

Synthesis and X-ray characterization of a monomeric Cp-free d¹-imido–vanadium(IV) complex

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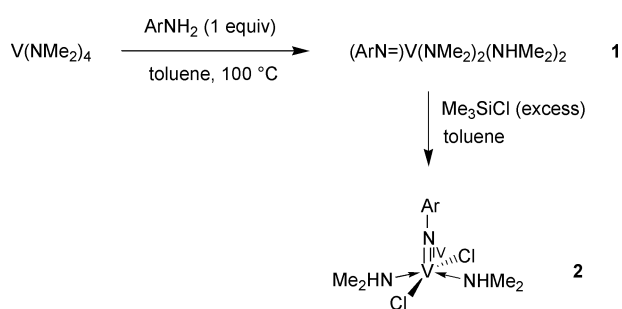
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Straightforward synthesis of the new imido complexes $\text{ArN}=\text{V}(\text{NMe}_2)_2(\text{HNMe}_2)_2$ **1 and $\text{ArN}=\text{VCl}_2(\text{HNMe}_2)_2$ **2** ($\text{Ar} = 2,6\text{-i-Pr}_2\text{C}_6\text{H}_3$) is described along with the molecular structure of **2** and preliminary ethylene polymerization catalytic studies.**

Few areas of inorganic chemistry have experienced such remarkable growth in recent years as that of transition metal imido chemistry.¹ In the case of imides of the Group 5 transition metals, the majority contain the d⁰-(V=NR) functional group. However, paramagnetic d¹-vanadium monoimides have been known for some time with stabilizing cyclopentadienyl ligands.² Nevertheless, there are only a few examples of isolated d¹-(V=NR) complexes that do not contain one or two cyclopentadienyl ligands, and, to our knowledge none of them have been structurally characterized,^{1,3–5} except the heptamer complex $[\text{t-BuNH}_3]_2[(\text{t-BuN})_7\text{V}_7(\mu\text{-Cl})_{14}\text{Cl}_2]$.⁶ Within this context, and in relation to our recent studies on bis-amido–vanadium(IV) complexes as catalysts for olefin polymerization,^{7,8} we wish to report the unprecedented synthesis and single crystal X-ray structure of a monomeric vanadium(IV) imido complex without Cp co-ligands.

Reaction of 2,6-isopropylaniline, ArNH_2 ($\text{Ar} = 2,6\text{-i-Pr}_2\text{C}_6\text{H}_3$), with $\text{V}(\text{NMe}_2)_4$ in toluene at 100 °C for 1 day affords the arylimido complex $\text{ArN}=\text{V}(\text{NMe}_2)_2(\text{NHMe}_2)_2$ **1** (see Scheme 1)



Scheme 1 Synthesis of complexes **1** and **2**.

as a dark red solid, after recrystallization from cold pentane. The bis-chloro analogue, namely $\text{ArN}=\text{VCl}_2(\text{HNMe}_2)_2$ **2**, is obtained by treatment of **1** with an excess of chlorotrimethylsilane at room temperature in toluene (Scheme 1, yield = 82% with respect to $\text{V}(\text{NMe}_2)_4$).^{†‡}

The crystal structure of **2** is presented in Fig. 1 and confirms the presence of a bis-dimethylamine adduct.§ The molecular structure is best described as distorted square-pyramid with an axial arylimido that exhibits a V–N distance of 1.654(3) Å, the imido linkage being almost linear (V1–N1–C1 angle = 178.4(3)°). The two chlorine atoms are mutually *trans* with a Cl1–V1–Cl2 angle of 136.48(5)° and V1–Cl bonds of 2.3202(14) and 2.3257(12) Å. The two *trans* dimethylamino ligands form the base of the square-pyramid with the chlorine

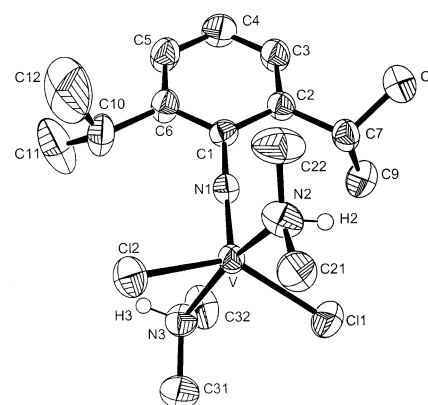


Fig. 1 Molecular structure of **2** with selected bond distances (Å) and angles (°): V1–N1 1.654(3), V1–N2 2.018(5), V1–N3 2.171(4), V1–Cl1 2.3202(14), V1–Cl2 2.3257(12), N1–C1 1.388(5), C1–N1–V1 178.4(3), Cl1–V1–Cl2 136.48(5), N2–V1–N3 165.23(16).

atoms, and have V–N_{amine} bonds of 2.018(5) and 2.171(4) Å and a N_{amine}–V–N_{amine} angle of 165.23(16)°.

Both **1** and **2** are paramagnetic; EPR spectra of toluene solutions at room temperature show typical eight band patterns for d¹-vanadium(IV) species, with $g = 1.985$, $a(^{51}\text{V}) = 65.3$ G and $g = 1.989$, $a(^{51}\text{V}) = 92.5$ G, respectively. Additionally, magnetic susceptibility measurements have been carried out on **2**. The effective moment μ_{eff} is 1.744 μ_{B} at 300 K, which is consistent with a V^{IV} unit.

Compound **2** may be the key starting material in a whole new coordination chemistry of vanadium. Preliminary studies show that HNMe_2 ligands in **2** can be displaced by amines, phosphines, and 2,2'-bipyridine. For example, after dissolution of **2** in pyridine, the complex $\text{ArN}=\text{VCl}_2(\text{Py})_3$ has been quantitatively synthesized and characterized by single crystal X-ray crystallography.⁸

The complexes **1** and **2** will provide interesting comparisons with their titanium/zirconium analogues⁹ and with very recent amido-,¹⁰ bis-amido-^{7,8} or linked Cp-amido-¹¹ vanadium(IV) complexes that are active olefin polymerization catalysts. Indeed, preliminary studies demonstrate that **1** is an active catalyst for ethylene polymerization when 10 equivalents of EtAlCl_2 are used as co-catalyst (activity = 2.2×10^4 (g of PE) (mol catalyst)^{–1} h^{–1} atm^{–1}). Similarly, a combination of **2**/MAO (ratio 1/500, MAO = methylaluminoxane) or **2**/EtAlCl₂ (ratio 1/10) catalyzes ethylene polymerization at room temperature and under 1 atm of ethylene (with activities of 5.9×10^4 and 1.2×10^5 (g of PE) (mol catalyst)^{–1} h^{–1} atm^{–1}, respectively). Solid polyethylene is obtained in all cases, and in the tests run with EtAlCl_2 as activator high molecular weight polyethylene was formed (see Table 1).

In summary, we have described a convenient synthesis of imido–vanadium(IV) complexes **1** and **2**, and, to our knowledge,

Table 1 Results of ethylene polymerisation runs using pre-catalysts **1** and **2**^a

Run	Precatalyst	Activator	Time/min	Polymer yield/mg	Productivity/(g of PE) (mol catalyst) ⁻¹ h ⁻¹ atm ⁻¹	M_n^d	M_w^d	M_w/M_n
1	1 ^b	EtAlCl ₂	3	35	2.2×10^4	6.285×10^5	1.795×10^6	2.875
2	2 ^c	MAO	10	254	5.9×10^4	6.490×10^3	2.146×10^4	3.306
3	2 ^b	EtAlCl ₂	2	98	1.2×10^5	6.505×10^5	1.933×10^6	2.972

^a General conditions: 0.025 mmol of pre-catalyst dissolved in 5 g of toluene, 20 °C, 1 atm of ethylene. ^b EtAlCl₂ 10 equiv. ^c MAO 500 equiv. ^d Determined by SEC analysis.

the first structurally characterized monomeric d¹-(V=NR) complex that does not possess Cp ligands. We have also demonstrated the activity of such imido complexes as catalysts for ethylene polymerization. Full details of the coordination chemistry, reactivity, and olefin polymerization activity of **1** and **2** and other analogues will be published shortly.

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Notes and references

† Direct synthesis of **2**: A toluene solution (4 g) of 1.00 g of V(NMe₂)₄ (4.40 mmol) and 780 mg of 2,6-i-Pr₂C₆H₃NH₂ (4.40 mmol) was heated at 100 °C for 1 day giving a dark red solution. After removal of the volatiles, toluene (4 g) and Me₃SiCl (4.00 g) were added and the solution was stirred for 2 days. Removal of the volatiles and recrystallization in cold pentane afforded 1.40 g of dark red-purple crystals of **2** (82%). These compounds release NHMe₂ during standard combustion analysis and therefore satisfactory (C,H,N) elemental analyses have not been obtained (and correspond to partial (for **2**) or total (for **1**) loss of the NHMe₂ ligands). When **2** is reacted with pyridine however, an exchange reaction occurs leading to red microcrystals of ArN=VCl₂(Py)₃. The pyridine ligands of this material are more strongly bound to the metal, permitting a satisfactory elemental analysis to be obtained.

‡ Selected data for **1** and **2**: **1** EPR (PhCH₃, 20 °C) $g = 1.985$, $a(^{51}\text{V}) = 65.3$ G, IR (KBr) $\nu_{\text{V=N}} = 952$ cm⁻¹. **2** EPR (PhCH₃, 20 °C) $g = 1.989$, $a(^{51}\text{V}) = 92.5$ G. $\mu_{\text{eff}} = 1.744\mu_{\text{B}}$ (300 K), IR (KBr) $\nu_{\text{V=N}} = 934$ cm⁻¹.

§ Crystallographic data for **2**: C₁₆H₃₁Cl₂N₃V, $M = 387.28$, monoclinic, space group $P2_1/n$, $a = 10.6595(13)$, $b = 9.7366(10)$, $c = 20.783(2)$ Å, $\beta = 103.099(14)^\circ$, $U = 2100.9(4)$ Å³, $Z = 4$, $D_c = 1.224$ Mg m⁻³, R (R_w) = 0.0905 (0.1374) for 2726 unique data and 213 parameters,

GOF = 1.067. Data collection were performed at ca. 160 K on a STOE Imaging Plate Diffraction System (I.P.D.S) diffractometer using graphite monochromatized Mo-K α radiation. The structure was solved by direct methods and subsequent difference Fourier maps. CCDC reference number 186/2265. See <http://www.rsc.org/suppdata/dt/b0/b008761k/> for crystallographic files in .cif format.

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