ORGANIC LETTERS

2007
Vol. 9, No. 21
4115−4118

Revisiting the Armed–Disarmed Concept: The Importance of Anomeric Configuration in the Activation of *S*-Benzoxazolyl Glycosides

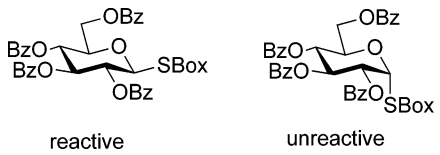
David Crich*,† and Ming Li†

Department of Chemistry, University of Illinois at Chicago, 845 West Taylor Street, Chicago, Illinois 60607-7061

dcrich@chem.wayne.edu

Received June 20, 2007

ABSTRACT



The unexpectedly high reactivity of a (2-benzoxazolyl) per-*O*-benzoyl-*β*-D-thioglucoside and related donors in reactions promoted by copper(II) trifluoromethanesulfonate is revealed, by comparison with the unreactive α-anomer, to be the result of neighboring group participation. Revision of the armed–disarmed concept for glycosyl donors is not required.

According to the armed−disarmed concept of Fraser−Reid glycosyl donors protected with ether protecting groups are armed and can be activated in the presence of disarmed donors carrying ester protecting groups.[1] This seminal concept builds on the observations of earlier workers in the field,[2] is widely applied in modern oligosaccharide synthesis,[3] and is at the heart of the one-pot oligosaccharide protocols developed by the Ley and Wong groups.[4] Mechanistically, the effect is readily understood in terms of the destabilization of positive charge, whether full or partial, at the anomeric center in the course of most glycosylation reactions by electron-withdrawing protecting groups. Recently, it was observed that while the perbenzyl donor **1** in the series of benzoxazolyl thioglycosides **1**−**5** was the most reactive as expected on the basis of the armed−disarmed rationale, the 2-*O*-benzyl *S*-benzoxazolyl glycosides **2** and **3** were less reactive than the corresponding 2-*O*-acyl donors **4** and **5** on activation with copper(II) triflate (Figure 1).[5]

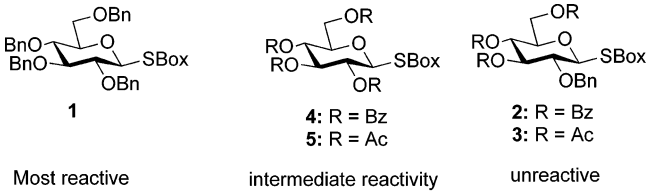


Figure 1. Donor reactivity sequence with $Cu(OTf)_2$.

Similarly, it was found inter alia that the disaccharide **6** could be activated with copper(II) triflate in the presence of

† Current address: Department of Chemistry, Wayne State University, 5101 Cass Avenue, Detroit, MI 48202.

(1) Mootoo, D. R.; Konradsson, P.; Udodong, U.; Fraser-Reid, B. *J. Am. Chem. Soc.* **1988**, *110*, 5583−5584.

(2) (a) Glaudemans, C. P. J.; Fletcher, H. G. *J. Am. Chem. Soc.* **1965**, *87*, 4636−4641. (b) Green, L. G.; Ley, S. V. In *Carbohydrates in Chemistry and Biology*; Ernst, B., Hart, G. W., Sinay, P., Eds.; Wiley-VCH: Weinheim, Germany, 2000; Vol. 1, pp 427−448. (c) Horenstein, N. A. *Adv. Phys. Org. Chem.* **2006**, *41*, 275−314.

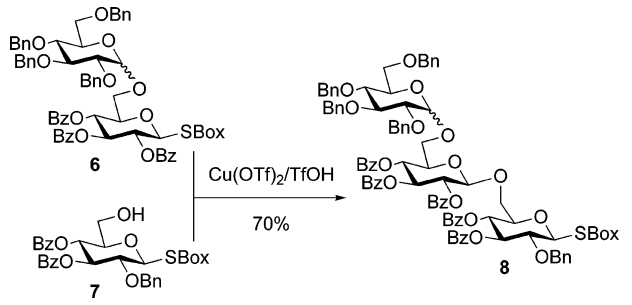
(3) (a) Garegg, P. J. *Adv. Carbohydr. Chem. Biochem.* **2004**, *59*, 69−134. (b) Fraser-Reid, B.; Tatsuta, K.; Thiem, J., Eds *Glycoscience*: *Chemistry and Biology*; Springer: Berlin, Germany, 2001. (c) Toshima, K.; Tatsuta, K. *Chem. Rev.* **1993**, *93*, 1503−1531.

(4) (a) Douglas, N. L.; Ley, S. V.; Lucking, U.; Warriner, S. L. *J. Chem. Soc., Perkin Trans. 1* **1998**, 51−65. (b) Zhang, Z.; Ollmann, I. R.; Ye, X.-S.; Wischnat, R.; Baasov, T.; Wong, C.-H. *J. Am. Chem. Soc.* **1999**, *121*, 734−753.

10.1021/ol701466u CCC: $37.00
Published on Web 09/22/2007

the acceptor **7** giving a 70% yield of trisaccahride **8** (Scheme 1) in apparent contravention of the armed−disarmed concept.

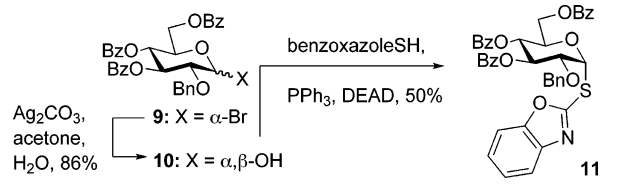
Scheme 1. Trisaccharide Synthesis



A rationale was advanced based on the differential stabilization of a fully developed positive charge (oxacarbenium ion) at the anomeric center by the protecting groups at O2 and/or O6, and on the changing ability of the "O5 lone pair" to eject the leaving group from the anomeric center according to the nature of the O6 protecting group.[5] It occurred to us that the at first sight unusual reactivity patterns of Figure 1 and Scheme 1 are better explained on the basis of classical anchimeric assistance, and, if so, that they would be dependent on the anomeric configuration. We report here on the results of our investigation and confirm that neighboring group participation is the underlying cause of the observed reactivity sequence.

The α-analogue (**11**) of **2** was synthesized from the known bromide **9**[6] by hydrolysis to the pyranose and subsequent Mitsunobu reaction with benzoxazolethiol (Scheme 2).

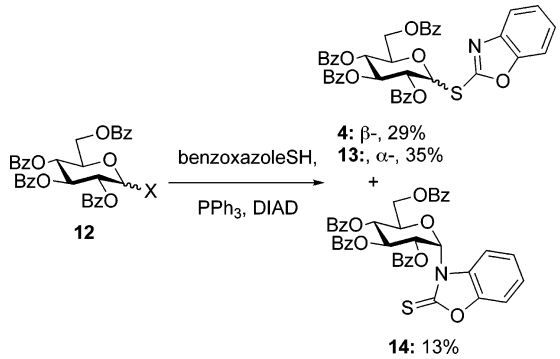
Scheme 2. Synthesis of the α-SBox Donor **8**



In a similar manner 2,3,4,6-tetra-*O*-benzoyl-D-glucopyranose **12**[7] was transformed into a separable mixture of the α- and *β*-SBox glycosides **13** and **4**, and the α-NBox glycoside **14** (Scheme 3). The *β*-SBox donors **1** and **2** were prepared according to the literature protocols.[8]

The identity of donors **11** and **13** as the α-SBox derivatives, as opposed to the corresponding NBox derivatives (e.g., **14**), was determined by inspection of the ^{13}C NMR spectra in which C2 of the benzoxazole is found around δ 160 as

Scheme 3. Synthesis of the α-SBox Donor **13**



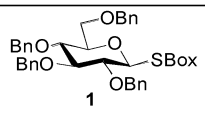
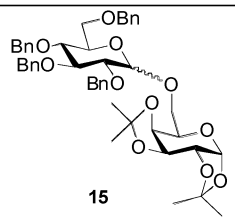
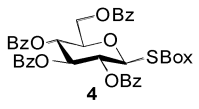
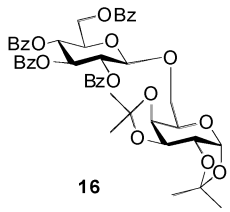
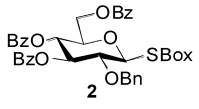
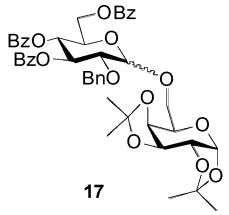
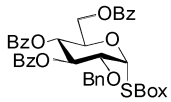
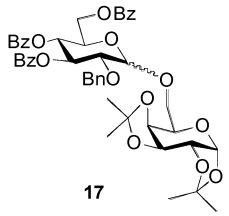
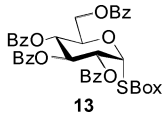
compared to the more downfield shift of this carbon resonance in the N-glycosides (δ ∼178), consistent with the thiocarbonyl nature of the latter system. These assignments were confirmed by the UV spectra, which showed the anticipated differences between 2-alkylthiobenzoxazoles and *N*-alkylbenzoxazol-2-thiones.[9]

A series of reactions were then conducted in which donors **1**, **2**, **4**, **11**, and **13** were activated with $Cu(OTf)_2$ in dichloromethane at room temperature in the presence of 5 Å molecular sieves[10] and 1,2;3,4-di-*O*-isopropylidene-α-D-galactopyranose, essentially according to the original conditions. The results obtained with donors **1**, **2**, and **4** (Table 1) conform to the pattern reported previously (Figure 1). Thus, the perbenzyl donor **1** was the most reactive and was consumed within 2 h (Table 1, entry 1), whereas the perbenzoyl system **4** required 14 h (Table 1, entry 2). The 2-*O*-benzyl-3,4,6-tri-*O*-benzoyl-*β*-SBox derivative **2** was recovered in 85% yield after 14 h in a reaction that afforded only 14% of the anticipated glycoside (Table 1, entry 3). Turning to the α-SBox donors **11** and **13**, the 2-*O*-benzyl-3,4,6-tri-*O*-benzoyl-α-SBox system **11** showed comparably poor reactivity to its *β*-anomer (Table 1, entry 4), whereas the per-*O*-benzoyl α-SBox donor **13** was completely unreactive under these conditions (Table 1, entry 5).

The complete contrast in reactivity of the α- and *β*-anomers of the perbenzoyl donors **4** and **13** (Table 1, entries 2 and 5) is informative, especially when viewed alongside the very similar reactivity of the two anomers of the 2-*O*-benzyl-3,4,6-*O*-benzyl system **2** and **11** (Table 1, entries 3 and 4). We conclude that the enhanced reactivity of donor **4** with respect to its α-isomer **13** is simply a manifestation of the weak promoting system (copper(II) triflate) being assisted by participation of the 2-*O*-benzoate ester. Thus, unlike the case of the fully armed per-*O*-benzyl ether, the promoter is unable to cause departure of the anomeric leaving group from the formally more disarmed donor **4** without participation by the ester group. When participation is stereoelectronically prevented, as in the case of **13**, no reaction occurs at all. The 2-*O*-benzyl-3,4,6-*O*-benzoyl systems, both of which are

(5) Kamat, M. N.; Demchenko, A. V. *Org. Lett.* **2005**, *7*, 3215−3218.
(6) Lichtenthaler, F. W.; Köhler, B. *Carbohydr. Res.* **1994**, *258*, 77−85.
(7) Sebesan, S.; Neira, S. *Carbohydr. Res.* **1992**, *223*, 169−185.
(8) Demchenko, A. V.; Kamat, M. N.; De Meo, C. *Synlett* **2003**, 1287−1290.
(9) Zinner, H.; Peseke, K. *Chem. Ber.* **1965**, *98*, 3515−3519.
(10) Without commenting on the reason, we note that reactions conducted in the presence of 4 rather than 5 Å molecular sieves were much slower with donors **1** and **4**, and prohibitively so with donors **2**, **11**, and **13**.

Table 1. Coupling Reactions with 1,2;3,4-Di-*O*-isopropylidene-α-D-galactopyranose[a]

entry	donor	time (h)	product	yield	α:β ratio	recovered donor
1	 **1**	2	 **15**	93%	1:1.4	0%
2	 **4**	14	 **16**	95%	β-only	0%
3	 **2**	14	 **17**	14%	α-only	85%
4	 **11**	14	 **17**	13%	α-only	88%
5	 **13**	14	-	No reaction	-	95%

[a] All reactions were conducted in dichloromethane at room temperature with stoichiometric amounts of donor and acceptor, 2.4 equiv of $Cu(OTf)_2$, and in the presence of 5 Å molecular sieves

moderately disarmed, but which cannot benefit from neighboring group participation from the 2-position, show the anticipated reactivity between **1** and **13**. It is anticipated that with stronger Lewis acids capable of promoting ionization without the need for neighboring group participation the reactivity order of the equatorial donors **1**−**5**, and of donors **6** and **7**, will revert to that predicted by the armed/disarmed hypothesis.

Comparable effects have been observed previously in carbohydrate chemistry. For example, Paulsen and Herold reported that *β*-glucosyl pentacetate **18** underwent reaction with antimony pentachloride to give a dioxolenium ion **19** and subsequent rearrangement products, under conditions in which the corresponding α-anomer **20** was unreactive (Figure 2).[11] Under the same conditions, the more reactive chlorides **21** and **22** were consumed irrespective of anomeric configuration. Konradsson studied the relative rates of reaction of a series of pentenyl glycosides and reported, inter alia, that the per-*O*-benzyl systems **23** and **24** showed only a 1.7-fold

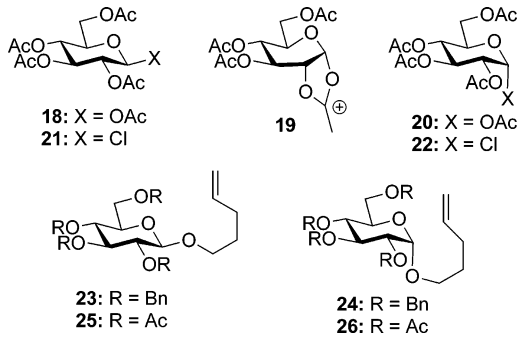


Figure 2. Literature examples of the phenomenon.

difference in reactivity favoring the *β*-isomer. In contrast, in the peracetylated systems **25** and **26** the *β*-anomer was

(11) Paulsen, H.; Herold, C.-P. *Chem. Ber.* **1970**, *103*, 2450−2462.

found to be 5.2 times more reactive than its α-isomer (Figure 2).[12] Both groups attributed the enhanced reactivity of the 2-*O*-acetate protected *β*-configured donors to neighboring group participation by the ester.

We conclude that the unusual reactivity sequence reported for donors **1**−**5** on promotion with copper(II) trifluoromethanesulfonate is the result of a weakly promoted system being facilitated by neighboring group assistance from an ester. The acceleration of solvolysis reactions by antiperiplanar vicinal ester groups is one of the classical concepts of physical organic chemistry from which carbohydrates are not immune.[13] Revision of the seminal armed−disarmed concept is not required.

Acknowledgment. We thank the NIH (GM62160) for support of our work on the mechanisms of glycosylation reactions, and Professor A. V. Demchenko (University of Missouri St. Louis) for helpful discussions.

Supporting Information Available: Full experimental details and characterization data for new compounds. This material is available free of charge via the Internet at http://pubs.acs.org.

OL701466U

(12) Konradsson, P. In *Glycoscience*: *Chemistry and Biolog*; Fraser-Reid, B., Tatsuta, K., Thiem, J., Eds.; Springer: Berlin, Germany, 2001; Vol. 1, pp 535−549.

(13) (a) Winstein, S.; Grunwald, E.; Ingraham, L. L. *J. Am. Chem. Soc.* **1948**, *70*, 821−828. (b) Capon, B.; McManus, S. P. *Neighboring Group Participation*; Plenum Press; New York, 1976. (c) Anslyn, E. V.; Dougherty, D. A. *Modern Physical Organic Chemistry*; University Science Books: Sausalito, CA, 2006.