

# Bidentate metalloligands containing the phosphinocyclopentadienyl ligand. New heterobimetallic complexes and crystal structure of $(C_5Me_5)(C_5H_4PPh_2)Ta(CO)(\mu-PPh_2)Cr(CO)_4$

Cyril Poulard, Gilles Boni, Philippe Richard and Claude Moïse\*

Laboratoire de Synthèse et d'Electrosynthèse Organométalliques (LSEO - UMR CNRS 5632),  
Faculté des Sciences Gabriel, 6 Bd Gabriel, 21000 Dijon, France.  
E-mail: claudemoise@u-bourgogne.fr

Received 6th May 1999, Accepted 30th June 1999

Four new kinds of metalloligands  $Cp^*(C_5H_4PPh_2)TaL_n$  [ $L_n = (H)_3$ ,  $(H)(CO)$ ,  $(H)_2(PPh_2)$ ,  $(CO)(PPh_2)$ ] **2**, **3**, **4b**, **5b** containing the diphenylphosphinocyclopentadienyl ligand have been prepared and characterized. The trihydride complex **2** was obtained by reduction of the dichloride compound  $Cp^*(C_5H_4PPh_2)TaCl_2$  **1**. Treatment of **2** under a CO atmosphere at 140 °C in decane affords the monohydride complex **3**. Compounds **2** and **3** easily react with chlorophosphines leading to phosphido derivatives **4b** and **5b** respectively. All these metalloligands act as monodentate and/or chelating bidentate ligands. Their reactivity toward unsaturated organometallic fragments  $[M'(CO)_5]$  or  $[M'(CO)_4]$  ( $M' = Cr, W$ ) has been studied. The X-ray crystal structure of  $Cp^*(C_5H_4PPh_2)Ta(CO)(\mu-PPh_2)Cr(CO)_4$  **10Cr** is reported.

The study of heterobimetallic complexes is of great interest in the main fields of the application of organometallic chemistry, namely catalysis.<sup>1</sup> Part of the increasing interest in this research is the hope that such heterobimetallic structures might exhibit unusual catalytic activity due to a cooperative reactivity between electron-deficient and electron-rich metals.<sup>2</sup>

Bridging ligands have been widely developed to maintain the two metal sites in a close proximity and to prevent breaking of the bimetallic framework. Among these linking ligands, the heterodifunctional diphenylphosphinocyclopentadienyl has received significant attention in the last few years and numerous examples of its ability to join two different metallic centers have been reported.<sup>3</sup>

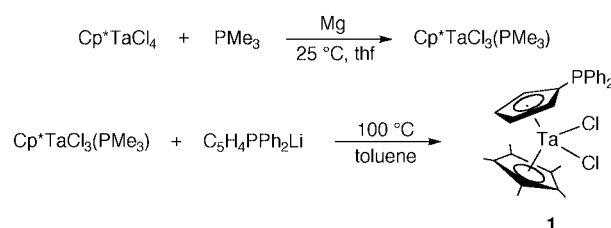
In the course of our work on binuclear complexes, we have already reported various hydrido and/or phosphido bridged systems. A series of bimetallic polyhydrides<sup>4</sup> and of  $\mu$ -hydrido  $\mu$ -phosphido cyclic structures<sup>5</sup> has been obtained which prove hydrido and phosphido bridges to be efficient in tying two metal centers.

In this paper, we report on the combination of the coordinating ability of a diphenylphosphinocyclopentadienyl ligand and that of a terminal phosphido or a hydrido functionality in the synthesis of new bimetallic tantalum–chromium and tantalum–tungsten complexes. The synthesis of the bidentate metalloligands proceeds in two main steps: firstly the diphenylphosphinocyclopentadienyl ligand is involved in the preparation of a Cp-phosphino functionalised tantalocene dichloride. Then, the hydrido or phosphino coordination site is introduced on the tantalum center by chemical transformations. The X-ray crystal structure of a tantalum–chromium compound is reported.

## Results and discussion

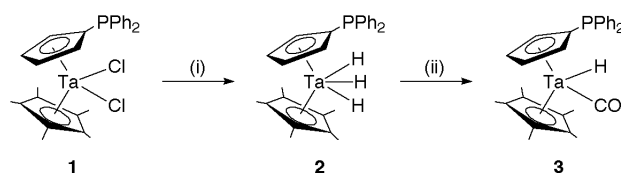
The synthesis of complex **1**  $Cp^*(C_5H_4PPh_2)TaCl_2$  is adapted from Bercaw's method<sup>6</sup> as outlined in Scheme 1.

Complex **1** is isolated in good yield (95%) as a green solid soluble in halogenated solvents. It displays the eight-line ESR



Scheme 1

spectrum characteristic of bis(cyclopentadienyl) tantalum(IV) derivatives.<sup>7</sup> Treatment of **1** with  $NaAl(H)_2(OCH_2CH_2OMe)_2$  in toluene over a period of three hours affords a red solution which, on hydrolysis, gives the trihydride metallocene **2**. Heating of **2** in decane under an atmosphere of carbon monoxide leads to the purple carbonyl derivative **3** (Scheme 2).

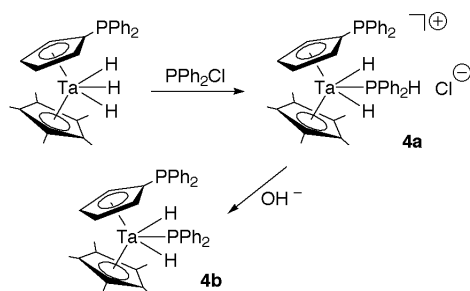


Scheme 2 (i)  $NaAl(H)_2(OCH_2CH_2OMe)_2$ , toluene; (ii) CO, decane, 130 °C.

The  $^1H$  NMR spectrum of **2** exhibits, in the region of the metal hydrides, a doublet of triplets and a doublet of doublets [ $J(H-H) = 11.2$  Hz,  $J(P-H) = 1.9$  Hz;  $J(H-H) = 11.2$  Hz,  $J(P-H) = 5.9$  Hz] corresponding to the central and lateral hydrides respectively. Noteworthy for the monohydride complex **3** is the doublet resonance at  $\delta -5.9$  [ $J(P-H) = 5.8$  Hz] as well as the  $\nu_{CO}$  absorption at  $1882\text{ cm}^{-1}$ .

The reaction of the trihydride **2** with  $PPh_2Cl$  in toluene results in the rapid insertion of the chlorophosphine into the central Ta–H bond affording quantitatively the ionic com-

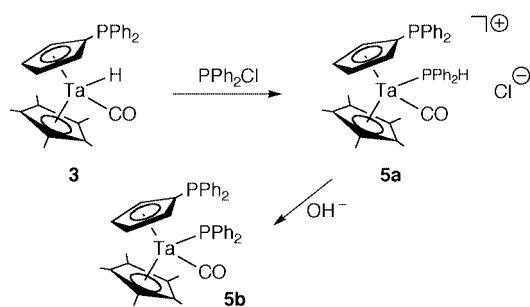
plex  $[\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\text{PPh}_2\text{H})]^+\text{Cl}^-$  **4a** as depicted in Scheme 3.<sup>5,8</sup> Complex **4a** is very sparingly soluble in hydro-



Scheme 3

carbon and aromatic solvents. However, it is sufficiently soluble in acetone to allow NMR spectra to be recorded. The  $^1\text{H}$  NMR spectrum exhibits a unique resonance for the two Ta–H protons (doublet of doublets of doublets) indicative of the central position of the  $\text{PPh}_2\text{H}$  group. The hydrogen signal of the P–H bond is displayed as a doublet of triplets at  $\delta$  7.8 [ $J(\text{H}–\text{H}) = 4.9$  Hz,  $J(\text{P}–\text{H}) = 393$  Hz]. Notably, no exchange process with the phosphorus of the cyclopentadienyl ring is detected. This trend confirms previous observations on the enhanced basicity of a metal-bonded phosphido group with respect to a classical organic phosphine.<sup>9</sup> The salt **4a** is quantitatively deprotonated by sodium hydroxide leading to the corresponding dihydride phosphido complex  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\text{PPh}_2)$  **4b**. The formation of the phosphido derivative **4b** rather than the phosphino one suggests that the P–H bond is a more acidic site than the Ta–H bond.

According to a procedure previously described, the carbonyl monohydride  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{TaH}(\text{CO})$  **3** quickly reacts with chlorodiphenylphosphine.<sup>10</sup> In a first step, the metallophosphonium salt  $[\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\text{PPh}_2\text{H})]^+\text{Cl}^-$  **5a** is formed, probably by a nucleophilic substitution at the phosphorus atom followed by  $\text{H}^+$  transfer onto the basic phosphorus center; deprotonation of this salt is achieved by hydroxide anion and leads to the corresponding neutral phosphido derivative  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\text{PPh}_2)$  **5b** (Scheme 4).



Scheme 4

The four previous kinds of metalloligand derivatives (**2**, **3**, **4b**, **5b**) are able to act as monodentate and/or chelating bidentate ligands. The trihydride complex  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3$  **2** may bond to other metal fragments in different ways, through phosphorus or through both hydrogen and phosphorus. In order to check its coordinating abilities, complex **2** was reacted with the tungsten pentacarbonyl fragment  $[\text{W}(\text{CO})_5]$ . The reaction proceeds at room temperature over a period of 20 min, affording  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3\text{W}(\text{CO})_5$  **6** (Scheme 5). The structure of this complex was determined by a combination of elemental analysis, IR and NMR spectral data. The  $^{31}\text{P}$  NMR spectrum shows a deshielded resonance [ $\delta = 16.1$ ;  $J(\text{W}–\text{P}) = 256$  Hz] with respect to that observed for the starting

material **2** indicating the presence of a tungsten coordinated phosphorus. The IR spectrum in thf solution shows the expected three  $\nu_{\text{CO}}$  bands for the monosubstituted octahedral  $\text{W}(\text{CO})_5$  moiety.

When exposed to  $\text{M}'(\text{CO})_4\text{L}_2$  fragments ( $\text{M}' = \text{Cr}$ ,  $\text{L}_2 = \text{nbdc}$ ;  $\text{M}' = \text{W}$ ,  $\text{L} = \text{MeCN}$ ), the trihydride complex **2** behaves as a bidentate ligand through the hydrogen and phosphorus atoms, leading to the substitution of  $\text{L}_2$  ligands with formation of cyclic compounds  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\mu\text{-H})\text{M}'(\text{CO})_4$  **7** ( $\text{Cr}$ ,  $\text{W}$ ). Very characteristic of these structures is the bridging central hydride resonance which exhibits a high-field signal.<sup>5</sup> As for complex **6**, the deshielded  $^{31}\text{P}$  resonance is indicative of the coordinated phosphorus. The IR spectra of **7** ( $\text{Cr}$ ,  $\text{W}$ ) show the typical four band pattern of the *cis*- $\text{M}'(\text{CO})_4$  fragments.

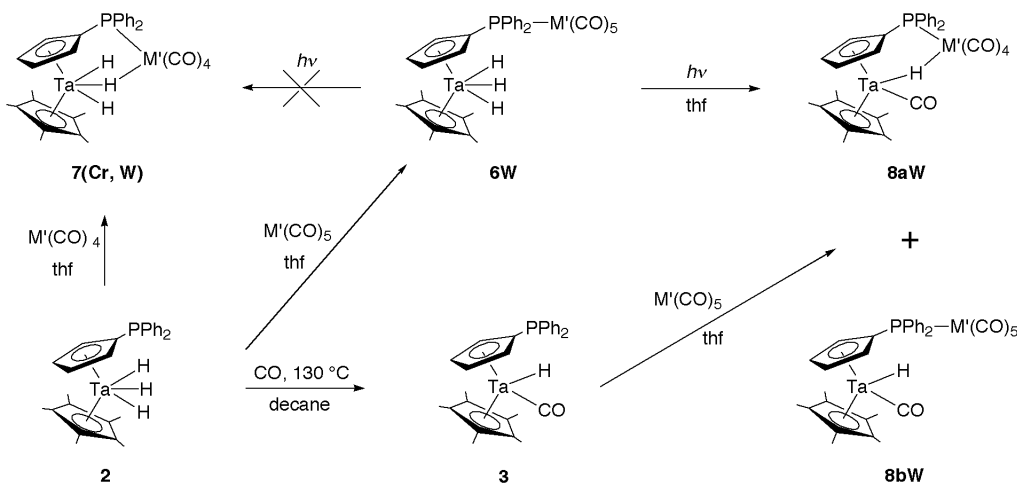
A photochemical process was carried out in order to convert complex **6W** to the cyclic compound **7W**. Surprisingly this attempt did not lead to the expected complex since the reaction product is the tantalum(III) derivative  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\mu\text{-H})\text{W}(\text{CO})_4$  **8aW** which results from the migration of a carbon monoxide fragment to the Ta center. The structure of this complex is proposed on the basis of IR,  $^1\text{H}$  and  $^{31}\text{P}$  NMR data. The  $^1\text{H}$  NMR spectrum of **8aW** shows a unique shielded hydride resonance at  $\delta$   $-15.0$  [ $J(\text{H}–\text{W}) = 40$  Hz] characteristic of a hydrido bridge. The tungsten coordinated phosphorus is evidenced by the  $^{31}\text{P}$  NMR singlet resonance at  $\delta$  8.2 [ $J(\text{P}–\text{W}) = 226$  Hz]. The IR spectrum of **8aW** exhibits five CO absorptions: four attributed to the *cis* disubstituted  $[\text{W}(\text{CO})_4]$  fragment and one for the tantalum coordinated carbonyl.

In a similar way, the reactivity of **3** was investigated. When opposed to  $[\text{W}(\text{CO})_5]$ , the monohydride **3** leads to a mixture of the phosphido linked complex **8bW** and the cyclic structure **8aW**. The formation of the hydrido bridge in **8aW** results from an easy loss of one CO ligand at the tungsten site.

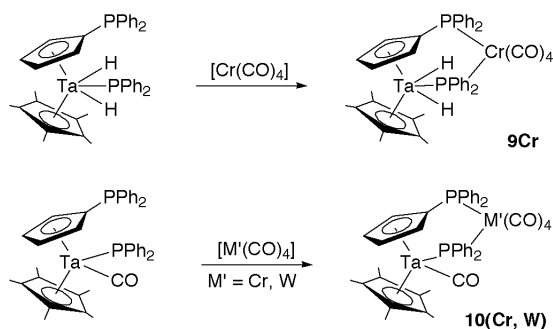
Unfortunately, complexes **8aW** and **8bW** are not separable from their mixture either by crystallisation or by column chromatography. However, when monohydride **3** is reacted with the  $[\text{W}(\text{CO})_4]$  moiety, the cyclic structure **8aW** is obtained as the sole product.

The metalloligands **4b** and **5b** combine the coordinating abilities of a metalloligand with those of an organic phosphine. Toward  $[\text{M}'(\text{CO})_4]$  moieties, they behave as chelating ligands affording two types of Ta(v) **9Cr** or Ta(III) **10** ( $\text{Cr}$ ,  $\text{W}$ ) heterobimetallic systems (Scheme 6).

The formation of **9Cr** and **10** ( $\text{Cr}$ ,  $\text{W}$ ) is ascertained by the noticeable deshielding found for the  $^{31}\text{P}$  resonances with respect to that observed in the parent metallodiphosphines. Furthermore, in the two chromium derivatives **9Cr** and **10Cr**, the low field signals fall in the same range ( $\delta$  59.7 and 65.5 respectively) and therefore can be attributed to the phosphorus nuclei of the  $\text{CpPPh}_2$  ligand. In **9Cr** and **10W**, the weak  $J(\text{P}–\text{P})$  value (10 Hz) is indicative of a *cis* configuration for the two phosphorus nuclei.<sup>11</sup> As expected, the two  $^{31}\text{P}$  resonances of **10W** are surrounded by the  $^{183}\text{W}$  satellites [ $J(\text{P}–\text{W}) = 249$  Hz and  $J(\text{P}–\text{W}) = 151$  Hz]. No coupling effect between the two phosphorus atoms is detected in **10Cr**. The crystal structure of the bimetallic complex **10Cr** is shown in Fig. 1. The compound crystallized with two chloroform solvate molecules. Selected bond lengths and bond angles are reported in Table 1. The structure consists of two fragments: a tantalocene moiety linked to a chromium-carbonyl unit by two phosphorus atoms leading to a five membered chelate ring. The same structural features were reported for the Re, Rh diphosphido bridged complex  $(\text{PPh}_3)(\text{C}_5\text{H}_4\text{PPh}_2)\text{Re}(\text{NO})(\mu\text{-PPh}_2)\text{Rh}(\text{nbdc})$ .<sup>12</sup> In both cases the five membered ring adopts a similar distorted “envelope” conformation<sup>13</sup> the tantalum atom and the *ipso*-cyclopentadienyl carbon C(1) lying 0.207 Å and 0.664 Å below the P(1)–Cr–P(2) plane. The structural parameters of the tantalocene moiety are similar to those reported for several



Scheme 5



Scheme 6

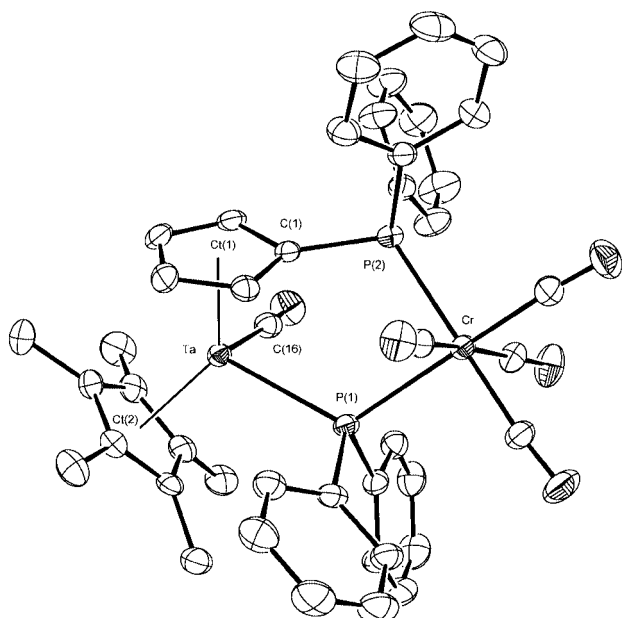


Fig. 1 ORTEP<sup>24</sup> view of  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\mu\text{-PPh}_2)\text{Cr}(\text{CO})_4$  **10Cr**. The chloroform solvate molecules are omitted for clarity.

monophosphido bridged derivatives, *i.e.*  $\text{Cp}'\text{CpTa}(\text{CO})(\mu\text{-PMe}_2)\text{W}(\text{CO})_4\text{L}$  ( $\text{L} = (o\text{-anisyl})\text{methylphenylphosphine}$ ),  $\text{Cp}_2\text{-Ta}(\text{CO})(\mu, \eta^1: \eta^1\text{-PMe}_2\text{S})\text{W}(\text{CO})_5$  and  $\text{Cp}^*\text{CpTa}(\text{CO})(\mu\text{-PPhOMe})\text{Cr}(\text{CO})_5$ .<sup>14</sup> The main differences observed between the mono bridged complex  $\text{Cp}^*\text{CpTa}(\text{CO})(\mu\text{-PPhOMe})\text{Cr}(\text{CO})_5$  and **10Cr** result from the cyclic nature of the bridge in **10Cr**, thus the Ta–P(1)–Cr angle is smaller in **10Cr** than in the mono bridged compound ( $119.68(4)^\circ$  vs.  $127.1(1)^\circ$ ) leading to a smaller Ta...Cr distance in **10Cr** than in the mono bridged complex ( $4.495 \text{ \AA}$  for **10Cr** vs.  $4.646 \text{ \AA}$ ). The P(2)–Cr bond

Table 1 Selected bond distances (Å) and bond angles (°) for **10Cr**

|                |            |                |            |
|----------------|------------|----------------|------------|
| Cr–C(17)       | 1.843(4)   | Ta–Ct(1)       | 2.06       |
| Cr–C(18)       | 1.887(4)   | Ta–Ct(2)       | 2.10       |
| Cr–C(19)       | 1.838(5)   | Ta–C(16)       | 2.025(4)   |
| Cr–C(20)       | 1.885(5)   | Ta–P(1)        | 2.6695(11) |
| Cr–P(2)        | 2.3671(12) | P(2)–C(1)      | 1.807(4)   |
| Cr–P(1)        | 2.5288(14) |                |            |
| C(16)–Ta–P(1)  | 90.65(12)  | C(18)–Cr–P(1)  | 87.92(12)  |
| Ct(1)–Ta–Ct(2) | 135.1      | C(18)–Cr–P(2)  | 88.00(13)  |
| Ct(1)–Ta–C(16) | 103.1      | C(18)–Cr–C(19) | 88.5(2)    |
| Ct(2)–Ta–C(16) | 101.5      | C(18)–Cr–C(20) | 172.2(2)   |
| Ct(1)–Ta–P(1)  | 104.4      | C(19)–Cr–P(1)  | 176.1(2)   |
| Ct(2)–Ta–P(1)  | 112.4      | C(19)–Cr–P(2)  | 93.51(14)  |
| P(1)–Cr–P(2)   | 87.95(4)   | C(19)–Cr–C(20) | 87.4(2)    |
| C(17)–Cr–P(1)  | 92.79(14)  | C(20)–Cr–P(1)  | 95.97(14)  |
| C(17)–Cr–P(2)  | 177.87(14) | C(20)–Cr–P(2)  | 98.82(14)  |
| C(17)–Cr–C(18) | 90.0(2)    | Cr(1)–P(1)–Ta  | 119.68(4)  |
| C(17)–Cr–C(19) | 85.6(2)    | C(1)–P(2)–Cr   | 113.13(12) |
| C(17)–Cr–C(20) | 83.1(2)    |                |            |

length is  $2.367(1) \text{ \AA}$  and compares to those usually observed for  $\text{PCr}(\text{CO})_n$  groups.<sup>15</sup> In contrast, the P(1)–Cr distance is remarkably long:  $2.5288(14) \text{ \AA}$ . This distance is close to the sum of the covalent radii of Cr ( $1.48 \text{ \AA}$ )<sup>16</sup> and P ( $1.10 \text{ \AA}$ ).<sup>17</sup> Such long distances ( $>2.50 \text{ \AA}$ ) have already been reported for some crystal structures involving a  $\text{M–P–Cr}(\text{CO})_n$  bridge.<sup>18</sup>

## Experimental

### General

Nuclear magnetic resonance spectra were recorded on a Bruker AC 200 instrument at 200 MHz for  $^1\text{H}$  and 81 MHz for  $^{31}\text{P}$ . The chemical shifts are reported in ppm relative to  $\text{SiMe}_4$  ( $^1\text{H}$ ) and external  $\text{H}_3\text{PO}_4$  ( $^{31}\text{P}\{-^1\text{H}\}$ ). Infrared spectra were obtained with a Nicolet 250 spectrometer, with the sample in a solution of thf. Elemental analyses were performed on an EA 1108 CHNS-O Fisons Instruments apparatus.

All reactions were carried out under an atmosphere of argon using standard Schlenk line techniques and conventional glass vessels. The solvents were dried over suitable reagents and freshly distilled under argon before use.  $\text{Li}(\text{C}_5\text{H}_4\text{PPh}_2)$ ,<sup>19</sup>  $\text{Cp}^*\text{TaCl}_4$ ,<sup>6</sup>  $\text{Cr}(\text{CO})_4(\text{nbd})$ ,<sup>20</sup>  $\text{W}(\text{CO})_4(\text{MeCN})_2$ <sup>21</sup> and  $\text{M}'(\text{CO})_5(\text{thf})$ <sup>22</sup> ( $\text{M}' = \text{Cr}, \text{W}$ ) were prepared according to literature methods.

### Syntheses

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{TaCl}_2$  1.** Compound **1** was synthesized according to Bercaw's method reported for  $\text{Cp}^*\text{CpTaCl}_2$ <sup>6</sup> with minor modifications.  $4.58 \text{ g}$  ( $10 \text{ mmol}$ ) of  $\text{Cp}^*\text{TaCl}_4$  was

reduced with 0.5 equiv. of Mg in thf in the presence of 1 equiv. of  $\text{PMe}_3$ . Removal of the solvent afforded a red residue. This crude product was treated with 2.49 g (10 mmol)  $\text{Li}(\text{C}_5\text{H}_4\text{PPh}_2)$  in toluene at 100 °C. After 12 h of stirring, extraction and crystallization in acetone afforded 6.04 g (95%) of the green paramagnetic product **1** (Found: C, 50.69; H, 4.52.  $\text{C}_{27}\text{H}_{29}\text{Cl}_2\text{PTa}$  requires C, 50.96; H, 4.59%). The ESR spectrum in toluene showed an eight-line spectrum:  $g_{\text{iso}} = 1.9306$ ;  $A_{\text{Ta, iso}} = 122.18$  G.

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3$  **2**.** 3.18 g of **1** (5 mmol) was reduced by 5  $\text{cm}^3$  of  $\text{NaAlH}_2(\text{OCH}_2\text{CH}_2\text{OMe})_2$  (2 M in toluene, 10 mmol) in 50  $\text{cm}^3$  of toluene at 0 °C. After 3 h of stirring, the brown-red solution was hydrolyzed slowly by 10  $\text{cm}^3$  of degassed water. The organic layer was decanted and extracted. Evaporation of the solvent, followed by washing with pentane gave 1.48 g (52%) of complex **2** (Found: C, 56.86; H, 5.71.  $\text{C}_{27}\text{H}_{32}\text{PTa}$  requires C, 57.05; H, 5.67%). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  7.47–7.15 (m, Ph), 5.19 (t,  $J = 2.4$  Hz,  $\text{C}_5\text{H}_4$ ), 4.56 (q,  $J = 2.4$  Hz [ $J_{\text{H-H}} = J_{\text{P-H}}$ ],  $\text{C}_5\text{H}_4$ ), 1.99 (s, Me),  $-0.54$  (dt,  $J = 1.9$ , 11.2 Hz, H),  $-1.77$  (dd,  $J = 5.9$ , 11.2 Hz, H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$   $-11.4$  (s).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})(\text{CO})$  **3**.** 2 g of **2** was stirred in 30  $\text{cm}^3$  of decane at 140 °C under a carbon monoxide atmosphere for 3 h. The solvent was then removed under reduced pressure and the residue was dried *in vacuo*. Crystallization from pentane yielded 1.25 g (60%) of the green compound **3** (Found: C, 56.27; H, 5.02.  $\text{C}_{28}\text{H}_{30}\text{POTa}$  requires C, 56.57; H, 5.09%). FT IR (thf,  $\text{cm}^{-1}$ ): 1882s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  7.63–7.19 (m, Ph), 5.25 (m,  $\text{C}_5\text{H}_4$ ), 4.80 (m,  $\text{C}_5\text{H}_4$ ), 4.20 (m,  $\text{C}_5\text{H}_4$ ), 4.00 (m,  $\text{C}_5\text{H}_4$ ), 1.88 (s, Me),  $-5.94$  (d,  $J = 5.8$  Hz, H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$   $-12.5$  (s).

**$[\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\text{PPh}_2\text{H})]^+\text{Cl}^-$  **4a**.** To a toluene solution (30  $\text{cm}^3$ ) of **2** (1 g, 1.76 mmol) was added  $\text{PPh}_2\text{Cl}$  (0.39 g, 1.76 mmol). A precipitate was immediately formed. After 30 min of stirring, it was filtered off, washed with pentane, and dried under vacuum affording 0.83 g (80%) of a white powder (Found: C, 58.96; H, 5.58.  $\text{C}_{39}\text{H}_{42}\text{P}_2\text{ClTa}$  requires C, 59.36; H, 5.36%). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.24–7.11 (m, Ph), 7.85 (td,  $J = 4.9$ , 393 Hz,  $\text{PPh}_2\text{H}$ ), 5.90 (m,  $\text{C}_5\text{H}_4$ ), 4.67 (m,  $\text{C}_5\text{H}_4$ ), 2.04 (s, Me), 0.30 (ddd,  $J = 5.1$ , 5.9 Hz,  $J = 72$  Hz, Ta–H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$  2.2 (s),  $-16.2$  (s).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\text{PPh}_2)$  **4b**.** A saturated aqueous solution of KOH (10  $\text{cm}^3$ ) was added to the dried salt  $[\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\text{PPh}_2\text{H})]^+\text{Cl}^-$  **4a** (1 g, 1.27 mmol). The mixture was stirred for 30 min. The product was extracted with  $2 \times 15$   $\text{cm}^3$  of toluene, and the organic layer was separated and evaporated giving 0.57 g (60%) of the white product **4b** (Found: C, 61.98; H, 5.22.  $\text{C}_{39}\text{H}_{41}\text{P}_2\text{Ta}$  requires C, 62.24; H, 5.49%). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  7.63–6.86 (m, Ph), 5.33 (t,  $J = 2.46$  Hz,  $\text{C}_5\text{H}_4$ ), 4.51 (m,  $\text{C}_5\text{H}_4$ ), 1.89 (s, Me), 0.63 (dd,  $J = 5.9$ , 61 Hz, Ta–H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$   $-12.0$  (s),  $-16.4$  (s).

**$[\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\text{PPh}_2\text{H})]^+\text{Cl}^-$  **5a**.** To a toluene solution (30  $\text{cm}^3$ ) of **3** (1 g, 1.68 mmol) was added 1 equiv. of  $\text{PPh}_2\text{Cl}$ . The white precipitate was filtered, washed with toluene (10  $\text{cm}^3$ ) and pentane ( $2 \times 10$   $\text{cm}^3$ ), and dried *in vacuo* to yield 1.16 g (85%) of **5a** (Found: C, 59.13; H, 4.64.  $\text{C}_{40}\text{H}_{40}\text{P}_2\text{OCITa}$  requires C, 58.94; H, 4.95%). FT IR (thf,  $\text{cm}^{-1}$ ): 1894 s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.22–7.05 (m, Ph), 7.77 (d,  $J = 378$  Hz,  $\text{PPh}_2\text{H}$ ), 6.48 (m,  $\text{C}_5\text{H}_4$ ), 5.28 (m,  $\text{C}_5\text{H}_4$ ), 4.99 (m,  $\text{C}_5\text{H}_4$ ), 3.24 (m,  $\text{C}_5\text{H}_4$ ), 1.9s (s, Me);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$  11.2 (s),  $-13.4$  (s).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\text{PPh}_2)$  **5b**.** To the dried salt **5a** (1 g, 1.23 mmol) was added a saturated aqueous solution of KOH (10  $\text{cm}^3$ ) followed by 15  $\text{cm}^3$  of toluene. The green organic layer was separated and the aqueous layer was extracted with 15  $\text{cm}^3$

of toluene. The solvent was removed and the green product was dried *in vacuo* to yield 0.62 g of **5b** (65%) (Found: C, 61.47; H, 5.12.  $\text{C}_{40}\text{H}_{39}\text{P}_2\text{OTa}$  requires C, 61.70; H, 5.05%). FT IR (thf,  $\text{cm}^{-1}$ ): 1910s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.20–8.05 (m, Ph), 7.56–6.77 (m, Ph), 5.35 (m,  $\text{C}_5\text{H}_4$ ), 5.23 (m,  $\text{C}_5\text{H}_4$ ), 4.21 (m,  $\text{C}_5\text{H}_4$ ), 2.91 (m,  $\text{C}_5\text{H}_4$ ), 1.87 (s, Me);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$   $-5.6$  (s),  $-12.3$  (s).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3\text{W}(\text{CO})_5$  **6**.** To a thf solution (20  $\text{cm}^3$ ) of  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3$  **2** (0.57 g, 1 mmol) was added an excess (20%) of  $\text{W}(\text{CO})_5(\text{thf})$ . The mixture was stirred for 20 min at room temperature (the reaction was monitored by IR spectroscopy). The solvent was removed under vacuum, and the crude reaction product was crystallized from acetone affording 0.66 g (75%) of **6** (Found: C, 43.11; H, 3.66.  $\text{C}_{32}\text{H}_{32}\text{O}_5\text{PTaW}$  requires C, 43.07; H, 3.61%). FT IR (thf,  $\text{cm}^{-1}$ ): 2070m, 1940vs, 1895s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.00–7.28 (m, Ph), 5.23 (m,  $\text{C}_5\text{H}_4$ ), 5.08 (m,  $\text{C}_5\text{H}_4$ ), 1.95 (s, Me),  $-0.46$  (td, 5.9, 9.8 Hz, H),  $-1.91$  (dd, 4.9, 9.8 Hz, H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$  16.07 [s,  $J(\text{W-P}) = 256$  Hz].

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\mu\text{-H})\text{M}'(\text{CO})_4$  ( $\text{M}' = \text{Cr}, \text{W}$ ) **7(Cr, W)**.**  $\text{Cr}(\text{CO})_4(\text{nbd})$  or  $[\text{W}(\text{CO})_4(\text{MeCN})_2]$  was added to a thf solution (30  $\text{cm}^3$ ) of **2** (0.3 g, 0.53 mmol). The orange solution was stirred for 2 h at 50 °C. The solvent was removed under reduced pressure and the resulting residue washed with pentane (10  $\text{cm}^3$ ) and dried *in vacuo*. The solid was crystallized from acetone affording the product **7** (70–74%).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\mu\text{-H})\text{Cr}(\text{CO})_4$  **7Cr**.** (Found: C, 50.83; H, 4.46.  $\text{C}_{31}\text{H}_{32}\text{O}_4\text{PTaCr}$  requires C, 50.83; H, 4.40%). FT IR (thf,  $\text{cm}^{-1}$ ): 2006s, 1898s, 1844s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.00–7.77 (m, Ph), 7.54–7.32 (m, Ph), 5.78 (m,  $\text{C}_5\text{H}_4$ ), 5.02 (m,  $\text{C}_5\text{H}_4$ ), 2.21 (s, Me), 0.03 (d,  $J = 11.8$  Hz, H),  $-15.35$  (t,  $J = 11.8$  Hz, H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$  55.06.

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_2(\mu\text{-H})\text{W}(\text{CO})_4$  **7W**.** (Found: C, 42.80; H, 3.91.  $\text{C}_{31}\text{H}_{32}\text{O}_4\text{PTaW}$  requires C, 43.08; H, 3.73%). FT IR (thf,  $\text{cm}^{-1}$ ): 2011s, 1937s, 1894s, 1847s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.00–7.78 (m, Ph), 7.61–7.33 (m, Ph), 5.86 (m,  $\text{C}_5\text{H}_4$ ), 5.09 (m,  $\text{C}_5\text{H}_4$ ), 2.21 (s, Me), 0.03 (d,  $J = 12$  Hz, H),  $-12.56$  (td,  $J = 12$ , 5 Hz, H);  $^{31}\text{P}$ - $\{^1\text{H}\}$ , 11.43.

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\mu\text{-H})\text{W}(\text{CO})_4$  **8aW**.** *Procedure A.* A thf solution (100  $\text{cm}^3$ ) of  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})_3\text{W}(\text{CO})_5$  **6** (0.50 g, 0.56 mmol) was irradiated with a HANAU TQ 150 lamp for 3 h at room temperature. The solvent was evaporated to give a green powder. The crude product was chromatographed over a silica gel (70–230 mesh) column (toluene eluant) affording 0.30 g (60%) of **8aW**. *Procedure B.* To a thf solution of  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})(\text{CO})$  **3** (0.25 g, 0.42 mmol) was added 1 equiv. of  $\text{W}(\text{CO})_4(\text{MeCN})_2$ . The mixture was stirred for 2 h at 40 °C and the reaction was monitored by IR spectroscopy. The solvent was removed and the crude product was crystallised from acetone leading to an analytically pure sample of **8aW** (57%). *Procedure C.* To a thf solution of  $\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{H})(\text{CO})$  **3** (0.25 g, 0.42 mmol) was added an excess (20%) of  $\text{W}(\text{CO})_5(\text{thf})$ . The mixture was stirred for 2 h at room temperature and the reaction was monitored by IR spectroscopy. The solvent was removed *in vacuo* and the crude reaction product was chromatographed over a silica gel column (70–230 mesh). Impurities were first eliminated by eluting with toluene–pentane (90:10). Further elution with toluene gave a large band corresponding to a mixture of **8aW** and **8bW** (42:58).

**$\text{Cp}^*(\text{C}_5\text{H}_4\text{PPh}_2)\text{Ta}(\text{CO})(\mu\text{-H})\text{W}(\text{CO})_4$  **8aW**.** (Found: C, 42.73; H, 3.40.  $\text{C}_{32}\text{H}_{30}\text{O}_5\text{PTaW}$  requires C, 43.14; H, 3.40%). FT IR (thf,  $\text{cm}^{-1}$ ): 2007s, 1919s, 1889s, 1883s, 1847s ( $\nu_{\text{CO}}$ ). NMR ( $\text{CD}_3\text{COCD}_3$ ):  $^1\text{H}$ ,  $\delta$  8.04–7.34 (m, Ph), 5.36 (m,  $\text{C}_5\text{H}_4$ ), 4.99 (m,  $\text{C}_5\text{H}_4$ ), 4.82 (m,  $\text{C}_5\text{H}_4$ ), 2.13 (s, Me),  $-15.05$  [s,  $J(\text{H-W}) = 40$  Hz, H];  $^{31}\text{P}$ - $\{^1\text{H}\}$ ,  $\delta$  8.22 [ $J(\text{W-P}) = 226$  Hz].



**Table 2** Crystal data and structure refinement for **10Cr**

|  |   |
|--|---|
| Formula                                      | C <sub>44</sub> H <sub>39</sub> O <sub>5</sub> P <sub>2</sub> CrTa·2CHCl <sub>3</sub> |
| <i>M</i>                                     | 1181.38   |
| <i>T</i> /K                                  | 293(2)  |
| Crystal system                               | Triclinic   |
| Space group                                  | <i>P</i> $\bar{1}$  |
| <i>a</i> /Å                                  | 11.095(2)   |
| <i>b</i> /Å                                  | 11.395(2)   |
| <i>c</i> /Å                                  | 20.220(2)   |
| $\alpha$ /°                                  | 79.86(3)  |
| $\beta$ /°                                   | 80.42(3)  |
| $\gamma$ /°                                  | 72.31(3)  |
| <i>V</i> /Å <sup>3</sup>                     | 2380.0(6)   |
| <i>Z</i>                                     | 2   |
| $\mu$ /mm <sup>-1</sup>                      | 2.974   |
| RC = Reflections Collected                   | 10879   |
| IRC = Independent RC                         | 9653 [ <i>R</i> <sub>int</sub> = 0.0141]  |
| IRCGT = IRC and [ <i>I</i> > 2σ( <i>I</i> )] | 8407  |
| <i>R</i> for IRCGT                           | <i>R</i> <sup>a</sup> = 0.0318, <i>wR</i> <sup>b</sup> = 0.0800                       |
| <i>R</i> for IRC                             | <i>R</i> <sup>a</sup> = 0.0451, <i>wR</i> <sup>b</sup> = 0.0855                       |

<sup>a</sup> *R*1 = Σ(|*F*<sub>o</sub>| - |*F*<sub>c</sub>|)/Σ|*F*<sub>o</sub>|. <sup>b</sup> *wR*2 = [Σ(*wF*<sub>o</sub><sup>2</sup> - *F*<sub>c</sub><sup>2</sup>)/Σ(*wF*<sub>o</sub><sup>2</sup>)]<sup>1/2</sup> where *w* = 1/[σ<sup>2</sup>(*F*<sub>o</sub><sup>2</sup>) + (0.0491*P*)<sup>2</sup> + 2.03*P*] where *P* = (Max(*F*<sub>o</sub><sup>2</sup>, 0) + 2*F*<sub>c</sub><sup>2</sup>)/3.

**Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(H)(CO)W(CO)<sub>5</sub> 8bW.** (Identified from the mixture of **8aW** and **8bW**.) FT IR (thf, cm<sup>-1</sup>): 2071s, 1939s, 1913s, 1892s (ν<sub>CO</sub>). NMR (CD<sub>3</sub>COCD<sub>3</sub>): <sup>1</sup>H, δ 7.76–6.76 (m, Ph), 4.96 (m, C<sub>5</sub>H<sub>4</sub>), 4.41 (m, C<sub>5</sub>H<sub>4</sub>), 4.00 (m, C<sub>5</sub>H<sub>4</sub>), 3.83 (m, C<sub>5</sub>H<sub>4</sub>), 1.65 (s, Me), -5.52 (d, *J* = 14.8 Hz, H); <sup>31</sup>P-{<sup>1</sup>H}, δ 11.83 [*J*(W–P) = 261 Hz].

**Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(H)<sub>2</sub>(μ-PPh<sub>2</sub>)Cr(CO)<sub>4</sub> 9Cr.** The procedure described earlier for complex **7Cr** was performed but starting from (0.20 g, 0.26 mmol) Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(H)<sub>2</sub>(PPh<sub>2</sub>) and 1 equiv. of Cr(CO)<sub>4</sub>(nbd). The crude product obtained after removal of solvent was recrystallised from acetone leading to 0.16 g of **9Cr** (65%) (Found: C, 56.01; H, 4.38. C<sub>43</sub>H<sub>41</sub>O<sub>4</sub>P<sub>2</sub>TaCr requires C, 56.32; H, 4.51%). FT IR (thf, cm<sup>-1</sup>): 1993s, 1908s, 1872s, 1858s (ν<sub>CO</sub>). NMR (CDCl<sub>3</sub>): <sup>1</sup>H, δ 8.22–7.96 (m, Ph), 7.57–7.02 (m, Ph), 6.26 (m, C<sub>5</sub>H<sub>4</sub>), 4.67 (m, C<sub>5</sub>H<sub>4</sub>), 2.47 (d, *J* = 54 Hz, H), 1.69 (s, Me); <sup>31</sup>P-{<sup>1</sup>H}, δ 59.7 (d, *J* = 10 Hz), -1.4 (d, *J* = 10 Hz).

**Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(CO)(μ-PPh<sub>2</sub>)M'(CO)<sub>4</sub> (M' = Cr, W) 10(Cr, W).** The procedure described earlier for complex **7Cr** was performed but starting from 0.2 g (0.26 mmol) Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(CO)(PPh<sub>2</sub>) and 1 equiv. of Cr(CO)<sub>4</sub>(nbd) or W(CO)<sub>4</sub>(MeCN)<sub>2</sub>. The crude product was chromatographed over a silica gel (70–230 mesh) column (toluene eluant) affording a deep red powder (70–75%) of **10(Cr, W)**.

**Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(CO)(μ-PPh<sub>2</sub>)Cr(CO)<sub>4</sub> 10Cr.** (Found: C, 55.54; H, 4.41. C<sub>44</sub>H<sub>39</sub>O<sub>5</sub>P<sub>2</sub>TaCr requires C, 56.06; H, 4.17%). FT IR (thf, cm<sup>-1</sup>): 1993s, 1906s, 1900s, 1879s, 1854s (ν<sub>CO</sub>). NMR (CDCl<sub>3</sub>): <sup>1</sup>H, δ 7.96–7.67 (m, Ph), 7.60–6.96 (m, Ph), 5.66 (m, C<sub>5</sub>H<sub>4</sub>), 5.05 (m, C<sub>5</sub>H<sub>4</sub>), 4.19 (m, C<sub>5</sub>H<sub>4</sub>), 1.57 (s, Me); <sup>31</sup>P-{<sup>1</sup>H}, δ 65.5 (s), 18.8 (s).

**Cp\*(C<sub>5</sub>H<sub>4</sub>PPh<sub>2</sub>)Ta(CO)(μ-PPh<sub>2</sub>)W(CO)<sub>4</sub> 10W.** (Found: C, 49.04; H, 3.76. C<sub>44</sub>H<sub>39</sub>O<sub>5</sub>P<sub>2</sub>TaW requires C, 49.18; H, 3.66%). FT IR (thf, cm<sup>-1</sup>): 2003s, 1909s, 1898s, 1880s, 1853s (ν<sub>CO</sub>). NMR (CDCl<sub>3</sub>): <sup>1</sup>H, δ 7.91–7.57 (m, Ph), 7.52–6.96 (m, Ph), 5.62 (m, C<sub>5</sub>H<sub>4</sub>), 5.09 (m, C<sub>5</sub>H<sub>4</sub>), 4.40 (m, C<sub>5</sub>H<sub>4</sub>), 1.58 (s, Me); <sup>31</sup>P-{<sup>1</sup>H}, δ 32.2 [d, *J* = 10, *J*(P–W) = 249 Hz], -9.0 [d, *J* = 10, *J*(P–W) = 151 Hz].

### Crystal structure analysis

A crystal of **10Cr** (red; 0.30 × 0.25 × 0.20 mm<sup>3</sup>) was mounted on an Enraf Nonius CAD4 diffractometer. Experimental details are summarized in Table 2. The structure was solved via a Patterson search program<sup>23</sup> and refined (space group *P* $\bar{1}$ )

with full-matrix least-squares methods<sup>23</sup> based on |*F*<sup>2</sup>|. All non-hydrogen atoms were refined with anisotropic thermal parameters. Hydrogen atoms of the complex were included in their calculated positions and refined with a riding model. The rather large residual peak and hole in the final Fourier difference map are located close (<1 Å) to the Ta atom. Selected bond distances and angles are collected in Table 1.

CCDC reference number 186/1549.

See <http://www.rsc.org/suppdata/dt/1999/2725/> for crystallographic files in .cif format.

### References

- R. Choukroun, D. Gervais, J. Jaud, P. Kalck and F. Senocq, *Organometallics*, 1986, **5**, 67; M. J. Hostetler, M. D. Butts and R. G. Bergman, *J. Am. Chem. Soc.*, 1993, **115**, 2743; Y. Yamaguchi, N. Suzuki, T. Mise and Y. Wakatsuki, *Organometallics*, 1999, **18**, 996 and references therein.
- D. W. Stephan, *Coord. Chem. Rev.*, 1989, **95**, 41; R. T. Baker, W. C. Fultz, T. B. Morder and I. D. Williams, *Organometallics*, 1990, **9**, 2357; R. M. Bullock and C. P. Casey, *Acc. Chem. Res.*, 1987, **20**, 167.
- X. Dong He, A. Maisonnat, F. Dahan and R. Poilblanc, *Organometallics*, 1989, **8**, 2618; W. Tikkanen, Y. Fujita and J. L. Petersen, *Organometallics*, 1986, **5**, 888; R. M. Bullock and C. P. Casey, *Acc. Chem. Res.*, 1987, **20**, 167; F. T. Ladipo, G. K. Anderson and N. P. Rath, *Organometallics*, 1994, **13**, 4741; J. Szymoniak, M. M. Kubicki, J. Besançon and C. Moïse, *Inorg. Chim. Acta*, 1991, **180**, 153.
- J. F. Reynoud, J. C. Leblanc and C. Moïse, *J. Organomet. Chem.*, 1985, **296**, 377; V. I. Bakhmutov, E. V. Vorontsov, G. Boni and C. Moïse, *Inorg. Chem.*, 1997, **36**, 4055.
- G. Bonnet, O. Lavastre, G. Boni, J. C. Leblanc and C. Moïse, *C.R. Acad. Sci. Paris, Ser. II*, 1994, **319**, 1293; O. Lavastre, G. Bonnet, G. Boni, M. M. Kubicki and C. Moïse, *J. Organomet. Chem.*, 1997, **547**, 141; C. Barré, P. Boudot, M. M. Kubicki and C. Moïse, *Inorg. Chem.*, 1995, **34**, 284; V. Comte, O. Blacque, M. M. Kubicki and C. Moïse, *Organometallics*, 1997, **16**, 5763.
- V. C. Gibson, J. E. Bercaw, W. J. Burton, Jr. and R. D. Sanner, *Organometallics*, 1986, **5**, 976.
- M. L. H. Green, J. A. McCleverty, L. Pratt and G. Wilkinson, *J. Chem. Soc.*, 1961, 4854; F. N. Tebbe and G. W. Parshall, *J. Am. Chem. Soc.*, 1971, **93**, 3793.
- G. I. Nikonov, D. A. Lemenovskii and J. Lorberth, *Organometallics*, 1994, **13**, 3127; G. I. Nikonov, Y. K. Grishin, D. A. Lemenovskii, N. B. Kazennova, L. G. Kuzmina and J. A. K. Howard, *J. Organomet. Chem.*, 1997, **547**, 183.
- G. Bonnet, M. M. Kubicki, C. Moïse, R. Lazzaroni, P. Salvadori and G. Vitulli, *Organometallics*, 1992, **11**, 964; W. E. Buhro, B. D. Zwick, S. Georgiou, J. P. Hutchinson and J. A. Gladysz, *J. Am. Chem. Soc.*, 1988, **110**, 2427; G. Boni, M. M. Kubicki and C. Moïse, *Bull. Soc. Chim. Fr.*, 1994, **131**, 895.
- G. Bonnet, M. M. Kubicki, J. C. Leblanc and C. Moïse, *New J. Chem.*, 1988, **12**, 551; G. Boni, P. Sauvageot and C. Moïse, *J. Organomet. Chem.*, 1995, **489**, C32; G. Boni, P. Sauvageot, E. Marpeaux and C. Moïse, *Organometallics*, 1995, **14**, 5652.
- F. Mercier, C. Hugel le Goff and F. Mathey, *Organometallics*, 1988, **7**, 955; P. S. Pregosin and R. W. Kunz, in *<sup>31</sup>P and <sup>13</sup>C NMR of transition metal phosphine complexes*, Springer-Verlag, Berlin, 1979, p. 116.
- B. D. Zwick, A. M. Arif, A. T. Patton and J. A. Gladysz, *Angew. Chem., Int. Ed. Engl.*, 1987, **26**, 910.
- B. Fuchs, *Top. Stereochem.*, 1978, **10**, 1.
- P. Sauvageot, O. Blacque, M. M. Kubicki, S. Jugé and C. Moïse, *Organometallics*, 1996, **15**, 2399; S. Challet, M. M. Kubicki, J. C. Leblanc, C. Moïse and B. Nuber, *J. Organomet. Chem.*, 1994, **483**, 47; S. Challet, O. Lavastre, C. Moïse, J. C. Leblanc and B. Nuber, *New J. Chem.*, 1994, **18**, 1155.
- Data obtained for 20 crystal structures from the Cambridge Crystallographic Data base.
- F. A. Cotton and D. C. Richardson, *Inorg. Chem.*, 1986, **5**, 1851.
- L. Pauling, in *Grundlagen der Chemie*, Verlag Chemie, Weinheim, 1973, p. 170.
- E. Hey, A. C. Willis and S. B. Wild, *Z. Naturforsch., Teil B*, 1989, **44**, 1041; L. Weber, R. Kirchhoff and R. Boese, *J. Chem. Soc., Chem. Commun.*, 1992, 1182; F. Lindenberg, T. Gelbrich and E. Hey-Hawkins, *Z. Anorg. Allg. Chem.*, 1995, **621**, 771; P. Oudet, D. Perrey, G. Bonnet, C. Moïse and M. M. Kubicki, *Inorg. Chem. Acta*, 1995, **237**, 79.

- 19 F. Mathey and J. P. Lampin, *Tetrahedron*, 1975, **31**, 2685.
- 20 M. A. Bennet, L. Pratt and G. J. Wilkinson, *J. Chem. Soc. A*, 1961, 2037.
- 21 I. W. Stolz, G. R. Dobson and R. K. Sheline, *Inorg. Chem.*, 1963, **2**, 323.
- 22 W. Strohmeier and F. J. Mueller, *Chem. Ber.*, 1969, **102**, 3608.
- 23 G. M. Sheldrick, SHELXS and SHELX97, University of Göttingen, Göttingen, Federal Republic of Germany, 1997.
- 24 C. K. Johnson, ORTEP, Report ORNL-5138, Oak Ridge National Laboratory, Oak Ridge, TN, 1976.

*Paper 9/03603B*