Gonzalez, M. Challacombe, P. M. W. Gill, B. G. Johnson, W. Chen, M. W. Wong, J. L. Andres, M. Head-Gordon, E. S. Replogle, J. A. Pople, Gaussian, Inc., Pittsburgh, PA, **1998**.

[15] C. Gonzalez, T. C. Allison, E. C. Lim, *J. Phys. Chem. A* **2001**, *105*, 10583; C. Gonzalez, E. C. Lim, *J. Phys. Chem. A* **1999**, *103*, 1437; C. Gonzalez, E. C. Lim, *Chem. Phys. Lett.* **2002**, *357*, 161.

[16] F. Piuzzi, I. Dimicoli, M. Mons, P. Millié, V. Brenner, Q. Zhao, B. Soep, A. Tramer, *Chem. Phys.* **2002**, *275*, 123.

[17] S. Y. Han, J. K. Song, J. H. Kim, H. B. Oh, S. K. Kim, *J. Chem. Phys.* **1999**, *111*, 4041.

[18] J. K. Song, N. K. Lee, S. K. Kim, *J. Chem. Phys.* **2002**, *117*, 1589.

[19] J. Schiedt, R. Weinkauf, D. M. Neumark, E. W. Schlag, *Chem. Phys.* **1998**, *239*, 511.

Migration/Redistribution

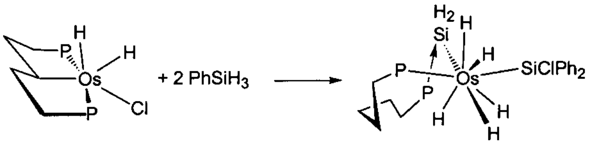
Polyhydrido(silylene)osmium and Silyl(dinitrogen)ruthenium Products Through Redistribution of Phenylsilane with Osmium and Ruthenium Pincer Complexes**

Dmitry G. Gusev, Frédéric-Georges Fontaine, Alan J. Lough, and Davit Zargarian**

Transition-metal silylene complexes, silicon analogues of metal carbenes, are reactive species believed to be important intermediates in the industrial synthesis of chloromethylsilanes,[1] the catalytic oligomerization of hydrosilanes,[2] and the redistribution chemistry of organosilanes.[3] The high reactivity of metal silylenes is related to the electron deficiency of the silicon center, which is presumably due to the very weak π bonding between the silicon atom and the transition metal.[4] Consistent with this notion, in most of the complexes reported to date[5] the silylene moiety requires stabilization from coordinating Lewis bases and/or heteroatom-based substituents (e.g., OR, SR, NR_2),[6] whereas nonstabilized silylene complexes are difficult to isolate.[7]

The tendency of Os and Ru pincer complexes to give coordinatively unsaturated alkylidene and vinylidene products in reactions with alkynes[8a] prompted us to investigate their reactivity towards hydrosilanes with the expectation that addition or 1,2 migration of the Si–H bonds might afford silylene products. Here we describe silylene and silyl complexes obtained from the reaction of $PhSiH_3$ with $[OsH_2Cl\{CH(C_2H_4PtBu_2)_2\}]$ (**1**),[8b] $[OsH_2Cl\{2,6\text{-}(CH_2PtBu_2)_2C_6H_3\}]$ (**2**),[8c] and $[RuHCl\{1,3\text{-}(CH_2PtBu_2)_2\text{-}C_6H_4\}]$ (**3**).[8d]

Addition of 2 equiv of $PhSiH_3$ to **1** in toluene afforded the pentahydrido complex **4** in 83% yield [Eq. (1), P = P*t*Bu$_2$]. A



1 **4** (1)

single-crystal X-ray analysis of **4** showed that the Os center adopts an 8-coordinate, dodecahedral geometry consisting of two orthogonal, trapezoidal planes defined by the atoms Si1,$H1_{Os}$,$H5_{Os}$,P2 and Si2,$H3_{Os}$,$H4_{Os}$,$H2_{Os}$ (Figure 1, left). The Os–Si distances for the silyl (2.374 Å) and the base-stabilized silylene (2.386 Å) ligand are in the normal range for Os–Si single bonds in hydrido(silyl)osmium complexes (3.34 to 2.49 Å).[9] The anticipated uncertainty associated with locating the hydride ligands in the solid-state structure of **4** prompted a study of its structure using ONIOM (DFT/HF) geometry optimization (Figure 1, right);[10] the closeness of the computational and experimental data indicates that the structure of this complex has been established reliably. Solution NMR data are also consistent with the proposed structure for **4** and show that the hydride ligands are highly fluxional. For example, the averaged OsH_5 resonance did not decoalesce down to −100 °C (^{1}H NMR, [D_8]toluene) and its

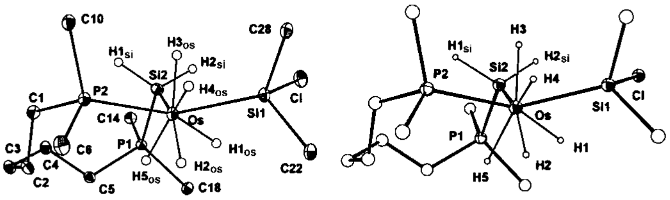


Figure 1. Partial experimental (left) and calculated (right) structures of **4** with the methyl groups and the co-crystallized toluene molecule omitted for clarity. Key experimental parameters [Å, deg]: Os-Si1 2.3739(8), Os-Si2 2.3857(8), Os-P2 2.4044(8), P1-Si2 2.361(1); P1-Si2-Os 120.29(3), P2-Os-Si1 140.63(3), P2-Os-Si2 100.96(3), Si1-Os-Si2 104.84(3), $H1_{Si}$-Si2-$H2_{Si}$ 102.9(19), H_{Si}-Si2-Os (av.) 120.2(14). Calculated parameters [Å, deg]: Os-Si1 2.417, Os-Si2 2.409, Os-P2 2.454, P1-Si2 2.363, Si1⋯H1 2.071, Si2⋯H3 2.246, Os-H1–Os-H5 1.638, 1.635, 1.670, 1.642, 1.653; P1-Si2-Os 124.3, P2-Os-Si1 142.2, P2-Os-Si2 99.8, Si1-Os-Si2 101.3, $H1_{Si}$-Si2-$H2_{Si}$ 101.0, H_{Si}-Si2-Os (av.) 120.9, H2⋯H5 1.883 (the shortest H⋯H separation).

---

[*] Prof. D. G. Gusev
Department of Chemistry
Wilfrid Laurier University
Waterloo, ON, N2L3C5 (Canada)
Fax: (+1) 519-746-0677
E-mail: dgoussev@wlu.ca

Dr. A. J. Lough
Department of Chemistry (X-Ray Laboratory)
University of Toronto
Toronto, ON, M5S3H6 (Canada)

Prof. D. Zargarian, F.-G. Fontaine
Département de chimie
Université de Montréal
Montréal, PQ, H3C 3J7 (Canada)
Fax: (+1) 514-343-7586
E-mail: zargarian.davit@umontreal.ca

[**] We gratefully acknowledge research funding from NSERC (D.G. and D.Z.), Research Corporation (D.G.), and FCAR (D.Z.). F.-G.F. is grateful to NSERC for PGS(A&B) scholarships.

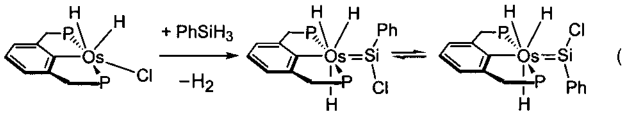
Supporting information for this article is available on the WWW under http://www.angewandte.org or from the author.

1433-7851/03/4202-0216 $ 20.00+.50/0

average relaxation time $T1_{min}$ was long (215 ms at −20 °C and 300 MHz), which is typical of polyhydrides with no short H···H contacts.

Evidently, complex **1** and two molecules of $PhSiH_3$ have undergone a series of redistribution reactions culminating in the net hydrogenation of the Os−C bond and the generation of one silyl, one base-stabilized silylene, and five hydride ligands. Complex **4** is the first compound bearing a base-stabilized SiH_2 ligand. We have briefly probed the relative stability of the P→Si interaction in the presence of other Lewis bases. Thus, addition of 1 equiv of NEt_3 to a C_6D_6 solution of **4** reduced the intensities of the ^{31}P resonances attributed to the P nuclei coordinated to the SiH_2 moiety ($\delta = 14.5$ ppm) and the Os center ($\delta = 56.6$ ppm), and gave rise to two new resonances at $\delta = 27.4$ and 50.3 ppm (conversion ca. 50 %). The chemical shift of one of the new resonances is very close to that of the uncoordinated bis(di-*tert*-butylphosphanyl)pentane ligand ($\delta = 27.7$ ppm in C_6D_6), indicating that this ligand is monodentate in the resulting complex. Similar observations were made when 1 equiv of PCy_3 was reacted with **4**: new ^{31}P resonances emerged at $\delta = 27.5$, 48.9, and −4.9 ppm (conversion ca. 80 %); decoupling experiments have allowed us to assign the latter resonance to the $Cy_3P \rightarrow SiH_2$ moiety. These preliminary observations demonstrate that good Lewis bases can displace the $PtBu_2$ group, thereby generating new donor-stabilized SiH_2 complexes.

To better understand the influence of the ligand on the reactivity of **1,** we investigated the reaction of $PhSiH_3$ with complex **2,** a direct analogue of **1** that has a pincer ligand with a more rigid backbone. In this case, the trihydrido(silylene) product **5** was obtained in 60 % yield [Eq. (2), P = $PtBu_2$].



(2)

2 **5a** **5b**

Complex **5** undergoes a slow decomposition at room temperature, both in the solid and in solution, as evident from a gradual color change from yellow to dark brown. The thermal instability of **5** prevented crystallographic characterization, but NMR spectra collected at low temperature using samples prepared at −70 °C established the identity of the product as a trihydrido(silylene) complex that exists as two rotamers in solution, as described below. In addition, a computational study using the ONIOM (DFT/HF) methodology supported the structural assignments for both isomers (the optimized structure of **5a** is shown in Figure 2).

The $^{31}P\{^1H\}$ NMR spectrum of the isolated solid showed a single resonance ($\delta = 85.7$ ppm), whereas the $^{29}Si\{^1H\}$ NMR spectrum showed two triplets (6:1 signal intensity ratio) in a chemical shift region characteristic of silylenes ($\delta = 239.3$ (**5a**) and 246.5 ppm (**5b**), $^2J_{PSi} = 9.4$ Hz). Two sets of resonances were also observed in the hydride region of the ^1H NMR spectrum (1:2 intensity ratio within each set), implying two different trihydrido substructures. We interpret these observations to indicate that **5** is a silylene complex with a plane of symmetry that renders the two phosphorus nuclei equivalent; the rotation of the silylene moiety about the Os−Si bond gives rise to two rotamers, which were assigned the structures **5a** and **5b** on the basis of the difference NOE spectra. For example, isomer **5a** showed a strong NOE between the SiPh and the H2/H3 signals, whereas no NOE was detected between the signals for the *ortho* hydrogen atoms of SiPh and H1 (numbering according to Figure 2). In contrast, a strong NOE was detected in isomer **5b** between the *ortho* hydrogen atoms of SiPh and H1. The NOE experiments also indicated a slow exchange at −70 °C between **5a** and **5b**: irradiation of the H1 resonance of one isomer resulted in a partial saturation transfer to the corresponding hydride resonance of the other. Integration of the hydride signals allowed us to determine that the **5b**/**5a** equilibrium has a small ΔG (−0.7 kcal mol^{-1}); consistent with this, the computational study showed a very small energy difference (−0.01 kcal mol^{-1}) between the two isomers.

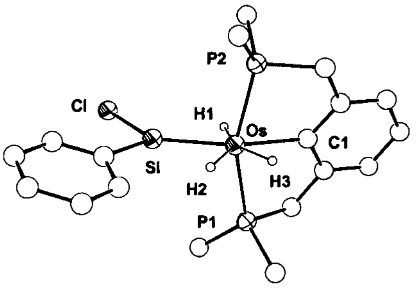
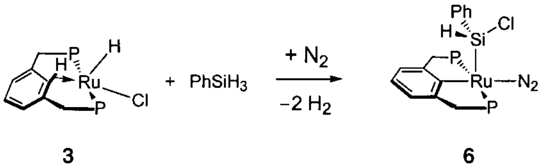


Figure 2. Partial calculated structure of **5a** with the methyl groups omitted for clarity. Key parameters [Å, deg]: Os-Si 2.281, Os-C1 2.171, Os-P1 2.387, Os-P2 2.394, Os-H1 1.670, Os-H2 1.641, Os-H3 1.651, H2···H3 1.652; P1-Os-P2 157.3, Si-Os-C1 161.9, Si-Os-H1 82.4, Si-Os-H2 63.9, C1-Os-H1 79.9, C1-Os-H3 74.3, H2-Os-H3 60.2.

The different reactivities of **1** and **2** underline the importance of the auxiliary ligand in determining the course and outcome of the reaction with $PhSiH_3$. We have also probed the importance of the metal atom by reacting $PhSiH_3$ with **3**, the Ru analogue of **2** that displays an agostic interaction between the central C−H moiety and the Ru center. This reaction gave the dinitrogen(silyl) complex **6** in 75 % yield [Eq. (3), P = $PtBu_2$]. The identity of **6** was established by spectroscopic studies and X-ray crystallography (Figure 3), as described below.



(3)

3 **6**

The IR spectrum of complex **6** features strong N≡N and Si−H stretches at 2142 and 2105 cm^{-1}, respectively. The presence of the Si−H group was also evident from the NMR spectra, which showed a ^1H signal ($\delta = 5.37$ ppm, d, $^3J_{PH} =$

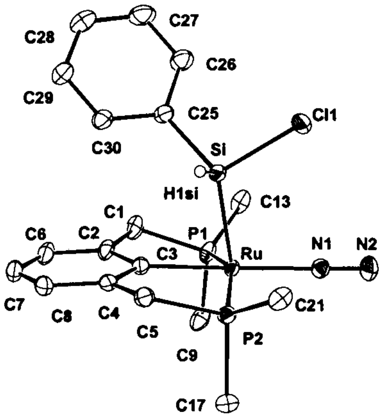


Figure 3. Partial experimental structures of **6** with the methyl groups omitted for clarity. Key parameters [Å, deg]: Ru-C3 2.068(3), Ru-Si 2.2821(7), Ru-N1 2.014(2), Ru-P1 2.3512(7), Ru-P2 2.3632(7), N1-N2 1.099(3); P1-Ru-P2 157.35(3), C3-Ru-Si 78.23(7), C3-Ru-N1 178.5(1), Si-Ru-N1 100.33(7).

17.7 Hz) featuring the anticipated Si satellites ($^1J_{SiH} = 206$ Hz) and a ^{29}Si{^{1}H} signal with the appropriate multiplicity ($\delta = 31.5$ ppm, t, $^2J_{PSi} = 24$ Hz). The asymmetric nature of this silyl ligand renders the CH_2 and P*t*Bu$_2$ groups inequivalent in the NMR spectra of **6**. According to the X-ray data, the geometry around Ru in **6** is square-pyramidal with the silyl ligand occupying the apical position; the absence of a ligand *trans* to the silyl moiety is presumably responsible for the very short Ru–Si bond (2.282 Å). The Ru–N (2.014 Å) and N≡N (1.099 Å) distances are similar to the corresponding distances in related neutral dinitrogen complexes: [RuH(N$_2$){CH(C$_2$H$_4$P*t*Bu$_2$)$_2$}] (1.965/1.117 Å),[11a] [RuH$_2$(N$_2$){κ^3(*P,P,P*)-PhP(C$_3$H$_6$PCy$_2$)$_2$}] (2.005/1.093 Å),[11b] and [(PPh$_3$)$_2$RuH-(μ-H$_3$)Ru(N$_2$)(PPh$_3$)$_2$] (2.003/1.086 Å).[11c]

A comparison of the reactivities observed for complexes **2** and **3** provides some insight into the mechanism of their reactions with PhSiH$_3$. The following sequence of steps can be envisaged: 1) addition of PhSiH$_3$ to **2**/**3** generates the initial silyl ligand PhSiH$_2$ and releases H$_2$; 2) a redistributive SiH/MCl exchange gives the 16-electron intermediate [(PCP-ligand)MH$_2$(PhSi(H)Cl)]; 3) the Os^{IV} intermediate undergoes a 1,2 migration of SiH to give the trihydrido(silylene) complex **5**, whereas its Ru counterpart loses H$_2$ to give the dinitrogen(silyl) complex **6**. The difference in the reactivities of **2** and **3** is presumably due to the weaker Ru–H bonds (compared to Os–H bonds) and the prevalence of the +2 oxidation state in the chemistry of Ru complexes stabilized by phosphanes.

In summary, the described reactions of PhSiH$_3$ with complexes **1**–**3** have provided an opportunity to study the direct and unassisted conversion of a hydrosilane to transition-metal silylene derivatives, a transformation believed to be important in many catalytic processes involving organosilicon compounds.[12] The spontaneous silyl-to-silylene conversions observed here are rather rare, though they can be induced by photolysis[13] or assisted by Lewis acids.[14] On the other hand, silylene formation by Si–SiR$_3$ migrations has been observed with Rh^{I},[15] Ir^{I},[16] and Ni^{II}[17] systems. The transformations with the Os complex **1** are even more interesting, since they involve complex redistribution reactions that convert two molecules of PhSiH$_3$ to the silyl ligand Ph$_2$SiCl and the unprecedented parent silylene moiety SiH$_2$.

Experimental Section

Spectroscopic data for compounds **4**–**6** are given below; all other information pertaining to the synthetic and computational work is provided in the Supporting Information. Complete crystallographic data for the structural analyses have been deposited with the Cambridge Crystallographic Data Centre. CCDC-190307 (**4**) and CCDC-190308 (**6**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/conts/retrieving.html (or from the Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB21EZ, UK; fax: (+44) 1223-336-033; or deposit@ccdc.cam.ac.uk).

Complex **4**: elemental analysis calcd for $C_{33}H_{63}ClOsP_2Si_2 \cdot$toluene (895.811): C 53.63, H 7.99; found: C 53.61, H 8.40. IR (Nujol): $\tilde{\nu}$(ν_{SiH}, ν_{OsH}) = 2079, 2066, 2001, 1913 cm^{-1}. NMR (20 °C, C_6D_6, *J* in Hz): ^{1}H NMR: $\delta = -9.70$ (ddt, $^2J_{HP} = 5.9$, $^3J_{HP} = 8.1$, $^3J_{HH} = 1.6$, 5H, OsH$_5$), 1.00 (d, $^3J_{HP} = 12.9$, 18H, CH$_3$), 1.16 (d, $^3J_{HP} = 12.3$, 18H, CH$_3$), 1.1–1.8 (m, 10H, CH$_2$), 4.74 (dd, $^2J_{HP} = 19.0$, $^3J_{HP} = 3.0$, $^1J_{HSi} = 179.3$, SiH$_2$), 7.12, 7.30, 8.34 ppm (m, 10H, Ph). ^{13}C{^{1}H} NMR: $\delta = 14.0$ (d, $J_{CP} = 21.7$, CH$_2$), 20.3 (s, CH$_2$), 21.0 (d, $J_{CP} = 17.7$, CH$_2$), 22.9 (d, $J_{CP} = 1.6$, CH$_2$), 29.6 (dd, $J_{CP} = 8.5$, 10.2, CH$_2$), 29.4 (s, CH$_3$), 30.0 (s, CH$_3$), 34.6 (d, $J_{CP} = 17.1$, PC), 35.3 (d, $J_{CP} = 20.7$, PC), 127.0 (s, CH), 127.5 (s, CH), 135.4 (s, CH), 151.5 ppm (d, $J_{CP} = 2.7$, C). ^{31}P{^{1}H} NMR: $\delta = 56.6$ ($^2J_{PSi} = 27.8$), 14.5 ppm ($^1J_{PSi} = 44.6$). ^{29}Si{^{1}H} NMR: $\delta = 30.4$ (dd, $^3J_{SiP} = 2.8$, $^2J_{SiP} = 27.8$), −67.4 ppm (dd, $^2J_{SiP} = 5.5$, $^1J_{SiP} = 44.7$).

Complex **5**: IR (Nujol): $\tilde{\nu}$(ν_{OsH}) = 2177, 2089, 1870 cm^{-1}. NMR (−70 °C, [D$_8$]toluene, *J* in Hz): ^{1}H NMR: $\delta = -7.50$ (**5a**), −7.34 (**5b**) (br, 2H, OsH$_2$), −6.98 (**5b**) (t, $^2J_{HP} = 11.4$, OsH), −6.33 ppm (**5a**) (t, $^2J_{HP} = 13.8$, OsH). ^{1}H{^{31}P} NMR: $\delta = 1.01$ (s, 18H, CH$_3$), 1.18 (s, 18H, CH$_3$), 3.24 (d, $^2J_{HH} = 15.9$, 2H, CH$_2$), 3.41 (d, $^2J_{HH} = 15.9$, 2H, CH$_2$), 7.0–7.3 (m, 6H, Ar), 8.66 ppm (d, $^2J_{HP} = 7.5$, 2H, Ph). ^{31}P{^{1}H} NMR: $\delta = 85.7$ ppm (s). ^{13}C{^{1}H} NMR: $\delta = 29.2$ (br, CH$_3$), 34.0 (vt, $^vJ = 10.5$, PC), 35.9 (vt, $^vJ = 12.0$, PC), 42.8 (vt, $^vJ = 14.8$, CH$_2$), 120.0 (vt, $^vJ = 7.8$, CH, Ar), 123.3 (s, CH, Ar), 132.4 (s, CH, Ph), 128.8 (s, CH, Ph), 135.2 (s, CH, Ph), 149.3 (vt, $^vJ = 8.5$, C, Ar), 149.9 (s, C, Ph), 155.2 ppm (t, $^2J_{CP} = 3.7$, RuC). ^{29}Si{^{1}H} NMR: $\delta = 239.3$ (**5a**), 246.5 ppm (**5b**) (t, $^2J_{SiP} = 9.4$).

Complex **6**: elemental analysis calcd for $C_{30}H_{49}ClN_2P_2RuSi$ (664.29): C 54.24, H 7.43, N 4.22; found: C 53.90, H 7.64, N 4.02. IR (Nujol): $\tilde{\nu} = 2142$ ($\nu_{N=N}$), 2105 cm^{-1} (ν_{SiH}). NMR ([D$_8$]toluene, *J* in Hz): ^{1}H NMR: $\delta = 5.37$ ppm (d, $^3J_{HP} = 17.7$, $^1J_{HSi} = 206$, SiH). ^{1}H{^{31}P} NMR: $\delta = 0.90$, 0.94, 1.20, 1.42 (s, 36H, CH$_3$), 2.56 (d, $^2J_{HH} = 16.4$, 1H, CH$_2$), 2.78 (d, $^2J_{HH} = 16.4$, 1H, CH$_2$), 3.09 (d, $^2J_{HH} = 16.7$, 1H, CH$_2$), 3.73 (d, $^2J_{HH} = 16.7$, 1H, CH$_2$), 6.9–7.1 ppm (m, 8H, Ar). ^{31}P{^{1}H} NMR: $\delta = 74.8$, 80.7 ppm (d, $^2J_{PP} = 217$). ^{13}C{^{1}H} NMR: $\delta = 30.5$, 30.6, 32.1, 32.9 (d, CH$_3$), 35.3 (d, $^1J_{CP} = 21.9$, CH$_2$), 36.3 (d, $^1J_{CP} = 19.3$, CH$_2$), 35.0, 36.6, 37.5, 38.1 (d, PC), 121.9 (d, $^3J_{CP} = 17.0$, CH, Ar), 122.5 (d, $^3J_{CP} = 16.7$, CH, Ar), 124.2 (s, CH, Ar), 126.5, 127.8, 135.6 (s, CH, Ph), 144.6 (d, $^3J_{CP} = 2.3$, SiC, Ph), 156.0 (m, C, Ar), 178.1 ppm (t, $^2J_{CP} = 2.9$, RuC). ^{29}Si{^{1}H} NMR: $\delta = 31.5$ ppm (t, $^2J_{SiP} = 24$).

Received: July 25, 2002 [Z19819]

---

[1] For details on the preparation of chloromethylsilanes from elemental silicon and chloromethane, the so-called Direct Process, see: a) M. P. Clarke, *J. Organomet. Chem.* **1989**, *376*, 165; b) M. P. Clarke, *J. Organomet. Chem.* **1991**, *408*, 149; c) for information on the reactivities of silylenes with chloroalkanes in solution see: P. W. Wanandi, P. B. Glaser, T. D. Tilley, *J. Am. Chem. Soc.* **2000**, *122*, 972.

[2] a) K. A. Brown-Wensley, *Organometallics* **1987**, *6*, 1590; b) E. A. Zarate, C. Tessier-Youngs, W. J. Youngs, *J. Am. Chem. Soc.* **1988**, *110*, 4068.
[3] H. Yamashita, M. Tanaka, M. Goto, *Organometallics* **1992**, *11*, 3227.
[4] H. Jacobsen, T. Ziegler, *Organometallics* **1995**, *14*, 224, and references therein.
[5] The first transition-metal silylene complexes were reported by the groups of Tilley and Zybill: a) D. A. Straus, T. D. Tilley, A. L. Rheingold, S. J. Geib, *J. Am. Chem. Soc.* **1987**, *109*, 5872; b) C. Zybill, G. Müller, *Angew. Chem.* **1987**, *99*, 683; *Angew. Chem. Int. Ed. Engl.* **1987**, *26*, 669.
[6] For reviews on transition-metal silylene complexes see: a) C. Zybill, *Top. Curr. Chem.* **1991**, *160*, 1; b) C. Zybill, H. Handwerker, H. Friedrich, *Adv. Organomet. Chem.* **1994**, *36*, 229.
[7] a) J. D. Feldman, G. P. Mitchell, J.-O. Nolte, T. D. Tilley, *J. Am. Chem. Soc.* **1998**, *120*, 11184; b) S. K. Grumbine, T. D. Tilley, F. P. Arnold, A. L. Rheingold, *J. Am. Chem. Soc.* **1994**, *116*, 5495.
[8] a) D. G. Gusev, T. Maxwell, F. M. Dolgushin, M. Lyssenko, A. J. Lough, *Organometallics* **2002**, *21*, 1095; b) D. G. Gusev, A. J. Lough, *Organometallics* **2002**, *21*, 2601; c) D. G. Gusev, F. M. Dolgushin, M. Y. Antipin, *Organometallics* **2001**, *20*, 1001; d) D. G. Gusev, M. Madott, F. M. Dolgushin, K. A. Lyssenko, M. Y. Antipin, *Organometallics* **2000**, *19*, 1734.
[9] a) C. E. F. Rickard, W. R. Roper, S. D. Woodgate, L. J. Wright, *J. Organomet. Chem.* **2000**, *609*, 177; b) G. R. Clark, K. R. Flower, C. E. F. Rickard, W. R. Roper, D. M. Salter, L. J. Wright, *J. Organomet. Chem.* **1993**, *462*, 331; c) M. L. Buil, P. Espinet, M. A. Esteruelas, F. J. Lahoz, A. Lledos, J. M. Martinez-Ilarduya, F. Maseras, J. Modrego, E. Onate, L. A. Oro, E. Sola, C. Valero, *Inorg. Chem.* **1996**, *35*, 1250.
[10] a) The computational studies were performed on a 2.2 GHz Pentium IV PC using Gaussian 98 (Revision A.11), M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, V. G. Zakrzewski, J. A. Montgomery, R. E. Stratmann, J. C. Burant, S. Dapprich, J. M. Millam, A. D. Daniels, K. N. Kudin, M. C. Strain, O. Farkas, J. Tomasi, V. Barone, M. Cossi, R. Cammi, B. Mennucci, C. Pomelli, C. Adamo, S. Clifford, J. Ochterski, G. A. Petersson, P. Y. Ayala, Q. Cui, K. Morokuma, D. K. Malick, A. D. Rabuck, K. Raghavachari, J. B. Foresman, J. Cioslowski, J. V. Ortiz, B. B. Stefanov, G. Liu, A. Liashenko, P. Piskorz, I. Komaromi, R. Gomperts, R. L. Martin, D. J. Fox, T. Keith, M. A. Al-Laham, C. Y. Peng, A. Nanayakkara, C. Gonzalez, M. Challacombe, P. M. W. Gill, B. G. Johnson, W. Chen, M. W. Wong, J. L. Andres, M. Head-Gordon, E. S. Replogle, J. A. Pople, Gaussian, Inc., Pittsburgh, PA, **1998**. The geometry optimizations were carried out using the ONIOM(B3PW91:HF/LANL2MB) method[10b], and the nature of the stationary points as a minimum of energy was verified in frequency calculations. For the optimized structures of **4**, **5a**, and **5b** only real frequencies are obtained. b) M. Svensson, S. Humbel, R. D. J. Froese, T. Matsubara, S. Sieber, K. Morokuma, *J. Phys. Chem.* **1996**, *100*, 19357; F. Maseras, *Chem. Commun.* **2000**, 1821.
[11] a) D. G. Gusev, A. J. Lough, *Organometallics* **2002**, *21*, 5019; b) G. Jia, D. W. Meek, J. C. Gallicci, *Inorg. Chem.* **1991**, *30*, 403; c) B. Chaudret, J. Devillers, R. Poilblanc, *Organometallics* **1985**, *4*, 1727.
[12] a) M. D. Curtis, P. S. Epstein, *Adv. Organomet. Chem.* **1981**, *19*, 213; b) T. D. Tilley in *The Chemistry of Organic Silicon Compounds* (Eds.: S. Patai, Z. Rappaport), Wiley, New York, **1989**, Chap. 24, p. 1415; c) T. D. Tilley, *Comments Inorg. Chem.* **1990**, *10*, 37; d) T. D. Tilley in *The Silicon–Heteroatom Bond* (Eds.: S. Patai, Z. Rappaport), Wiley, New York, **1991**, Chap. 9 and 10, pp. 245 and 309; e) K. H. Pannel, H. L. Sharma, *Chem. Rev.* **1995**, *95*, 1351.
[13] R. J. P. Corriu, G. F. Lanneau, B. P. Chauhan, *Organometallics* **1993**, *12*, 2001.
[14] a) G. P. Mitchell, T. D. Tilley, *J. Am. Chem. Soc.* **1998**, *120*, 7635; b) G. P. Mitchell, T. D. Tilley, *Angew. Chem.* **1998**, *110*, 2602; *Angew. Chem. Int. Ed.* **1998**, *37*, 2524; c) J. C. Peters, J. D. Feldman, T. D. Tilley, *J. Am. Chem. Soc.* **1999**, *121*, 9871.
[15] G. P. Mitchell, T. D. Tilley, *Organometallics* **1996**, *15*, 3477.
[16] G. P. Mitchell, T. D. Tilley, G. P. A. Yap, A. L. Rheingold, *Organometallics* **1995**, *14*, 5472.
[17] S. Nlate, E. Herdtweck, R. A. Fischer, *Angew. Chem.* **1996**, *108*, 1957; *Angew. Chem. Int. Ed. Engl.* **1996**, *35*, 1861.

Asbestos Decontamination

Soil Fungal Hyphae Bind and Attack Asbestos Fibers**

Elena Martino, Laura Prandi, Ivana Fenoglio, Paola Bonfante, Silvia Perotto, and Bice Fubini**

Once a major issue in occupational health and safety, asbestos has now become a general environmental problem.[1] Several mountain areas—from the western Alps in Italy to the Sierra Nevada in the USA—are rich in asbestos and asbestiform minerals, and many defunct asbestos industries and mines have left substantial amounts of asbestos fibers on the abandoned sites. Exposure to airborne asbestos fibrils causes a severe pneumoconiosis (asbestosis) and malignancies such as bronchogenic carcinoma and pleural mesothelioma.[2–4] The decontamination of asbestos fibers dispersed over wide areas of soil and in waters obviously requires a different approach from what was proposed for asbestos localized in buildings.[5] The fibers cannot be removed but have to be inactivated in situ without damaging the environment. New, environmentally friendly techniques are thus required for

---

[*] Prof. Dr. S. Perotto, Dr. E. Martino,[+] Prof. Dr. P. Bonfante
Interdepartmental Center “G. Scansetti” for Studies on Asbestos and Other Toxic Particulates
University of Torino, Dipartimento Biologia Vegetale
via le Mattioli 25, 10125 Torino (Italy)
Fax: (+39) 11-670-7459
E-mail: silvia.perotto@unito.it

Prof. Dr. B. Fubini, Dr. L. Prandi,[+] Dr. I. Fenoglio
Interdepartmental Center “G. Scansetti” for Studies on Asbestos and Other Toxic Particulates
University of Torino, Dipartimento di Chimica IFM
via P. Giuria 7, 10125 Torino (Italy)
Fax: (+39) 11-670-7566
E-mail: bice.fubini@unito.it

[+] E.M. and L.P. contributed equally to the work.

[**] The research was funded in part by CNR, Finalized Project on Biotechnology (Subproject 2) and by Regione Piemonte, Italy.

Supporting information for this article is available on the WWW under http://www.angewandte.org or from the author.