

Six-Coordinate Thallium(III) Porphyrin Triflate: Synthesis, Physiochemical Characteristics, and X-ray Crystal Structure of (Trifluoromethanesulfonato)(tetrahydrofuran)-(meso-tetraphenylporphyrinato)thallium(III) Tetrahydrofuran Solvate, [Tl(tpp)(OSO₂CF₃)(THF)·THF]

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Introduction

As widely recognized, the trifluoromethanesulfonate anion (CF₃SO₃[−], triflate) is an excellent leaving group in nucleophilic substitution reactions in organic chemistry.¹ Lemke *et al.*² investigated the ¹H, ¹⁹F, and X-ray structure of bis(trifluoromethanesulfonato)(tetra-*p*-porphyrinato)silicon(IV), Si(ttp)-(OSO₂CF₃)₂, complex. Arnold *et al.*³ elucidated the diaquo complex [Sn(tpp)(H₂O)₂](OSO₂CF₃)₂, indicating that the triflates were disordered in the unit cell with reasonably close O—O contacts. Goff *et al.*⁴ closely examined the synthesis and ¹⁹F, ¹H, and ¹³C characterization of (trifluoromethanesulfonato)-(tetraphenylporphyrinato)iron(III), Fe(tpp)(CF₃SO₃). Guillard *et al.*⁵ investigated the synthesis and ¹H characterization of (trifluoromethanesulfonato)(tetraphenylporphyrinato)gallium(III), Ga(tpp)(CF₃SO₃), and (2,3,7,8,12,13,17,18-octaethylporphyrinato)(trifluoromethanesulfonato)gallium(III), Ga(oep)(CF₃SO₃). Although the triflate (CF₃SO₃[−]) of Si(ttp)(CF₃SO₃)₂ is unidentately coordinated to the silicon atom in solid-state, according to a previous investigation, the complex Si(ttp)(CF₃SO₃)₂ is converted into Si(ttp)(OH)₂ and triflic acid (CF₃SO₂OH) when exposed to water in a solution.² Similar to Si(ttp)(CF₃SO₃)₂, many metallocporphyrin complexes [M(por)(CF₃SO₃)_{*n*}, *n* = 1, 2] do not exist for a long time outside the solid state and dissociate or react readily in a solution. Most previous studies neglect the ¹³C and ¹⁹F NMR spectra of triflate in the complexes. In this work, we largely focus on the coordination of the triflate onto the thallium porphyrin complexes. The distinction between the covalently bound CF₃SO₃

and ionic CF₃SO₃[−] for a novel complex (trifluoromethanesulfonato)(tetrahydrofuran)(*meso*-tetraphenylporphyrinato)thallium(III) tetrahydrofuran solvate, [Tl(tpp)(OSO₂CF₃)(THF)·THF] (**1**), in acetone-*d*₆ can be drawn from the ¹³C and ¹⁹F NMR spectroscopy studies of triflate ligand at 24 and −80 °C.

In light of above developments, this work synthesizes and characterizes a six-coordinate thallium(III) porphyrin complex (**1**) which appears to be the first complex with a transoid geometry. The structure of this complex and comparisons with related cisoid thallium porphyrin structures provide further insight into the area of main group porphyrin chemistry. Meanwhile, the ¹⁹F and ¹³C NMR spectra of (**1**) in a solid state at 24 °C and in acetone-*d*₆ at low temperature are examined to identify the bound triflate in a solid and the ionic CF₃SO₃[−] in solution.

Experimental Section

Preparation of [Tl(tpp)(OSO₂CF₃)(THF)·THF] (1**).** Tetrahydrofuran (15 mL) was added to a flask containing Tl(tpp)Cl (100 mg, 0.117 mmol) and Ag(CF₃SO₃) (60 mg, 0.234 mmol). The mixture was stirred for 2 h and then filtered through Celite. The deep red filtrate was evaporated to dryness to yield (**1**) (88.2 mg, 0.098 mmol, 84%). **1** was dissolved in THF-*d*₈ (99.5% from Aldrich) for ¹H NMR measurement at 24 °C. ¹H NMR, δ (ppm): 9.11 [d, β -pyrrole, ⁴J(Tl—H) = 88.4 Hz], 8.26 (m, *o*-H), 7.80 (m, *m*, *p*-H). The crystals were grown from THF solution. MS, *m/z* (assignment, rel intensity): 966 ([Tl(tpp)-(SO₃CF₃)]⁺, 11.54), 817 ([Tl(tpp)]⁺, 100.00), 615 (H₂tpp⁺, 70.57). UV/visible spectrum λ (nm) ($\epsilon \times 10^{-3}$ (M^{−1} cm^{−1})) in acetone (or THF): 332 (11.2), 410 (19.8), 430 (206.6), 530 (2.2), 564 (7.5), 604 (3.9).

Spectroscopy. Proton and ¹³C NMR spectra in acetone-*d*₆ (99.96% from Aldrich) or THF-*d*₈ were recorded at 400.13 (or 600.20) and 100.61 (or 150.92) MHz, respectively, on Bruker AM-400 (or DMX-600) spectrometer locked on solvent deuterium, and referenced to the solvent peak. ¹⁹F NMR spectra were measured in acetone-*d*₆ (or THF-*d*₈) at 282.40 MHz on Bruker MSL-300 spectrometer. Proton and ¹³C NMR are relative to acetone-*d*₆ or THF-*d*₈ at δ = 2.04 or 3.58 (the downfield resonance) and the center line of acetone-*d*₆ or THF-*d*₈ at δ = 29.8 (CH₃) or 67.4 (the downfield resonance). ¹⁹F data are externally relative to CFCl₃. Next, the temperature of the spectrometer probe was calibrated by the shift difference of methanol resonance in the ¹H NMR spectrum. The solid-state ¹³C CP/MAS and ¹⁹F MAS spectra were recorded at 24 °C at 50.33 and 470.6 MHz, respectively, on Bruker MSL-200 and Bruker MSL-500 spectrometer. In addition, dry nitrogen gas was used to drive the MAS rates of 3.7 kHz for ¹³C and 12 kHz for ¹⁹F.

The positive-ion fast atom bombardment mass spectrum (FABMS) was obtained in a nitrobenzyl alcohol (NBA) matrix using a JEOL JMS-SX/SX 102A mass spectrometer. UV/visible spectra were recorded at 24 °C on a Hitachi U-3210 spectrophotometer.

Crystallography. Table 1 presents the crystal data and other information for (**1**). X-ray structure was measured on a Siemens SMART CCD diffractometer using monochromatized Mo K α radiation (λ = 0.710 73 Å). Absorption corrections were based on 3127 symmetry-equivalent reflections using the SHELXTL-PC program package with (*T*_{min,max} = 0.647, 0.943). The structures were solved by direct methods (SHELXTL PLUS) and refined by full-matrix least-squares. All non-hydrogen atoms were refined with anisotropic thermal parameters, whereas all hydrogen atom positions were calculated using a riding model and included in the structure factor calculation. Table 2 lists selected bond distances and angles.

Results and Discussion

Molecular structure of [Tl(tpp)(OSO₂CF₃)(THF)·THF] (1**).** Figure 1 illustrates the skeletal framework of complex **1**.

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Table 1. Crystal Data for Compound $\text{Ti}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})\cdot\text{THF}$ (**1**)

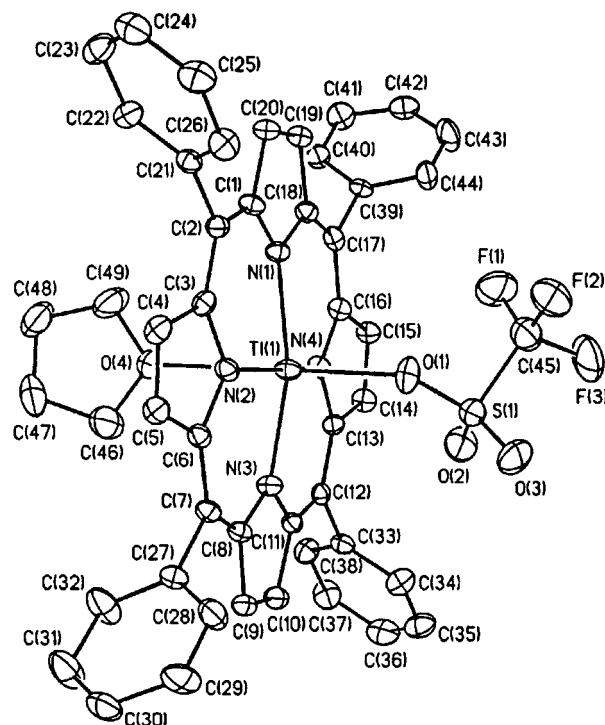
empirical formula	$\text{C}_{49}\text{H}_{36}\text{F}_3\text{N}_4\text{O}_4\text{STl}\cdot(\text{C}_4\text{H}_8\text{O})$
fw	1110.4
space group	<i>Cc</i>
cryst syst	monoclinic
cryst color, habit	blue, columnar
<i>a</i> , Å	19.2967(1)
<i>b</i> , Å	11.0937(1)
<i>c</i> , Å	23.1386(3)
β , deg	111.702(1)
<i>V</i> , Å ³	4602(1)
<i>Z</i>	4
<i>D</i> _{calcd} , g cm ⁻³	1.602
$\mu(\text{Mo K}\alpha)$, cm ⁻¹	36.21
<i>S</i>	1.07
cryst size, mm ³	0.28 × 0.38 × 0.58
$2\theta_{\text{max}}$, deg	51.1
<i>T</i> , K	296
no. of reflns measd	9817
no. of reflns obsd (<i>I</i> ≥ 3.0σ(<i>I</i>))	4641
<i>R</i> ^a (%)	2.98
<i>R</i> _w ^b (%)	4.13

^a $R = [\sum ||F_o| - |F_c||] / \sum |F_o|$. ^b $R_w = [\sum w(|F_o| - |F_c|)^2]^{1/2} / \sum w(|F_o|)^2$; $w = 1/(\sigma^2 F_o + BF_o^2)$.

Table 2. Selected Bond Distances (Å) and Angles (deg) for $\text{Ti}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})\cdot\text{THF}$ (**1**)

Distances			
Tl(1)–O(1)	2.387(6)	Tl(1)–N(1)	2.137(8)
Tl(1)–O(4)	2.778(7)	Tl(1)–N(2)	2.140(7)
O(1)–S(1)	1.455(8)	Tl(1)–N(3)	2.150(8)
O(2)–S(1)	1.42(1)	Tl(1)–N(4)	2.153(6)
O(3)–S(1)	1.416(9)	C(45)–F(1)	1.31(2)
S(1)–C(45)	1.77(2)	C(45)–F(2)	1.32(3)
		C(45)–F(3)	1.38(2)
Angles			
Tl(1)–O(1)–S(1)	129.6(5)	S(1)–C(45)–F(1)	113(1)
O(1)–S(1)–O(2)	113.8(5)	S(1)–C(45)–F(2)	112(1)
O(1)–S(1)–O(3)	116.0(5)	S(1)–C(45)–F(3)	108(2)
O(1)–S(1)–C(45)	100.7(8)	F(1)–C(45)–F(2)	109(2)
O(2)–S(1)–C(45)	105.2(8)	F(1)–C(45)–F(3)	108(1)
O(3)–S(1)–C(45)	104.1(7)	F(2)–C(45)–F(3)	107(1)
O(2)–S(1)–O(3)	114.7(7)	N(1)–Ti(1)–N(2)	89.2(3)
O(4)–Ti(1)–N(1)	83.0(3)	N(1)–Ti(1)–N(3)	164.6(2)
O(4)–Ti(1)–N(2)	78.6(2)	N(1)–Ti(1)–N(4)	88.2(3)
O(4)–Ti(1)–N(3)	81.7(3)	N(2)–Ti(1)–N(3)	89.2(3)
O(4)–Ti(1)–N(4)	87.5(2)	N(2)–Ti(1)–N(4)	166.1(2)
O(1)–Ti(1)–N(1)	97.2(3)	N(3)–Ti(1)–N(4)	89.7(3)
O(1)–Ti(1)–N(2)	95.0(2)	O(1)–Ti(1)–O(4)	173.6(3)
O(1)–Ti(1)–N(3)	98.2(3)	Tl(1)–O(4)–C(46)	124(1)
O(1)–Ti(1)–N(4)	98.9(2)	Tl(1)–O(4)–C(49)	124.7(8)

It is an octahedral complex of the porphyrin N_4 , one triflate oxygen, and one THF oxygen. Bond distances are $\text{Ti}(1)\text{--O}(1) = 2.387(6)$ Å, $\text{Ti}(1)\text{--O}(4) = 2.778(7)$ Å, and the mean $\text{Ti}(1)\text{--N} = 2.145(7)$ Å. The angle between the two axial oxygen ligands is nearly linear with $\text{O}(1)\text{--Ti}(1)\text{--O}(4) = 173.6(3)^\circ$. The thallium–O(4) distance of $2.778(7)$ Å is long, but just within the sum of van der Waals radii of thallium and oxygen, ~ 3.5 Å. This longer $\text{Ti}\cdots\text{O}(4)$ contact is described as the effective coordination (or as “semicoordinated”).^{6,7} The short $\text{Ti}(1)\text{--O}(1)$ distance and the displacement of the thallium atom toward the triflate are consistent with a much stronger metal ligand bond to the triflate anion than to the THF molecule. This type of asymmetry in bond lengths was also observed for $\text{Fe}(\text{tpp})(\text{OTeF}_5)(\text{THF})$ with $\text{Fe}\text{--O}(\text{OTeF}_5) = 1.967$ Å and $\text{Fe}\text{--O}(\text{THF}) = 2.334$ Å.⁸ In compound **1**, the interaction of

**Figure 1.** Molecular configuration and atom labeling scheme for $\text{Ti}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})\cdot\text{THF}$ (**1**). Hydrogen atoms and the free THF (or noncoordinated THF) are omitted for clarity.

CF_3SO_3^- with thallium is purely unidentate, the second (or the third) triflate oxygen i.e. O(2) [or O(3)] being 3.592 (or 4.594 Å) from the thallium atom. The triflate is monodentate both in **1** and in $\text{Si}(\text{tpp})(\text{OSO}_2\text{CF}_3)_2$.²

The displacements of the metal center from the planes of the macrocycle atoms (C_{20}N_4) and the four porphyrin nitrogens are labeled as $\Delta 24$ and $\Delta 4\text{N}$, respectively. The Tl atom is 0.295/0.273 (i.e. $\Delta 24/\Delta 4\text{N}$) out of the porphyrin or N_4 plane toward the triflate oxygen. The dihedral angles between the mean plane (C_{20}N_4) and the plane of the phenyl group are 69.4° [C(24)], 82.7° [C(30)], 104.2° [C(36)], and 100.3° [C(42)]. The radius of the central “hole” ($\text{C}_1'\cdots\text{N}$, the distance from the geometrical center (C_1') of the mean plane of the 24-atom core to the porphyrinato-core N atoms) in **1** is 2.128 Å, which is larger than 2.01 Å as suggested by Collin and Hoard.⁹ The thallium(III) is bonded in a highly expanded porphyrinato core (C_{20}N_4) and the porphyrin (C_{20}N_4 and Tl) can be viewed as a MOOP (metal out-of-plane) shape.¹⁰ Figure 2 illustrates the displacement (in Å) of each atom of the porphyrin (C_{20}N_4 and Tl) from the porphyrin mean plane (C_{20}N_4).

In general, the coordination sphere of the Ti^{3+} ion for the thallium porphyrins is an approximate square-based pyramid in which the apical site is occupied either by a bidentate group for complexes $\text{Ti}(\text{tpp})(\text{O}_2\text{CCF}_3)$, $\text{Ti}(\text{tmpp})(\text{C}_6\text{H}_5\text{CO}_2)$, and $\text{Ti}(\text{por})(\text{OAc})$, with por = tpp, tmpp (5,10,15,20-tetra(4-methoxyphenyl)porphyrinate), tpyp (5,10,15,20-tetra(4-pyridyl)porphyrinate), or by a monodentate group for complexes $[\text{Ti}(\text{por})\text{X}]$ with the axial ligand X = CN, Cl, CH_3 and por = tpp, tmpp, tpyp (Table 3). These thallium porphyrins generally show $\Delta 24$ of 0.74–1.06 Å. However, the compound **1** is a distorted octahedron with O(1) (of triflate) and O(4) (of THF) in the trans position. To coordinate the O(1) of triflate and semicoordinate

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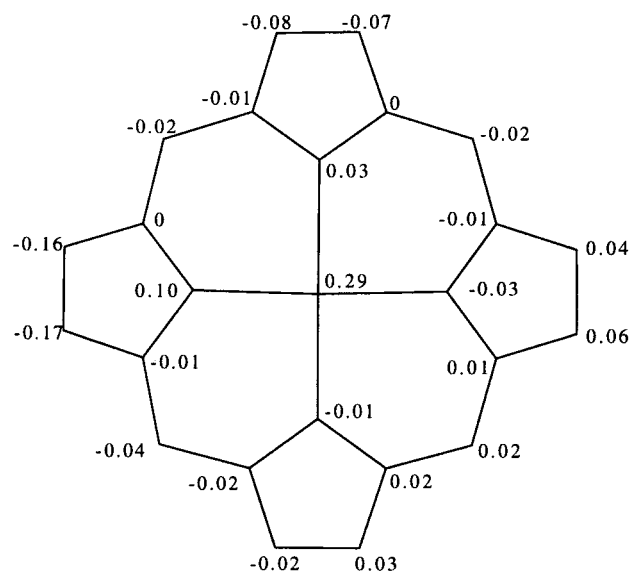
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Table 3. Stereochemical Parameters for Thallium(III) Porphyrin Complexes

compound	$\Delta 24$ (Å)	binding of axial ligand (found)	distortion ($C_{20}N_4$ and Tl)	core size $C_t \cdots N$ (Å)
Tl(tpp)(O ₂ CCF ₃) ¹²	0.741	bidentate	domed	2.092
Tl(tpp)(OAc) ^{19,20}	0.842	bidentate	domed	—
Tl(tmpp)(OAc) ²⁴	0.808	bidentate	domed	2.084
Tl(tpyp)(OAc) ^{25,26}	0.858	bidentate	domed	—
Tl(tmpp)(C ₆ H ₅ CO ₂) ²⁷	0.811	bidentate	domed	2.096
Tl(tpp)(CN) ²⁸	0.908	monodentate	domed	—
	1.027			
Tl(tmpp)(CN) ¹⁷	0.86	monodentate	domed	2.085
Tl(tpyp)(CN) ²⁹	0.825	monodentate	domed	—
	0.900			
Tl(tpp)Cl ³⁰	0.737	—	domed	—
Tl(tmpp)Cl ¹⁷	0.82	—	domed	2.095
Tl(tpyp)Cl ³¹	0.826	—	domed	2.091
	0.85			
Tl(tpp)CH ₃ ³⁰	0.979	—	domed	—
Tl(tpyp)CH ₃ ³²	1.059	—	domed	—
	0.975			
Tl(tpp)(OSO ₂ CF ₃)(THF)·THF	0.295	monodentate	MOOP	2.128

**Figure 2.** Diagram of the porphyrin core ($C_{20}N_4$ and Tl) of compound **1** showing the displacement (in Å) of the atoms from the mean plane of the porphyrin ($C_{20}N_4$).

the O(4) of THF simultaneously, the Tl^{3+} atom closely approaches the C_t' of $C_{20}N_4$, resulting in the smaller $\Delta 24$ of 0.295 Å for compound **1**. The displacement of the thallium atom depends at least partially on the relative importance of the binding of the two dissimilar axial ligands for the six-coordinate metal porphyrin complexes.⁷ Some measurements are taken of the asymmetry in axial ligand binding with respect to the difference in axial bond lengths (Δl). Interestingly, increasing Δl from 0 (for the equivalent axial ligand, i.e. [Fe(oep)(THF)₂]⁺ClO₄[−])¹¹ and 0.367 (for Fe(tpp)(OTeF₅)(THF))⁸ to 0.391 Å (Tl(tpp)(OSO₂CF₃)(THF)) causes an increase of the $\Delta 24$ from 0 and 0.20 to 0.295 Å.

¹³C and ¹⁹F for Tl(tpp)(OSO₂CF₃)(THF)·THF (1**) in Acetone-*d*₆.** Results obtained from the ¹³C NMR studies of **1** in acetone-*d*₆ indicate that **1** ionizes completely and simply *via*

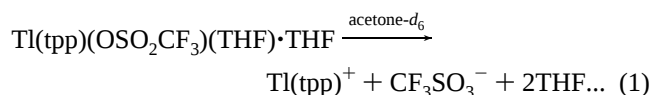


Table 4 summarizes ¹³C data for **1** in acetone-*d*₆ at 24 and −80 °C. Eight major resonances with a α carbon (C_α) at δ = 149.8, C_1 at δ = 141.3, $C_{2,6}$ at δ = 135.5, C_β at δ = 134.1, C_4 at δ = 129.2, $C_{3,5}$ at δ = 127.9, *meso* carbon (C_m) at δ = 122.7 and the CF₃SO₃ carbon at δ = 121.2 ppm were observed at −80 °C. Figure 3 shows the ¹³C spectrum for **1** in acetone-*d*₆ at −100 °C. The ¹³C resonance of CF₃SO₃ for compound **1**, as obtained by averaging over 24 and −80 °C at δ = 121.5 ± 0.3 ppm with ¹J(C–F) coupling constant 321 Hz, resembles that of ionic CF₃SO₃[−] (or free CF₃SO₃[−]) with δ = 120.8 ± 0.2 and ¹J(C–F) = 322 Hz at room temperature (Table 4). ¹⁹F NMR spectroscopy was used to examine the binding of CF₃SO₃[−] to Tl(tpp)⁺. The fluorine resonance for [Bu₄N⁺][CF₃SO₃[−]] in chloroform is found at −78.7 ppm (where CFCl₃ is used as an external reference) at room temperature.⁴ The ¹⁹F signal for compound **1** at 24 °C is −78.7 ppm (in THF-*d*₈) or −78.1 ppm (in acetone-*d*₆) resembles that in ionic CF₃SO₃[−]. Both ¹³C and ¹⁹F data indicate that only the free CF₃SO₃[−] is detected for **1** in acetone-*d*₆. Although **1** is a THF solvate, dissolution in acetone-*d*₆ could produce Tl(tpp)(THF)₂⁺, or the acetone adduct Tl(tpp)(THF)(OC(CD₃)₂)⁺. However, when **1** is dissolved in acetone-*d*₆, the NMR resonances of THF downfield from Me₄Si are observed with similar chemical shifts at 67.7 and 25.9 ppm for ¹³C at −100 °C (Figure 3a), at 67.9 and 26.3 ppm for ¹³C at 24 °C, and at 3.59 and 1.77 ppm for ¹H at −100 and 24 °C. These resonances confirm that all THF is free THF. In addition, these data do not support the exchange between a small amount of coordinated THF and free THF for **1** in acetone-*d*₆. Meanwhile, the ¹³C and ¹H NMR resonances of acetone downfield from TMS for **1** in acetone-*d*₆ at −100 and 24 °C are observed with the same chemical shifts at 206.2 and 29.8 ppm for ¹³C (Figure 3a) and at 2.04 ppm for ¹H, respectively. These latter resonances indicate that all the acetone-*d*₆ is free deuterated acetone. Notably, all THF and acetone-*d*₆ are free when **1** is dissolved in acetone-*d*₆, implying that the cationic thallium species Tl(tpp)⁺ is four-coordinate. The results above support the proposed dissociation of compound **1** in acetone-*d*₆ as shown in eq 1.

Solid-State CP/MAS ¹³C and MAS ¹⁹F for Compound (**1**).

Table 4 presents the solid-state ¹³C CP/MAS data of compound **1** at 24 °C. Five major resonances with C_α at δ = 150.5, C_1 at 141.2, $C_{2,6}$ and C_β at δ = 133.5, C_4 and $C_{3,5}$ at 128.6, and C_m and CF₃SO₃ at 121.2–122.7 ppm were observed. Due to the overlap of CF₃SO₃ with C_m , the unambiguous assignment of the carbon chemical shift for a coordinated triflate is impossible. However, the ¹⁹F chemical shifts of the triflate ions are detected

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Table 4. ^{19}F and ^{13}C Chemical Shifts (δ) and Tl– ^{13}C Coupling Constant (J in Hz) for Tl(tpp)(OSO₂CF₃)(THF)·THF (**1**) in Acetone-*d*₆^a

compound	solvent (carbon freq.)	<i>T</i> (°C)	<i>C</i> _α	<i>C</i> ₁	<i>C</i> _{2,6}	<i>C</i> _β	<i>C</i> ₄	<i>C</i> _{3,5}	<i>C</i> _m	CF ₃ SO ₃	^{19}F
Tl(tpp)(OSO ₂ CF ₃)(THF)·THF (1)	acetone- <i>d</i> ₆ (100.61 MHz)	24	150.5 (7)	141.9 (19)	135.5; 135.4	133.7 (170)	129.2	127.9 (166)	123.2	121.7 (320) ^b	−78.1
	acetone- <i>d</i> ₆ (150.89 MHz)	−80	149.8	141.3	135.5	134.1 (179)	129.2	127.9 (166)	122.7	121.2 (321) ^b	−78.4
(1)	solid (50.33 MHz)	24	150.5	141.2	133.5		128.6		122.7–121.1 (m)		−74.0 ^d
CF ₃ SO ₃ [−] Na ⁺	DMSO- <i>d</i> ₆	room temp	—	—	—	—	—	—	—	120.6 ³³ (322)	—
Bu ₄ N ⁺ CF ₃ SO ₃ [−]	CDCl ₃	room temp	—	—	—	—	—	—	—	121.0 ³³ (321)	−78.7 ⁴

^a Chemical shifts in ppm relative to the methyl carbon of acetone-*d*₆ at δ = 29.8. Values in parentheses beneath are $J(\text{Tl}–^{13}\text{C})$ coupling constants in Hz unless specified. ¹⁹F was measured at 282.40 MHz unless specified. ^b $^1J(\text{C}–\text{F})$ coupling constant. ^c It was measured in THF-*d*₈. ^d Solid-state ¹⁹F was measured at 470.6 MHz.

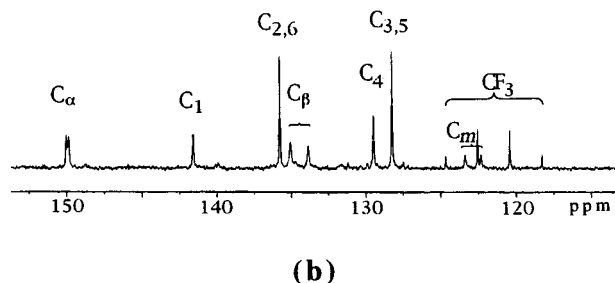
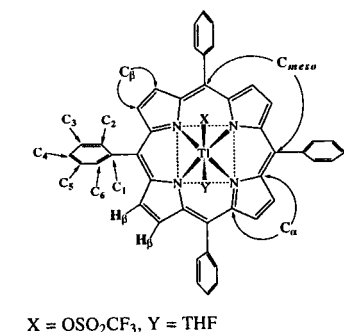
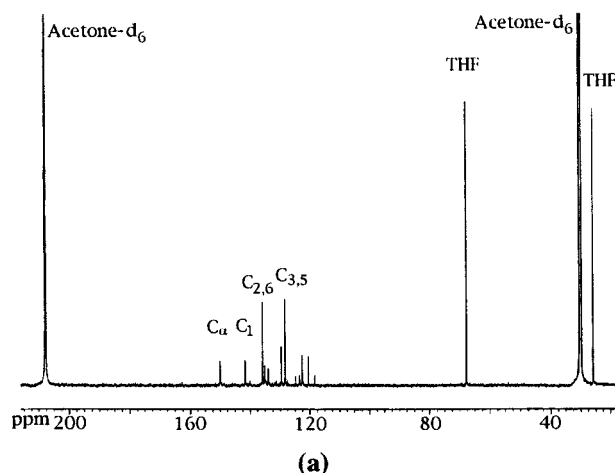


Figure 3. ^{13}C broad band NMR spectra for Tl(tpp)(OSO₂CF₃)(THF)·THF (**1**) in acetone-*d*₆ (150.92 MHz) at $-100\text{ }^{\circ}\text{C}$: (a) entire spectrum; (b) expansion of the region of 115–155 ppm in (a). TMS is used as internal reference.

when they are coordinated by the solid state ^{19}F NMR measurement. The resonance of the solid ^{19}F of compound **1** occurs as a broad singlet (line width 1700 Hz, ~ 3.6 ppm) at -74.0 ppm at $24\text{ }^{\circ}\text{C}$, which is ~ 4.7 ppm downfield of the fluorine resonance for [Bu₄N⁺][CF₃SO₃[−]]. Moreover, we cannot resolve the long-range coupling of ^{19}F to thallium, with

a typical $^4J(\text{Tl}–\text{F})$ value of $\sim 29.2 \pm 0.6$ Hz,¹² for a coordinated triflate. This is despite the fact that the CF₃ group in **1** lies within the shielding region of **1**, which normally leads to upfield shifts. However, the downfield shift of ~ 4.7 ppm for F in the solid-state NMR is due to the shift of ^{19}F in **1** being controlled by the paramagnetic term,^{13–16} not by the ring current effect. This paramagnetic contribution is a shift toward the lower field, which is the largest for covalent bonds of CF₃SO₃ in **1**. Our earlier work observed a similar downfield shift of ~ 3 ppm [from -78.07 (obtained from CF₃COOH) to -75.09 ppm] for ^{19}F in a 0.02 M solution of Tl(tmpp)(O₂CCF₃) in CD₂Cl₂ at $-70\text{ }^{\circ}\text{C}$.¹²

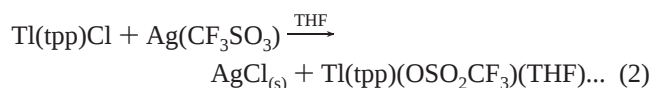
UV/Visible Absorption Spectra for Compound 1. The electronic spectrum of **1** in acetone (or THF) and the cisoid six-coordinate thallium porphyrin complexes (e.g. Tl(tpp)(O₂CCF₃) and Tl(tpp)(OAc)) closely resemble with electronic absorption at 430 (B(0,0)), 564 (Q(1,0)), and 604 (Q(0,0)) for **1** versus 432, 566, and 606 nm for Tl(tpp)(O₂CCF₃)^{12,17} and 433, 567, and 607 nm for Tl(tpp)(OAc).^{18–20} This finding is due to the former spectrum being the spectrum of Tl(tpp)⁺. Notably, the four-coordinate Tl(tpp)⁺ and the cisoid six-coordinate thallium porphyrin complexes belongs to regular (or “normal”) metalloporphyrins containing closed-shell Tl³⁺ ions (d^{10}), in which the d_{π} (d_{xz} , d_{yz}) metal-based orbitals are relatively low in energy.^{21,22} These low-energy orbitals only slightly influence the porphyrin π -to- π^* energy gap in porphyrin electronic spectra. In contrast to most transition metal complexes, their color is due to absorption within the porphyrin ligand involving the excitation of electrons from π -to- π^* porphyrin ring orbitals. When Tl binds to the porphyrin in Tl(tpp)⁺ or in cisoid six-coordinate thallium porphyrin complexes, the absorption spectrum changes due to symmetric effects. However, the π -to- π^* energy gap is negligibly affected

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and a regular metallorporphyrin spectrum results. This phenomenon accounts for why variations of the axial ligand in $\text{Tl}(\text{tpp})^+$ or in cisoid six-coordinate thallium complexes do not significantly affect the electronic spectra, although minor change are apparent.¹⁸

Conclusions

This investigation synthesizes and characterizes a new thallium(III) porphyrin complex $\text{Tl}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})\cdot\text{THF}$ (**1**). Metathesis reaction of $\text{Tl}(\text{tpp})\text{Cl}$ with $\text{Ag}(\text{CF}_3\text{SO}_3)$ in THF solvent provides the convenient route to $\text{Tl}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})$:



$\text{Tl}(\text{tpp})(\text{OSO}_2\text{CF}_3)(\text{THF})$ is formed by eq 2 and dissociates by eq 1. This compound has little or no stability outside the crystal lattice; its existence is directly attributed to the solid-state lattice effects.²³

The typical thallium(III) porphyrin is cisoid five- or six-coordinate with a large thallium atom displacement (0.74–1.06 Å) from a modestly domed porphyrinato core (with an average $\text{C}_1'\cdots\text{N}$ of 2.091 Å). The complex **1** studied herein is the first structurally characterized transoid six-coordinate thallium(III) porphyrin complex with the thallium atom locating (0.295 Å) slightly out of the plane toward $\text{OSO}_2\text{CF}_3^-$ and having a radially expanded core (with an average $\text{C}_1'\cdots\text{N}$ of 2.128 Å). ^1H , ^{13}C , and ^{19}F of the triflate and THF for compound **1** in acetone- d_6 (THF- d_8) at -80 and 24 °C confirm that complex **1** ionizes completely to $\text{Tl}(\text{tpp})^+$ and ionic CF_3SO_3^- in acetone- d_6 (or THF- d_8).

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Supporting Information Available: Tables of structure determination summary, atomic coordinates, bond lengths, bond angles, anisotropic displacement, and H-atom coordinates coefficients and ORTEP diagrams for compound **1** (14 pages). An X-ray crystallographic file, in CIF format, for compound **1** is also available on the Internet only. Ordering and access information is given on any current masthead page.

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