Charge Delocalization in a Homologous Series of α, α' -Bis(dianisylamino)-Substituted Thiophene Monocations

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Supporting Information

ABSTRACT: A homologous series of three molecules containing thiophene, bithiophene, and terthiophene bridges between two redox-active tertiary amino groups was synthesized and explored. Charge delocalization in the one-electron-oxidized forms of these molecules was investigated by a combination of cyclic voltammetry, near-infrared optical absorption spectroscopy, and EPR spectroscopy. All three cation radicals can be described as organic mixed-valence species, and for all of them the experimental data are consistent with strong delocalization of the unpaired electron.



Depending on what model is used for analysis of the optical absorption data, estimates for the electronic coupling matrix element (H_{AB}) range from ~5000 to ~7000 cm⁻¹ for the shortest member of the homologous series. According to optical absorption and EPR spectroscopy, even the terthiophene radical appears to belong either to Robin–Day class III or to a category of radicals commonly denominated as borderline class II/class III systems. The finding of such a large extent of charge delocalization over up to three adjacent thiophene units is remarkable.

INTRODUCTION

The one-electron-oxidized forms of oligothiophenes exhibit a remarkable electrical conductivity, and this property makes them well suited as hole transport materials in a variety of different (opto)electronic applications.¹⁻⁵ Understanding distance-dependent charge delocalization in thiophene monocations is therefore of significant interest. Numerous experimental and theoretical studies already explored the electronic structure of oxidized thiophenes at a fundamental level, but mixed-valence approaches are comparatively rare.⁶⁻¹⁴ In a purely computational study, Lacroix and co-workers arrived at the conclusion that oligothiophene monocations may be described as organic mixed-valence species.¹⁵ Indeed, the one-electron-oxidized or one-electron-reduced forms of several oligo-p-phenylene and oligo-p-phenylenevinylene systems could be described well as organic mixed-valence compounds in various experimental studies, $^{16-25}$ but until now there has been comparatively little conceptually analogous work on thiophene systems.²⁶⁻³⁵ A notable exception is the recent study of the oxidized forms of bis(4-(alkoxyphenyl)amino) derivatives of dithienylethene and bithiophene and the finding that a mixed-valence description of these cations is meaningful.³⁶ Another exception is our own recent investigation of photoswitchable organic mixed valence in dithienylethene systems.37,38

Here, we report on a detailed experimental investigation of the homologous series of α, α' -bis(diphenylamino)-substituted

thiophene molecules (1-3) from Scheme 1. Further included in the study was a reference compound (4) comprised of a thiophene unit containing only one diphenylamino substituent. From a combined electrochemical, optical spectroscopic, and EPR study of the monocationic forms of these molecules, we obtain detailed insight into charge-delocalization phenomena in thiophene monocations. Of key interest in this study was the question to which Robin–Day class the three monocations 1^+ , $\hat{2}^+$, and 3^+ belong.³⁹ We anticipated that 1^+ would be a fully delocalized (class III) radical and wanted to explore whether the longer bithiophene (2^+) and terthiophene (3^+) systems would have a similar electronic structure, or whether there would be a transition to partial charge localization (class II) or even complete charge localization (class I).^{40,41} For phenylenebridged organic mixed-valence systems this question had been answered previously,^{16,19} but to the best of our knowledge such a length dependence of charge delocalization has not yet been investigated in a homologous series of thiophene-bridged organic mixed-valence compounds.^{34,35} In view of the wide interest in thiophene materials mentioned above, it seemed relevant to explore over how many adjacent thiophene units an unpaired electron can be delocalized.

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Scheme 1. Chemical Structures of the Molecules Investigated in This Work



RESULTS AND DISCUSSION

Synthesis. Compound 1 was obtained using a palladium(0)-catalyzed reaction between 2,5-dibromothiophene and 4,4'-dimethoxydiphenylamine. The longer congeners (2 and 3) were synthesized in an analogous manner using 5,5'-dibromo-2,2'-bithiophene and 5,5"-dibromo-2,2':5',2"-terthiophene as reaction partners with 4,4'-dimethoxydiphenylamine. Reference compound 4 was available from a prior investigation.³⁷ Detailed synthetic protocols and product characterization data for all new compounds are found in the Supporting Information.

Electrochemistry. Figure 1 shows the results of cyclic voltammetry experiments with compounds 1-4 in deoxy-



Figure 1. Cyclic voltammograms of the four compounds from Scheme 1 in dry and deoxygenated acetonitrile measured in the presence of 0.1 M tetrabutylammonium hexafluorophosphate (TBAPF₆) as a supporting electrolyte: (a) compound 1; (b) compound 2; (c) compound 3; (d) compound 4. The waves at -0.51 V vs Fc⁺/Fc (dashed vertical line) are due to decamethylferrocene, which was added for internal voltage calibration.

genated dry acetonitrile solution in the presence of tetrabutylammonium hexafluorophosphate (TBAPF₆) as a supporting electrolyte. The waves at -0.51 V vs Fc⁺/Fc (vertical dashed line) are due to decamethylferrocene, which was added in small quantities as an internal reference for voltage calibration. The voltammograms of compounds 1 (Figure 1a) and 2 (Figure 1b) are comprised of two clearly discernible reversible waves which are separated by 450 and 250

mV, respectively, and they are attributed to one-electronoxidation processes involving predominantly the two different tertiary amino groups. The decrease in the voltage splitting (ΔE) between the two waves from 450 mV in 1 to 250 mV in 2 is a manifestation of the increasing N–N distance and a decrease in electronic communication between the two redoxactive moieties. For compound 3 only one reversible wave is observable at first glance, but digital fitting of the relevant part of the cyclic voltammogram from Figure 1 c yields $\Delta E = 70$ mV (see Supporting Information). For reference compound 4 there is only one tertiary amino group; hence the observation of a single oxidation wave in the potential range considered here is no surprise (Figure 1d). Moreover, this particular wave is irreversible, presumably due to electrochemically induced polymerization chemistry of the 4⁺ species.

The electrochemical potentials for one- and two-electron oxidation of molecules 1-4 are summarized in Table 1. The

Table 1. Electrochemical Potentials for One- and Two-Electron Oxidation of the Four Compounds from Scheme 1 (in volts) vs Fc^+/Fc^a

compd	$E_{1/2}(1)$	$E_{1/2}(2)$	$\Delta E \ [mV]$	K _c
1	-0.23	0.22	450	4.2×10^{7}
2	-0.17	0.08	250	1.7×10^{4}
3	-0.08	-0.01	70	1.5×10^{1}
4	0.16			

^{*a*} ΔE is the difference between the first ($E_{1/2}(1)$) and second oxidation potentials ($E_{1/2}(2)$) in mV; K_c is the comproportionation constant.

fact that the first one-electron oxidation in the bis-(diphenylamino) systems occurs at significantly more negative potentials than the first one-electron oxidation of reference compound 4 is attributed to the more electron-rich nature of the 2-fold amino-substituted systems with respect to the singly amino-substituted reference compound. We note that in the α,α' -bis(diphenylamino)-substituted systems 1-3 the first oxidation potential $(E_{1/2}(1))$ increases to more positive values with increasing thiophene bridge length, while the second oxidation potential $(E_{1/2}(2))$ is decreasing simultaneously to more negative values.⁴² This leads to the above-mentioned decrease in the overall oxidation potential splitting (ΔE) , which may be interpreted as a manifestation of decreasing electronic communication with increasing length. In mixed-valence chemistry the comproportionation constant $(K_c = 10^{\Delta E/59} \text{ mV})$ is often used as a measure for the stability of the mixed-valence state,^{43,44} and we calculate values of 4.2×10^7 for 1⁺, 1.7×10^4 for 2⁺, and 1.5×10^1 for 3⁺ (Table 1). Classification of mixedvalence systems based on K_c values is possible,³⁹ but it requires some caution.^{43–46} With our thiophene systems from Scheme 1 we are in the fortunate situation of having the possibility to gain complementary insight into the electronic structures of the monocations by optical absorption and EPR spectroscopies. Therefore, we refrain from premature conclusions on the exclusive basis of the cyclic voltammetry studies.

Optical Absorption Spectroscopy. The black traces in Figure 2 are the optical absorption spectra of the charge-neutral



Figure 2. Optical absorption spectra of the four compounds from Scheme 1 in acetonitrile solution in the presence of increasing amounts of $Cu(ClO_4)_2$ as a chemical oxidant: (a) compound 1; (b) compound 2; (c) compound 3; (d) compound 4. Black lines, neat compounds before adding any oxidant; blue lines, after addition of 0.5 equiv of oxidant; green traces, after addition of 1 equiv of oxidant; red traces, after addition of 2 equiv of oxidant.

forms of molecules 1-4 in acetonitrile solution at room temperature. The data show that none of the four compounds from Scheme 1 has any significant absorption features below 17 000 cm⁻¹ as long as the molecules are in their charge-neutral forms. Absorption spectra of molecules 1-4 with extinction coefficients are shown in Figure S1 of the Supporting Information, and we find that the lowest energetic absorption maxima are at 29 700 cm⁻¹ for 1, 25 000 cm⁻¹ for 2, 22 200 cm⁻¹ for 3, and 33 300 cm⁻¹ for 4. Thus, there is a red shift of the lowest absorption band with increasing thiophene bridge length, which is likely a manifestation of increasing π conjugation.⁴⁷⁻⁴⁹

The colored traces in Figure 2 were measured after the addition of increasing amounts of copper(II) perchlorate as an oxidant. The blue traces represent the spectra detected after addition of 0.5 equiv, the green traces are those measured after adding 1 equiv, and the red traces are the spectra obtained after 2 equiv of $Cu(ClO_4)_2$ had been added. We note that in each of the three bis(dianisylamino) systems (1–3; Figure 2a–c) absorption bands between 5000 and 15 000 cm⁻¹ appear upon oxidation. Importantly, in each case the lowest energetic absorptions are only present in situations in which less than 2 equiv of oxidant has been added (blue and green traces), indicating that the respective absorption bands are due to the one-electron-oxidized (mixed-valence) species. The absence of

any low-energy absorptions in the oxidized forms of reference molecule 4 (Figure 2d) is consistent with the interpretation of the lowest energy absorptions of 1^+ , 2^+ , and 3^+ as intervalence absorption bands. However, we also note that (at least in cyclic voltammetry, vide supra) compound 4 undergoes irreversible oxidation.

Figure 3 shows the lowest energetic absorptions of the monocations in more detail and on a properly calibrated



Figure 3. Black traces, low-energy portions of optical absorption spectra of monocationic forms of the molecules from Scheme 1. (a) 1^+ ; (b) 2^+ ; (c) 3^+ ; (d) 4^+ . Colored traces, Gaussian fits to the experimental data. The parameters of the Gaussian functions represented by the dotted green traces are reported in Table 2.

extinction scale. For this purpose, the low-energy portions of the optical absorption spectra measured after addition of 0.25 equiv of oxidant were employed. The use of data obtained in the presence of a shortfall of chemical oxidant is a common procedure to avoid disproportionation of the mixed-valence state.³⁶ The black traces in Figure 3 represent the actual experimental data. While 1⁺ has only one absorption band in the relevant spectral range (Figure 3a), the longer congeners $(2^+$ in Figure 3b and 3^+ in Figure 3c) exhibit two bands, and the reference compound (4⁺) is spectroscopically innocent (Figure 3d). By analogy to prior investigations on bithiophene and dithienyl systems with diphenylamino substituents, it appears plausible to interpret the lowest energetic bands in 1^+-3^+ as intervalence absorptions.^{36,37,50} Indeed, these absorptions fall into the spectral range in which intervalence absorptions of bis(triarylamine) radical cations are commonly ob-served.^{16,20,34,35,51-61} An important difference from the intervalence absorptions of coordination compounds^{40,62-65} is the comparatively large extinction coefficients of these bands, but this is a (favorable) peculiarity of organic mixed-valence compounds which has been noted many times before.^{34,35}

In order to perform a quantitative analysis of the intervalence absorptions, the experimental spectra in Figure 3 were fitted to Gaussian functions. Due to the increasing complexity of the absorption spectra along the series $1^+ < 2^+ < 3^+$, an increasing number of Gaussians was required to obtain satisfactory fits: Three Gaussians were sufficient for Figure 3a, four were needed for the spectrum in Figure 3b, and five were necessary in the case of the spectrum in Figure 3c. In all cases the lowest energetic absorption band had to be fitted with two Gaussian functions (dashed green traces). The characteristics of the

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Table 2. Parameters Obtained from Analysis of the Lowest Energetic Near-Infrared Absorptions in Figure 3: Twice the Bandwidth on the High-Energy Side of the Lowest Absorption $(\nu_{1/2[high]})$, Twice the Bandwidth on the Low-Energy Side of the Lowest Absorption $(\nu_{1/2[high]})$, and Ratio between the Two $(\nu_{1/2[high]})^{a}$

compound	$ u_{1/2[\text{high}]} \ [\text{cm}^{-1}] $	$\nu_{1/2[low]} [cm^{-1}]$	$\nu_{1/2[\rm high]}/\nu_{1/2[\rm low]}$	$\nu_{ m max,G1} \ [m cm^{-1}]$	$\nu_{1/2,{ m G1}}~[{ m cm}^{-1}]$	$ u_{\mathrm{max,G2}} [\mathrm{cm}^{-1}] $	$\nu_{1/2,{ m G2}}~[{ m cm}^{-1}]$	$\nu_{ m class~II}~[m cm^{-1}]$
1+	5700	4300	1.33	12 945	3231	15 276	3723	5656
2 ⁺	3550	3060	1.16	10 112	2426	11 671	2232	4953
3+	2640	1940	1.36	7 652	1800	9 0 5 1	1850	4237

 $^{a}\nu_{\text{max,G1}}$ and $\nu_{\text{max,G2}}$ are the energetic positions of the maxima of the two Gaussian functions required to fit the lowest energetic near-infrared absorption. $\nu_{1/2,\text{G1}}$ and $\nu_{1/2,\text{G2}}$ are the full widths at half-maximum values of these two Gaussian fit functions. $\nu_{\text{class II}}$ is the bandwidth which is expected based on eq 4.

respective Gaussian functions are summarized in Table 2: $\nu_{\max,Gi}$ is the energetic position of the maximum of the *i*th Gaussian while $\nu_{1/2,Gi}$ is its full width at half-maximum. Given the fact that all of the lowest energetic absorptions have to be fitted by two Gaussian functions, it is obvious that the shapes of these bands are asymmetric, an observation that is commonly made for class III systems or mixed-valence compounds that are near the borderline of class II and class III.^{16,36,63,66} In order to capture the asymmetry of an intervalence band in one single number, it is customary to determine the bandwidth on the high-energy side of the intervalence band ($\nu_{1/2[high]}$ in Table 2) as well as on its low-energy side ($\nu_{1/2[low]}$ in Table 2), and to build the ratio between the two measures. In our specific cases this procedure yields $\nu_{1/2[high]}/\nu_{1/2[low]}$ ratios between 1.16 and 1.36 (fourth column of Table 2), which are typical values for systems near the borderline of class II and class III and class III.^{16,36,63,66}

There is no clear trend of $\nu_{1/2[high]}/\nu_{1/2[low]}$ along the series of compounds 1⁺, 2⁺, and 3⁺. One might expect the respective ratio to decrease as the systems get longer and the class II character of the relevant mixed-valence species increases. From the finding that $\nu_{1/2[high]}/\nu_{1/2[low]} \neq 1$, one can merely conclude that all three systems deviate from pure class II behavior, but it seems impossible to draw any quantitative conclusions regarding the percentage of class III character.

From the Gaussian fits to the intervalence bands in Figure 3 it is possible to determine the dipole moment $(\mu_{\rm ge})$ associated with the intervalence transition by employing eq 1.^{67,68}

$$\mu_{\rm ge} = 0.09584 \sqrt{\frac{\int \varepsilon(\nu) \, d\nu}{\nu_{\rm max}}} \tag{1}$$

Equation 1 yields the transition dipole moment in units of debyes and requires $\nu_{\rm max}$ and $\varepsilon(\nu)$ input values in units of cm⁻¹ and M⁻¹ cm⁻¹, respectively. Here, $\nu_{\rm max}$ represents the experimentally determined maximum of the intervalence absorptions (fourth column in Table 3) while the integral over $\varepsilon(\nu)$ was taken as the integral over the two Gaussians needed to fit the lowest energetic near-infrared absorptions. Thus we obtain $\mu_{\rm ge}$ values ranging from 5.7 D for 1⁺ to 7.8 D for 3⁺ (Table 3).

The transition dipole moment μ_{ge} is in direct relation with the electronic coupling matrix element H_{AB} describing the electronic interaction between the two redox-active units of a mixed-valence species.⁶⁹ This direct relationship is captured by eq 2, which was used here to estimate H_{AB} on the basis of the experimentally determined μ_{ge} values.

$$H_{AB} = \frac{\mu_{ge} \nu_{max}}{eR}$$
(2)

In eq 2, e is the elemental charge and R is the effective distance between redox centers. A question of central

Table 3. Nitrogen–Nitrogen Distance $(d_{\rm NN})$,^{*a*} Number of σ -Bonds between Nitrogen Atoms (n_{σ}) , and Parameters Extracted from the Near-Infrared Absorption Spectra of the Monocations in Figure 3: Absorption Band Maximum $(\nu_{\rm max})$, Transition Dipole Moment $(\mu_{\rm ge})$ in Units of debyes),

(μ_{max}), Hanshton Dipote Moment (μ_{ge}) in Onte of debyes), and Electronic Coupling Matrix Elements (H_{AB}) Calculated Using Different Models (See Text)

					$H_{AB} [\mathrm{cm}^{-1}]$		
compd	$d_{\mathrm{NN}}^{}a}$ [Å]	n _o	$\nu_{\rm max} [{\rm cm}^{-1}]$	ν _{ge} [D]	class II	class II; (2/3) $\times d_{\rm NN}$	class III
1+	5.2	4	13 850	5.7	3160	4790	6925
2 ⁺	9.1	7	10 620	7.7	1870	2835	5310
3+	13.0	10	7 770	7.8	970	1470	3885
^a Estimated from molecular modeling of the charge-neutral parent							

compounds. importance is usually what to take for R. The effective electron transfer distance can be substantially shorter than what might be expected from purely geometric considerations. For instance, the effective electron transfer distance in selected bis(triarylamine) mixed-valence systems has been found to be only roughly two-thirds of the geometric distance between nitrogen atoms,^{34,61} and analogous investigations of mixedvalent dinitroaromatic anions have come to similar conclusions.^{70,71} In principle it is possible to estimate R with various computational methods, 72-76 but here we intend to rely exclusively on experimental observations. Therefore, we give in Table 3 the results of two calculations with eq 2: Once we report H_{AB} calculated on the basis of the geometric N-N distance $(d_{NN};$ second column of Table 3), and once we give H_{AB} calculated on the basis of $R = (2/3)d_{NN}$. For compound 1⁺ we obtain $H_{AB} = 3160 \text{ cm}^{-1} (d_{NN})$ and $H_{AB} = 4790 \text{ cm}^{-1} ((2/100) \text{ cm}^{-1})$ 3) $d_{\rm NN}$). For reference, the N,N,N',N'-tetraanisyl-p-phenylenediamine monocation (TAPD⁺) has $H_{AB} = 3240 \text{ cm}^{-1}$, and this is a mixed-valence species which was classified as a borderline class II/class III system.¹⁶ For our systems we consider the analysis on the basis of $R = (2/3)d_{NN}$ more adequate than analysis based on $R = d_{NN}$, inter alia because our EPR data (see below) suggest that there is significant unpaired spin density on the

For class III systems, it is possible to estimate H_{AB} directly from the intervalence absorption band maximum (ν_{max}):

$$H_{AB} = \nu_{\rm max}/2 \tag{3}$$

The $\nu_{\rm max}$ values are given in the fourth column of Table 3, and using eq 3 one obtains H_{AB} values of 6925 (1⁺), 5310 (2⁺), and 3885 cm⁻¹ (3⁺) (last column of Table 3). These are relatively high H_{AB} values even when compared to some of the most strongly coupled organic mixed-valence systems known to date.^{34,35} For reference, the $N_{,N}N'_{,N}N'$ -tetramethyl-*p*-phenyl-

thiophene bridges.

enediamine cation (Wurster's blue) has $H_{AB} \approx 8100 \text{ cm}^{-1}$, while $N_{,}N_{,}N_{,}N_{,}N_{,}$ tetraanisyl-*p*-phenylenediamine cation has $H_{AB} = 3240 \text{ cm}^{-1}$.^{34,35} Here we find H_{AB} values which are close to 4000 cm⁻¹ even for the system comprised of *three* adjacent thiophene units.

When attempting to classify mixed-valence species according to the Robin–Day scheme,³⁹ the solvent dependence of the intervalence absorptions may be helpful. We have therefore explored the solvent dependence of the near-infrared absorptions of 1^+ , 2^+ , and 3^+ , and the results of our efforts are shown in Figure 4. Due to the poor solubility of our



Figure 4. Optical absorption spectra of (a) 1^+ , (b) 2^+ , (c) 3^+ , and (d) 4^+ in acetonitrile (solid black traces) and in dichloromethane (red dashed traces).

monocations in most common solvents, the solvent dependence study remained restricted to acetonitrile (black traces) and dichloromethane (red dashed traces). The key observation is the same in all three cases: The lowest energy absorption band is red shifted when going from acetonitrile to dichloromethane. However, the spectral shifts are relatively modest and range from 850 (1⁺) to 560 cm⁻¹ (3⁺). For class II systems a red shift upon solvent change from acetonitrile to dichloromethane is to be expected because of a decrease of the solvent reorganization energy, whereas for pure class III systems one would expect $\nu_{\rm max}$ to be independent of solvent. In our case there is a solvent dependence, but it is a comparatively weak one, suggesting that at least the shorter members of our cation series (1⁺ and 2⁺) are indeed near the borderline of class II and class III. The comproportionation constants (K_c) calculated above (4.2 × 10⁷ for 1⁺ and 1.7 × 10⁴ for 2⁺) do not contradict this classification. For the longest congener (3⁺) the electrochemical data suggest class I ($K_c = 15$), but the optical spectroscopic data are more consistent with borderline class II/class III. We note that comproportionation constants have often been employed as measures for electronic couplings, but it has been demonstrated that they can actually provide a poor guide and should be used with caution.⁷⁷ Therefore, we think that the optical absorption (and EPR) data are more reliable for classifying our mixed-valence compounds.

An additional argument which supports interpretation of 2^+ and 3^+ as borderline class II/class III systems is the relatively narrow width of the intervalence absorption bands observed for these compounds: For a pure class II system one would expect the bandwidth to satisfy eq 4, but in the present cases the calculated bandwidths ($\nu_{class II}$ last column of Table 2; 4953 and 4237 cm⁻¹) are significantly larger than the experimentally observed bandwidths ($\nu_{1/2[high]}$, second column of Table 2; 3550 and 2640 cm⁻¹).

$$\nu_{\rm class\,II} = \sqrt{2310}\nu_{\rm max} \tag{4}$$

EPR Spectroscopy. Given the challenges in classifying the monocations 1^+ , 2^+ , and 3^+ according to the Robin–Day scheme based on electrochemical results and optical absorption spectroscopy, we decided to perform EPR spectroscopy with these systems in order to gain complementary insight into their electronic structures. Furthermore, EPR spectroscopy can potentially shed some light on the question of to what extent these monocations may indeed be considered diamino mixed-valence compounds, or to what extent they have to be regarded as thiophene-bridge oxidized species.³⁶

The solid lines in the upper half of Figure 5 are the experimental X-band EPR spectra obtained at room temperature from dilute dichloromethane solutions of 1^+ (a), 2^+ (b), and 3^+ (c). In all three cases, the EPR signals are centered at a *g*-value of 2.005, which is typical for triarylamine radical cations.⁷⁸ The appearance of the spectra in Figure 5 resembles that observed for related bithiophene and dithienylethene



Figure 5. Solid traces, experimental X-band EPR spectra obtained from radicals (a) 1^+ , (b) 2^+ , and (c) 3^+ in dilute dichloromethane solution at room temperature. Dashed traces, simulated EPR spectra using the hyperfine coupling parameters given in Table 4. The simulations occurred using the WINEPR SimFonia Software.

radicals which were end-capped with 4-alkoxyphenylamino groups;³⁶ the EPR spectrum of 2^+ had even been reported before.³⁶ The hyperfine structure of our EPR spectra can be understood on the basis of an interaction of the unpaired electron spin with the nuclear spins of ¹⁴N (I = 1) and ¹H (I = 1/2), whereby the magnitudes of the coupling constants (for interaction with ¹⁴N and ¹H) are similar.^{31,36} This statement is corroborated by the simulated EPR spectra shown as dashed traces in the lower half of Figure 5. In performing these simulations, we attempted to use a minimum of adjustable parameters. We found that when going from 1^+ to 2^+ and finally to 3^+ , the experimental EPR data can only be simulated satisfactorily when increasing the number of coupling constants along this series of radicals (Table 4).

Table 4. Hyperfine Coupling Constants (in Units of gauss) Used To Simulate the Experimental X-Band EPR Spectra Represented by the Solid Traces in the Upper Half of Figure 5^a

radical	$a_{\rm N}$ [G]	$a_{\rm H}$ [G]	$a_{\mathrm{H}'}[\mathrm{G}]$	$a_{\mathrm{H}''}$ [G]
1+	4.3	2.6		
2^+	2.9	2.4	2.0	
3 ⁺	2.2	2.1	2.0	1.8
-				

^{*a*}The dashed lines in the lower half of Figure 5 illustrate the outcomes of these simulations.

For 1⁺, it is sufficient to invoke coupling to two nitrogen nuclei with a hyperfine coupling constant $a_{\rm N} = 4.3$ G and simultaneous coupling to two hydrogen nuclei with $a_{\rm H} = 2.6$ G. In other words, the experimental data for 1⁺ are in line with uniform spin distribution over the entire molecule, consistent with class III mixed-valence behavior on the EPR time scale. This observation is similar to what has been reported for other bis(triarylamine) radical cations.^{31,36,60,61} It appears plausible that the two chemically equivalent hydrogen nuclei which couple to the unpaired electron spin with $a_{\rm H} = 2.6$ G are in fact those attached to the thiophene bridging unit of molecule 1.⁷⁹

The experimental EPR spectrum of 2^+ in Figure 5b can be simulated using three different hyperfine coupling constants: a_N = 2.9 G, a_H = 2.4 G, and $a_{H'}$ = 2.0 G (second row of Table 4). Importantly, all of these couplings involve two equivalent nuclei, suggesting that the electron is again symmetrically delocalized on the EPR time scale. The additional hydrogen hyperfine constant in this case ($a_{H'}$) is most likely due to the fact that there are now two chemically distinct hydrogen atoms attached to the thiophene bridge; indeed, this interpretation is in line with a prior experimental and theoretical study of 2^+ and a series of related thiophene cations.³⁶

For radical 3^+ , four hyperfine coupling constants must be invoked (third row of Table 4), namely one for two equivalent nitrogen atoms ($a_N = 2.2$) and three for pairs of equivalent hydrogen nuclei ($a_H = 2.1$ G, $a_{H'} = 2.0$ G, $a_{H'} = 1.8$ G). In analogy to the shorter 1^+ and 2^+ radicals, the EPR simulation in Figure 5c thus leads us to the conclusion that the unpaired electron spin is symmetrically delocalized over the 3^+ species on the EPR time scale.

From the EPR spectra alone, one would thus arrive at the conclusion that 1^+ , 2^+ , and 3^+ are all class III mixed-valence species. However, it is to be noted that even class II compounds can in principle exhibit EPR spectra that suggest complete delocalization of the unpaired electron spin.³⁴ Specifically, this is the case in systems in which thermal electron transfer

between the two ends of the molecule is rapid on the EPR time scale, i.e., faster than ~ 10^{-7} s. There have been several comparative investigations of optical and thermal electron transfer in organic mixed-valence systems (mostly on bis-(hydrazine) radical cations),^{80–82} which have come to the conclusion that the results obtained by the two fundamentally different experimental techniques (i.e., optical absorption and EPR spectroscopies) need not necessarily lead to the same conclusion, mostly because these techniques probe electron transfer on significantly different time scales. Be that as it may, in our specific case the EPR data are consistent with the optical absorption results in that both sets of data can be analyzed adequately by assuming complete (in EPR) or nearly complete (in absorption) delocalization of the unpaired electron.

On a final note in this section we point out that the hyperfine constant for coupling to the two equivalent nitrogen nuclei decreases from $a_{\rm N} = 4.3$ G in 1⁺ to $a_{\rm N} = 2.9$ G in 2⁺, and finally to $a_{\rm N} = 2.2$ G in 3⁺. It appears plausible that this decrease is a manifestation of increasing spin delocalization away from the nitrogen atoms toward the center of the thiophene bridge. The expected increase in π -conjugation with increasing thiophene bridge length seems compatible with this interpretation.^{83,95}

SUMMARY AND CONCLUSIONS

One of the key findings from the research presented in this paper is that a mixed-valence description of α, α' -bis-(dianisylamino)-substituted thiophene monocations is meaningful from an experimental point of view, as suggested by the theoretical work performed by Lacroix and co-workers.¹⁵ Our experimental work demonstrates that the mixed-valence approach is in fact reasonable for analysis of electrochemical, optical absorption, and EPR spectroscopic data obtained from amine-decorated thiophene monocations.

The optical absorption and EPR data are consistent with interpretation of the 1^+ , 2^+ , and 3^+ radical species as strongly delocalized systems: EPR spectroscopy suggests them to be class III systems, while optical absorption spectroscopy is compatible with interpretation as both class III and borderline class II/class III species. The computational work by Lacroix and co-workers clearly shows that a transition between different Robin–Day classes is expected for long oligothiophenes.¹⁵ Our current experimental study suggests that this is the case only for molecules comprised of more than three adjacent thiophene units.

EXPERIMENTAL SECTION

Synthetic protocols and product characterization data for 1, 2, 3, and all relevant isolable reaction intermediates are provided in the Supporting Information. ¹H and ¹³C NMR spectroscopy was performed using Bruker Avance DRX300 and Bruker B-ACS-120 instruments, electron ionization mass spectrometry (EI-MS) occurred on a Finnigan MAT8200 instrument, and elemental analyses were conducted on a Vario EL III CHNS analyzer from Elementar. Chemical oxidation of the charge-neutral molecules from Scheme 1 to their monocationic and dicationic forms was effected with Cu(ClO₄)₂, which represents a standard procedure for generating triarylamine radical cations.^{37,84,85} Cyclic voltammetry was performed using a potentiostat from Princeton Applied Research (Versastat3-200) with a glassy carbon disk working electrode and two separate silver wires as counter and quasi-reference electrodes. Optical absorption spectra were measured on a Cary 5000 instrument

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from Varian. The EPR spectra were recorded on a ELEXSYS CW-EPR spectrometer E500 from Bruker, using a microwave frequency of about 9.42 GHz. The modulation frequency was 100 kHz, the modulation amplitude was 1×10^{-4} G, and the power of the microwave was 5–6 mW. The spectrometer was equipped with a digital temperature control system ER 4131 VT, covering a temperature range from 110 to 350 K. UV–vis and cyclic voltammetry data were analyzed using the Igor Pro software (version 6.0.0.0) from WaveMetrics.

ASSOCIATED CONTENT

Supporting Information

Synthetic protocols and characterization data for molecules 1-3 and all intermediate reaction products. Additional optical absorption data. Digital fit of the cyclic voltammogram from Figure 1c. This material is available free of charge via the Internet at http://pubs.acs.org.

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Notes

The authors declare no competing financial interest.

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