

# Unusual isolation of a hemiaminal product from 4-cyclohexyl-3-thiosemicarbazide and di-2-pyridyl ketone: Structural and spectral investigations

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Received 20 December 2004; revised 25 March 2005; accepted 29 March 2005

Available online 31 May 2005

## Abstract

Stable hemiaminal product was isolated as single crystals from the condensation reaction of di-2-pyridyl ketone with 4-cyclohexyl-3-thiosemicarbazide and structurally and spectrochemically characterized. The compound is the first report of a stable hemiaminal product of the condensation reaction of a substituted thiosemicarbazone. Spectral characterization was carried out using IR, UV–vis., <sup>1</sup>H NMR, <sup>13</sup>C NMR, COSY and HMQC. The compound crystallizes into a triclinic lattice with space group *P*-1. The molecule reveals a highly strained *Z**E* conformation and a one-dimensional packing is effected by a network of intermolecular hydrogen bonding interactions in the unit cell.

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**Keywords:** Di-2-pyridyl ketone; Thiosemicarbazone; Hemiaminal; COSY; HMQC; Crystal structure

## 1. Introduction

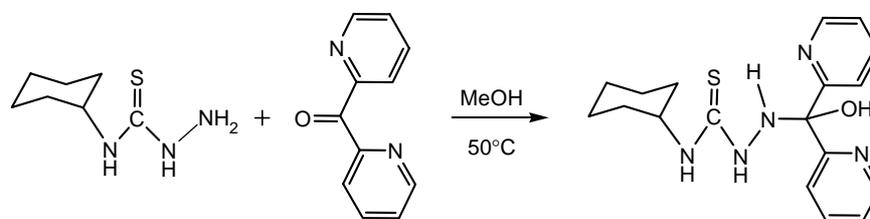
The mechanistic pathway for the condensation of amines with carbonyl compounds involves the formation of a tetrahedral addition intermediate followed by its dehydration to yield imines [1]. These reaction intermediates, formerly referred to as ‘carbinolamines’, are thermodynamically unstable and are presently termed as ‘hemiaminals’ by IUPAC. The rate of formation of hemiaminals with the variation of substituents and pH has been the subject of several investigations and many conclusions were established with regard to their mechanisms. Isolation of the hemiaminal product was immaterial in the above studies and the products of the reactions of aldehydes with thiosemicarbazides were the corresponding thiosemicarbazones. However, here we report the isolation of the hemiaminal derivative, HL of a substituted thiosemicarbazone as crystalline product from a reaction between 4-cyclohexyl-3-thiosemicarbazide and di-2-pyridyl ketone

(Scheme 1). The structure of the compound (Fig. 1) is established by single crystal X-ray diffraction and spectral studies.

Unsubstituted or alkyl substituted imines are highly reactive and imines are stabilized by one or more aryl groups attached to carbon or nitrogen. In such cases, the compounds are easily isolated and are called Schiff bases. Enhanced stability is imparted by aryl substituents due to the conjugative effect of the aryl group and the carbon–nitrogen double bond. For Schiff bases such as hydrazones and semicarbazones, additional stability is imparted by a second nitrogen attached to the azomethine nitrogen [1]. In the light of these facts, a stable thiosemicarbazone is always expected as the final product of the condensation between a thiosemicarbazide and a carbonyl compound as observed from our previous findings [2–13]. From this standpoint also, the present study is interesting, since in spite of the stability factors such as aromatic substitution and hydrazinic nitrogen linkage, the imine product is not separated out here. This is contrary to the observations so far, since stable thiosemicarbazones are the commonly and easily obtained products after the condensation of a thiosemicarbazide with any carbonyl compound.

Previous reports of stable hemiaminal crystals were in the form of covalent intermediates during the reaction of an

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Scheme 1.

aldehyde with lysine [14]. However, to the best of our knowledge, the title compound is the first report of the isolation of a hemiaminal product as single crystals from the condensation of a substituted thiosemicarbazide with any carbonyl compound.

## 2. Experimental

### 2.1. Materials

Cyclohexyl isothiocyanate (Fluka), hydrazine hydrate (Lancaster) and di-2-pyridyl ketone (Aldrich) were used as received. Infrared spectrum was recorded on ABB Bomem FTIR instrument as KBr pellets in the range 4000–500  $\text{cm}^{-1}$ . The  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, COSY and HMQC spectra were recorded using Bruker DRX 500, with  $\text{CDCl}_3$  as solvent and TMS as the standard at Sophisticated Instruments Facility, Indian Institute of Science, Bangalore, India. The solid state reflectance spectrum was recorded on Ocean Optics, SD2000 Fibre Optic Spectrometer.

### 2.2. X-ray data collection, structure solution and refinement

The X-ray diffraction data were measured at room temperature (293 K) using a Bruker SMART APEX CCD diffractometer, equipped a fine focus sealed tube. Selected crystal data and data collection parameters are given in Table 1. The SMART software was used for data acquisition and the SAINT software for data extraction [15]. Empirical absorption corrections were made on the intensity data and the structure was refined by full-matrix least squares using the SHELXL system of programs [16] and the graphics tool was DIAMOND [17]. All non-hydrogen atoms were refined

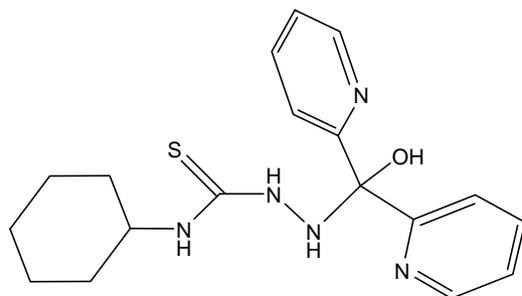


Fig. 1. Structure of compound HL.

anisotropically. The hydrogen atoms were treated with a mixture of independent and constrained refinement and used for structure factor calculation only. The selected bond lengths and bond angles are listed in Table 2.

### 2.3. Synthesis of HL

The hemiaminal, HL was prepared by adapting the following procedure. Ethanolic solutions of cyclohexyl isothiocyanate (0.706 g, 5 mmol) and hydrazine hydrate (0.250 g, 5 mmol) were mixed with constant stirring. The stirring was continued for one more hour and the white product, *N*(4)-cyclohexylthiosemicarbazide formed was washed, dried and recrystallised from ethanol. A methanolic solution of the *N*(4)-cyclohexyl thiosemicarbazide (0.865 g, 5 mmol) was then refluxed with di-2-pyridyl ketone (0.921 g, 5 mmol) in 5 ml methanol continuously for 4 h after adding 1–2 drops of acetic acid and the pH of

Table 1  
Crystal data and structural refinement for HL

Parameters	HL
Empirical formula	$\text{C}_{18}\text{H}_{25}\text{N}_5\text{O}_2\text{S}$
Formula weight, $M$	375.49
Temperature, $T$ (K)	293(2)
Wavelength, Mo $K\alpha$ ( $\text{\AA}$ )	0.71073
Crystal system	Triclinic
Space group	$P-1$
Lattice constants $a$ ( $\text{\AA}$ ), $b$ ( $\text{\AA}$ ), $c$ ( $\text{\AA}$ ), $\alpha$ ( $^\circ$ ), $\beta$ ( $^\circ$ ), $\gamma$ ( $^\circ$ )	8.571(5), 10.700(6), 11.104(6), 95.089(9), 104.996(8), 91.481(9)
Volume $V$ ( $\text{\AA}^3$ )	978.5(9)
$Z$ calculated density, $\rho$ ( $\text{Mg m}^{-3}$ )	2, 1.274
Absorption coefficient, $\mu$ ( $\text{mm}^{-1}$ )	0.188
$F(000)$	400
Crystal size (mm)	$0.35 \times 0.33 \times 0.30$
$\theta$ Range for data collection	1.91–26.25
Limiting indices	$-10 \leq h \leq 10$ , $-13 \leq k \leq 13$ , $-13 \leq l \leq 13$
Reflections collected	10253
Unique reflections	3978 [ $R_{\text{int}} = 0.0145$ ]
Completeness to $\theta$	26.25 (91.1%)
Absorption correction	Multi-scan
Max. and min. transmission	0.9459 and 0.9373
Refinement method	Full-matrix least-squares on $F^2$
Data/restraints/parameters	3978/0/335
Goodness-of-fit on $F^2$	1.029
Final $R$ indices [ $I > 2\sigma(I)$ ]	$R_1 = 0.0443$ , $wR_2 = 0.1165$
$R$ indices (all data)	$R_1 = 0.0496$ , $wR_2 = 0.1211$
Largest difference peak and hole ( $\text{e \AA}^{-3}$ )	0.304 and $-0.265$

Table 2  
Selected bond lengths (Å) and bond angles (°) of HL

	HL
O(1)–C(6)	1.410(1)
C(6)–N(3)	1.464(2)
N(3)–N(4)	1.406(1)
N(4)–C(12)	1.340(2)
C(12)–S(1)	1.702(1)
N(3)–N(4)	1.371(4)
C(12)–N(5)	1.327(2)
N(5)–C(13)	1.461(2)
C(6)–C(7)	1.537(2)
C(5)–C(6)	1.527(2)
O(1)–C(6)–N(3)	111.50(13)
O(1)–C(6)–C(7)	110.03(13)
C(6)–N(3)–N(4)	110.06(12)
N(3)–N(4)–C(12)	120.68(13)
N(4)–C(12)–S(1)	118.67(12)
N(4)–C(12)–N(5)	117.22(15)
S(1)–C(12)–N(5)	124.10(13)
C(12)–N(5)–C(13)	125.38(16)

the solution was measured to be 4.7. On cooling the solution pale yellow crystals separated, which were filtered and washed with methanol. The crystals were recrystallized from ethanol and dried over  $P_4O_{10}$  in vacuo. Single crystals suitable for X-ray diffraction were obtained by slow evaporation of a methanolic solution of HL. Mp 124 °C.

### 3. Results and discussion

The carbonyl addition of both weakly and moderately basic amines at weakly acidic pH values corresponds to a stepwise mechanism involving the formation of an unstable zwitterionic form of hemiaminal, which undergoes an intramolecular ‘proton switch’ or a proton transfer from a molecule of acid in order to convert into the hemiaminal product [18]. Usually, dehydration of the hemiaminal readily occurs as the next step yielding imines. One mechanistic approach towards the hemiaminal formation can be made by a possible hydrolysis of the imine. The hemiaminal is then obtained as the intermediate product, which finally yields the corresponding amine and carbonyl compounds. However, this mechanistic prediction seems to be unreliable in the present case, since once the imine, i.e., the di-2-pyridyl ketone 4-cyclohexyl-3-thiosemicarbazone is formed, it is expected to be stabilized by substituent effects and electron delocalization, which contradicts the possibility for getting hydrolysed. Hence, here the stabilization of the hemiaminal derivative can be rationalized only by taking into account the different electron transfer processes in the transition state during the condensation reaction. A suggestive mechanistic approach would be that the double bond formation and the rehybridisation of carbon and nitrogen atoms from tetrahedral  $sp^3$  structure of the hemiaminal to planar  $sp^2$  structure of the imine product lag behind other processes such as proton removal and electron

delocalization in the transition state. As stated by Sayer and Jencks [19], the energetic advantage of maintaining significant overlap and bonding with the departing oxygen and hydrogen atoms in an  $sp^3$  hybridised transition state is more than enough to offset the stabilization achieved through a carbon-nitrogen double bond by rehybridisation towards a planar  $sp^2$  structure.

#### 3.1. Crystal structure of HL

The compound crystallizes into a triclinic lattice with space group  $P-1$  and the molecular structure of HL along with atom numbering scheme is given in Fig. 2. The molecule is non-planar as a whole and exists in the  $ZE$  conformation since  $Z$  and  $E$  configurations are perceived about C6–N3 and C12–N4 bonds respectively. The central moiety comprising of atoms N3, N4, C12, S1 and N5 is almost planar with a maximum deviation of 0.0211(3) Å from the mean plane. A torsion angle value of 177.86(11)° corresponding to the S1–C12–N4–N3 moiety confirms the trans configuration of the thiocarbonyl S1 atom [27]. Unlike the azomethine double bond in thiosemicarbazones, the C6–N3 bond distance (1.464(2) Å) is more close to that of a C–N single bond [28], supporting the hemiaminal structure of the compound. In order to minimize the steric strain, the bulky pyridyl groups are deviated farthest with a dihedral angle of 75.31(1)° between each other. The intramolecular hydrogen bonding interactions, viz., N(5)–H(401)⋯N(3) and O(1)–H(101)⋯N(2) lead to the formation of two five membered rings comprising of atoms N(3), N(4), C(12), N(5), H(401) and N(2), C(7), C(6), O(1) and H(101), respectively. Ring puckering analysis and least square planes calculations show that the cyclohexyl ring, Cg(3) adopts a chair conformation ( $Q_T=0.5404$  Å).

The molecular packing of compound HL is shown in Fig. 3. The unit cell is viewed down the ‘a’ axis and the repeating unit consists of two molecules set in an offset fashion, arranged one-dimensionally in the crystal lattice. A residual water molecule in the crystal lattice is involved in extensive hydrogen bonding interactions with three of its

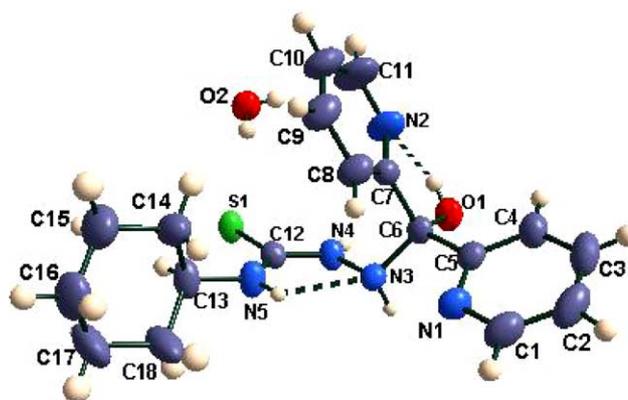


Fig. 2. Molecular structure of HL. Intramolecular hydrogen bonding interactions are shown as dashed lines.

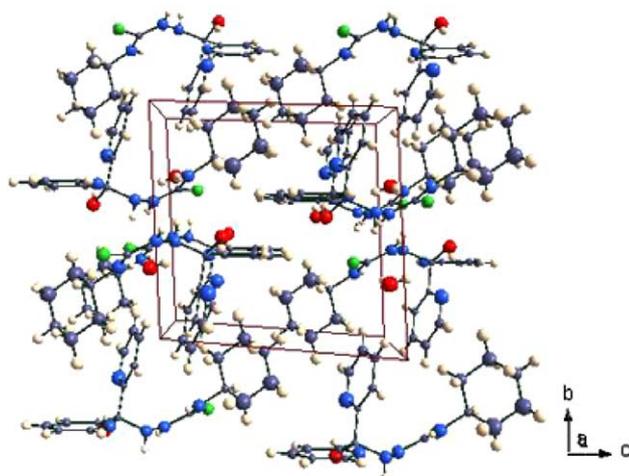


Fig. 3. Molecular packing diagram of HL, the unit cell is viewed down the 'a' axis.

adjacent molecules. Each thiocarbonyl sulphur atom is involved in two intermolecular hydrogen bonding interactions,  $N(4)-H(301)\cdots S(1)^i$  [ $d_{D\cdots A}=3.410(1)$  Å;  $i = -x+1, -y+1, -z+2$ ] and  $O(2)-H(202)\cdots S(1)^{ii}$  [ $d_{D\cdots A}=3.490(1)$  Å;  $ii = +x, +y, +z$ ]. The water molecule residue reveals two more intermolecular hydrogen bonds, viz.,  $N(5)-H(401)\cdots O(2)^{iii}$  [ $d_{D\cdots A}=3.035(2)$  Å;  $iii = +x, +y, +z$ ] and  $O(2)-H(201)\cdots N(1)^{iv}$  [ $d_{D\cdots A}=2.823(2)$  Å;  $iv = x-1, +y, +z$ ]. These intermolecular hydrogen bonding interactions are mainly operational in the packing of the molecules in the crystal lattice, since  $\pi-\pi$  and  $C-H\cdots\pi$  interactions are weak, perceived at rather longer distances.

### 3.2. Spectral studies

The compound reveals well-defined molecular vibrations in the IR region. The  $-OH$  stretching vibration of the hydroxyl group at C6 is observed as a distinct peak at  $3516\text{ cm}^{-1}$  in the IR spectrum. The intramolecular hydrogen bonding interactions of C6–OH with N2 observed in the crystal structure is confirmed by an additional broad, shallow  $-OH$  stretching peak observed at  $3050\text{ cm}^{-1}$  [20]. The  $-NH$  symmetrical and asymmetrical stretching vibrations are perceived as two distinct peaks at  $3726$  and  $3187\text{ cm}^{-1}$ , respectively, while the  $-CH$  stretching vibrations of the cyclohexyl moiety are observed as sharp peaks at  $2930$  and  $2853\text{ cm}^{-1}$ . The absence of  $\nu(SH)$  band around  $2600\text{ cm}^{-1}$  suggest the existence of the thiosemicarbazone in the thione form and the  $C=S$  stretching vibration is observed as a sharp peak at  $1237\text{ cm}^{-1}$ . In the solid state electronic spectrum of the compound, the  $\pi^* \leftarrow \pi$  transitions of the pyridyl ring and the thiosemicarbazone moiety are observed at  $38,610$  and  $37,310\text{ cm}^{-1}$ , respectively, while the  $\pi^* \leftarrow n$  transition of the pyridyl nitrogen is observed as a broad band at  $28,330\text{ cm}^{-1}$ .

Structural studies reveal a highly strained conformation for the hemiaminal under study. A wide network of

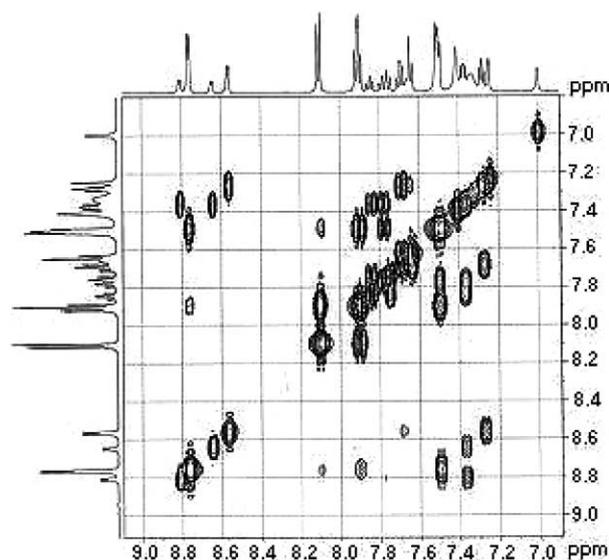


Fig. 4.  $^1H-^1H$  COSY spectrum of HL.

intra- and intermolecular hydrogen bonding interactions contribute towards the stability of packing in the crystal lattice. An interesting aspect of these interactions is that, unlike the previous reports of compounds of similar structures [21–25], in the present case, the similar protons of the two pyridyl rings are found to resonate at slightly varied frequencies. For instance, while the C1–H resonances are observed as double doublets at  $8.60$  ppm, an identical set of peaks are observed at  $8.79$  ppm, which can be assigned to C11–H protons. Assignment of these two double doublets to C1–H and C11–H protons is confirmed by the  $^1H-^1H$  COSY spectrum of HL (Fig. 4). From the COSY, the C8–H proton ( $8.10$  ppm) is found to give off diagonal peaks at  $8.79$  ppm, resultant of its interaction with the proton on C11. Similarly, the off diagonal peaks around  $8.60$  ppm correspond to the coupling interaction between C1–H and C4–H protons. The prominent intramolecular interaction  $O1-H101-N2$  observed at a donor-acceptor distance of  $2.577(1)$  Å from the crystal structure is instrumental in bringing about the different resonant frequencies. An increased charge density is experienced at N2 as a result of the hydrogen bonding, which shifts the C11–H resonance to more downfield compared to that of the C1–H proton. The molecule as a whole adapts a conformation in which the pyridyl ring with N2 is highly strained with active participation in the hydrogen bonding interactions. This results in a redistribution of electron density over the entire ring causing slight variations in the resonating frequencies of similar protons on the two pyridyl rings. Hence the C2–H and C10–H peaks are observed as triplets at  $7.28$  and  $7.38$  ppm respectively, while the C3–H and C9–H resonances are perceived as multiplets centered around  $7.71$  and  $7.50$  ppm, respectively. Similarly, the doublets at  $8.10$  ppm and the doublet of a triplet at  $7.90$  ppm are assigned to C8 and C4 protons respectively. Assignment of these peaks to the respective

protons is carried out with the help of  $^1\text{H}$ – $^1\text{H}$  correlation data available from the COSY spectrum.

Two sets of double doublets observed at 8.79 and 8.60 ppm for protons  $\alpha$  to pyridyl nitrogens attribute interesting features to the compound under study. The  $^1\text{H}$  NMR spectrum of the present compound follows strictly first order with well-separated chemical shift positions for protons with different environments. For example, the signal for C1–H proton appears as a double doublet around 8.60 ppm due to the following coupling pattern observed between adjacent protons. At first, the proton at C1 couples with C2 proton, which splits the C1–H signal into a doublet. But C1–H also couples with C3–H so that each line of the C1–H is further split into two, giving a double doublet [20]. Although doublets and double doublets are reported previously [24,25], two well-defined sets of double doublets corresponding to two pyridyl protons in identical chemical environment are observed for the first time in a compound prepared from di-2-pyridyl ketone.

More distinctive features are associated with the resonance peaks of the protons on N3 and N4 atoms, which appear as a triplet at 7.84 ppm corresponding to the two protons of identical chemical environments. Contrary to this, in all the previous observations of structurally related compounds, a sharp singlet at appreciable downfield values ( $\delta$  13–15 ppm) is observed for the N3–H proton [22]. Hence, in order to confirm the assignment of these triplets to –NH protons, the  $^1\text{H}$ – $^{13}\text{C}$  correlation spectra were recorded. Absence of coupling of the triplets at 7.84 ppm with the  $^{13}\text{C}$  peaks in the HMQC experiments (Fig. 5) authenticated its assignment to –NH protons. The triplet is in accordance with the  $2I+1$  splitting ( $I=1$  for  $^{14}\text{N}$ ) of the resonance peak of the proton attached to the  $^{14}\text{N}$  nucleus. The appreciable electrical quadrupole moment at N3 and N4 nitrogens is able to induce an efficient spin relaxation to observe three

well-defined spin states of the nitrogen nucleus, to which the proton responds by giving a triplet in the  $^1\text{H}$  NMR spectrum.

The singlets at appreciable downfield in the previous reports were the result of rapid exchange and resultant decoupling of the –NH protons, which is found to be absent for the N3 and N4 protons in the present compound. However, the N5H proton shows exchangeability at a low rate and turns up as a broad peak at 7.33 ppm. At low –NH exchange rates,  $^{14}\text{N}$  nucleus imparts moderate spin relaxation resulting in an intermediate lifetime for the spin states of nitrogen. Response of the –NH proton to these three spin states which are changing at a moderate rate is usually observed as a broad peak [26]. As the cyclohexyl moiety is observed in the chair conformation, the axial and equatorial hydrogens show slightly varied chemical shift values with the equatorial protons observed at more downfield (2.10–1.73 ppm) compared to their axial counterparts (1.42–1.22 ppm). However, the C13–H is deshielded by the adjacent electronegative nitrogen resulting in a multiplet at 4.21 ppm, which couples with the proton on N5. Assignment of the broad peak to N5 proton is also confirmed by the  $^1\text{H}$ – $^1\text{H}$  correlation experiments, since coupling of the N5–H proton with the C13–H proton of the cyclohexyl ring is clearly evident from the COSY.

#### 4. Supplementary data

Crystallographic data for structural analysis has been deposited with the Cambridge Crystallographic Data Center, CCDC for compound HL. Copies of this information maybe obtained free of charge from The Director, CCDC 271996, 12 Union Road, Cambridge, CB2, IEZ, UK (fax: +44 1223 336 033; e-mail: [deposit@ccdc.cam.ac.uk](mailto:deposit@ccdc.cam.ac.uk) or <http://www.ccdc.cam.ac.uk>).

#### Acknowledgements

The authors are thankful to Sophisticated Instruments Facility, IISc Bangalore, India for NMR measurements. DST IRPHA is kindly acknowledged for X-ray data collection. V. Suni is thankful to the State Council for Science, Technology and Environment, Kerala, India for financial support in the form of a fellowship.

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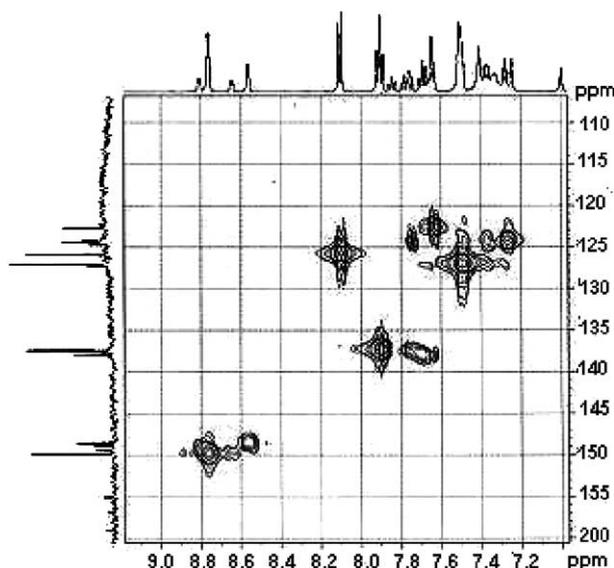


Fig. 5.  $^1\text{H}$ – $^{13}\text{C}$  HMQC spectrum of HL.

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