

Stack Exchange Strategies for the Synthesis of Covalent Double-Channel Photosystems by Self-Organizing Surface-Initiated Polymerization

Naomi Sakai* and Stefan Matile*

Department of Organic Chemistry, University of Geneva, Geneva, Switzerland

Supporting Information

ABSTRACT: Ring-opening disulfide exchange polymerization has recently been identified as ideal for synthesizing single-channel photosystems by self-organizing surface-initiated polymerization (SOSIP). Here we introduce chemoorthogonal hydrazone exchange chemistry to engineer additional channels into single-channel photosystems. Post-SOSIP stack exchange is shown to provide facile access to complete supramolecular n/p-heterojunction architectures with high activity and freely variable composition, including oriented antiparallel redox gradients. With appropriate templation from the surface, post-SOSIP stack exchange is nearly quantitative.

lthough they are essential for building the materials of the Afuture, general approaches to covalent multichannel architectures on solid surfaces that are both oriented and ordered have remained beyond reach for decades.¹ This is understandable because of the formidable challenges in synthetic organic, supramolecular, and analytical chemistry involved. Current solution-processing methods for building optoelectronic devices naturally lack directionality.² Available approaches to build directly on the surface, such as layer-by-layer assembly, give little axial and no lateral organization.³ The alternative zipper assembly has been successful in building supramolecular n/p-heterojunctions with oriented double-channel gradients,⁴ but the synthetic efforts involved are much too demanding for general use, and covalent capture studies have failed.⁵ Classical polymer brushes seem to stop growing as soon as the involved chemistry becomes more ambitious.⁶

Self-organizing surface-initiated polymerization (SOSIP) has recently been introduced as a promising method that applies lessons from nature to secure facile access to oriented and ordered single-channel architectures on solid surfaces.⁷ However, supramolecular n/p-heterojunction (SHJ) photosystems require the coaxial alignment of two molecular-level pathways that can transport holes and electrons in opposite directions after their separation with light at the maximized n/p interface.^{1,4} Lateral self-sorting during co-SOSIP is currently being explored as a supramolecular approach to SHJs.⁸ This attractive approach is likely to be successful also in practice. However, it is less likely to obtain the experimental evidence to prove the occurrence of perfectly alternate lateral and uniform axial self-sorting during co-SOSIP. To build SOSIP architectures with confirmed coaxial channels at molecular-level resolution, covalent approaches will



Figure 1. The concept of post-SOSIP stack exchange. Chemoorthogonal to the disulfide exchange chemistry used for SOSIP, hydrazone exchange chemistry is used to replace noninterfering and inactive domains (shaded circles in A) with active π -stacks (red boxes in C). For full structures, see Scheme 1.

be unavoidable. In this communication, post-SOSIP stack exchange is introduced to tackle this challenge. In this general strategy, chemoorthogonal hydrazone exchange is employed after SOSIP to drill large pores next to the existing π -stacks and fill them with the complementary π -stacks (Figure 1).

In initiator 1 and propagators 2, all that is needed for SOSIP is preserved (Scheme 1).⁷ Both contain central naphthalenediimides (NDIs).⁹ To direct their assembly into charge-transporting π -stacks, flanking hydrogen-bonded networks were inserted on both sides before the polymerizing subunits. Acyl hydrazones were placed in the self-organizing subunits of propagators 2 to place exchangeable subunits for the installation of additional charge-transporting channels by post-SOSIP stack exchange (Figure 1 and Scheme 1). In initiator 1, the same hydrazones were used to link the NDI templates and diphosphonate "feet" to the central unit.

Initiator 1 and propagators 2 were prepared by reacting hydrazides 3 and 4 with aldehydes 5 and 6, respectively. The synthesis of building blocks 3-5 from commercial starting materials is described in the Supporting Information (SI).¹⁰

Initiator 1 contains four diphosphonate feet for "tetravalent" binding to the ITO surface to firmly preorganize both SOSIP and the exchangeable subunit. The protecting groups on the surface of monolayer 7 (Figure S2 in the SI) were removed with dithiothreitol (DTT) to afford thiolates on the surface of monolayer 8

Received:August 11, 2011Published:October 25, 2011



Scheme 1. Synthesis of Covalent Double-Channel Photosystems by SOSIP and Stack Exchange^a

^{*a*} Conditions: (a) TFA. (b) (1) **1**, DMSO, ITO; (2) 1 h, 120 °C. (c) DTT. (d) **2a**, 1:1 CHCl₃/MeOH, iPr₂NEt. (e) NH₂OH HCl, H₂O, (f) aldehyde, TFA or AcOH. See Figure S1 for the complete chemical structures. Dotted lines from the diphosphonate groups indicate the bonds to the ITO surface. **12b**, **12c**, **14**, and **15** are mixtures of 2,6- and 3,7-regioisomers; polymer structures are speculative generalized structures that are consistent with the experimental results. They naturally contain some defects, and the nature of the presumably oxidized thiol termini (\mathbb{R}^5) at the surface is unknown.

(Scheme 1c). Propagator 2a and a base catalyst were added next to initiate SOSIP (Scheme 1d). Molecular recognition by NDI stacking and hydrogen bonding was expected to place the strained disulfides from asparagusic acid in propagator 2a on top of the thiolates of 8 to initiate ring-opening disulfide exchange^{11,12} polymerization and produce SOSIP photosystem 9.

SOSIP was followed by the increase in the 470 nm absorption of the yellow NDI from the propagator **2a** bound to the ITO surface. Under the optimized conditions, the onset of SOSIP occurred at a critical concentration $c_{\text{SOSIP}} \approx 7 \text{ mM}$ (Figure S3A). Competing polymerization in solution was not observed under these conditions. Replacement of the benzaldehyde hydrazones in propagator **2a** by acetone hydrazones made the eventual SOSIP indistinguishable from polymerization in solution (Scheme S2 and Figure S3C). The same was true for initiator **10** containing only two diphosphonate feet and primary amides in place of the templating NDI diphosphonates (Figure S3A, × symbols).

Stack exchange within the resulting SOSIP photosystem 9 was initiated by hydrazone—oxime exchange using excess hydroxylamine (Scheme 1e). The chemoorthogonality of disulfide and hydrazone exchange has been demonstrated extensively.^{11,13} The hydrazide-rich pores drilled into the resulting architecture 11 were then filled by reversible covalent capture of various alde-hydes (e.g., 12a-e) to give hydrazones 13-17 (Scheme 1f).

Post-SOSIP stack exchange was monitored by the change in absorption of the coated ITO electrode (Figure 2). On the functional level, post-SOSIP stack exchange was characterized by photocurrent generation upon irradiation of the photosystem



Figure 2. (A) Absorption spectra of photosystem 9 (dotted) and 13 (solid). (B) Molar ratio of colorless HH-NDIs from 12a and yellow OO-NDIs from 2a as a function of the thickness of photosystem 13 (\bullet , $A_{470} = 0.1$ au \approx 30 nm or 110 π -stacks).¹⁰ For comparison, data for the control system 13' made from template-free initiator 10 are also shown (O). (C) Changes in absorption spectra in response to stack exchange from 9 (black solid line) to 16 (blue dashed line), 13 (purple dotted line), and 17 (red solid line).

used as a working electrode in the presence of mobile triethanolamine hole acceptors, a Pt wire as a counter electrode, and Ag/AgCl as a reference electrode (Figure 3). The unchanged photocurrent and absorption after several on-off cycles of irradiation confirmed the stability of the photosystems under these conditions.

The exchange of the benzaldehyde hydrazones in photosystem 9 with the NDI stacks in photosystem 13 was correctly reflected by the appearance of a strong NDI absorption at 380 nm (Figure 2A). According to their extinction coefficients



Figure 3. (A) Photocurrent generated by photosystems 9, 13, and 14 upon irradiation by a solar simulator (66 mW/cm²). (B) Photocurrent action spectra of photosystems 9 (\diamond), 13 (\blacktriangle), and 14 (\bigcirc). (C, D) HOMO (bold) and LUMO (dotted) energy levels for photosystems (C) 13 and (D) 14 (in eV vs vacuum).

in solution, the molar ratio of unsubstituted and yellow NDIs could be calculated from changes in A_{380} and A_{470} (\bullet data in Figure 2B). The obtained value of 1.7 \pm 0.2 indicated that subunit exchange from photosystem 9 to photosystem 13 was almost quantitative.

Control experiments with SOSIP photosystems grown from initiator 10 revealed very little binding of NDI aldehyde 12a (\bigcirc data in Figure 2B; also see Figure S7). Quite remarkably, the yield of templated post-SOSIP stack exchange changed little with increasing thickness of the photosystem, whereas that of untemplated post-SOSIP stack exchange became even poorer (Figure 2B). This sensitivity to the structure of the initiator demonstrates the central importance of the NDI diphosphonate hydrazones in initiator 1 as a template on the surface to secure the space needed for SOSIP of bulky propagators such as 2a as well as for operational post-SOSIP stack exchange. These results are in excellent agreement with the formation of ordered ladderphanelike¹⁴ structures by SOSIP (Scheme 1).

Consistent with the well-known chemoorthogonality of disulfide and hydrazone exchange reactions, post-SOSIP stack exchange could be repeated many times without significant loss of polymeric material (Figure 2C and Figure S6).

Post-SOSIP stack exchange from photosystem **9** to photosystem **13** caused a large increase in photocurrent generation (Figure 3A). This improved performance was as expected for a formal SHJ architecture, where coaxial stacks of yellow and unsubstituted NDIs can transport holes and electrons, respectively, and photoinduced charge separation can occur at the maximized molecular-level interface between the stacks.^{1,4} The increased activity of yellow NDIs found at 470 nm in the photocurrent action spectra upon exchange from **9** to **13** (Figure 3B) is in agreement with this interpretation.

The exchange of stacks of unsubstituted NDIs in photosystem 13 with stacks of red NDIs in photosystem 14 nearly doubled the photocurrent generation (Figure 3A). In double-channel SHJ 14, the red NDI stacks were expected to serve as p channels, whereas the yellow stacks should transport electrons (Figures 3D and 4A). Post-SOSIP stack exchange from SHJ 13 to 14 should occur with umpolung of the yellow NDI stacks from p to n transporters. This "ambipolar" behavior of yellow NDI stacks is consistent with their ability to undergo symmetry-breaking photoinduced charge separation (i.e., disproportionation into oxidized and reduced radicals) in response to irradiation with light.⁷



Figure 4. Schematic structures of (A) SHJ 14, (B) three-component OMARG SHJ 18, (C) four-component OMARG SHJ 19, and (D) anti-OMARG SHJ 20 (a–e show the sequence of construction; see Schemes S6 and S7 for details). (E) HOMO and LUMO levels for 19. (F) Dependence of the short-circuit current density (I_{SC}) on the light intensity (I) for 19 (\bullet), 14 (\bigcirc), 18 (\blacktriangle), and 20 (\bigtriangleup). Bimolecular charge recombination efficiencies η_{BR} are also shown.

Post-SOSIP stack exchange with other aldehydes, including more electron-rich NDIs 12c, known hole conductors 12d and 12e, and electron-poor cyano NDIs gave clearly less active photosystems (Figures S9 and S10).

The three-component photosystem 18 was constructed from SHJ photosystem 13 with a central yellow π -stack of length n (Figures 2 and 4B). About half of the adjacent π -stacks formed by the unsubstitued NDIs 12a was removed with hydroxylamine. The resulting pores were filled with red NDI aldehyde 12b to give red π -stacks (from 12b) of length l on top of colorless π -stacks (from 12a) of length m (Figure 4B). All processes were followed quantitatively by absorption spectroscopy ($l \approx m, l + m \approx n$; Figure 4B and Figure S8A).¹⁰

The four-component system 19 was produced analogously from 8 by successive SOSIP of 2b and 2a (with $n \approx m$) followed first by complete stack exchange with 12b and then by partial stack exchange with 12c (with $l \approx 20k$ because of the poor intrinsic activity of the ON-NDIs in 12c; Figure 4C and Figure S8B).¹⁰ 19 possesses clearly separated electron- and holeconducting pathways, both equipped with redox gradients (Figure 4E). This system resembles the previously reported SHJ with oriented multicolored antiparallel redox gradients (OMARG), which was shown to generate more photocurrent than the one without gradients because of the minimized charge recombination.⁴ On the other hand, the three-component system 18 is a new type of OMARG in which, as a result of the ambipolar nature of the yellow NDI stacks, two conducting pathways partially overlap. The misaligned three-component anti-OMARG SHJ **20**, in which the position of the red and unsubstituted NDIs in **18** is switched, was also assembled by partial stack exchange from **14** to assess the significance of this novel construct ($l \approx m$; Figure 4D and Figure S8C).

The effect of redox gradients was clearly seen in the dependence of the short-circuit current density J_{SC} on the light intensity I (Figure 4F). Their relationship is often described by the equation $J_{\rm SC} \propto I^{\alpha}$. The exponent α has been shown to be related to the bimolecular recombination loss efficiency η_{BR} (the fraction of charges lost by recombination) by the equation η_{BR} = $\alpha^{-1} - 1.^{15}$ As expected, the lowest $\eta_{\rm BR}$ was found for the systems with correctly oriented antiparallel redox gradients in the two channels (22 \pm 3% for 19; • in Figure 4F). The lower η_{BR} obtained for 18 (53 \pm 8%; \blacktriangle in Figure 4F) in comparison with its constitutional isomer 20 (76 \pm 8%; \bigtriangleup in Figure 4F) is in agreement with the favorable effect of the gradient in minimizing charge recombination and supports the occurrence of partial hydrazone exchange from the surface of the film. On the other hand, the $\eta_{
m BR}$ obtained for three-component OMARG SHJ 18 was significantly higher than that for four-component OMARG SHJ 19 and almost the same as that of simple SHJ 14 ($50 \pm 3\%$; \bigcirc in Figure 4F). These results demonstrate that double-channel redox gradients are needed to minimize the loss of charges.

Taken together, our results show that stack exchange in SOSIP architectures can be considered as a unique, reliable, and general approach for building ordered and oriented multicomponent systems of appreciable sophistication in a straightforward, user-friendly manner. Our current objective is to maximize the complexity accessible by subunit exchange on the one hand and to explore the potential for practical applications on the other.

ASSOCIATED CONTENT

Supporting Information. Details of experimental procedures and additional results. This material is available free of charge via the Internet at http://pubs.acs.org.

AUTHOR INFORMATION

Corresponding Author

Naomi.Sakai@unige.ch; Stefan.Matile@unige.ch

ACKNOWLEDGMENT

We thank D. Jeannerat, A. Pinto, and S. Grass for NMR measurements; the Sciences Mass Spectrometry (SMS) Platform for mass spectrometry services; and the University of Geneva, the European Research Council (ERC Advanced Investigator), the National Centre of Competence in Research (NCCR) in Chemical Biology, and the Swiss NSF for financial support.

REFERENCES

(a) Würthner, F.; Meerholz, K. Chem.—Eur. J. 2010, 16, 9366.
 (b) Wasielewski, M. R. Acc. Chem. Res. 2009, 42, 1910.
 (c) Bassani, D. M.; Jonusauskaite, L.; Lavie-Cambot, A.; McClenaghan, N. D.; Pozzo, J.-L.; Ray, D.; Vives, G. Coord. Chem. Rev. 2010, 254, 2429.
 (d) Bhosale, R.; Míšek, J.; Sakai, N.; Matile, S. Chem. Soc. Rev. 2010, 39, 138.

(2) (a) Sugiyasu, K.; Kawano, S.; Fujita, N.; Shinkai, S. *Chem. Mater.* **2008**, *20*, 2863. (b) Jonkheijm, P.; Stutzmann, N.; Chen, Z.; de Leeuw, D. M.; Meijer, E. W.; Schenning, A. P. H. J.; Würthner, F. *J. Am. Chem. Soc.* **2006**, *128*, 9535. (c) Li, W.; Saeki, A.; Yamamoto, Y.; Fukushima, T.; Seki, S.; Ishii, N.; Kato, K.; Takata, M.; Aida, T. *Chem.—Asian J.* **2010**, 5, 1566. (d) Bu, L.; Guo, X.; Yu, B.; Qu, Y.; Xie, Z.; Yan, D.; Geng, Y.; Wang, F. J. Am. Chem. Soc. **2009**, 131, 13242.

(3) (a) Decher, G. Science **1997**, 277, 1232. (b) Mwaura, J. K.; Pinto, M. R.; Witker, D.; Ananthakrishnan, N.; Schanze, K. S.; Reynolds, J. R. Langmuir **2005**, 21, 10119. (c) Guldi, D. M. J. Phys. Chem. B **2005**, 109, 11432. (d) Kira, A.; Umeyama, T.; Matano, Y.; Yoshida, K.; Isoda, S.; Park, J. K.; Kim, D.; Imahori, H. J. Am. Chem. Soc. **2009**, 131, 3198.

(4) Sakai, N.; Bhosale, R.; Emery, D.; Mareda, J.; Matile, S. J. Am. Chem. Soc. 2010, 132, 6923.

(5) Sakurai, S.; Areephong, J.; Bertone, L.; Lin, N.-T.; Sakai, N.; Matile, S. *Energy Environ. Sci.* **2011**, *4*, 2409.

(6) (a) Snaith, H. J.; Whiting, G. L.; Sun, B.; Greenham, N. C.; Huck, W. T. S.; Friend, R. H. *Nano Lett.* **2005**, *5*, 1653. (b) Foster, S.; Finlayson, C. E.; Keivanidis, P. E.; Huang, Y.-S.; Hwang, I.; Friend, R. H.; Otten, M. B. J.; Lu, L.-P.; Schwartz, E.; Nolte, R. J. M.; Rowan, A. E. *Macromolecules* **2009**, *42*, 2023.

(7) Sakai, N.; Lista, M.; Kel, O.; Sakurai, S.; Emery, D.; Mareda, J.; Vauthey, E.; Matile, S. J. Am. Chem. Soc. **2011**, 133, 15224.

(8) Lista, M.; Areephong, J.; Sakai, N.; Matile, S. J. Am. Chem. Soc. 2011, 133, 15228.

(9) Sakai, N.; Mareda, J.; Vauthey, E.; Matile, S. *Chem. Commun.* 2010, 46, 4225.

(10) See the SI.

(11) (a) Corbett, P. T.; Leclaire, J.; Vial, L.; West, K. R.; Wietor, J.-L.; Sanders, J. K. M.; Otto, S. *Chem. Rev.* **2006**, *106*, 3652. (b) Skene, W. G.; Lehn, J.-M. *Proc. Natl. Acad. Sci. U.S.A.* **2004**, *101*, 8270.

(12) (a) Krisovitch, S. M.; Regen, S. L. J. Am. Chem. Soc. 1992, 114, 9828. (b) Carnall, J. M. A.; Waudby, C. A.; Belenguer, A. M.; Stuart, M. C. A.; Peyralans, J. J.; Otto, S. Science 2010, 327, 1502. (c) Otto, S.; Furlan, R. L. E.; Sanders, J. K. M. Science 2002, 297, 590. (d) Anfinsen, C. B. Science 1973, 181, 223. (e) Mindell, J. A.; Zhan, H.; Huynh, P. D.; Collier, R. J.; Finkelstein, A. Proc. Natl. Acad. Sci. U.S.A. 1994, 91, 5272. (f) Prins, L. J.; Scrimin, P. Angew. Chem., Int. Ed. 2009, 48, 2288.

(13) (a) Rodriguez-Docampo, Z.; Otto, S. Chem. Commun. 2008, 5301. (b) von Delius, M.; Geertsema, E. M.; Leigh, D. A. Nat. Chem. 2010, 2, 96.

(14) Chou, C.-M.; Lee, S.-L.; Chen, C.-H.; Biju, A. T.; Wang, H.-W.; Wu, Y.-L.; Zhang, G.-F.; Yang, K.-W.; Lim, T.-S.; Huang, M.-J.; Tsai, P.-Y.; Lin, K.-C.; Huang, S.-L.; Chen, C.; Luh, T.-Y. *J. Am. Chem. Soc.* **2009**, *131*, 12579.

(15) Koster, L. J. A.; Kemerink, M.; Wienk, M. M.; Maturová, K.; Janssen, R. A. J. *Adv. Mater.* **2011**, *23*, 1670.