex (*vide supra*). The in-plane CF_2 conformation, with the π^* -perp interaction, is the ground state, with two transition states at 45 degree and 90 degree (perpendicular to the coordination plane) lying 2.0 and 4.8 kcal/mol higher in energy, respectively. The barriers are consistent with the experimental observation of fast rotation on the NMR timescale. In contrast, the rotation of the corresponding CH_2 ligand is even more facile, with the perpendicular, 45 degree and in-plane conformations lying at essentially equal energies.

Comparison of CH^+ , CF^+ and CCF_3^+ ligands.

Figure 6 presents the corresponding NLMOs for [(PNP)Rh= CF^+]. As with the carbene complexes, those for [(PNP)Rh= CH^+] are similar except for enhanced “tailing” in the antibonding NLMOs for the CF^+ complex. In contrast to the alkene and carbene complexes (*vide supra*) there is now a second fully engaged π -component for the Rh-C bond involving the in-plane d-orbital and a second p-orbital on the CF^+ (or CH^+) ligand. In the antibonding NLMOs we see the expected σ -donation from the *trans*-N in σ^* , a small π -donation from *trans*-N in π^* (perp), and a small donation from the Rh-P bonds in π^* (in-plane). But once again the largest delocalizations in the π^* NLMOs comes from the F lone pairs, interactions that cannot occur in

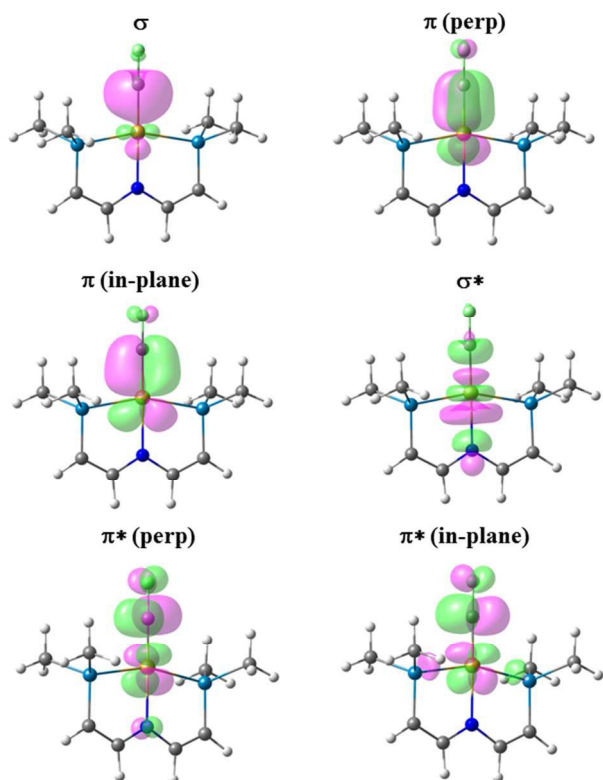


Figure 6. NLMOs for the bonding interactions between the CF^+ ligand and the (PNP)Rh fragment. For clarity the $PiPr_2$ groups have been replaced by PMe_2 groups and the aryl part of the pincer truncated to $P-CH=CH-N$ linkers.



[(PNP)Rh≡CH]⁺. Consequently the Rh-C WBI for [(PNP)Rh≡CH]⁺ is substantially larger than that for the CF⁺ analogue, consistent with the shorter Rh-C distance in the former; as before, a smaller Rh-C WBI in [(PNP)Rh≡CF]⁺ leads to a larger WBI for the trans Rh-N bond. In the CCF₃ analogue, in which F lone pair participation with the carbyne carbon is removed, a larger Rh-C WBI is calculated, with a correspondingly smaller Rh-N WBI.

Degree of Rh-C multiple bonding.

The Rh-C bonding in these carbene and carbyne complexes is inextricably linked to interactions with the *trans* N since there are necessarily shared bonding components between *trans* ligands, with additional complexities introduced by the fluorine substituents on carbon. So we cannot expect the Rh=CX₂ interaction to be a true double bond, or that in the cationic Rh≡CX to be a triple bond, even though we may draw resonance structures that reflect these prejudices. However, the WBI values for both Rh=CX₂ bonds (CH₂ 1.250; CF₂ 1.168) are significantly larger than unity, though not close to two, while those for the Rh≡CX cations are significantly larger still (CH 1.714; CF 1.587; CCF₃ 1.770), though not close to the bond order of three. Clearly there is significant multiple bonding between Rh and these unsaturated carbon ligands with higher bond orders to these ligands being reflected in lower bond orders to the *trans* N.

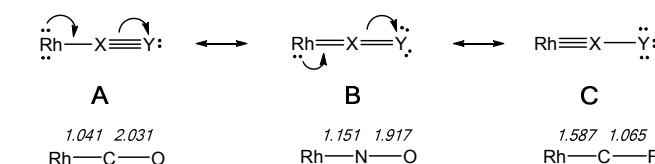


Figure 7. Resonance forms for the π -system in a linear Rh-X-Y ligand array, with WBI values for the bonds in Rh-C-O, Rh-N-O, and Rh-C-F complexes. All three complexes are isoelectronic and no formal charges are shown.

Comparison of CF⁺, NO⁺, and CO ligands.

The NLMOs for the CO and NO⁺ complexes are similar to those of the CF⁺ compounds discussed previously and are not shown here. Considering this series of isoelectronic complexes as involving a linear Rh-X-Y array the three resonance forms (A, B, C) for the contiguous π -system are shown in Figure 7, along with the WBI values for the appropriate bonds in the Rh-C-O, Rh-N-O and Rh-C-F complexes. The WBI values are consistent with progressively increased contributions of resonance forms C>B>A on changing the ligand from CO to NO⁺ to CF⁺, as expected from their relative π -acceptor abilities. Similar conclusions were reached for the fragments M(CO)₂(XY) [M = Cr, Mo, W; XY = CO, NO⁺, CF⁺] in a previous study.^{10b}

Relative energies of isomeric fluorocarbon ligands.

It was of interest to compare the relative energies of the (PNP)Rh(C₂F₄) complex with its carbene isomer (PNP)Rh(CFCF₃). At the DFT/M06/LACV3P**++ level the free energy of the carbene isomer is found to be 2.4 kcal/mol uphill from its alkene analogue. Interestingly the carbene CFCF₃ ligand

lies perpendicular to the (PNP)Rh plane, in contrast to the CF₂ analogue described above, presumably due to steric interactions between the CF₃ and the *cis* PR₂ groups.

Potential products arising from fluoride abstraction from these isomeric complexes were also subjected to DFT evaluation. Abstraction of fluoride from (PNP)Rh(CFCF₃) could occur from the α -position to yield a carbyne complex (PNP)Rh(CCF₃)⁺, analogous to the characterized CF⁺ complex described above, or from the β -position to afford the corresponding isomeric η^1 -perfluorovinyl cation (PNP)Rh(CF=CF₂)⁺. These are found to have almost identical free energies, with the carbyne complex lying only 0.4 kcal/mol higher than its perfluorovinyl isomer. An η^2 -perfluorovinyl isomer, the potential initial product of fluoride abstraction from the (PNP)Rh(C₂F₄) was found to lie 12.1 kcal/mol above its η^1 -perfluorovinyl analogue. Structures of all these compounds and their relative energies are provided in Figure 8.

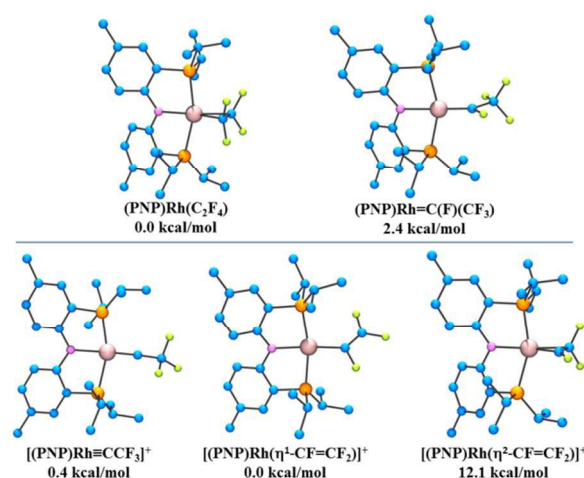


Figure 8. (Top) Calculated structures and relative energies for (PNP)Rh C₂F₄ isomers. (Bottom) Calculated structures and relative energies for proposed structures resulting from fluoride abstraction from (PNP)Rh=C(F)(CF₃).

Conclusions

In summary, we have shown that (PNP)Rh perfluorocarbene complexes can be synthesized by treating a Rh(I) precursor with Rupert's reagent or a fluoroalkane containing a C-H bond. Using silylium reagents, a fluoride can be abstracted from (PNP)Rh=CF₂ to form a cationic fluoromethylidyne. Thus the (PNP)Rh system conveniently allows synthesis and comparison of perfluoroolefin, perfluorocarbene, and perfluorocarbyne complexes. Using DFT calculations we were able to compare the Natural Localized Molecular Orbitals of these fluoroorganic complexes to their hypothetical hydrocarbon analogues. We established that the fluorine atoms on the carbene and carbyne ligands participate in π donation to the acceptor orbitals on carbon to compete with back donation from the metal. This resulted in a longer Rh-C bond in the fluorinated complexes compared to their hydrocarbon analogues. However, C₂F₄ was calculated to form a shorter Rh-C bond than the C₂H₄ complex. Calculated Wiberg Bond Indices also showed that although the unsaturated fluorocarbon ligands have bond orders greater than



one to rhodium, the nitrogen *trans* to these ligands interacts with their antibonding orbital and decreases the bond order to less than a true double and triple bond.

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Notes and references

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