out. A similar mechanism is suggested for anodic oxidation of endo alcohol 1c.

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Richard S. Glass,\* John R. Duchek Jacob T. Klug, George S. Wilson

Department of Chemistry, The University of Arizona Tucson, Arizona 85721

## Trifluoromethylated "Dewar" Pyrroles: 5-Azabicyclo[2.1.0]pentene-2 Ring System<sup>1</sup>

Sir:

In previous papers,<sup>2</sup> we reported that 1,2,3,4-tetrakis(trifluoromethyl)-5-thiabicyclo[2.1.0]pentene-2, a "Dewar" thiophene (1), reacted as a dienophile with the conjugated dienes. In this report, we discuss the 1,3-dipolar cycloaddition of 1 and transformation of the adducts to "Dewar" pyrroles (Chart I).

Treatment of 1 with phenyl or cyclohexyl azide (2a or 2b) in CH<sub>2</sub>Cl<sub>2</sub> at room temperature gave the corresponding adduct (**3a** or **3b**).<sup>3</sup> **3a**: 28.1% yield, mp 62–64 °C; IR (CHCl<sub>3</sub>) 3020,  $1600, 1500, 1290, 1185 \text{ cm}^{-1}$  (no absorption near 1700 cm<sup>-1</sup> ascribable to the cyclobutenic double bond<sup>2</sup>); <sup>1</sup>H NMR (CDCl<sub>3</sub>) § 7.43; <sup>19</sup>F NMR (CDCl<sub>3</sub>) -3.6 (3 F, m), -2.6 (3 F, m), -2.0 (3 F, m), 7.4 (3 F, m) ppm;<sup>4</sup> m/e 447 (M - N<sub>2</sub>). 3b: 81% yield, mp 59-60 °C; IR (CHCl<sub>3</sub>) 2940, 2860, 1480, 1300, 1180, 1165 cm<sup>-1</sup>; <sup>1</sup>H NMR (CDCl<sub>3</sub>)  $\delta$  3.68 (1 H, m, N-CH), 1.2-2.36 [10 H, m, (CH<sub>2</sub>)<sub>5</sub>]; <sup>19</sup>F NMR (CDCl<sub>3</sub>) -4.6 (3 F, m), -2.4 (3 F, m), 1.6 (3 F, m), 6.8 (3 F, m) ppm; m/e 453 (M – N<sub>2</sub>).

Irradiation of 3a or 3b in *n*-pentane with a low-pressure mercury lamp caused elimination of a nitrogen molecule to give the valence-bond isomers of 1,4-thiazine, compounds of a new ring system (4a or 4b). 4a: 55% yield, colorless oil; IR (CHCl<sub>3</sub>) 3040, 1600, 1290, 1170 cm<sup>-1</sup>; <sup>1</sup>H NMR (CDCl<sub>3</sub>) δ 7.18; <sup>19</sup>F NMR (CDCl<sub>3</sub>) -5.6 (6 F, s), 2.8 (6 F, s) ppm; m/e 447 (M<sup>+</sup>); high resolution, m/e 446.997 (calcd 446.995). 4b: quantitative yield, colorless oil; <sup>19</sup>F NMR spectrum is quite similar to that of 4a.

Chart I





IR spectrum. Compound 5a rapidly isomerized to 1,2,2a,7btetrakis(trifluoromethyl)-2a,7b-dihydro-3*H*-cyclobuta[*b*]indole (6): 35.6% yield, colorless oil; IR (n-pentane) 3430 (N-H), 1715 (C=C), 1210 (C-F) cm<sup>-1</sup>; <sup>1</sup>H NMR (CCl<sub>4</sub>) δ 7.2-7.5 (2 H, m), 6.75-7.04 (2 H, m), 4.72 (1 H, b); <sup>19</sup>F NMR (CDCl<sub>3</sub>) -0.8 (3 F, m), 0.0 (3 F, m), 5.2 (3 F, m), 10.8 (3 F, m) ppm; m/e 415 (M<sup>+</sup>); high resolution, m/e 415.024 (calcd 415.023). 5b: 60% yield, colorless oil; IR (CCl<sub>4</sub>) 2940, 2860, 1705, 1190, 1160 cm<sup>-1</sup>; <sup>1</sup>H NMR (CCl<sub>4</sub>)  $\delta$  2.6 (1 H, N-CH), 1.15-1.95 [10 H, m, (CH<sub>2</sub>)<sub>5</sub>]; <sup>19</sup>F NMR (CCl<sub>4</sub>) 1.52 (6 F, s) (bridgehead CF<sub>3</sub>), 2.72 (6 F, s) (olefinic CF<sub>3</sub>) ppm; <sup>13</sup>C NMR (CDCl<sub>3</sub>) 23.78 (3' and 5'), 25.88 (4'), 31.97 (2' and 6'), 49.95 (1 and 4), 55.11 (1'), 118.08, and 121.24 (CF<sub>3</sub> carbon), and 140.44 (2 and 3) ppm; m/e 421 (M<sup>+</sup>); half-life at 100 °C in benzene, about 8.8 h.

Treatment of **5b** with furan and cyclopentadiene gave the Diels-Alder adducts (7 and 8).<sup>3</sup> 7: 83% yield, mp 123-125 °C; IR (CCl<sub>4</sub>) 2940, 2860, 1210, 1170 cm<sup>-1</sup>; <sup>1</sup>H NMR (CCl<sub>4</sub>)  $\delta$ 6.6 (2 H, b, =-CH-), 5.36 (2 H, s, O-CH<), 3.0 (1 H, b, N-CH<), 1.0-2.0 [10 H, m, (CH<sub>2</sub>)<sub>5</sub>]; <sup>19</sup>F NMR (CCl<sub>4</sub>) -3.8 (6 F, s), -0.2 (6 F, s) ppm; m/e 421 (5b) and 68 (furan). 8: 73% yield, mp 101-104 °C; IR (CCl<sub>4</sub>) 2940, 2860, 1200, 1180  $cm^{-1}$ ; <sup>1</sup>H NMR (CCl<sub>4</sub>)  $\delta$  6.40 (2 H, b, ==CH-), 3.44 (2 H, b, >CH), 3.04 (1 H, b, N-CH<), 1.1-2.3 (12 H, m, methylenes); <sup>19</sup>F NMR (CCl<sub>4</sub>) -3.5 (6 F, s), -1.8 (6 F, s); *m/e* 487  $(M^{+}).$ 

In this research, some quite novel compounds of new ring systems, too highly strained to be isolated without the trifluoromethyl groups, were obtained. The formation of these might be due to the stabilization of strained molecules by the perfluoroalkyl effect.

## **References and Notes**

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Yoshiro Kobayashi,\* Itsumaro Kumadaki, Akio Ohsawa, Akira Ando

Tokyo College of Pharmacy Horinouchi, Hachioji-shi, Tokyo, Japan Received May 11, 1977

## Threshold Photoionization and Dissociation of Toluene and Cycloheptatriene

Sir:

There has been some conjecture as to the threshold structure of the  $C_7H_7^+$  ion formed by ionizing toluene and cycloheptatriene. From investigations using collisional activation mass spectra, McLafferty and Winkler<sup>1</sup> concluded that approximately equal abundances of benzyl and tropylium cations are formed from threshold energy decomposition, following a rapid equilibration between  $C_7H_8^+$  ions with toluene and cycloheptatriene structures. Dunbar<sup>2</sup> also proposed that such a dynamic equilibrium was consistent with his photodissociation results for the toluene parent ion. However, the fraction of benzyl ion was observed to decrease to zero near the photodissociation threshold. This he interpreted in terms of the activation energy for dissociation being lower for the cycloheptatriene form of  $C_7H_8^+$  than the toluene form. More recently, a MINDO calculation<sup>3</sup> has proposed that the easiest route for H loss from  $C_7H_8^+$  for both toluene and cycloheptatriene is via the cycloheptatriene molecular ion to tropylium.

Appearance potentials can be used in conjunction with other thermochemical data as a means of determining possible structures of ions at threshold. However, in the case of  $C_7H_7^+$ formed from toluene and cycloheptatriene there is a large kinetic shift associated with the decompositions which has resulted in previous overestimates of the appearance energies. This energy in excess of threshold, necessary for the ion to react sufficiently fast to be detected, can be minimized by increasing the ion source residence time<sup>4</sup> and/or increasing the detection sensitivity of the instrument. We wish to report here the use of a photoionization mass spectrometer and associated signal-averaging techniques which has enabled the measurement of AP(C<sub>7</sub>H<sub>7</sub><sup>+</sup>) from toluene and cycloheptatriene at high sensitivity.

The instrument used in the present experiments consists of a windowless hydrogen discharge lamp and 1-m Seya-Namioka monochromator<sup>5</sup> attached to a 15-cm single-focusing mass spectrometer. A microprocessor is used for the on-line control and extended signal averaging of experimental data needed to achieve a high sensitivity of detection. A detailed description of the photoionization mass spectrometer will form part of a later publication.<sup>6</sup>

The photoionization efficiency curve for  $C_7H_7^+$  produced from toluene is shown in Figure 1. Using the method of interpretation of photoionization data described by Guyon and Berkowitz,<sup>7</sup> we obtain an appearance energy of  $10.71 \pm 0.03$ eV. This is in excellent agreement with the results of Gordon and Reid<sup>4</sup> who obtained a limiting value of 10.70 eV at ion source residence times in excess of 900  $\mu$ s. The corresponding appearance energy for cycloheptatriene (Figure 2) is found to be 9.36  $\pm$  0.02 eV. Combining the heats of formation for the neutral species<sup>8</sup> in the reaction  $C_7H_8 + h\nu \rightarrow C_7H_7^+ + H +$ e + (excess energy) with the observed appearance energies, and neglecting any excess energy contribution, gives heats of formation for  $C_7H_7^+$  of 206.9  $\pm$  0.7 kcal/mol and 207.6  $\pm$  0.5



Figure 1. Threshold photoionization efficiency curve for formation of  $C_7H_7^+$  ion from toluene.



Figure 2. Threshold photoionization efficiency curve for formation of  $C_7H_7^+$  from cycloheptatriene.

kcal/mol from toluene and cycloheptatriene, respectively. The kinetic energy release accompanying both processes has been measured from metastable studies<sup>9</sup> to be 4.0 kcal/mol which probably represents an upper limit to the reverse activation energy. A recent MINDO calculation<sup>3</sup> indicates that the reverse activation energy is indeed small, giving an estimate of only 1.4 kcal/mol.

Lossing<sup>10</sup> has measured the adiabatic ionization energy for the benzyl radical by electron impact to be  $\leq$ 7.27 eV which is in good agreement with a recent corresponding photoelectron spectroscopic value of 7.20 ± 0.03 eV.<sup>11</sup> If  $\Delta H_f$ (benzyl) is taken as 45 kcal/mol,<sup>12</sup> one obtains  $\Delta H_f$ (benzyl<sup>+</sup>) = 211 kcal/mol which is lower than either a recent ab initio calculation of 217.1 kcal/mol<sup>13</sup> or a MINDO calculation of 220.4 kcal/mol.<sup>14</sup> It is also lower than the value of 219 ± 4 kcal/mol obtained by Jackson, Lias, and Ausloos<sup>15</sup> who used ion cyclotron resonance spectroscopy to measure the rate constants for reaction of benzyl ion with a number of alkanes and cycloalkanes. It thus appears that from an energetic viewpoint the C<sub>7</sub>H<sub>7</sub><sup>+</sup> ion formed at threshold for toluene and cycloheptatriene cannot have the benzyl structure.

Our value of 207 kcal/mol is in good agreement with the theoretical calculation for tropylium ion by Abboud, Hehr, and Taft<sup>13</sup> (207.9 kcal/mol) and ionization energy measurements of the tropyl radical<sup>16,17</sup> (209 kcal/mol), but is at variance with the MINDO value of 195.6 kcal/mol.<sup>14</sup> Dewar and coworkers<sup>14</sup> have suggested that the heat of formation derived from the tropyl radical ionization potential could be due to a large