

# Allylboronates from Vinyl Triflates and $\alpha$ -Chloroboronates by Reductive Nickel Catalysis

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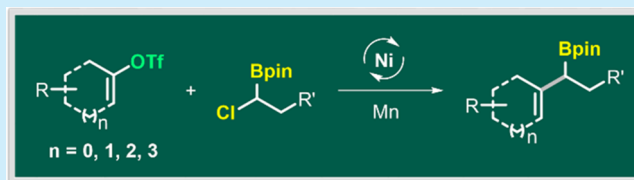


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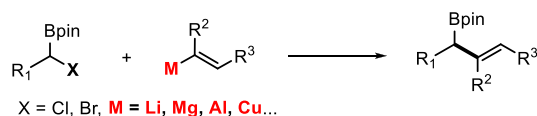
**ABSTRACT:** Allylboronates are unique building blocks widely used in organic synthesis, but the construction of cyclic allylboronates remains a challenging subject. We demonstrate here a mild and efficient access to this type of compound through the cross-electrophile coupling of vinyl triflates and  $\alpha$ -chloroboronates. The reaction proceeded with a good substrate scope and good functional group compatibility. The ready availability of vinyl triflates from ketones, as well as the rich chemistry of allylboronates, makes our method suitable for the divergent modification of biologically active compounds. Preliminary mechanistic studies revealed that  $\alpha$ -chloroboronates were activated via a radical process.



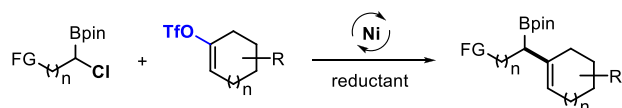
Allylboronates are essential building blocks frequently used in the synthesis of natural compounds and pharmaceuticals.<sup>1</sup> They have widely served as coupling partners in the Suzuki–Miyaura reaction for the precise installation of an allyl group.<sup>2</sup> They are also used as nucleophiles for additions to aldehydes and imines to afford homoallylic alcohols and amines.<sup>1</sup> Therefore, the synthesis of allylboronates has received considerable attention over the past years.<sup>3</sup> Approaches to these compounds typically include,<sup>4</sup> but are not limited to, (1) the boration reactions of allylic substrates such as allyl nucleophiles,<sup>5</sup> electrophiles (e.g., halides, acetates, alcohols),<sup>6</sup> and C–H bonds,<sup>7</sup> (2) the hydroboration or diboration reactions of 1,3-dienes,<sup>8</sup> and (3) the hydroboration reactions of allenes.<sup>9</sup> Despite these promising advances, these processes are generally effective for the synthesis of acyclic allylboronates. To date, a method for the construction of cyclic allylboronates, with the structure shown in Scheme 1b, remains to be developed.

## Scheme 1. Synthesis of Allylboronates from $\alpha$ -Haloboronates

(a) **Previous approach:** The nucleophilic substitution with vinyl metals



(b) **This work:** The cross-electrophile coupling with vinyl triflates



One strategy that could be used to overcome this limitation would be the reaction of well-defined vinyl species with  $\alpha$ -haloboronates (Scheme 1).<sup>10</sup> In this context, the nucleophilic substitution reactions using vinyl metallic species (e.g., vinyl–M, M = Li, Mg, Al, Cu) represent the current state-of-the-art (Scheme 1a). However, partially because of the difficulty in accessing cyclic vinyl metals, these reactions have rarely been investigated for the synthesis of cyclic allylboronates. The use of vinyl electrophiles instead of vinyl metals is advantageous in availability, and it would provide access to a complementary scope of synthetically useful products. However, to our knowledge, there has been no report on the reaction between vinyl electrophiles and  $\alpha$ -haloboronates.

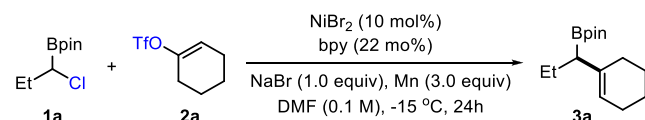
The cross-electrophile coupling has recently emerged as a powerful tool for forging C–C bonds between electrophiles.<sup>11</sup> Among the various coupling partners, the amphoteric reagents such as  $\alpha$ -haloboronates and  $\alpha$ -halosilanes are particularly attractive.<sup>12</sup> These reagents possess nucleophilic and electrophilic sites, thus offering a versatile platform for chemoselective transformations. Recent elegant work by the Martin group has demonstrated the possibility of chemoselective arylation and alkylation of  $\alpha$ -haloboronates with aryl bromides and alkenes, respectively.<sup>12b</sup> However, the cross-electrophile vinylation reaction is still unknown. As part of our ongoing interest in reductive vinylation reactions,<sup>13</sup> here we demonstrate a cross-electrophile coupling of vinyl triflates<sup>14</sup> with  $\alpha$ -haloboronates (Scheme 1b). This method provides a mild and efficient

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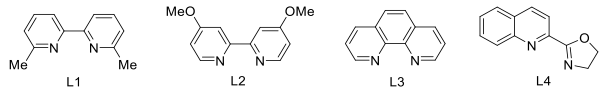
approach to cyclic allylboronates that can be otherwise difficult to access. It thus serves as a good complementary to the existing methodologies for the construction of allylboronates.

We began our investigations by studying the reaction of  $\alpha$ -chloroboronate **1a** with vinyl triflate **2a** (Table 1). After

Table 1. Optimization of Reaction Conditions<sup>a</sup>



entry	change of conditions	3a (%)
1	none	86 (84) <sup>b</sup>
2	NiCl <sub>2</sub>	80
3	NiI <sub>2</sub>	82
4	Ni(dme)Cl <sub>2</sub>	74
5	Ni(diglyme)Br <sub>2</sub>	77
6	L1 instead of bpy	61
7	L2 instead of bpy	4
8	L3 instead of bpy	71
9	L4 instead of bpy	35
10	no NaBr	78
11	r.t. instead of -15 °C	61
12	Zn instead of Mn	22
13	no Ni or Mn	0

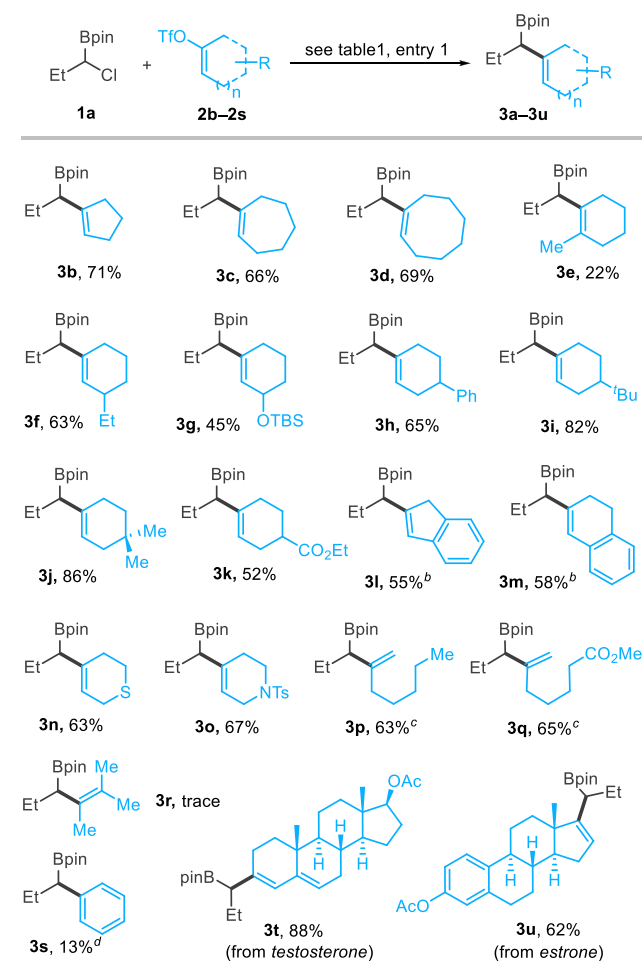


<sup>a</sup>**1a** (0.18 mmol) and **2a** (0.1 mmol) were used. The yields were determined by GC using dodecane as an internal standard. <sup>b</sup>**1a** (0.36 mmol) and **2a** (0.2 mmol) were used. Isolated yield.

screening a range of reaction conditions (see Tables S1–S7), we found that the combination of NiBr<sub>2</sub> (10 mol %), bpy (22 mol %), NaBr (1.0 equiv), and Mn (3.0 equiv) in DMF at -15 °C gave the best result, affording **3a** in 84% isolated yield (entry 1). Compatible results were obtained when NiCl<sub>2</sub> or NiI<sub>2</sub> was used, whereas the reactions with Ni(dme)Cl<sub>2</sub> or Ni(glyme)Br<sub>2</sub> gave decreased yields (entries 2–5). The inferior results were obtained when other nitrogen ligands were used (entries 6–9). The presence of NaBr has a positive effect on the yield, in which the formation of a vinyl dimer byproduct was partly inhibited (entry 10). It is possible that the in situ halide exchange of NaBr with  $\alpha$ -chloroboronate generates  $\alpha$ -bromoboronate, which is more reactive and may couple to vinyl triflate more efficiently. The reaction at room temperature afforded a significant amount of vinyl dimer, leading to a decreased yield (entry 11). The reaction with Zn as a reductant was highly ineffective (entry 12). In the absence of a nickel catalyst or reductant, no desired product was observed (entry 13).

With the optimized reaction conditions in hand, we studied the reaction with respect to the scope of vinyl triflates (Table 2). The cyclic vinyl triflates, ranging from five- to eight-membered rings, coupled to **1a** efficiently to afford the desired products in moderate to good yields (**3a–d**). Whereas the 2-substituted vinyl triflate resulted in a low yield of product **3e**, substitution at the 3- and 4-position was tolerated (**3f–k**). The moderate yields of coupling products were obtained when indenyl triflate (**3l**) and 3,4-dihydronaphthalenyl triflate (**3m**) were used. Nonaromatic heterocycles are essential structural

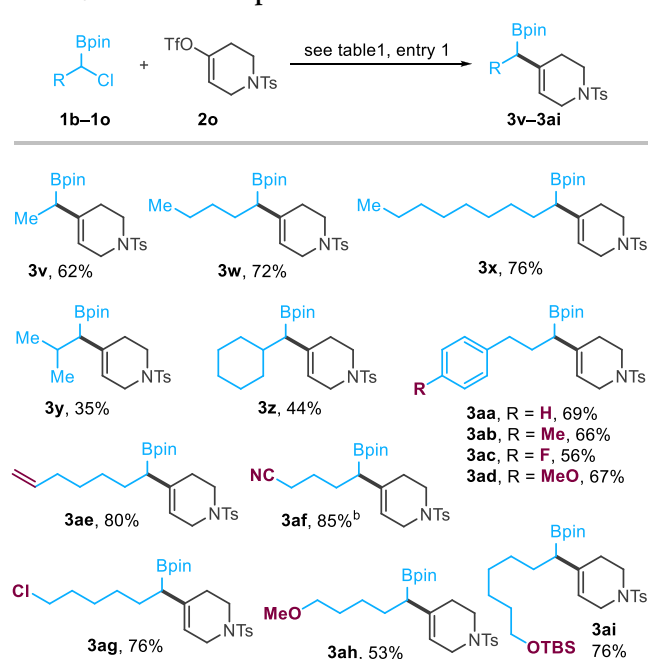
Table 2. Substrate Scope of Vinyl Triflates<sup>a</sup>



<sup>a</sup>**1a** (0.36 mmol) and vinyl triflates (0.2 mmol) were used. Isolated yields. <sup>b</sup>NMR yield was used because of the difficulty in purification with the dechloro byproduct of  $\alpha$ -chloroboronates. <sup>c</sup>4,7-Diphenyl-1,10-phenanthroline was used instead of 2,2'-bipyridine. <sup>d</sup>Phenyl triflate (0.2 mmol) was used.

motifs found in various pharmaceuticals.<sup>15</sup> Our method provides an efficient approach to produce heterocyclic allylboronates, including those bearing 3,6-dihydro-2H-thiopyran (**3n**) and 3-piperidine (**3o**) heterocycles. Reactions with acyclic vinyl substrates were less effective under the standard conditions. By changing the ligand to 4,7-diphenyl-1,10-phenanthroline, the reactions with **1a** afforded the desired products in moderate yields (**3p**, **3q**). Only a trace of product was observed when all-substituted acyclic vinyl triflate was used (**3r**). The use of phenyl triflate afforded the desired product in 13% yield (**3s**). The abundance of ketones in nature prompted us to investigate the potential of our method for the functionalization of biologically active molecules. Testosterone- and estrone-derived vinyl triflates coupled to **1a** efficiently, and allylboronate **3t** and **3u** were produced in moderate to good yield.

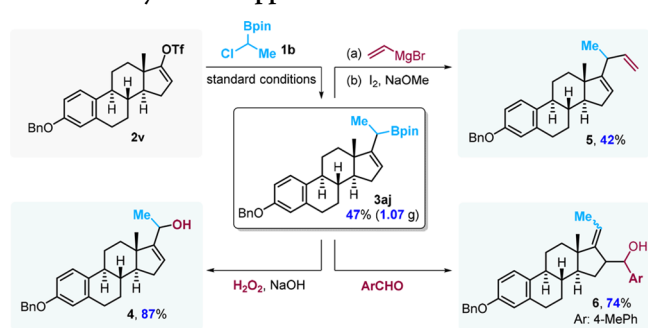
The substrate scope of  $\alpha$ -chloroboronates is shown in Table 3.  $\alpha$ -Chloroboronates with different chain lengths of the alkyl group were tolerated (**3v**, **3w**, **3x**). The reactions with sterically hindered substrates were less effective (**3y**, **3z**). The presence of an aryl group, bearing either electron-rich or electron-poor substituents, at the alkyl chain was tolerated (**3aa–ad**). The reaction has shown good functional group compatibility.  $\alpha$ -

Table 3. Substrate Scope of  $\alpha$ -Chloroboronates<sup>a</sup>

Chloroboronates with functionalities, such as alkene (3ae), nitrile (3af), alkyl chloride (3ag), ether (3ah), and silyl ether (3ai), could be selectively vinyllated to afford the products in moderate to good yields.

The modification of biologically active compounds represents a promising approach to alter their pharmacological profiles. Our method offers the opportunity for the divergent modification of these compounds. For example, estrone-derived vinyl triflates could be functionalized with 1b on the gram scale to afford allylboronate 3aj (Scheme 2). Because of

Scheme 2. Synthetic Application

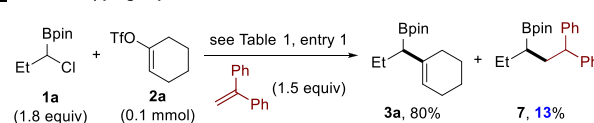


the rich chemistry of allylboronate, compound 3aj could undergo divergent transformation to provide the hydration product 4, the vinylation product 5, and the addition product 6.

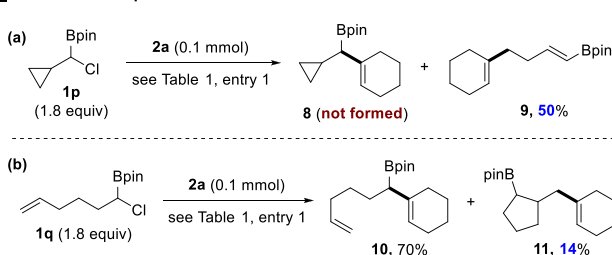
In the presence of 1,1-diphenylethylene, the reaction of 1a with 2a afforded 3a in 80% yield, along with 13% of alkene trapping product 7 (Scheme 3 (1)). This result suggests that  $\alpha$ -chloroboronates might be activated via a radical mechanism. To confirm this hypothesis, several control experiments were performed. (1) The reaction of cyclopropyl substrate 1p with 2a exclusively produced the ring-opening product (Scheme 3

Scheme 3. Mechanistic Studies

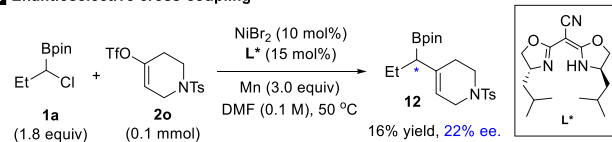
## 1 Radical trapping experiments



## 2 Radical clock experiments



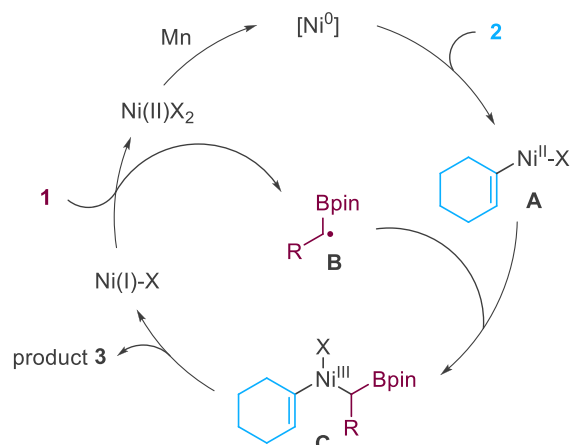
## 3 Enantioselective cross-coupling



(2a)).<sup>16</sup> (2) Under the standard conditions, the reaction of alkene substrate 1q with 2a afforded the cyclized product 11 in 14% yield (Scheme 3 (2b)). (3) The reaction of racemic 1a afforded allylboronate 12 in 22% ee when chiral ligand L\* was used (Scheme 3 (3)). These results are consistent with a pathway in which a radical process might be involved in the activation of  $\alpha$ -chloroboronates.

Although the detailed mechanism for this reaction requires further investigation, on the basis of the above results and on reports presented by other investigators,<sup>17</sup> we tentatively proposed a catalytic cycle, as shown in Scheme 4. The

Scheme 4. Proposed Mechanism



oxidative addition of vinyl triflate to Ni(0) would afford vinyl-Ni(II) (A), which may trap an alkyl radical to afford complex C.<sup>17a</sup> The reductive elimination of compound C would give the desired product 3 and generate Ni(I). Alkyl radical B might be generated via the single-electron reaction of  $\alpha$ -chloroboronates 1 with Ni(I).<sup>17c</sup>

In summary, we have demonstrated the cross-electrophile coupling of vinyl triflates with  $\alpha$ -chloroboronates. This

reaction offers a mild and efficient approach for the synthesis of cyclic allylboronates, which are difficult to access otherwise. The synthetic utility of this method was demonstrated by its application to the divergent modification of biologically active molecules. Preliminary mechanistic experiments revealed that a radical process was involved in the activation of  $\alpha$ -chloroboronates. Further studies on the reductive vinylation reactions are ongoing in our laboratories.

## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.orglett.0c01683>.

Detailed optimization of reaction conditions, part of the mechanistic studies, detailed experimental procedures, characterization data, and copies of  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra for new compounds (PDF)

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### Notes

The authors declare no competing financial interest.

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