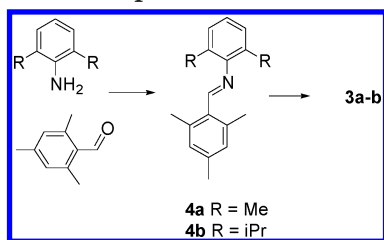


Scheme 1. Preparation of Palladacycles**Results and Discussion**

Preparation of Ligands and Complexes. Substituted trimethylbenzylideneamines were prepared by the condensation of 2,4,6-trimethylbenzaldehyde with a substituted aniline (Scheme 1). Both infrared absorptions near 1626 cm^{-1} and ^{13}C NMR shifts around 162–163 ppm of imines **4a,b** are characteristic of the functionality of $\text{C}=\text{N}$. Cyclopalladation reactions were carried out in a mixture of $(\text{CH}_3\text{CN})_2\text{PdCl}_2$, sodium acetate, and the corresponding ligand in tetrahydrofuran at ambient temperature for 38 h. The desired palladium products **3a,b** were isolated as air-stable solids upon crystallization. In both instances, the cyclopalladation readily occurs at the benzylic position to form an endo six-membered chelating ring, and the product is in a chloro-bridged dipalladium structure, which is consistent with most of the related species.¹⁰ The C–H activation at the benzylic position is established by their ^1H NMR and X-ray single-crystal determination. The appearance of a signal at 3.19 ppm for **3a** and 3.24 for **3b** with the integration of 2H in the ^1H NMR spectra was assigned to the methylene protons of $\text{Pd}-\text{CH}_2-$ resulting from the C–H activation with the palladium complex.

The detailed structure of palladacycle **3a** was proved by an X-ray diffraction study on single crystals grown from a hexane/dichloromethane solution, whereas the structural characterization of **3b** has been published previously.⁹ Figure 1 displays the ORTEP plot of **3a**, and Table 1 summarizes the selected bond distances,

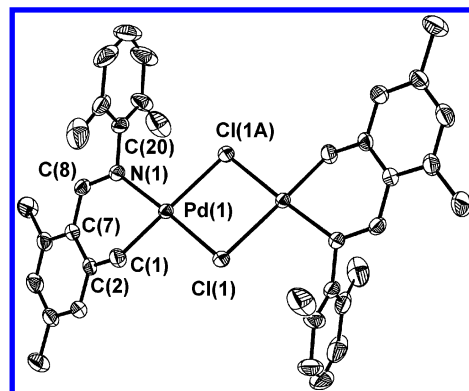


Figure 1. ORTEP drawing of **3a** (drawn with 30% probability ellipsoids; labeling of phenyl groups is omitted for clarity).

Table 1. Selected Bond Distances (Å) and Bond Angles (deg) of **3a and **3b****

	3a	3b^a
Pd(1)–N(1)	2.024(3)	2.043(3)
Pd(1)–C(1)	2.002(3)	2.003(4)
Pd(1)–Cl(1)	2.3283(7)	2.498(1)
Pd(1)–Cl(1A)	2.5051(8)	2.325(1)
N(1)–C(7)	1.298(4)	1.280(5)
N(1)–Pd(1)–C(1)	86.0(1)	86.2(1)
Pd(1)–N(1)–C(7)	126.5(2)	126.6(3)
Pd(1)–C(1)–C(2)	115.8(2)	116.0(3)
N(1)–Pd(1)–Cl(1)	176.95(8)	97.28(8)
N(1)–Pd(1)–Cl(1A)	96.96(7)	176.42(9)
C(1)–Pd(1)–Cl(1)	91.11(9)	90.3(1)
C(1)–Pd(1)–Cl(1A)	174.7(1)	171.2(1)
Cl(1)–Pd(1)–Cl(1A)	85.81(3)	86.29(3)

Torsion Angles (deg) along the Chelate Ring of 3a			
Pd(1)–N(1)–C(7)–C(8)	–2.0	C(8)–C(2)–C(1)–Pd(1)	–47.2
N(1)–C(7)–C(8)–C(2)	22.8	C(2)–C(1)–Pd(1)–N(1)	49.6
C(7)–C(8)–C(2)–C(1)	4.6	C(1)–Pd(1)–N(1)–C(7)	–29.5

^a Reference 9.

bond angles, and torsion angles. In both instances, the palladium metal displays a slightly distorted square-planar geometry with nitrogen and carbon donors in *cis*-fashion. All bond distances and bond angles lie within normal ranges, which are essentially similar to those for **3b** except Pd(1)–N(1). The distance of Pd(1)–N(1) [2.024(3) Å] for **3a** is slightly shorter than that for **3b** [2.043(3) Å], which is attributed to steric differences between methyl and isopropyl groups. The length of C(7)–N(1) [1.298(4) Å] is characteristic for a $\text{C}=\text{N}$ double bond. The small bite angle N(1)–Pd(1)–C(1) [86.0(1)°] is similar to those benzylic-type palladacycles reported by Sales and co-workers.^{10e} It is noticed that the angle Pd(1)–C(1)–C(2) [115.8(2)°] deviating from the angle for a tetrahedral geometry, similar to that for **3b**, is presumably due to the strain forced by the chelating rings. The dihedral angles along the chelating ring of Pd(1)–N(1)–C(8)–C(7)–C(2)–C(1) in **3a** (Table 1) indicate that this ring is adopted into a twist half-chair conformation, apparently resulting from the occurrence of two double bonds in the chelate ring.

Phosphine-Substituted Complexes. Both palladacycles **3a** and **3b** readily underwent ligand substitution reaction with phosphines (eq 1) to give the corresponding complex. Both complexes were in air-stable crystalline forms. The appearance of only one signal around 40 ppm in the ^{31}P NMR spectra for both

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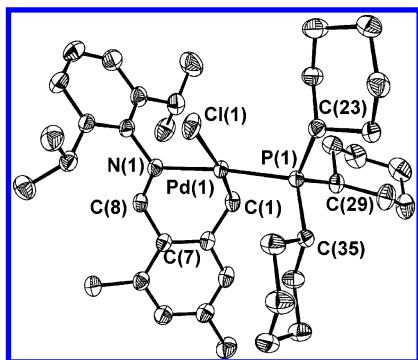
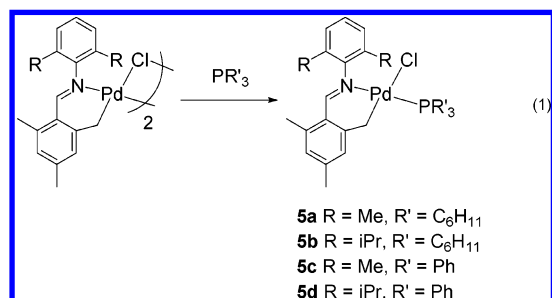


Figure 2. Molecular structure of **5b** (30% probability ellipsoids).

Table 2. Selected Bond Distances (Å), Bond Angles (deg), and Torsional Angles (deg) of **5b**, **5d**, and **7**

	5b	5d	7
Pd(1)–N(1)	2.148(2)	2.144(2)	2.109(2)
Pd(1)–C(1)	2.037(3)	2.049(3)	2.016(2)
Pd(1)–Cl(1)	2.4136(9)	2.3682(8)	2.3904(7)
Pd(1)–P(1)	2.2683(7)	2.2388(7)	2.3049(7)
N(1)–C(8)	1.275(4)	1.282(3)	1.279(3)
P(1)–Pd(1)–N(1)	174.30(7)	172.67(6)	171.69(6)
C(1)–Pd(1)–N(1)	82.7(1)	84.2(1)	80.22(9)
P(1)–Pd(1)–Cl(1)	94.62(9)	95.25(3)	89.87(2)
C(1)–Pd(1)–Cl(1)	171.62(9)	170.19(6)	168.73(7)
P(1)–Pd(1)–C(1)	93.1(1)	88.57(9)	101.38(7)
Pd(1)–N(1)–C(8)	118.9(2)	120.5(2)	112.1(2)
Pd(1)–N(1)–C(8)–C(7)	–8.94	2.72	
N(1)–C(8)–C(7)–C(2)	–29.07	–30.28	
C(8)–C(7)–C(2)–C(1)	6.41	–6.86	
C(7)–C(2)–C(1)–Pd(1)	47.88	60.11	
C(2)–C(1)–Pd(1)–N(1)	–61.32	–62.74	
C(1)–Pd(1)–N(1)–C(8)	47.31	39.23	

complexes **5a** and **5b** suggests the formation of a single trialkylphosphine-substituted isomer out of the two possibilities. The benzylic methylene unit *cis* to the phosphorus in **5a** and **5b** was established from the ^1H NMR spectra, where the methylene group bound to the palladium appears as a doublet with a coupling constant $J_{\text{P-H}} \approx 2\text{--}5$ Hz. This value is in the typical range reported for the *cis*-arrangement of the methylene group and phosphine in related species.^{10a,e} Even though both complexes could be easily characterized by means of spectral and elemental analysis, the crystal structure of **5b** was determined to confirm the details. Figure 2 displays the ORTEP drawing of **5b**, and selected bond distances and bond angles are collected in Table 2. A method similar to that for the preparation of **5a,b** was employed to produce the triphenylphosphine adduct **5c,d**. Both spectroscopic and crystallographic analyses of **5d** verify the structure of **5c,d** (Figure 3).



Both **5b** and **5d** show the square planar arrangement around the metal center with the phosphine and imine

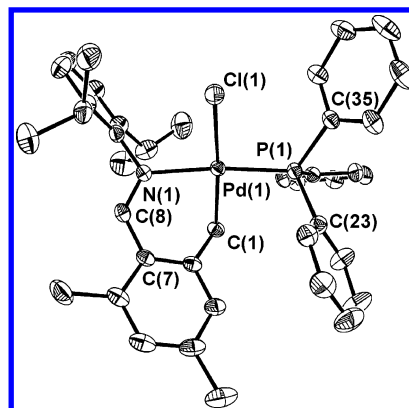


Figure 3. ORTEP plot of **5d** (30% probability ellipsoids).

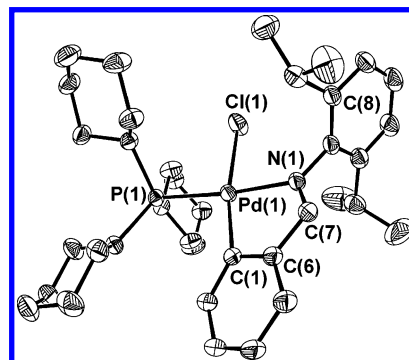


Figure 4. Molecular structure of **7** (30% probability ellipsoids).

donor *trans* to each other, presumably due to the *trans* influence of these donors. All bond distances and bond angles lie in the normal range of closely related complexes such as $[\text{Pd}(\text{TFA})(\kappa^2\text{N},\text{C}-\text{C}_6\text{H}_4\text{CH}_2\text{NMe}_2)(\text{PCy}_3)]$ reported by Bedford and co-workers.^{7a} When two structures are compared, it is found that all bond distances and bond angles around the metal center are very close except the bond angle of P(1)–Pd(1)–C(1) and the bond length of Pd–P. The P(1)–Pd(1)–C(1) in **5b** [93.1(1)°] is considerably larger than that of the analogous angle in **5d** [88.57(9)°], reflecting a result of steric relief between the carbon donor and tricyclohexylphosphine. The Pd–P bond distance of **5b** [2.2683(7) Å] is slightly longer than that of **5d** [2.2388(7) Å], which is due to the contribution of palladium back-donation to the π -acidic nature of triphenylphosphine. The conformation of chelating rings associated with the metal center in **5b** and **5d** remains a twist half-chair form, as illustrated by the torsional angles along the ring consisting of Pd(1)–N(1)–C(8)–C(7)–C(2)–C(1) (Table 2).

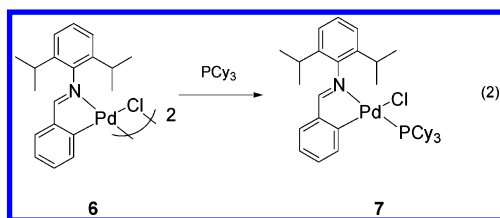
Another phosphine-substituted palladacycle, **7**, with a five-membered chelating ring was prepared via the substitution of **6** with tricyclohexylphosphine (eq 2). The ^{31}P NMR chemical shift of **7** appeared at 43.1 ppm, which is in the typical range of trialkylphosphine palladium complexes.^{7a,c} The crystal structure of **7** was determined (Figure 4). As expected, the molecular structure of **7** shows a square planar geometry around the metal center. Out of four coordination sites, two are occupied by an N–C ligand in *cis*-fashion and the other two by phosphine and chloride with nitrogen and phosphorus donors in *trans* arrangement. The bond lengths between the Pd and C, N, P, and chloride donor atoms for **7** (Table 2) are within the range and compa-

Table 3. Results of Coupling Reactions Catalyzed by Palladacycles^a

entry	cat. (Pd mmol)	ArBr	[ArBr]/[Pd]	atm	t (h)	conv ^b (%)
1	3a (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	100
2	3a (2×10^{-5})	<i>p</i> -MeOC ₆ H ₄ Br	100 000	air	10	91
3	3a (2×10^{-6})	<i>p</i> -MeOC ₆ H ₄ Br	1 000 000	air	18	62
4	3a (2×10^{-3})	<i>p</i> -MeOC ₆ H ₄ Br	1000	air	20 ^c	100
5	3b (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	100
6 ^d	3b (1×10^{-4}) ^e	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	75
7 ^e	3b (1×10^{-4}) ^f	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	28
8 ^f	3b (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	100
9	3b (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	64
10	3b (2×10^{-5})	<i>p</i> -MeOC ₆ H ₄ Br	100 000	air	10	86
11	3b (2×10^{-6})	<i>p</i> -MeOC ₆ H ₄ Br	1 000 000	air	20	57
12	3b (2×10^{-3})	<i>p</i> -MeOC ₆ H ₄ Br	1000	air	14 ^c	100
13	5c (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	85
14	5d (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	95
15	5a (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	100
16	5b (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	92
17	6 (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	86
18	6 (2×10^{-6})	<i>p</i> -MeOC ₆ H ₄ Br	1 000 000	air	16	26
19	7 (1×10^{-4})	<i>p</i> -MeOC ₆ H ₄ Br	20 000	air	3	83
20	3b (2×10^{-5})	<i>p</i> -MeCOC ₆ H ₄ Br	100 000	air	5	77
21	3b (2×10^{-6})	<i>p</i> -MeCOC ₆ H ₄ Br	1 000 000	air	3	40
22	3b (2×10^{-6})	<i>p</i> -MeCOC ₆ H ₄ Br	1 000 000	air	20	100
23	3a (2×10^{-6})	<i>p</i> -MeCOC ₆ H ₄ Br	1 000 000	air	20	89
24	5a (2×10^{-6})	<i>p</i> -MeCOC ₆ H ₄ Br	1 000 000	air	20	100
25	5b (2×10^{-6})	<i>p</i> -MeCOC ₆ H ₄ Br	1 000 000	air	20	100
26 ^f	3b (1×10^{-4})	<i>p</i> -MeCOC ₆ H ₄ Br	20 000	air	3	100
27	3a (1×10^{-3})	2,4,6-tri-MeC ₆ H ₂ Br	2000	air	3	76
28	3b (1×10^{-3})	2,4,6-tri-MeC ₆ H ₂ Br	2000	air	3	69
29	5b (1×10^{-3})	<i>p</i> -MeOC ₆ H ₄ Cl	2000	air	17	72
30 ^f	3a (2×10^{-4})	<i>p</i> -MeCOC ₆ H ₄ Cl	10 000	air	1	96
31	5b (2×10^{-5})	<i>p</i> -MeCOC ₆ H ₄ Cl	10 000	air	3	80

^a Reaction conditions: ArX (2 mmol), PhB(OH)₂ (3 mmol), catalysts, K₂CO₃ (4 mmol), EtOH (5 mL), reflux. ^b Determined by ¹H NMR spectroscopy based on ArX. ^c At 30 °C. ^d In DMF. ^e In CH₃OH. ^f In water and 0.5 equiv of TBAB (tetrabutylammonium bromide) was added.

table with complex **5b**. However, it is noticed that all bond angles around the metal center in **7** are much more derivated from square planar than in **5b**, due to the five-membered chelate ring.



Suzuki–Miyaura Coupling Reaction. All palladacyclic complexes prepared in this work were subjected to an evaluation of their catalytic activities on the coupling reaction of phenylboronic acid with aryl halides. In a typical experiment for the reaction, aryl bromide, phenylboronic acid, and K₂CO₃ in a ratio of 1:1.5:2 were placed in a flask, followed by the addition of the solvent and the catalyst. In all instances the solvent was used as obtained commercially without further purification, while deionized water was used for aqueous systems. The organic product was isolated by extraction and then analyzed by ¹H NMR spectroscopy.

The initial screen on solvents was performed using **3b** as the catalyst precursor (Table 3, entries 5–8). It appeared that the protic solvents including water gave much better results than any other organic solvents.⁹ This coupling reaction running in DMF only provided a moderate yield as compared with alcoholic solvents. This result is quite similar to that reported by Nolan and co-workers using a palladacycle carbene complex

in the coupling of aryl chloride with phenylboronic acid.^{3h} It is noticed that the reactions can be carried out under an atmosphere of air. Therefore, the activities of various palladium complexes were screened using ethanol as the solvent under aerobic conditions. Results are summarized in Table 3.

As the molar ratio of [substrate]/[Pd] remains ~20 000, all palladacyclic catalysts show good catalytic activities in Suzuki coupling. Among them, complexes **3a** and **3b** ought to be the best ones, even better than the phosphine-substituted complexes **5a–d**. For **3a** and **3b** it appears that the steric hindrance of the ligand has less influence on the catalysis. However, the yield drops dramatically within a reasonable period of reaction time (~20 h) when the molar ratio of [substrate]/[Pd] increases up to 10⁶ (entries 11, 18), indicating a concentration limitation of these catalysts. Among these studies, we found that the palladacycle with a six-membered chelate ring, **5a–d**, were slightly better than those with five-membered rings, **6**. This observation also applies for the phosphine-substituted palladacycles; That is, the catalytic activities of **5a–d** are better than that of **7** (entries 13–16 versus 19). The catalytic activities of **3a,b** are generally as good as those for **5a–d**, indicating that phosphine ligands are not necessary in this reaction. It is worthy to mention that the coupling reaction can be carried out at room temperature (entry 4). A quantitative conversion was observed for the coupling of *p*-bromoanisole with phenylboronic acid catalyzed by **3b** in water with the presence of tetraalkylammonium salt (entries 8, 26).

As for the activated substrate (*p*-acetylphenyl bromide), the turnover number can reach 10⁶ for **3b**, **5a**,

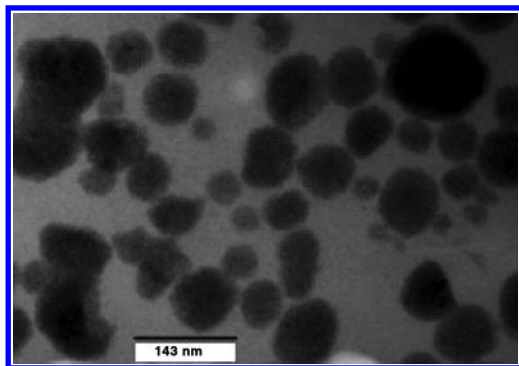


Figure 5. TEM micrograph of palladium nanoparticles generated from the coupling reaction catalyzed by **3b**.

and **5b** (entries 22–25). In addition, complexes **3a** and **3b** catalyzed the coupling of the sterically bulky substrate such as 2,4,6-(Me)₃C₆H₂Br with phenylboronic acid in good conversions (entries 27, 28).

Under similar reaction conditions, the catalytic activities of these palladacycles toward aryl chloride appear to be lower than the bromo substrates (entry 29). However, the conversion can be improved by carrying out the reaction in water medium and in the presence of tetrabutylammonium salt (entries 30, 31).⁹ Overall, the palladium complexes prepared in this work behave with high catalytic activity in the Suzuki–Miyaura coupling reaction.

Mechanistic Pathway of Catalysis. To probe the reaction pathway, the course of the coupling reaction was monitored by taking samples and analyzing them by ¹H NMR and TEM. First, an aliquot from the reaction of **3b** with phenylboronic acid (2 equiv) in the presence of K₂CO₃ and ethanol was examined by transmission electron microscopy, showing the formation of palladium particles with a diameter in the range 50–60 nm (Figure 5). Upon addition of *p*-bromoanisole to the above solution, the coupling reaction proceeded smoothly to yield *p*-methoxybiphenyl quantitatively, suggesting that the palladium nanoparticles might be the active catalyst for the reaction.

In further studies, we found that complex **3b** readily reacted with ethanol under basic conditions to generate acetaldehyde and the free ligand **4b** accompanied with palladium nanoparticles, but aggregate to produce palladium black without the presence of other substrates. The formation of acetaldehyde is presumably due to the substitution of chloride ligands by ethoxide under basic conditions followed by β -elimination.^{3h} Addition of phenylboronic acid to the above solution yielded benzene and biphenyl immediately (Table 4). On the other hand, *p*-acetophenyl bromide was converted into acetophenone and 4,4'-bisacetobiphenyl, as evidenced by ¹H NMR spectroscopy. However, the reaction of *p*-acetophenyl chloride was much slower than that of bromide (Table 4, entry 2). The production of biphenyl derivatives and the reduced compounds is presumably via the addition of phenylboronic acid or *p*-acetophenyl bromide to the palladium nanoparticles followed by reductive elimination (Scheme 2). We also found that the decomposition rate of aryl bromide is much slower than that of arylboronic acid, revealing that the oxidative addition of aryl halide to the metal is the rate-limiting step in the cross-coupling reaction. This also explains that the

use of excess arylboronic acid in the Suzuki–Miyaura reactions is generally required for better conversions. It should be mentioned that the reduction of aryl halides or arylboronic acid proceeded superiorly than that of the homo-coupling reaction.

The decomposition rate of palladacycles in basic ethanol solution follows the order **6** > **3b** > **5b** > **5d** (Table 5). Complex **6** appeared to be slower than **3b**, but aggregated into larger nanoparticles (100–200 nm), which explained the activity difference between **6** and **3b** due to the size effect.¹¹ As for the phosphine-substituted complexes **5b** and **5d**, the decomposition was much slower than that of **3b** presumably due to the stabilization of the coordinating phosphine. Another finding is that the surfactant-stabilized palladium nanoparticles appear to be less active than those generated in situ. In the presence of tetrabutylammonium bromide, palladium nanoparticles was obtained from the decomposition of complex **3b** in ethanol. The resulting palladium nanoparticles were used as the catalyst for the coupling reaction of PhB(OH)₂ and MeOC₆H₄Br to give the desired product but in a slower conversion rate by ca. 20%.

It has been demonstrated that palladium nanoparticles can catalyze the C–C bond coupling reaction,¹² particularly under Jeffery condition for Heck reaction.¹³ Unlike the palladium clusters stabilized by tetraalkylammonium salt, the nanoparticles generated from the reduction of palladacycles are presumably surrounded by the imine ligands, which tend to aggregate with the precipitation of palladium black. This reveals the weak stabilizing effect of these ligands toward nanoparticles. On the other hand, the loose protection makes these palladium particles highly active toward the substrates.¹²

Furthermore, the low concentrations of palladium complexes may also prevent the rapid formation of palladium black and the formation of larger nanoparticles, suggesting the extreme activity of metal clusters. However, the conversion decreases when the ratio [substrate]/[Pd] is higher than 10⁶ (Table 3 entries 3, 11), showing a limitation of these nanoparticles in catalysis.

Summary

We have synthesized and characterized a new series of air-stable palladacycles. These palladium complexes were successfully applied in the Suzuki–Miyaura coupling reaction and had good catalytic activities, for example, the turnover frequency of up to 10⁷ mol/mol-(Pd)·h for **3b** in the coupling of *p*-bromoanisole or *p*-bromoacetophenone with phenylboronic acid. Notable were the reaction conditions employed, i.e., the moisture and air insensitivity. Several observations confirm that the Suzuki–Miyaura coupling catalyzed by palladacyclic complexes is via the palladium nanoparticles. Reactions of *p*-bromoacetophenone or arylboronic acid individually on the nanoparticles leading to homo-coupled and reduced products were investigated, but the cross-

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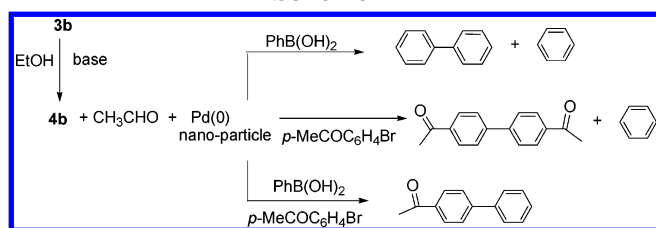
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Table 4. Decomposition of Aryl Halides or Arylboronic Acids on Palladium Nanoparticles Generated in Situ from Palladacyclic Complexes^a

entry	substrate	complex	conditions	time	conversion	reduced product:coupling product (ratio) ^b
1	<i>p</i> -MeCOC ₆ H ₄ Br	3b	reflux	2 h	100%	MeCOC ₆ H ₅ (63):(p-MeCOC ₆ H ₄) ₂ (37)
2	<i>p</i> -MeCOC ₆ H ₄ Cl	3b	reflux	10 h	46%	MeCOC ₆ H ₅ (67):(p-MeCOC ₆ H ₄) ₂ (33)
3	<i>p</i> -MeCOC ₆ H ₄ Br	6	reflux	2 h	77%	MeCOC ₆ H ₅ (80):(p-MeCOC ₆ H ₄) ₂ (20)
4	<i>p</i> -MeCOC ₆ H ₄ Br	5b	reflux	2 h	100%	MeCOC ₆ H ₅ (100):(p-MeCOC ₆ H ₄) ₂ (0)
5	PhB(OH) ₂	3b	reflux	1 h	100%	C ₆ H ₆ (98):C ₆ H ₅ -C ₆ H ₅ (2)
6	PhB(OH) ₂	3b	40 °C	8 h	100%	C ₆ H ₆ (91):C ₆ H ₅ -C ₆ H ₅ (9)
7	<i>p</i> -MeOC ₆ H ₄ B(OH) ₂	3b	40 °C	8 h	100%	MeOC ₆ H ₅ (98):(p-MeOC ₆ H ₄) ₂ (2)
8	<i>p</i> -MeC ₆ H ₄ B(OH) ₂	3b	40 °C	8 h	100%	MeC ₆ H ₅ (98):(p-MeC ₆ H ₄) ₂ (2)

^a Conditions: Aryl halides (1 mmol) [or ArB(OH)₂ (0.5 mmol)], K₂CO₃ (2 mmol), and Pd complex (mmol) in EtOH (2 mL). ^b Ratio is determined by ¹H NMR integration and is given in parentheses.

Scheme 2**Table 5. Relative Decomposition Rate of Palladacycle in Basic Etanol Solution^a**

	3a	3b	5b	5d
% of decomposition ^b	40%	70%	14%	trace

^a A solution of complex (2 × 10⁻² mmol) and K₂CO₃ (2 × 10⁻¹ mmol) in ethanol (2 mL) was stirred at 40 °C for 30 min. ^b Percent of decomposition was determined by ¹H NMR spectroscopy.

coupled ones were obtained when the aryl halide and arylboronic acid were present. However, the pathway for the cross-coupling and reduction processes on these particles remains unclear. More work is necessary in order to work out the nature of these nanoparticles.

Experimental Section

General Information. All reactions, manipulations, and purifications steps were performed under a dry nitrogen atmosphere. Tetrahydrofuran was distilled under nitrogen from sodium benzophenone ketyl. Dichloromethane and acetonitrile were dried with CaH₂ and distilled under nitrogen. Other chemicals and solvents were of analytical grade and were used as received unless otherwise stated.

Nuclear magnetic resonance spectra were recorded in CDCl₃ on either a Bruker AM-300 or AVANCE-400 spectrometer. Chemical shifts are given in parts per million relative to Me₄S for ¹H and relative to 85% H₃PO₄ for ³¹P NMR. Infrared spectra were measured on a Perkin-Elmer 983G spectrometer (Series-II) as KBr pellets, unless otherwise noted. Preparation of **6** followed the procedures published previously.¹⁴

General Procedure for Preparation of Schiff Base Ligands 4a,b. A methanol (15 mL) solution of the substituted benzylaldehyde (10 mmol) and the substituted aniline (10 mmol) was added into a single-necked round-bottomed flask equipped with a condenser. The reaction mixture was stirred at reflux temperature for 24 h. After the completion of the reaction, the yellow crystal product was precipitated upon cooling. The solid product was filtered, washed with the precooled hexane, and dried, whereas the liquid compound **1b** was purified by distilling off the starting aldehyde and chromatographed on silica gel with elution of hexane.

(2,6-Dimethylphenyl)(2,4,6-trimethylbenzylidene)-amine, 4a: yellow liquid (90%); IR (KBr) 1625 cm⁻¹ (ν_{C=N}); ¹H NMR (CDCl₃, 400 MHz) δ 8.68 (s, 1H, -HC=N), 7.15 (s, 2H, Ar H), 7.04–6.95 (m, 3H, Ar H), 2.66 (s, 6H, -Me), 2.40 (s, 3H, -Me), 2.28 (s, 6H, -Me); ¹³C NMR δ 163.0, 152.6, 139.9, 138.8, 130.0, 129.9, 128.1, 126.7, 123.4, 21.3, 21.1, 18.8; FAB *m/z* 251.1 (M⁺). Anal. Calcd for C₁₈H₂₁N: C, 86.01; H, 8.42; N, 5.57. Found: C, 86.11; H, 8.65; N, 5.55.

(2,6-Diisopropylphenyl)(2,4,6-trimethylbenzylidene)-amine, 4b: white solids (91%); mp 99–101 °C; IR (KBr) 1627 cm⁻¹ (ν_{C=N}); ¹H NMR (CDCl₃, 400 MHz) δ 8.64 (s, 1H, -HC=N), 7.24–7.22 (m, 2H, Ar H), 7.17–7.15 (m, 1H, Ar H), 7.01 (s, 1H, Ar H); 3.12 (sept., *J* = 6.8 Hz, 2H, -CH), 2.65 (s, 6H, -Me), 2.39 (s, 3H, -Me), 1.25 (d, *J* = 6.8 Hz, 12H, -Me); ¹³C NMR δ 162.3, 150.5, 140.1, 139.2, 137.4, 130.2, 129.7, 123.8, 122.9, 27.8, 23.7, 21.5, 21.1; FAB *m/z* 307.2 (M⁺). Anal. Calcd for C₂₂H₂₉N: C, 85.94; H, 9.51; N, 4.56. Found: C, 86.24; H, 9.64; N, 4.57.

General Procedure for Preparation of Palladacycle Complexes 3a,b. To a round-bottomed flask with a stir bar was placed palladium dichloride (118 mg, 0.67 mmol) under nitrogen. Predried acetonitrile (10 mL) was added, and the resulting mixture was stirred at room temperature. The mixture turned yellow immediately. After stirring for 2 days, the solvent was removed to dryness. Then ligand (0.71 mmol), excess sodium acetate (110 mg, 10.4 mmol), and tetrahydrofuran (15 mL) were added. The reaction mixture was stirred at room temperature for another 38 h. After removal of solvents, dichloromethane (20 mL) was added and the solution was filtered through Celite. The filtrate was concentrated, and the residue was washed with hexane (10 mL × 3) to give the desired product.

Complex 3a: brown solids (80%); mp 252 °C (dec); IR (KBr) 1600 cm⁻¹ (ν_{C=N}); ¹H NMR (CDCl₃, 400 MHz) δ 7.77 (s, 1H, -HC=N), 7.07–7.01 (m, 4H, Ar H), 6.80 (s, 1H, Ar H), 3.19 (s, 2H, -H₂C-Pd), 2.41 (s, 6H, -Me), 2.28 (s, 3H, -Me), 2.27 (s, 3H, -Me); ¹³C NMR δ 163.1, 150.2, 145.8, 143.8, 139.5, 130.9, 130.0, 128.2, 127.6, 126.2, 125.3, 23.3, 21.4, 19.5, 18.9; FAB *m/z* 749.0 (M⁺ - Cl⁻). Anal. Calcd for C₃₆H₄₀Cl₂N₂Pd₂: C, 55.12; H, 5.14; N, 3.57. Found: C, 55.34; H, 5.04; N, 3.44.

Complex 3b: orange solids (90%); mp 270 °C (dec); IR (KBr) 1599 cm⁻¹ (ν_{C=N}); ¹H NMR (CDCl₃, 400 MHz) δ 7.81 (s, 1H, -HC=N), 7.23 (t, *J* = 7.3 Hz, 1H, Ar H), 7.14 (d, *J* = 7.3 Hz, 2H, Ar H), 6.98 (s, 1H, Ar H), 6.82 (s, 1H, Ar H), 3.48–3.39 (m, 2H, -CH), 3.24 (s, 2H, -H₂C-Pd), 2.30 (s, 3H, -Me), 2.29 (s, 3H, -Me), 1.47 (d, *J* = 6.7 Hz, 6H, -Me), 1.11 (d, *J* = 6.8 Hz, 6H, -Me); ¹³C NMR δ 162.7, 148.0, 146.5, 143.7, 140.7, 139.8, 130.6, 127.6, 126.8, 125.4, 123.3, 28.2, 24.8, 23.2, 22.5, 21.4, 19.1; FAB *m/z* 861.1 (M⁺ - Cl⁻). Anal. Calcd for C₄₄H₅₆Cl₂N₂Pd₂: C, 58.94; H, 6.29; N, 3.12. Found: C, 58.89; H, 6.35; N, 3.08.

General Procedure for Preparation of Phosphine-Substituted Palladacycle Complexes. To a degassed flask loaded with the palladacycle dimer complex (1 mmol) and dichloromethane (5 mL) was added phosphine (2 mmol). Upon stirring at room temperature for 2 h, the solvent of the reaction

Table 6. Crystallographic Data of 3a, 5b, 5d, and 7

	3a	5b	5d	7
formula	C ₃₆ H ₄₀ Cl ₂ N ₂ Pd ₂	C ₄₀ H ₆₁ ClNPPd	C ₄₀ H ₄₃ ClNPPd	C ₃₇ H ₅₅ ClNPPd
fw	784.40	728.72	710.57	686.64
cryst syst	monoclinic	orthorhombic	monoclinic	monoclinic
space group	<i>P2₁/c</i>	<i>Pbca</i>	<i>P2₁/c</i>	<i>P2₁/n</i>
<i>a</i> , Å	12.7010(1)	19.3320(2)	17.0000(3)	11.7670(1)
<i>b</i> , Å	8.794(1)	17.3720(2)	11.6040(2)	17.0560(1)
<i>c</i> , Å	15.7000(2)	22.6590(2)	18.4260(3)	18.4730(2)
α, deg	90	90	90	90
β, deg	103.978(1)	90	104.528(1)	105.267(1)
γ, deg	90	90	90	90
<i>V</i> , Å ³	1701.65(3)	7609.7(1)	3518.6(1)	3576.65(5)
<i>Z</i>	2	8	4	4
<i>D</i> _{calcd} , Mg/m ³	1.531	1.272	1.341	1.275
<i>F</i> (000)	792	3088	1472	1448
cryst size, mm	0.25 × 0.20 × 0.15	0.30 × 0.25 × 0.15	0.25 × 0.20 × 0.15	0.30 × 0.20 × 0.15
θ range	1.65–25.00	1.80–27.48	2.09–27.47	1.65–27.47
no. of reflns collected	9728	50 044	25 239	26 619
no. of indep reflns	3006 (<i>R</i> _{int} = 0.0286)	8719 (<i>R</i> _{int} = 0.0467)	8026 (<i>R</i> _{int} = 0.0367)	8175 (<i>R</i> _{int} = 0.0336)
<i>R</i> [<i>I</i> > 2σ(<i>I</i>)]	<i>R</i> ₁ = 0.0302	<i>R</i> ₁ = 0.0474	<i>R</i> ₁ = 0.0404	<i>R</i> ₁ = 0.0402
	<i>wR</i> ₂ = 0.0803	<i>wR</i> ₂ = 0.1219	<i>wR</i> ₂ = 0.0973	<i>wR</i> ₂ = 0.1069
goodness-of-fit on <i>F</i> ²	1.187	1.111	1.173	1.122

mixture was evaporated. The residue was crystallized in dichloromethane/ether to give the desired complex in crystalline solids.

Complex 5a: yellow solids (99%); IR (KBr) 1607 cm⁻¹ ($\nu_{C=N}$); ¹H NMR (CDCl₃, 300 MHz) δ 8.06 (d, *J* = 11.0 Hz, 1H, -HC=N), 7.07–7.00 (m, 3H, Ar *H*), 6.89 (s, 1H, Ar *H*), 6.76 (s, 1H, Ar *H*), 2.90 (s, 2H, -H₂C-Pd), 2.53 (s, 6H, -Me), 2.34 (s, 3H, -Me), 2.36–2.26 (m, 3H, PCy₃), 1.94–1.42 (m, 21H, PCy₃), 1.28–1.22 (m, 9H, PCy₃); ¹³C NMR (CDCl₃, 100 MHz) δ 163.3, 151.3, 146.7, 142.5, 139.5, 130.9, 139.9, 127.9, 127.4, 126.2, 125.6, 33.4 (d, *J*_{C-P} = 22 Hz), 29.8, 27.7 (d, *J*_{C-P} = 10.9 Hz), 26.6, 22.4, 21.3, 20.4, 18.9; ³¹P NMR (CDCl₃, 161.9 MHz) 39.8. Anal. Calcd for C₃₆H₃₅ClNPPd: C, 64.28; H, 7.94; N, 2.08. Found: C, 63.97; H, 7.65; N, 1.80.

Complex 5b: yellow solids (99%); IR (KBr) 1603 cm⁻¹ ($\nu_{C=N}$); ¹H NMR (CDCl₃, 400 MHz) δ 8.03 (d, *J*_{H-P} = 10.9 Hz, -HC=N), 7.20 (bs, 3H, Ar *H*), 6.92 (s, 1H, Ar *H*), 6.80 (s, 1H, Ar *H*), 3.85–3.78 (m, 2H, -CH), 2.92 (d, *J* = 2.4 Hz, 2H, -H₂C-Pd), 2.37 (s, 3H, -Me), 2.37–2.30 (m, 3H, PCy₃), 2.30 (s, 3H, -Me), 1.96–1.94 (m, 6H, PCy₃), 1.71–1.50 (m, 15H, PCy₃), 1.46 (d, *J* = 6.5 Hz, 6H, -Me), 1.30–1.26 (m, 9H, PCy₃), 1.21 (d, *J* = 6.6 Hz, 6H, -Me); ¹³C NMR (CDCl₃, 100 MHz) δ 163.1, 148.7, 146.8, 142.5, 140.8, 139.6, 131.1, 127.4, 126.5, 126.2, 123.1, 33.6 (d, *J*_{C-P} = 22.4 Hz), 29.8, 28.5, 27.6 (d, *J*_{C-P} = 10.9 Hz), 26.4, 25.2, 23.6, 21.6, 21.3, 19.4; ³¹P NMR (CDCl₃, 161.9 MHz) 40.3. Anal. Calcd for C₄₀H₆₁ClNPPd: C, 65.92; H, 8.44; N, 1.92. Found: C, 65.86; H, 8.59; N, 1.83.

Complex 5c: yellow solids (99%); IR (KBr) 1602 cm⁻¹ ($\nu_{C=N}$); ¹H NMR (CDCl₃, 300 MHz) δ 8.09 (d, *J* = 12.3 Hz, 1H, -HC=N), 7.62–7.56 (m, 6H, Ar *H*), 7.42–7.32 (m, 9H, Ar *H*), 7.11–7.04 (m, 3H, Ar *H*), 6.71 (s, 1H, Ar *H*), 5.72 (s, 1H, Ar *H*), 2.81 (d, *J* = 5.1 Hz, 2H, -H₂C-Pd), 2.57 (s, 6H, -Me), 2.33 (s, 3H, -Me), 2.10 (s, 3H, -Me); ³¹P NMR (CDCl₃, 161.9 MHz) δ 36.4. Anal. Calcd for C₃₆H₃₅ClNPPd: C, 66.06; H, 5.39; N, 2.14. Found: C, 66.29; H, 5.39; N, 1.68.

Complex 5d: yellow solids (99%); IR (KBr) 1600 cm⁻¹ ($\nu_{C=N}$); ¹H NMR (CDCl₃, 400 MHz) δ 8.08 (d, *J*_{H-P} = 12 Hz, -HC=N), 7.67–7.60 (m, 6H, Ar *H*), 7.45–7.37 (m, 9H, Ar *H*), 7.24 (bs, 3H, Ar *H*), 6.75 (s, 1H, Ar *H*), 5.73 (s, 1H, Ar *H*), 3.85–3.78 (m, 2H, -CH), 2.83 (d, *J* = 5.1 Hz, 2H, -H₂C-Pd), 2.36 (s, 6H, -Me), 2.14 (s, 3H, -Me), 2.10 (s, 3H, -Me), 1.5 (d, *J* = 6.7 Hz, 6H, -Me), 1.28 (d, *J* = 6.8 Hz, 6H, -Me); ¹³C NMR (CDCl₃, 100 MHz) δ 163.6, 148.0, 144.7 (d, *J*_{C-P} = 2.3 Hz), 142.4, 140.8, 139.3, 134.6 (d, *J*_{C-P} = 11.4 Hz), 131.5, 131.0 (d, *J*_{C-P} = 48.6 Hz), 130.1 (d, *J*_{C-P} = 2.3 Hz), 128.0 (d, *J*_{C-P} = 10.5 Hz), 127.5, 126.5, 126.4, 123.3, 28.6, 27.9, 25.2, 23.7, 21.2, 19.3; ³¹P NMR (CDCl₃, 161.9 MHz) δ 37.7. Anal. Calcd for C₄₀H₄₃ClNPPd: C, 67.61; H, 6.10; N, 1.97. Found: C, 67.79; H, 5.97; N, 1.53.

Complex 7: yellow solids 99%; IR (KBr) 1617 cm⁻¹ ($\nu_{C=N}$); ¹H NMR (CDCl₃, 400 MHz) δ 7.98 (d, *J*_{H-P} = 6.8 Hz, 1H, -HC=N), 7.42 (dd, *J* = 7.5 Hz, *J* = 2.3 Hz, 1H, Ar *H*), 7.37 (dd, *J* = 7.1 Hz, *J* = 1.9 Hz, 1H, Ar *H*), 7.19–7.10 (m, 5H, Ar *H*), 3.34 (hept, *J* = 6.8 Hz, 2H, -CH), 1.30 (d, *J* = 6.8 Hz, 6H, -Me), 1.12 (d, *J* = 6.8 Hz, 6H, -Me); ¹³C NMR (CDCl₃, 100 MHz) δ 176.1 (d, *J*_{C-P} = 3.4 Hz), 159.0, 148.2, 145.6, 141.0, 137.4 (d, *J*_{C-P} = 5.1 Hz), 130.7 (d, *J*_{C-P} = 3.4 Hz), 129.1, 126.3, 124.1, 122.6, 34.1 (d, *J*_{C-P} = 21.7 Hz), 30.4, 28.4, 27.7 (d, *J*_{C-P} = 10.9 Hz), 26.4, 24.6, 22.9; ³¹P NMR (CDCl₃, 161.9 MHz) 43.1. Anal. Calcd for C₃₇H₅₅ClNPPd: C, 64.72; H, 8.07; N, 2.04. Found: C, 64.85; H, 8.04; N, 1.91.

Catalysis. To a 10 mL round-bottom flask equipped with condenser were placed palladium complex (5 × 10⁻³ mmol), aryl halide (2 mmol), K₂CO₃ (3 mmol), phenylboronic acid (3 mmol), and solvent (5 mL). The mixture was heated under reflux for a certain period of time. Upon cooling, the reaction mixture was poured into diluted hydrochloric acid (10%) and extracted with dichloromethane. The extract was dried, concentrated, and analyzed by ¹H NMR spectroscopy. Results of the coupling reactions are listed in Table 3.

Decomposition of Palladacycles in the Presence of Aryl Halides or Arylboronic Acids. To a 10 mL round-bottom flask equipped with condenser was placed palladium complex (1.0 × 10⁻² mmol), K₂CO₃ (2.0 mmol), and aryl halides or arylboronic acid in ethanol (2 mL). The resulting mixture was stirred with heating. An aliquot from the reaction mixture was analyzed by ¹H NMR. The results are summarized in Table 4.

Crystallography. Crystals suitable for X-ray determination were obtained for **3a**, **5b,d**, and **7** by slow diffusion of hexane into a dichloromethane solution at room temperature. Cell parameters were determined by a Siemens SMART CCD diffractometer. Crystal data of these complexes are listed in Table 5. Other crystallographic data are deposited as Supporting Information.

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Supporting Information Available: Complete description of the X-ray crystallographic structure determination of **3a**, **5b,d**, and **7** including tables of crystal data, atomic coordinates, isotropic and anisotropic thermal parameters, and bond distances and angles. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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