

acetate. An acetyl determination on the product showed two acetyl groups. Attempted ketone cleavage with sodium ethylate, or even with cold dilute aqueous barium hydroxide, removed two acetyl groups and gave β -phenyllactic acid as the end-product. The reaction had therefore proceeded exclusively in the sense of the acid cleavage with simultaneous saponification of the acetoxy group. Attempts at acid hydrolysis were also unsuccessful, where the compound showed great resistance to this treatment.

It is probable that the reaction of halides with α -acetoxyacetates may be generally useful in the preparation of α -hydroxy acids.

Experimental Part

Ethyl α -Acetoxyacetate.¹—Eighty-four grams of ethyl acetoacetate (0.65 mole) was dissolved in 320 ml. of benzene, and 88 g. of lead tetraacetate (0.6 mole calculated on a 100% basis) was added in several small portions with stirring, the temperature being kept below 35°. The mixture was stirred for one-half hour after the last addition, and then ice-water was added and the layers were separated. The benzene solution was washed with ice water and sodium bicarbonate solution, dried, and distilled from a modified Claisen flask. Forty-one grams (34%) of product was obtained boiling at 125–127° at 17.5 mm., n_D^{25} 1.4280.

The compound distills as a pale yellow liquid, which becomes colorless upon standing. It readily reduces Fehling solution in the cold. *Anal.* Calcd. for $C_8H_{12}O_5$: C_2H_5O , 23.9. Found: C_2H_5O , 23.8.

Ethyl α -Acetoxy- α -benzylacetoacetate.—Two and four-tenths grams (0.1 mole) of sodium hydride was placed in a nitrogen-swept flask with 100 ml. of dioxane, and 18.8 g. (0.1 mole) of the acetoxy compound in 40 ml. dioxane was added dropwise to the stirred mixture over a period of thirty minutes. After an additional thirty minutes 17.1 g. (0.1 mole) of benzyl bromide was added slowly, and the mixture was then refluxed for one and a half hours. The reaction mixture was poured into 1 liter of water containing sufficient acetic acid to neutralize the alkali. The heavy oil which separated was extracted twice with ether, and the ethereal solution was washed with water and then with saturated sodium chloride solution. The solution was dried, the solvent removed, and the product distilled from a modified Claisen flask. The yield was 18.6 g. (67%) of product boiling at 118–122° at 0.5 mm., n_D^{25} 1.4916.

Anal. Calcd. for $C_{15}H_{18}O_5$: C, 64.73; H, 6.52; C_2H_5O , 16.2; $2CH_3CO$, 30.9. Found: C, 64.02; H, 6.72; C_2H_5O , 16.0; CH_3CO , 26.7.

Hydrolysis of Ethyl α -Acetoxy- α -benzylacetoacetate.—The compound was unchanged by refluxing with water for one hour. Refluxing with glacial acetic acid containing a little concentrated hydrochloric acid also failed to hydrolyze the compound. Although some carbon dioxide was evolved by the process, the starting material was recovered almost quantitatively. The same was true when the compound was refluxed with 20% sulfuric acid. The products of these experiments were identified, by boiling points, refractive indices and ethoxyl analyses, as the original material.

Hydrolysis with Sodium Ethylate.—Five and one-half grams (0.02 mole) of ethyl α -acetoxy- α -benzylacetoacetate was added to 2 g. of sodium dissolved in 35 ml. of ethanol and the solution was refluxed for one hour. Most of the alcohol was removed under diminished pressure, and the residue was acidified to congo red with dilute hydrochloric acid. The solution was then saturated with sodium chloride and extracted several times with ether. On evaporation of the solvent on the steam-bath and subsequently under reduced pressure, 3.3 g. of crystalline material was obtained, which was recrystallized from ben-

zene. The product melted at 95–96°, neutralization equivalent 166, and was identified as β -phenyllactic acid.²

Hydrolysis with Barium Hydroxide.—Five and one-half grams of ethyl α -acetoxy- α -benzylacetoacetate was added to 250 ml. of a 2.5% aqueous solution of barium hydroxide, and the suspension was shaken for forty-eight hours. Carbon dioxide was then passed into the solution to remove the excess alkali. After the barium carbonate had been filtered off, most of the water was removed under reduced pressure. On addition of ethanol, 5.8 g. of crude barium salt of β -phenyllactic acid was obtained.

(2) *Ber.*, 42 [4], 4892 (1909).

U. S. DEPARTMENT OF AGRICULTURE
AGRICULTURAL RESEARCH ADMINISTRATION
BUREAU OF ENTOMOLOGY AND PLANT QUARANTINE
BELTSVILLE, MARYLAND RECEIVED MAY 17, 1948

Preparation of 1,2-Dichloropropene-1

BY ERNEST H. HUNTRESS AND F. SANCHEZ-NIEVA¹

In view of the extent of recent studies of chloro- and polychloro-olefins the paucity of reports on 1,2-dichloropropene is remarkable. This compound should exist in two geometrically stereoisomeric forms. The only record of the lower boiling stereomer is an ancient report² claiming its preparation from 1,2,2-trichloropropane by dehydrochlorination with alcoholic potassium hydroxide. A liquid regarded as the higher boiling stereomer was subsequently obtained³ from α,β,β -trichloro-*n*-butyric acid by solution and heating in aqueous sodium carbonate. None of the cited papers give satisfactory details.

We have developed the preparation of the lower boiling stereomer to give readily 55–58% yields, have established certain needed physical constants, have shown that upon ozonolysis the compound yields formic and acetic acids, and that with methanol it gives an azeotrope. We did not, in this work, study the preparation of the supposed higher boiling stereomer³ but all attempts to isomerize the lower boiling material were unsuccessful. The configuration of the two stereomers is still undetermined.

1,2-Dichloropropene-1 (lower boiling stereomer).—One mole of 1,2,2-trichloropropane (b. p. 123–125° at 762.4 mm., n_D^{25} 1.4609, obtained from 2-chloropropene-1 by addition of chlorine) was added dropwise with frequent shaking to a 25% solution of sodium methoxide (1.3–1.5 moles) in methanol kept cold in ice water. After completion of the vigorous reaction the solution was poured into a large volume of water, and the heavy insoluble layer washed, dried and distilled. Because of its much higher boiling point any unreacted trichloropropane was easily separated from the desired dichloropropene (55–58% yield).

Redistillation of the crude product through a ten theoretical plate column gave the lower boiling 1,2-dichloropropene-1, b. p. 76.8–77.0° at 757.0 mm.; d_4^{25} 1.1755, d_4^{20} 1.1818, n_D^{25} 1.4451, n_D^{20} 1.4471.

This product formed with anhydrous methanol an azeo-

(1) From part of a thesis submitted by Mr. Sanchez-Nieva in October, 1945, in partial fulfillment of the requirements for the degree of Master of Science at M. I. T.

(2) Friedel and Silva, *Compt. rend.*, 74, 807 (1872); 75, 81 (1872); *Bull. soc. chim.*, [2] 17, 386 (1872); *Jahresber.*, 322, 329 (1871).

(3) Szenic and Taggesell, *Ber.*, 28, 2667–2668 (1895).

trope, b. p. 56.5–56.8° at 760.5 mm., n_D^{25} 1.4030, found by comparison with a plot of $n_D^{25}/\%$ composition for known mixtures to contain 75 weight % of lower boiling 1,2-dichloropropene.

Ozonization of the above 1,2-dichloropropene-1 in carbon tetrachloride solution at -15° followed by treatment with boiling water gave a mixture of hydrochloric, formic and acetic acids. Treatment with silver hydroxide precipitated silver chloride, oxidized the formic to carbonic acid, and thus permitted identification by the Duclaux method of the acetate in the filtrate.

The structure of the 1,2-dichloropropene-1 was confirmed by addition of chlorine at 0° in light from a 200-watt clear glass lamp, giving the saturated 1,1,2,2-tetrachloropropane, b. p. 153–155° at 774 mm., n_D^{25} 1.4850.

DEPARTMENT OF CHEMISTRY
MASSACHUSETTS INSTITUTE OF TECHNOLOGY
CAMBRIDGE, MASSACHUSETTS RECEIVED APRIL 16, 1948

The Structure of C_3F_6

BY EDMOND G. YOUNG AND WILLIAM S. MURRAY

Studies in this Laboratory involving methods of commercial feasibility for the synthesis of CF_3-

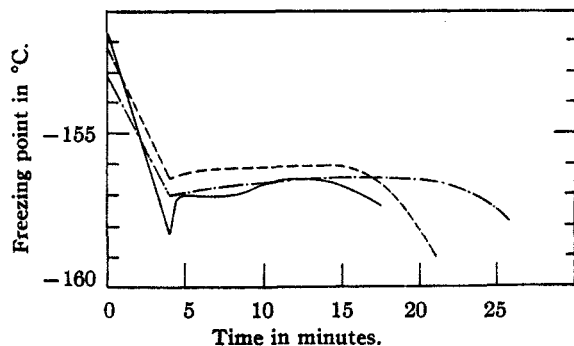


Fig. 1.— C_3F_6 products: ---, pyrolysis product; - - -, dechlorination product; —, mixed.

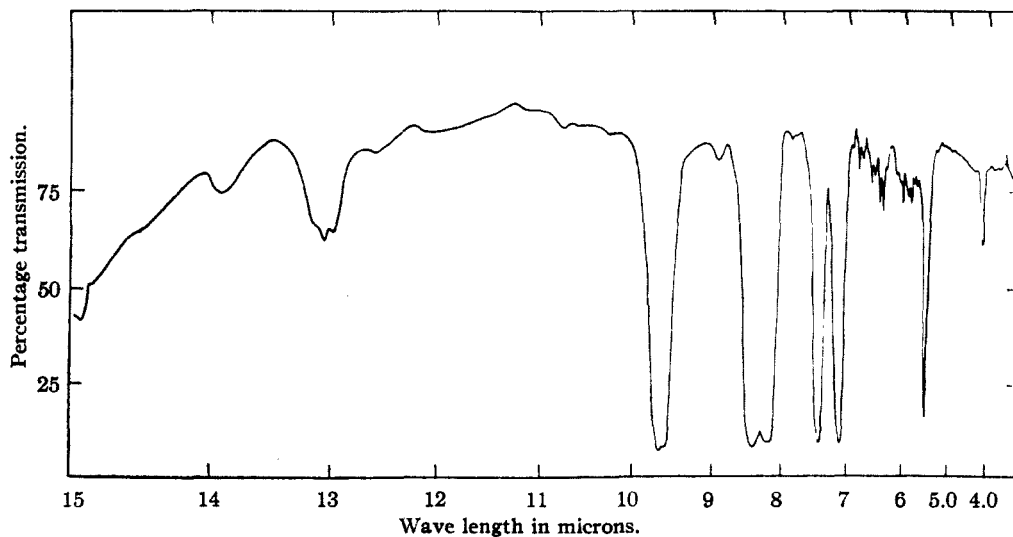


Fig. 2.— $CF_3CF=CF_2$.

$CF=CF_2$ brought to the fore the question of the proper structure of C_3F_6 as prepared by several methods. In particular, Benning, Downing and

Park¹ reported the structure of the C_3F_6 obtained from pyrolysis of "Teflon" tetrafluoroethylene polymer as being a cyclic compound. It is to be noted that Lewis and Naylor² report the difficult oxidation of this product to acidic substances which also suggests the cyclic structure.

It has been the purpose of this work now being reported, to establish the structure of C_3F_6 firmly and to correct any errors which have appeared in the literature.

A sample of C_3F_6 prepared by pyrolysis of tetrafluoroethylene polymer was compared with a sample of C_3F_6 prepared by the dechlorination of $CF_3CFCICF_2Cl$ with zinc and alcohol.³ The marked similarity of these products is shown in Table I.

TABLE I

Product source	"Teflon" pyrolysis	$CF_3CFCICF_2Cl$ dechlorination
Molecular weight from vapor density	151	153
Boiling point, $^\circ C$.	- 29.8	- 29.6
Freezing point, $^\circ C$.	-156.2	-156.5
Mixed freezing point, $^\circ C$.		156.6

Not only is the mixed freezing point datum excellent evidence for the identity of these products (see Fig. 1), but also the infrared absorption curves of these two preparations were found to be identical. The curve for C_3F_6 is given in Fig. 2. The maximum at 5.55 microns is a positive indication of the existence of a double bond in the molecule.

Chlorination in light of the two C_3F_6 products (A) from polymer and (B) from $CF_3CFCICF_2Cl$ gave the corresponding dichlorides which were identical to each other and different from the dichloride obtained (C) by chlorination of $H(CF_2)_3-$

- (1) Benning, Downing and Park, U. S. Patent 2,394,581.
- (2) Lewis and Naylor, THIS JOURNAL, **69**, 1968 (1947).
- (3) Henne and Waalkes, *ibid.*, **68**, 496 (1946).