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ARTICLE

Coordination polymers and polygons using di-pyridyl-thiadiazole spacers and substituted phosphorodithioato Ni^{II} complexes: potential and limitations for inorganic crystal engineering

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Coordinatively unsaturated P-substituted dithiophosphonato, dithiophosphato, and dithiophosphito complexes $[Ni((MeO)_2PS_2)_2]$ (**1**), $[Ni((EtO)_2PS_2)_2]$ (**2**), $[Ni(MeOdtp)_2]$ (**3**), and $[Ni((Ph)_2PS_2)_2]$ (**4**) were reacted with the *bis*-functional ligands 3,5-di-(4-pyridyl)-1,2,4-thiadiazole (**L1**) and 3,5-di-(3-pyridyl)-1,2,4-thiadiazole (**L2**) to give the coordination polymers $(\mathbf{1–4{\cdot}L1})_\infty$, $(\mathbf{3{\cdot}L2})_\infty$, $(\mathbf{4{\cdot}L2}{\cdot}2C_7H_8)_\infty$ and the discrete dimers $(\mathbf{1–2{\cdot}L2})_2$, all characterised by single crystal X-ray diffraction. The comparison of the structures show that **L1** can be exploited for the predictable assembly of undulating chains independent of the nature of the Ni^{II} complex while **L2** allows for the existence of different supramolecular constructs ensuing from different ligand conformations deriving from the rotation of the pyridyl rings.

Introduction

The last decades experienced an impressive number of novel multidentate bridging ligands designed to construct a variety of discrete metal–organic polygons and polyhedra (MOPs), or metal–organic frameworks (MOFs) with high dimensionality. These supramolecular assemblies are of great appeal for a combination of their intrinsic beauty and promising applications in fields as varied as gas storage, ion exchange, chemical sensing, catalysis, energy transfer, and separation.[1,‡]

The properties of these crystalline materials are critically dependent on their network structures, and the deliberate creation of a crystalline network, planned using properly designed building blocks remains, nowadays, a challenge. Several aspects need to be evaluated in developing a network based on coordination polymers: the building blocks, i.e. metal nodes and organic spacers, metal coordination environments, formation conditions, and weak secondary interactions. Multi-topic organic molecules are commonly used as spacers and an opportune choice of the number and position of donor atoms can be used to direct the network assembly, although factors such as flexibility of the ligand and different accessible conformations need to be examined and taken into account. The metal coordination environment is exceptionally difficult to control when "naked" metal ions are used as nodes and, in view of that, the use of neutral coordination complexes held together by additional donor molecules or secondary bonding interactions has gained striking importance.[2] In fact, by reducing the degrees of freedom of the system, for instance by using *cis*-protected metal blocks in place of the naked metal ions, or by using a neutral, coordinatively unsaturated metal complex, less uncertainty can be expected. Moreover, the use of neutral synthons leads to self-reliant supramolecular assemblies which do not require the presence of a counterion and thus reduce the occurrence of isomerism.

In this respect, we have been developing a synthetic program based on the ability of neutral dithiophosphonato Ni^{II} complexes $[Ni(ROdtp)_2]$ [ROdtp = $(RO)(4\text{-}MeOC_6H_4)PS_2^-$; R = alkyl substituent][3] to assemble coordination polymers in combination with a variety of polypyridyl donors, in particular 4,4'-bipyridine and its analogues.[§,4] This assembly process is based on the capability of the coordinatively unsaturated Ni^{II} ion of these square-planar complexes to axially bind monodentate ligands, such as pyridine, to yield discrete octahedral complexes.[5,6] In addition suitable N–L–N bidentate bipyridyl-based spacers yield 1D coordination polymers of the type $[Ni(ROdtp)_2(N–L–N)]_\infty$.[7,8] The primary structural motif of these polymers mainly depends on the features of the spacers such as length, rigidity and orientation of the donor atoms, as recently confirmed by the deliberate stereospecific generation of homochiral polymeric helices built from a designed enantiopure binaphthyl-based ligand.[9]

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Electronic Supplementary Information (ESI) available: crystallographic data and figures for $(\mathbf{1–4{\cdot}L1})_\infty$, $(\mathbf{1–2{\cdot}L1})_2$, $(\mathbf{3{\cdot}L2})_\infty$, and $(\mathbf{4{\cdot}L2}{\cdot}2C_7H_8)_\infty$. See DOI: 10.1039/x0xx00000x

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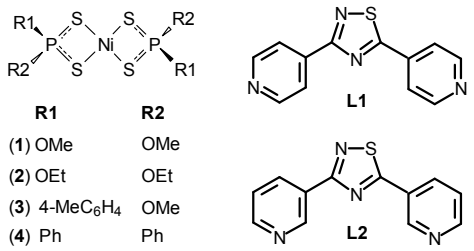

R1	R2
(**1**) OMe	OMe
(**2**) OEt	OEt
(**3**) 4-MeC$_6$H$_4$	OMe
(**4**) Ph	Ph

Scheme 1 General scheme of complexes **1–4** (left) and ligands **L1–L2** (right).

The substituents on the phosphorus atoms are responsible for the connection of the polymers through hydrogen bonds and face–to–face or edge–to–face π–π interactions, thus influencing the final 3D-architecture.[7] As a consequence, coordination polymers and 3D assemblies with different structures and architectures can be built up by varying either the bridging ligands or the substituents on the P atom of the initial Ni complexes. In order to better understand the process of molecular recognition between components and how the steric information contained in the P-substituents and in the orientation of ligand binding sites combine to give the final structure, differently P-substituted dithiophosphonato, dithiophosphato, and dithiophosphito complexes [Ni((MeO)$_2$PS$_2$)$_2$] (**1**), [Ni((EtO)$_2$PS$_2$)$_2$] (**2**), [Ni(MeOdtp)$_2$] (**3**), and [Ni((Ph)$_2$PS$_2$)$_2$] (**4**) were reacted with the *bis*-functional ligands 3,5-di-(4-pyridyl)-1,2,4-thiadiazole (**L1**)[10] and 3,5-di-(3-pyridyl)-1,2,4-thiadiazole (**L2**)[10] (Scheme 1).

We recently investigated the reactivity of **L1** and **L2** with I_2 and IBr, in both polar and apolar media, thereby elucidating the role of specific directional interactions, namely NH$^+$···N and N···I, combined with geometrical features of the molecules, in the formation of different supramolecular constructs.[10] The investigation of these ligands is extended here to the formation of coordination polymers.

Experimental

Materials

All commercially available compounds were used as received. Bis[*O*-alkyl-dithiophosphato]Ni complexes {[Ni((RO)$_2$PS$_2$)$_2$] R = Me (**1**), Et (**2**)},[5] the dithiophosphonato Ni^{II} complex [Ni(MeOdtp)$_2$] [MeOdtp = (RO)(4-MeOC$_6$H$_4$)PS$_2^-$],[3] and bis-functional ligands 3,5-di-(4-pyridyl)-1,2,4-thiadiazole (**L1**) and 3,5-di-(3-pyridyl)-1,2,4-thiadiazole (**L2**),[10] were synthesised according to previously reported procedures. The solvents used were freshly distilled over the appropriate drying agent and used directly from the stills.

Synthesis of bis(diphenyldithiophosphinato)nickel(II), [Ni((C$_6$H$_5$)$_2$PS$_2$)$_2$] (4). Complex **4** was synthesised and firstly X-ray characterised in 1968.[11] However, in this work we developed a new high yielding and clean synthetic route, which starts from primary phosphines: Ph$_2$PH (6.7 mL, 10% water in hexane) was added to KCH$_2$Ph (0.3265 g, 2.5 mmol) in 10 mL of freshly distilled toluene at room temperature. After 10 minutes, the solution was transferred onto dried powdered S_8 (0.1588 g, 5 mmol), and the newly-formed brown suspension was transferred onto NiI_2 (0.3906 g, 1.25 mmol) and refluxed for 4 hours. The solvent was removed under reduced pressure and the resulting brown solid was dissolved in CH_2Cl_2 and filtered through a small celite plug. The filtrate was concentrated under vacuum and **4** obtained as purple crystals, suitable for X-ray analysis. $C_{24}H_{20}P_2S_4Ni$ (formula mass = 557.30 Da, 0.9971 g, 1.8 mmol, 72% yield). M.p.: 253-255 °C (m). ^{1}H NMR (400 MHz, D$_6$-DMSO) δ = 7.22 (d, 3H, Ph), 7.94 (m, 2H, Ph); ^{31}P NMR (400 MHz, D$_6$-DMSO) δ = 60.22.

Synthesis of [Ni((MeO)$_2$PS$_2$)·L1)]$_\infty$, (1·L1)$_\infty$. Complex **1** (18.6 mg, 0.05 mmol) and **L1** (12.0 mg, 0.05 mmol) were reacted at 130 °C in a high pressure Aldrich tube in 30 mL of MeOH. After complete dissolving of the reagents, the reaction mixture was slowly cooled at room temperature. After a few days (**1·L1**)$_\infty$ (4.0 mg, 0.006 mmol, 13% yield) was obtained as green crystals suitable for X-ray analysis. M.p.: 170 °C (d). Elemental analysis found (calc. for $C_{16}H_{20}N_4O_4P_2S_5Ni$; formula mass = 611.9 Da): C, 31.63 (31.33); H, 2.38 (3.29); N, 9.31 (9.14); S, 21.75 (26.14). FT-IR (KBr, 3000-300 cm^{-1}): 2938 w, 2834 vw, 2361 vw, 1608 m, 1462 m, 1411 m, 1335 m, 1290 w, 1210 w, 1176 w, 1130 vs, 827 s, 798 s, 709 m, 691 s, 675 s, 665 m, 530 m, 439 vw, 398 w, 324 m cm^{-1}. FT-Raman (3500-100 cm^{-1}, 600 mW, solid state, relative intensities between parentheses related to the highest peak taken equal to 10.0): 1922 (6.0), 1894 (6.4), 1877 (6.0), 1811 (6.4), 1758 (4.6), 1612 (10.0), 1513 (8.6), 1459 (9.3), 1410 (9.3), 1335 (6.1), 1293 (6.0), 1020 (6.2) cm^{-1}.

Synthesis of [Ni((EtO)$_2$PS$_2$)·L1)]$_\infty$, (2·L1)$_\infty$. Complex **2** (21.4 mg, 0.05 mmol) and **L1** (12.0 mg, 0.05 mmol) were reacted at 160 °C in a high pressure Aldrich tube in 30 mL of EtOH. After complete dissolving of the reagents, the reaction mixture was slowly cooled at room temperature. After a few days (**2·L1**)$_\infty$ (4.1 mg, 0.006 mmol, 12% yield) was obtained as green crystals suitable for X-ray analysis by slow evaporation of the solvent. M.p.: 155 °C (d). Elemental analysis found (calc. for $C_{20}H_{28}N_4O_4P_2S_5Ni$; formula mass = 668.0 Da): C, 36.31 (35.89); H, 4.17 (4.22); N, 8.46 (8.37); S, 24.12 (23.95). FT-IR (KBr, 3000-350 cm^{-1}): 3054 vw, 3032 w, 2934 w, 2893 vw, 2459 vw, 2285 vw, 1931 vw, 1609 s, 1496 vs, 1440 m, 1412 m, 1336 m, 1121 m, 1019 vs, 945 vs, 848 w, 830 m, 805 m, 773 s, 713 m, 673 s, 657 s, 644 m, 620 w, 546w, 410 w cm^{-1}. FT-Raman (3500-0 cm^{-1}, 150 mW, solid state, relative intensities between parentheses related to the highest peak taken equal to 10.0): 3075 (4.8), 3027 (4.2), 2933 (5.1), 2888 (5.0), 1981 (4.1), 1611 (10.0), 1513 (5.7), 1413 (8.5), 1338 (5.9), 1229 (5.3), 1215 (5.5), 1095 (7.7), 1015 (7.7) 999 (5.7), 729 (4.2), 650 (6.1) 548 (8.8), 376 (7.5) cm^{-1}.

Synthesis of [Ni((MeO)(4-MeOC$_6$H$_4$)PS$_2$)$_2$·L1]$_\infty$, (3·L1)$_\infty$. Complex **3** (26.2 mg, 0.05 mmol) and **L1** (12.0 mg, 0.05 mmol) were reacted at 100 °C in a high pressure Aldrich tube in 30 mL of CH_3OH. After complete dissolving of the reagents, the reaction mixture was slowly cooled at room temperature. After a week, (**3·L1**)$_\infty$ (29.5 mg, 0.39 mmol, 77% yield) was

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obtained as green crystals suitable for X-ray analysis. M.p.: 160°C (d). Elemental analysis found (calc. for $C_{28}H_{28}N_4O_4P_2S_5Ni$; formula mass = 764.0 Da): C, 44.73 (43.98); H, 3.38 (3.69); N, 7.53 (7.32); S, 21.08 (20.94). FT-IR (KBr, 1800-300): 1214 w, 1179 mw, 1130 vw, 1114 s, 1065 w, 1029 vs, 1020 vs, 909 vw, 851 vw, 830 ms, 779 vs, 754 w, 733 vw, 709 w, 690 vw, 654 ms, 640 s, 625 mw, 546 vs, 520 mw, 508 w, 457 vw, 442 w, 398 vw, 369 vw, 326 ms cm^{-1}. FT-Raman (3500-100 cm^{-1}, 150 mW, solid state, relative intensities between parentheses related to the highest peak taken equal to 10.0): 3054 (2.8), 2924 (2.8), 1615 (10.0), 1582 (5.7), 1420 (7.8), 1310 (3.6), 1280 (4.2), 1110 (5.7), 1020 (5.0), 1000 (5.0), 547 (6.4), 102 (6.4) cm^{-1}.

Synthesis of $[Ni((C_6H_5)_2PS_2)_2{\cdot}L1]_\infty$, $(4{\cdot}L1)_\infty$. A mixture of **4** (5.5 mg, 0.01 mmol) and **L1** (2.4 mg, 0.01 mmol) in 3 mL of toluene was heated to 100 °C in a sealed 5 mL screw-top glass bottle for 3 days. The mixture was allowed to cool to room temperature and the resulting green crystals of $(\mathbf{4{\cdot}L1})_\infty$ filtered from the reaction mixture; (5.1 mg, 0.007 mmol, 64% yield) M.p.: >230 °C. Elemental analysis found (calc. for $C_{36}H_{28}N_4P_2S_5Ni$; formula mass = 797.57 Da): C, 44.88 (54.21); H, 3.45 (3.54); N, 4.93 (7.02)%. FT-IR (KBr, 4000- 400 cm^{-1}): 3448 vw, 3045 w, 1607 s, 1465 s, 1435,09 s, 1408 s, 1329 m, 1304 w, 1285 w, 1226 w, 1099 s, 1061 m, 999 m, 833 m, 823 m, 746 s, 707 vs, 699 vs, 612 vs, 571 s, 565 vs, 518 m, 483 m, 443 w, 420 w cm^{-1}.

Synthesis of $[Ni((MeO)_2PS_2){\cdot}L2)]_2$, $(1{\cdot}L2)_2$. Complex **1** (18.6 mg, 0.05 mmol) and **L2** (18.0 mg, 0.07 mmol) were reacted at 140 °C in a high pressure Aldrich tube in 30 mL of MeOH. After complete dissolving of the reagents, the reaction mixture was slowly cooled at room temperature. After a few days $(\mathbf{1{\cdot}L2})_2$ (5.5 mg, 0.008 mmol, 18% yield) was obtained as green crystals suitable for X-ray analysis by slow evaporation of the solvent. M.p.: 180 °C (d). Elemental analysis found (calc. for $C_{16}H_{20}N_4O_4P_2S_5Ni$; formula mass = 611.9 Da): C, 31.55 (31.33); H, 2.21 (3.29); N, 9.13 (9.14); S, 24.69 (26.14). FT-IR (KBr, 4000-400 cm^{-1}): 2938 w, 2834 vw, 2361 vw, 1608 m, 1462 vm, 1411 m, 1335 m, 1290 w, 1210 w, 1176 w, 1130 vs, 827 s, 798 s, 709 m, 691 s, 675 s, 665 m, 530 m, 439 vw, 398 w, 324 m cm^{-1}. FT-Raman (4000-0 cm^{-1}, 200 mW, solid in KBr, relative intensities between parentheses related to the highest peak taken equal to 10.0): 1922 (6), 1894 (6.4), 1877 (6), 1811 (6.4), 1758 (4.6), 1612 (10), 1513 (8.6), 1459 (9.3), 1410(9.3), 1335 (6.1), 1293 (6), 1020 (6.2) cm^{-1}.

Synthesis of $[Ni((EtO)_2PS_2){\cdot}L2)]_2$, $(2{\cdot}L2)_2$. A solution of **L2** (10 mg, 0.042 mmol) in 10 mL of EtOH was slowly diffused into a of CH_2Cl_2 solution of **2** (10 mg, 0.023 mmol, 0.5 mL) and left to stand at room temperature for several weeks. Green crystals of $(\mathbf{2{\cdot}L2})_2$ (16.9 mg, 0.025 mmol, 50% yield) suitable for X-ray analysis were obtained. M.p.: 159-162 °C (m). Elemental analysis found (calc. for $C_{20}H_{28}N_4O_4P_2S_5Ni$; formula mass = 668.0 Da): C, 34.84 (35.89); H, 4.53 (4.22); N, 8.11 (8.37); S, 24.15 (23.95). FT-IR (KBr, 4000-50 cm^{-1}): 3054 vw, 3032 w, 2934 w, 2893 vw, 2459 vw, 2285 vw, 1931 vw, 1609 s, 1496 vs, 1440 m, 1412 m, 1336 m, 11215 m, 1019 vs, 945 vs, 848 w, 830 m, 805 m, 773 s, 713 m, 673 s, 657 s, 644 m, 620 w, 546 w, 410 w cm^{-1}. FT-Raman (3500-0 cm^{-1}, 600 mW, solid state, relative intensities between parentheses related to the highest peak taken equal to 10.0): 3075 (4.8), 3027 (4.2), 2933 (5.1), 2888 (5), 1981 (4.1), 1611 (10), 1513 (5.7), 1413 (8.5), 1338 (5.9), 1229 (5.3), 1215 (5.5), 1095 (7.7), 1015 (7.7), 999 (5.7), 729 (4.2), 650 (6.1), 548 (8.8), 376 (7.5) cm^{-1}.

Synthesis of $[Ni((MeO)(4\text{-}MeOC_6H_4)PS_2)_2{\cdot}L2]_\infty$, $(3{\cdot}L2)_\infty$. Complex **3** (26.3 mg, 0.05 mmol) and **L2** (24.0 mg, 0.10 mmol) were reacted at 100 °C in a high pressure Aldrich tube in 25 mL of MeOH. After complete dissolving of the reagents, the reaction mixture was slowly cooled at room temperature. After a week, $(\mathbf{3{\cdot}L2})_\infty$ (29.5 mg, 0.031 mmol, 62% yield) was obtained as green crystals suitable for X-ray analysis. M.p.: 160 °C (d). Elemental analysis found (calc. for $C_{28}H_{28}N_4O_4P_2S_5Ni$; formula mass = 764.0 uma): C, 44.73 (43.93); H, 3.38 (3.69); N, 7.53 (7.32); S, 21.08 (20.94). FT-IR (KBr, 1600-350 cm^{-1}): 1593 s, 1569 mw, 1499 s, 1474 ms, 1432 w, 1406 m, 1331 w, 1294 s, 1254 vs, 1175 ms, 1113 vs, 1021 vs, 827 mw, 775 vs, 730 w, 654 vs, 623 s, 545 vs, 525 w, 436 w, 406 vw, 327 ms cm^{-1}. FT-Raman (4000-0 cm^{-1}, 100 mW, solid in KBr, relative intensities between parentheses related to the highest peak taken equal to 10.0): 3054 (0.7), 2850 (0.4), 2670 (0.3), 1618 (5.2), 1477 (2.6), 1418 (2.0), 1199 (6.0), 1156 (10.0), 1031 (2.6), 642 (0.8), 545 (1.2), 125 (4.2), 104 (2.4) cm^{-1}.

Synthesis of $[Ni((C_6H_5)_2PS_2)_2{\cdot}L2]_\infty$, $(4{\cdot}L2)_\infty$. A mixture of **4** (5.5 mg, 0.01 mmol) and **L2** (2.4 mg, 0.01 mmol) in 3 mL toluene was heated to 100 °C in a sealed 5 mL screw-top glass bottle for 3 days. The mixture was allowed to cool to room temperature and the resulting green crystals filtered from the reaction mixture; (4.3 mg, 0.0042 mmol, 42% yield). M.p.: 233 °C (d). Elemental analysis found (calc. for $C_{48}H_{36}N_8P_2S_6Ni$; formula mass = 1037.86 Da): C, 54.2 (55.6); H, 3.5 (3.5); N, 10.8 (7.1)%. FT-IR (KBr, 4000- 400 cm^{-1}): 3432 vw, 1637 w, 1587 m, 1497 s, 1471 vs, 1415 s, 1325 s, 1289 s, 1096 m, 1048 m, 1024 m, 998 m, 902 m, 815 s, 730 s, 693 s, 650 s, 611 m, 566 vs, 489 vs, 415 w cm^{-1}.

Characterisation

^1H- and ^{31}P-NMR spectra were recorded at 25°C in D_6-DMSO on a Bruker DPX400 NMR spectrometer with internal standards. Elemental analyses were performed with an EA1108 CHNS-O Fisons instrument. FT-Infrared spectra were recorded on a Thermo Nicolet 5700 spectrometer at room temperature using a flow of dried air. Middle IR spectra (resolution 4 cm^{-1}) were recorded as KBr pellets, with a KBr beam-splitter and KBr windows. X-ray structure determinations and crystallographic data for compounds $(\mathbf{1{\cdot}L1})_\infty$, $(\mathbf{2{\cdot}L1})_\infty$, $(\mathbf{3{\cdot}L1})_\infty$, $(\mathbf{1{\cdot}L2})_2$, and $(\mathbf{3{\cdot}L2})_\infty$, were collected at 120(2) K by means of combined phi and omega scans on a Bruker-Nonius Kappa CCD area detector, situated at the window of a FR591rotating anode (graphite Mo-K_α radiation, λ = 0.71073Å). Data for compound $(\mathbf{2{\cdot}L2})_2$, were collected at 120(2) K by means of fine-slice/omega scans on Bruker SMART APEX2 CCD diffractometer with Daresbury SRS station 9.8 synchrotron source (silicon 111, λ = 0.6893Å). Data for compound $(\mathbf{4{\cdot}L1})_\infty$, were collected at 173(2) K by means of combined phi and diffractometer with enhance X-ray source

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Table 1 Selected bond lengths (Å), bond and torsion angles (°), and angles between pyridyl (Py)/thiadiazole (Tdz) ring mean planes (°) for $(\mathbf{1}\cdot\mathbf{L1})_\infty$, $(\mathbf{2}\cdot\mathbf{L1})_\infty$, $(\mathbf{3}\cdot\mathbf{L1})_\infty$, and $(\mathbf{4}\cdot\mathbf{L1})_\infty$. Numbering scheme according to Fig. S1.[b]

	$(\mathbf{1}\cdot\mathbf{L1})_\infty$	$(\mathbf{2}\cdot\mathbf{L1})_\infty$	$(\mathbf{3}\cdot\mathbf{L1})_\infty$	$(\mathbf{4}\cdot\mathbf{L1})_\infty$
Ni–N1	2.095(3)	2.1042(16)	2.1343(18)	2.1376(16)
Ni–N4	2.100(3)	2.1042(16)	2.137(2)	2.1052(16)
Ni–S1	2.5041(12)[a]	2.4846(5)	2.4876(6)[a]	2.4836(5)[a]
Ni–S2	2.4585(11)[a]	2.4683(5)	2.4662(6)[a]	2.4940(6)[a]
P1–S1	1.9735(15)[a]	1.9693(8)	1.9993(9)[a]	2.0001(8)[a]
P1–S2	1.9814(15)[a]	1.9884(8)	1.9974(9)[a]	2.0025(7)[a]
N1–Ni–S1	91.02(8)	89.72(5)	90.94(5)	91.22(4)
N1–Ni–S2	89.58(9)	90.39(5)	89.12(5)	88.76(5)
S1–Ni–S2	82.04(4)[a]	82.010(17)	82.55(2)[a]	82.42(2)[a]
S1–P1–S2	110.95(7)[a]	110.39(3)	109.70(4)[a]	110.2(4)[a]
C2–C3–C6–N2	19(3)	31.9(3)	34.2(3)	4.7(3)
C9–C8–C7–N2	17(3)	31.9(3)	20.3(3)	11.0(3)
Py(N1)^Py(N4)	25.8	63.7	31.9	16.1
Py(N1)^Tdz	20.1	32.9	36.5	4.9
Py(N4)^Tdz	17.3	32.9	20.0	11.4

[a] Average of the bond parameters for the two symmetry independent fragments ($P_1S_1S_2Ni_1$) and ($P_2S_3S_4Ni_2$). [b] Due to the different space groups own by the polymers, as a consequence, to the different equivalent atoms imposed by symmetry, each structure had to be solved with different labels. In order to compare structural parameters of differently labelled polymers, Table 1 refers to the common artificial numbering scheme shown in Fig. S1.

(graphite Mo-K_α radiation). Data for compound $(\mathbf{4}\cdot\mathbf{L2}\cdot 2C_7H_8)_\infty$, were collected at 173(2) K by means of omega scans on an Oxford Diffraction Xcalibur PX Ultra diffractometer with enhance ultra (Cu) X-ray source (graphite Cu-K_α radiation, λ = 1.54184Å). The structures were solved by direct methods, SHELXS-97 and refined on F^2 using SHELXL-97.[12,13] Anisotropic displacement parameters were assigned to all non-hydrogen atoms. Hydrogen atoms were included in the refinement, but thermal parameters and geometry were constrained to ride on the atom to which they are bonded. The data of $(\mathbf{1}\cdot\mathbf{L1})_\infty$, $(\mathbf{2}\cdot\mathbf{L1})_\infty$ and $(\mathbf{3}\cdot\mathbf{L1})_\infty$ were corrected for absorption effects using SADABS V2.10.[14] The data for $(\mathbf{4}\cdot\mathbf{L1})_\infty$ and $(\mathbf{4}\cdot\mathbf{L2})_\infty$ were refined using CrysAlis RED,[15] implemented in SCALE3 ABSPACK scaling algorithm for empirical absorption correction using spherical harmonics. Structures have been deposited with the Cambridge Crystallographic Data Centre: deposition numbers

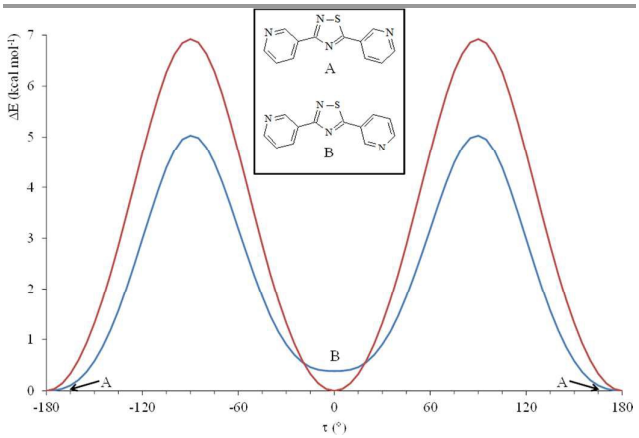

Fig. 1 Relative electronic energy variation ΔE as a function of the rotation angle τ of one pyridine ring (rotation step 5.0°) calculated at DFT level for **L1** (red) and **L2** (blue). In the inset the two possible *cisoids* (A) and *transoid* (B) **L2** conformers are depicted. $\Delta E_B - \Delta E_A$ = 0.39 kcal mol^{-1}.

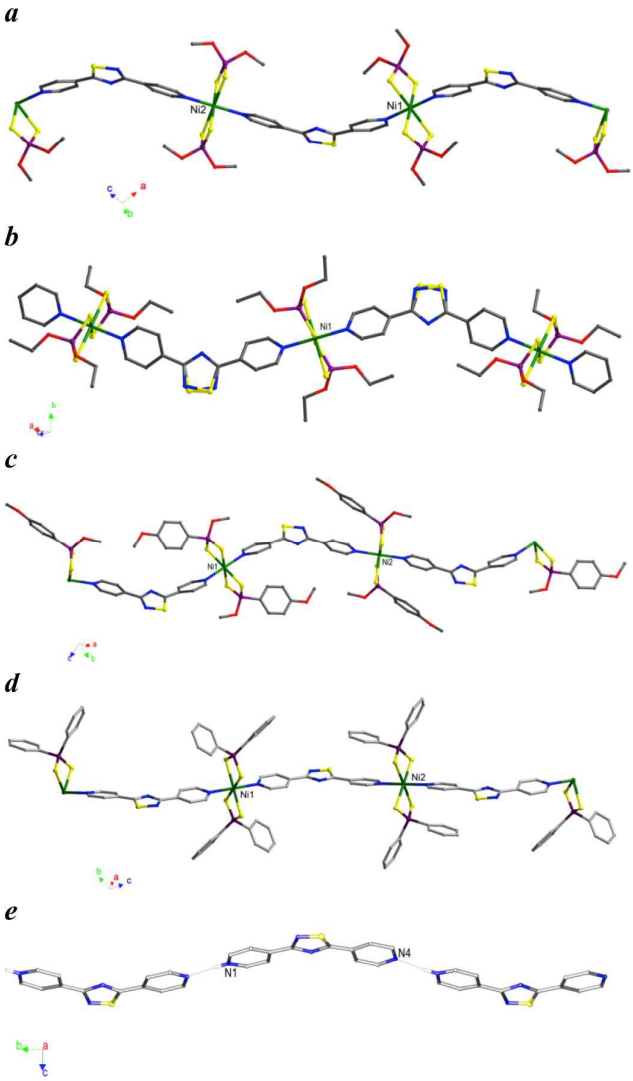

Fig. 2 Polymeric and protonated chains involving ligand **L1** in the compounds $(\mathbf{1}\cdot\mathbf{L1})_\infty$ (***a***), $(\mathbf{2}\cdot\mathbf{L1})_\infty$ (***b***), $(\mathbf{3}\cdot\mathbf{L1})_\infty$(***c***), $(\mathbf{4}\cdot\mathbf{L1})_\infty$ (***d***), and $(\mathbf{L1H}^+)_\infty$ (***e***; ref. 10).

CCDC 1474147, 1474143, 1474148, 1474145, 1474150, 1474144, 1474146, and 1474149, for $(\mathbf{1}\cdot\mathbf{L1})_\infty$, $(\mathbf{2}\cdot\mathbf{L1})_\infty$, $(\mathbf{3}\cdot\mathbf{L1})_\infty$, $(\mathbf{4}\cdot\mathbf{L1})_\infty$, $(\mathbf{1}\cdot\mathbf{L2})_2$, $(\mathbf{2}\cdot\mathbf{L2})_2$, $(\mathbf{3}\cdot\mathbf{L2})_\infty$, and $(\mathbf{4}\cdot\mathbf{L2}\cdot 2C_7H_8)_\infty$,, respectively. Theoretical calculations based on Density Functional Theory (DFT)[16] were carried out on **L1** and **L2** with the Gaussian09[17] commercial suite of software by adopting the mPW1PW[18] functional and Schäfer, Horn, and Ahlrichs double-zeta plus polarisation (pVDZ) all-electron basis sets (BSs) for all atomic species.[19] A potential energy surface was carried out by rotating by an angle τ one of the pyridine rings (-180.0° ≤ τ ≤ 180.0°; Δτ = 5.0°). The programmes GaussView 5.0.8 and Molden 5.2 were used to investigate the charge distributions and MO shapes.[20]

Results

The *bis*-functional ligands **L1** and **L2** (Scheme 1) were first isolated by Meltzer *et al.* in 1955,[21] but their complexing ability

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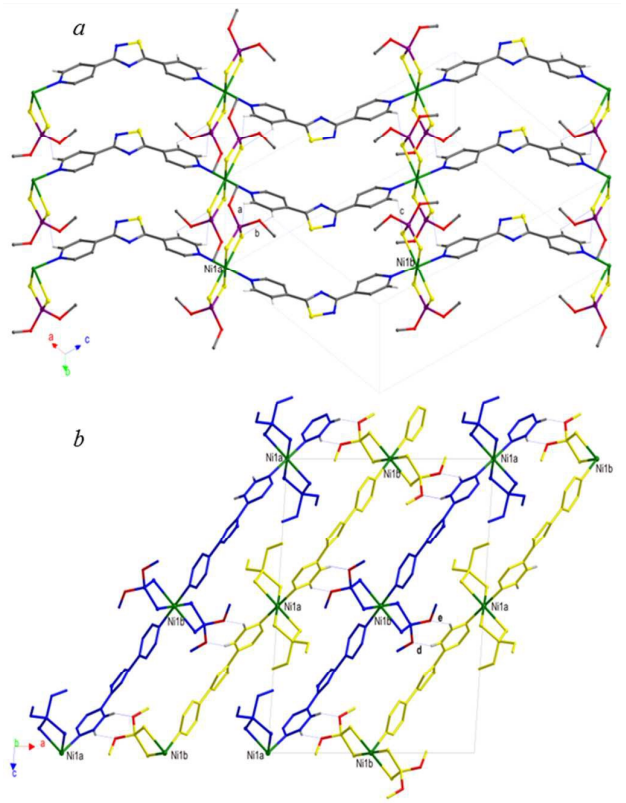

Fig. 3 Packing view of (**1**·**L1**)$_\infty$ showing parallel chains interacting through C–H···S interactions (***a***; H···S distances and C–H···S angles: a, C9^{i}–H9^{i}···S2 2.89(3), 120(2); b, C8^{i} –H8^{i} ···S5 2.85(3), 122(2); c, C15^{i}–H15^{i}···S3 2.91(4) Å, 126(2)°. ^{i} x, –1+y, z); (***b***) view along the 010 direction with polymeric chains coloured in blue and yellow according to their orientation. The Ni^{II} ions and the O and H atoms involved in the described H-bonds (d, C6^{ii}–H6^{ii}···O3 2.37(4), 3.288(5), 155(3); e, C5^{ii} –H5^{ii} ···O4 2.60(4) Å, 3.419(5) Å, 148(3)°; ^{ii} x, 1.5-y, 0.5+z) have been left of the conventional colours.

towards metal ions, or as Lewis donors, has not been investigated to date. In fact, a search in the Cambridge Crystallographic Data Base revealed only a recent work by Ondrejkovicova et al.[22] Both **L1** and **L2** contain two pyridyl groups linked by a 1,2,4-thiadiazole ring acting as a rigid, non-reactive spacer. Compared with the popular 4,4'-bipyridine linker, the pyridyl rings in **L1** and **L2** feature different geometry and separation lengths (9.95 and 9.60 Å, respectively). Moreover, due to the different position of the nitrogen atoms in **L1** and **L2** and to the different rotational conformations possible for the pyridyl rings, several orientations of the binding sites can be expected. A Potential Energy Surface (PES) analysis carried out on **L1** and **L2** at DFT level in the gas phase by rotating one pyridyl substituent by an angle τ ranging between -180 and 180° clearly shows that both donors display an energy minimum when the central 1,2,4-thiadiazole ring is coplanar with the pyridine moieties (Fig. 1). In the case of **L2**, two planar isomers, *cisoid* and *transoid*, are possible, differing in energy by less than 1 kcal mol^{-1} and showing similar metric parameters. As previously described, **L1** and **L2** feature the highest Kohn-Sham molecular orbitals localized on the negatively charged nitrogen atoms of the two pyridine moieties, which are therefore available to behave as donor sites towards Lewis unsaturated metal complexes.[10] The reactions of **L1** with nickel dithiophosphato [Ni((RO)$_2$PS$_2$)$_2$] [R = Me (**1**), Et (**2**); Scheme 1]; dithiophosphonato [Ni(MeOdtp)$_2$] (**3**) [MeOdtp = (CH$_3$O)(4-MeOC$_6$H$_4$)PS$_2^-$, Scheme 1], and dithiophosphito [Ni(Ph$_2$PS$_2$)$_2$] (**4**) complexes under solvothermal conditions afforded solid, crystalline compounds, which were isolated and identified by means of single crystal X-ray diffraction as coordination polymers of formula (**1**·**L1**)$_\infty$, (**2**·**L1**)$_\infty$, (**3**·**L1**)$_\infty$, and (**4**·**L1**)$_\infty$, respectively. Crystallographic data and selected bond lengths and angles are reported in Tables S1 and 1 respectively. With the exception of compound (**4**·**L1**)$_\infty$, the **L1** molecules are found in two different but essentially superimposable orientations, arising from a 180° rotation of the molecule about the direction passing through the midpoint of the N-S bond and through the remaining nitrogen atom of the penta-atomic ring (Figs. S1–S3). As a consequence, the nitrogen and the adjacent sulfur atoms show fractional occupancies refined at values of about 60% and 40%, respectively, and only the major orientation of the molecule is illustrated in the figures. In the case of compound (**2**·**L1**)$_\infty$ the **L1** molecule is located about a crystallographic two-fold axis and therefore the resulting disorder is modelled with occupancies of 50% for each orientation.

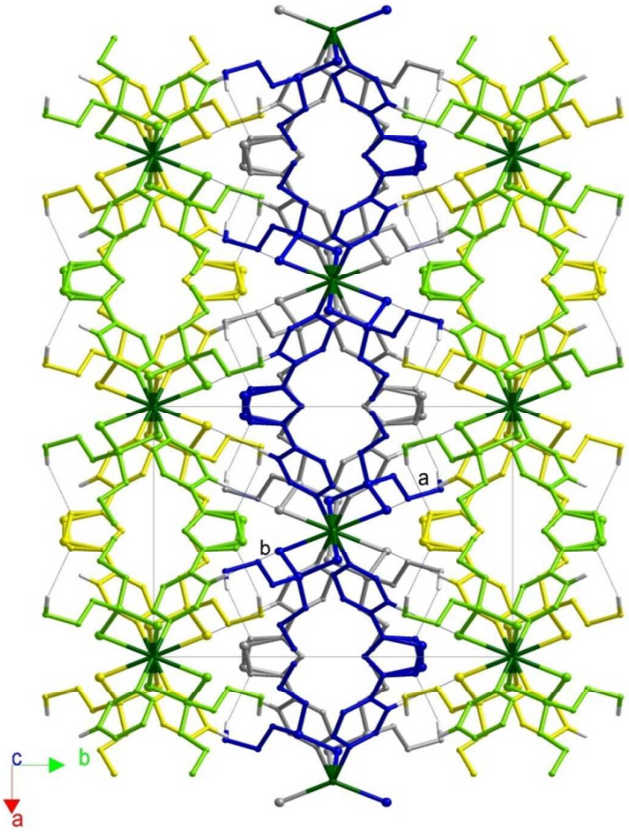

Fig. 4 Packing view along the 001 direction of (**2**·**L1**)$_\infty$ showing polymeric chains differently coloured according to their orientation. The Ni^{II} ions and the H atoms involved in the described H-bonds [a, C2–H2a···N3^{i}/S3^{i} 2.54(4)/2.73(3), 3.426(10)/3.617(5), 147(2)/148(2); C5–H5···S2^{ii} 2.85(2) Å, 3.364(2) Å, 115(2)°; ^{i} x, –y, 0.5+z; ^{ii} –0.5+x, 0.5–y, 1-z] have been left of the conventional colours. H-atoms not involved in showed interactions have been omitted for clarity. The labels used for the described interactions refer to the original numbering scheme (Fig. S2).

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Coordination polymers (**1**–**4**·**L1**)$_\infty$ show similar coordination environments: the nickel ions display distorted octahedral geometries with the equatorial plane occupied in iso-bidentate bonds with two dithiophosphoric ligands and the pyridine rings belonging to spacer **L1** axially bridging adjacent Ni^{II} ions to form infinite polymeric chains. The relevant bond lengths and angles (Table 1) are similar to those found in analogous coordination polymers.[7,8,9] The structures of the coordination polymers share the presence of neutral undulated polymeric chains with very similar pitches: Ni-Ni distances, through coordinate bonds, of 13.70, 14.15, 13.97, and 14.02 Å for (**1**·**L1**)$_\infty$, (**2**·**L1**)$_\infty$, (**3**·**L1**)$_\infty$, and (**4**·**L1**)$_\infty$, respectively (Fig. 2). These chains are closely analogous to the $C_1^1(12)$ chains of type [···(Py∩PyH$^+$···Py∩PyH$^+$)$_n$···] built up through the moderate NH$^+$···N bonds derived by the protonation of **L1** molecules, reported for comparison in Fig. 2e.[10] This confirms that the orientation of the nitrogen atoms *para*-positioned in the outwards pyridyl rings of **L1** self-govern the geometry of the resulting supramolecular aggregates, leading to similar shapes independent of the nature of the interacting Lewis acid. The asymmetric unit of compound (**1**·**L1**)$_\infty$ (Fig. S2, Fig 2a) contains two independent Ni^{II} ions, namely Ni1 and Ni2, lying on crystallographic inversion centres with coordination sites differing in the orientation of the methoxy substituents at the phosphorous atoms. The two coordination sites are bridged by **L1** molecules thus forming 1D parallel chains propagating along the 101 direction. These chains interact to each other through C–H···S short contacts, involving the pyridine rings and the coordinated sulphur atoms, leading to the 2D layers shown in Fig. 3a. The polymeric chains pack in two different orientations (coloured blue and yellow in Fig. 3b and S3) generated by inversion along the screw axes parallel to the 010 direction. As a consequence, symmetry related 2D layers formed by differently oriented chains alternate when packed and interact through synergic C–H···O bonds ("d" and "e" in Fig. 3b and S3a) involving the pyridine rings and the methoxy P-substituents belonging to Ni1b coordination sphere in a $R_2^2(7)$ motif. This packing arrangement enables the formation of "rippled" void channels of about 60 Å^3 corresponding to 2.2% of the unit cell volume (Fig. S3b). The presence of OEt P-substituents in place of the OMe ones in the compound (**2**·**L1**)$_\infty$ (Fig. 2b, Fig. S4) does not result in significant changes in either the Ni^{II} coordination environment or in the primary motif of the polymer and no relevant intramolecular interactions are worthy of note. The polymeric chains assume four symmetry related orientations as illustrated in the colour scheme in Fig. 4. Pairs of symmetry related chains propagating in the [10(1/2)] (yellow and green in Fig. 4) and [–10(1/2)] (blue and grey in Fig. 4) directions, respectively, interact through weak H-bonds involving the P-substituents and the central thiadiazole ring of **L1** ("a" in Fig. 4) form 2D layers which stack along the [010] direction leading to a very dense packing with no voids.

The coordination polymers (**3**·**L1**)$_\infty$ and (**4**·**L1**)$_\infty$ (Figs. 2c, 2d, S5 and S6) are characterised by the presence of aryl substituents at the phosphorous atoms which are expected to engender additional *intra*- and *inter*-molecular aromatic interactions. Similarly to those described previously herein, the asymmetric units of compounds (**3**·**L1**)$_\infty$ and (**4**·**L1**)$_\infty$ contain two independent Ni^{II} ions located on crystallographic inversion centres and differing only in the orientation of the aryl P-substituents. In the coordination polymer (**3**·**L1**)$_\infty$ the interactions involving the aromatic rings are intramolecular in nature (edge-to-face interaction "a" in Fig. 5). Weak interactions between the methoxy groups (interactions "b" and "c" in Fig. 5) connect the chains in layers which pack parallel to each other, forming a three-dimensional network through C–H···S interactions mainly involving the coordinated sulphur atoms and the methoxy substituents (interactions "d"-"g" in Fig. 5Sb).

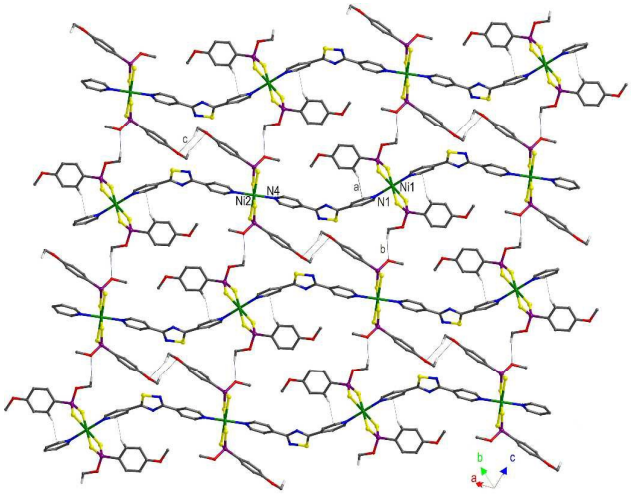

Fig. 5 Packing views of (**3**·**L1**)$_\infty$ showing layers of undulated chains; a, C2–H2···CntPy(N1) 2.62 Å; b, C8–H8c···O3^i 2.24, 3.171(3), 158.0; c, C15–H15a···O4^{ii} 2.53 Å, 3.237(3) Å, 129.0°; ^{i} –1+x, –1+y, z; ^{ii} 2–x, 2–y, –z. H-atoms not involved in showed interactions have been omitted for clarity. The labels used for the described interactions refer to the original numbering scheme (Fig. S5).

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Table 2 Selected bond lengths (Å), bond and torsion angles (°), and angles between pyridyl (Py)/thiadiazole (Tdz) ring mean planes (°) for $(\mathbf{1}\cdot\mathbf{L2})_2$, $(\mathbf{2}\cdot\mathbf{L2})_2$, $(\mathbf{3}\cdot\mathbf{L2})_\infty$, and $(\mathbf{4}\cdot\mathbf{L2}\cdot 2C_7H_8)_\infty$. Numbering scheme according to Fig. S7.

	$(\mathbf{1}\cdot\mathbf{L2})_2$	$(\mathbf{2}\cdot\mathbf{L2})_2$	$(\mathbf{3}\cdot\mathbf{L2})_\infty$	$(\mathbf{4}\cdot\mathbf{L2})_\infty\cdot 2C_7H_8)_\infty$
Ni–N1	2.093(5)	2.103(5)[b]	2.160(2)	2.129(6)
Ni–N4	2.099(5)	2.100(5)[b]	2.160(2)	2.129(6)
Ni–S1	2.4928(18)[a]	2.500(2)[b]	2.4915(6)	2.4809(15)
Ni–S2	2.4846(18)[a]	2.487(2)[b]	2.4871(6)	2.4995(17)
P1–S1	1.979(2)[a]	1.969(3)[b]	1.9974(8)	2.0059(19)
P1–S2	1.981(2)[a]	1.976(3)[b]	1.9915(9)	1.997(2)
N1–Ni–S1	90.66(15)	89.93(19)[b]	88.91(5)	90.82(11)
N1–Ni–S2	88.99(15)	89.76(18)[b]	90.52(5)	89.35(11)
S1–Ni–S2	81.89(10)[a]	81.98(8)[b]	81.86(2)	83.00(5)
S1–P1–S2	110.93(10)[a]	112.0(2)[b]	109.71(4)	111.08(8)
N1–Ni–N4	175.7(2)	176.3(2)[b]	180.00	180.00
N2–C6–C3–C2	9.6(3)	1.4(11)[b]	11.7(3)	0.0
N2–C7–C8–C9	17.9(3)	4.4(11)[b]	11.7(3)	0.0
Py(N1)^Py(N4)	12.3	5.0[b]	19.2	0.0
Py(N1)^Tdz	4.3	3.0[b]	14.7	0.0
Py(N4)^Tdz	20.3	7.2[b]	14.7	0.0

[a] Average of the bond parameters for the two fragments ($P_1S_1S_2Ni$) and ($P_2S_3S_4Ni$). [b] Average of the bond parameters for the symmetry independent coordination environment around Ni1 and Ni2.

Aromatic interactions become prevalent in polymer $(\mathbf{4}\cdot\mathbf{L1})_\infty$ due to the presence of phenyl substituents at the phosphorous atoms, so that the polymeric chains pack in layers built up by face-to-face and edge-to-face interactions only ("a"-"g" in Fig. 6). The layers pack parallel to the *a* axis through weak C–H···S interactions ("h"-"j" in Fig. 6Sb). Despite the fact that all the coordination polymers featuring **L1** as a spacer exhibit the same primary motif found in the supramolecular aggregates of $\mathbf{L1}H^+$ (Fig. 2e), the presence and orientation of the P-substituents influences the final architecture of the polymers via aromatic or C-H···S intramolecular interactions. The main consequence is the loss of planarity of the bridging **L1** ligands (Table 1), probably ascribable to the conformational arrangement the pyridine rings adopt to optimize these interactions. It is interesting to note that in the H-bonded chains built up by protonated **L1** molecules, where such interactions are not present, the ligands retain planarity.[10]

The reactions of **L2** with the nickel complexes **1–4** under solvothermal conditions afforded crystalline compounds recognised by means of single crystal X-ray diffraction as the dimers $(\mathbf{1}\cdot\mathbf{L2})_2$, $(\mathbf{2}\cdot\mathbf{L2})_2$, and the coordination polymers $(\mathbf{3}\cdot\mathbf{L2})_\infty$, and $(\mathbf{4}\cdot\mathbf{L2})_\infty$ in the compound $(\mathbf{4}\cdot\mathbf{L2}\cdot 2C_7H_8)_\infty$, respectively (Figs. 7–10, S8–S11). The two different constructs reflect the different conformations of **L2** behaving as a convergent linker in the dimeric structures and as divergent linker in the polymers. Crystallographic data and selected bond lengths and angles are reported in Tables S2 and 2. Similarly to the previously discussed cases of polymers $(\mathbf{1–3}\cdot\mathbf{L1})_\infty$, in the crystal structures of the dimer $(\mathbf{1}\cdot\mathbf{L2})_2$ and the polymers $(\mathbf{3–4}\cdot\mathbf{L2})_\infty$ the spacer **L2** is located in two essentially superimposable orientations. As a consequence, the nitrogen and the adjacent sulfur atoms show fractional occupancies refined at values of 68.8 and 31.2% for $(\mathbf{1}\cdot\mathbf{L2})_2$, and 61.5 and 31.5% for $(\mathbf{4}\cdot\mathbf{L2})_\infty$ (the figures represent the molecules in the major orientation only).

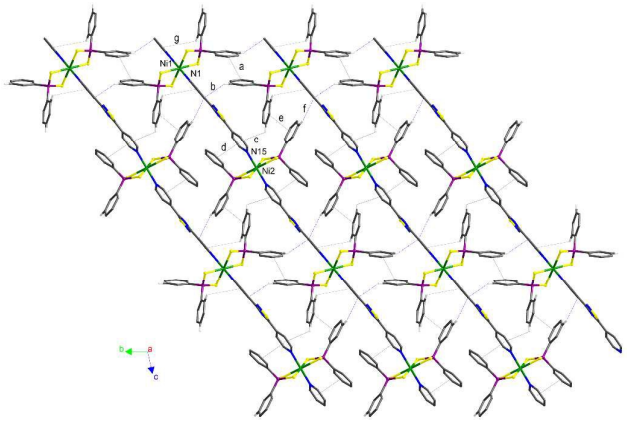

Fig. 6 Packing view of $(\mathbf{4}\cdot\mathbf{L1})_\infty$ showing layers built up by the aromatic interactions a-g: a, $Cnt_{Ph(C24\text{-}C29)}\cdots Cnt_{Ph(C24\text{-}C29)'}$ 4.16, 0; b, $Cnt_{Py(N1)}\cdots H26^{i}–C26^{i}$ 3.32, 49; c, $Cnt_{Py(N15)}\cdots H21a^{i}–C21^{i}$ 3.08, 133; d, $Cnt_{Py(N15)}\cdots Cnt^{ii}_{Ph(C36\text{-}C41)}$ 3.67, 11; e, $Cnt_{Ph(C30\text{-}C35)}\cdots Cnt^{i}_{Ph(C18\text{-}C23)}$ 3.90, 19; f, C34–H34a···$Cnt^{iii}_{Py(N1)}$ 3.58, 109; g, $C19^{iv}$–$H19a^{iv}\cdots Cnt_{Py(N1)}$ 3.90 Å, 129°. Symmetry codes: [i] –x, 1–y, 1–z; [ii] 1–x, 1–y, 2–z; [iii] x, –1+y, z; [iv] –x, 2–y, –z. H-atoms not involved in showed interactions have been omitted for clarity. The labels used for the described interactions refer to the original numbering scheme (Fig. S6).

In the case of compound $(\mathbf{3}\cdot\mathbf{L2})_\infty$ the **L2** molecule is located about a crystallographic two-fold axis and therefore the resulting disorder is modelled with occupancies of 50% for each orientation. The asymmetric unit of the dimer $(\mathbf{2}\cdot\mathbf{L2})_2$ contains two independent units featuring differently oriented ligands **L2** (Fig. 7b). Both dimers $(\mathbf{1–2}\cdot\mathbf{L2})_2$ and polymers $(\mathbf{3–4}\cdot\mathbf{L2})_\infty$ display octahedral Ni^{II} coordination environments similar to those previously discussed, with two iso-bidentate dithiophosphoric ligands in the equatorial plane and the axial positions occupied by the pyridine rings of **L2** bridging spacers. The bond lengths and angles (Table 2) are also similar to those found in the analogous coordination polymers $(\mathbf{1–4}\cdot\mathbf{L1})_\infty$. Notwithstanding similarities with the results obtained with **L1**, the use of **L2** as a linker leads to different constructs. In fact, as evidenced by DFT calculations (see above), **L2** can exist either as a *transoid* or *cisoids* isomer and a convergent or a divergent conformation can be distinguished depending on the orientation of N-atoms of the pyridyl rings that can point inwards or outwards with respect to the pinch angle. In $(\mathbf{1}\cdot\mathbf{L2})_2$ and $(\mathbf{2}\cdot\mathbf{L2})_2$ the pyridyl rings of ligand **L2** are oriented in a convergent fashion leading to closed rings rather than polymeric chains (Fig. 7a-b). It is interesting to note that this construct is far from predictable in its behaviour since a convergent fashion does not necessarily lead to closed rings. This is demonstrated by the formation of the mono-dimensional $C_1^1(10)$ spirals found in the two structures of $(H\mathbf{L2})I_3$ and $(H\mathbf{L2})I_3$ assembled through $NH^+\cdots N$ bonds between protonated **L2** molecules exhibiting a convergent *cisoids* conformation.[10] In dimers $(\mathbf{1}\cdot\mathbf{L2})_2$ and $(\mathbf{2}\cdot\mathbf{L2})_2$ two spacers bridge two dithiophosphato nickel complexes through axial coordination generating eicos-atomic planar wheels with openings of about 8×8 $Å^2$, and inner Ni···Ni distances of 7.73, 7.77, and 7.58 Å for $(\mathbf{1}\cdot\mathbf{L2})_2$ and the two independent units of $(\mathbf{2}\cdot\mathbf{L2})_2$, respectively.

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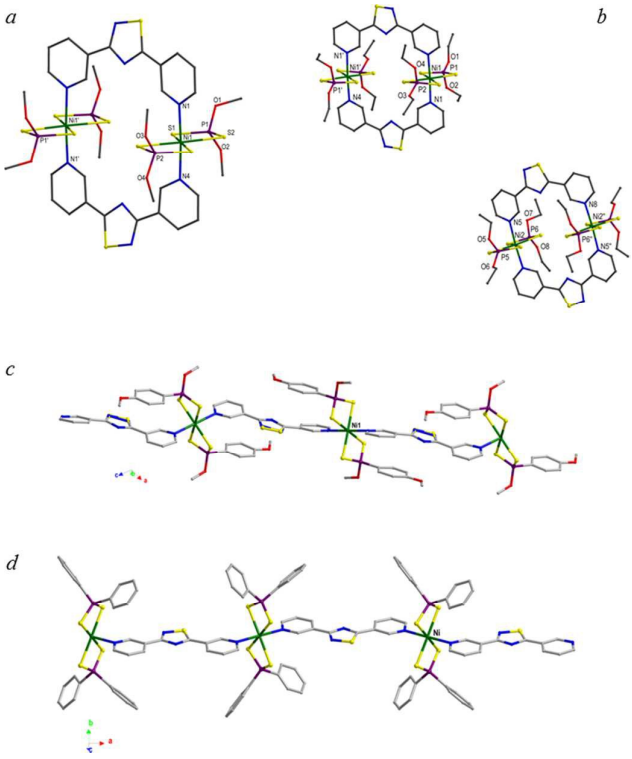

Fig. 7 Dimeric units and polymeric chains for $(\mathbf{1}\cdot\mathbf{L2})_2$ (***a***), $(\mathbf{2}\cdot\mathbf{L2})_2$ (***b***), $(\mathbf{3}\cdot\mathbf{L2})_\infty$ (***c***), and $(\mathbf{4}\cdot\mathbf{L2})_\infty$ in compound $(\mathbf{4}\cdot\mathbf{L2}\cdot 2C_7H_8)_\infty$ (***d***). The labels reported refer to the original numbering schemes (Figs. S8–S12). H-atoms have been omitted for clarity. Only the labels used in the discussion are here reported.

Despite the analogies in their structures, the dimers $(\mathbf{1}\cdot\mathbf{L2})_2$ and $(\mathbf{2}\cdot\mathbf{L2})_2$ show different packing arrangements. The dimers $(\mathbf{1}\cdot\mathbf{L2})_2$ are arrayed in regular perforated layers assembled by N···H and S···H interactions (Fig. 8a). Symmetry related parallel layers pack in an off-set compact arrangement along the *b* direction showed in Fig. 8.

In the crystal packing of $(\mathbf{2}\cdot\mathbf{L2})_2$ the two symmetrical independent units (containing Ni1 and Ni2 ions) that, as previously observed differ in the orientation of the thiadiazole ring (Tdz), interact with each other through π–π interactions involving the pyridine and Tdz rings forming puckered ribbons (Fig. 9a). A view along the *a* axis (Fig. 9b) shows the ribbons aligned in a parallel arrangement through H-bonds and weaker interactions mainly involving the sulphur atoms and the pyridine hydrogens. A *cisoid* but divergent conformation of **L2** leads to the coordination polymers $(\mathbf{3}\cdot\mathbf{L2})_\infty$ and $(\mathbf{4}\cdot\mathbf{L2})_\infty$ (Figs. 7c–d, S10–S12). The asymmetric units of both compounds contain one independent Ni^{II} ion situated on a crystallographic inversion centre in the first case and on a mirror plane in the latter. The polymeric supramolecular constructs present close analogies with the $C_1^1(10)$ chains formed by head-to-tail $NH^+\cdots N$ bonds between adjacent pyridine rings of protonated **L2** molecules featuring a *trans* conformation.[10] Both complexes **3** and **4** feature aromatic substituents that are involved in π-type interactions which govern the assembly of the relevant polymers. In $(\mathbf{3}\cdot\mathbf{L2})_\infty$ the aromatic P-substituents engage in intramolecular π-π interactions with the facing pyridine rings and intermolecular C–H···O interactions with the oxygen belonging to the MeO P-substituents which connect the chains in layers (interactions a and b in Fig. S12a). The layers pack in parallel through C–H···S interactions involving the *para*-methoxy group of the Ar-P substituents and the coordinated sulphur atoms (interactions c-e in Fig. S12b), thus forming an overall three-dimensional network. Solvent accessible voids of 69 $Å^3$ (in blue in the inset of Fig. S12), correspond to 4% of the unit cell volume. In polymer $(\mathbf{4}\cdot\mathbf{L2})_\infty$ the exclusive presence of phenyl substituents at the phosphorous atoms results in the aromatic interactions becoming prevalent such that the parallel polymeric chains pack through edge-to-face interactions only (Fig. 10a). The resulting network features empty channels suitable for the inclusion of toluene molecules which engage in π-interactions with the phenyl rings (Fig. 10b–c). It is interesting to note that, in contrast to the previously discussed structures (see also Tables 1 and 2), $(\mathbf{4}\cdot\mathbf{L2})_\infty$ contains perfectly planar **L2** ligands.

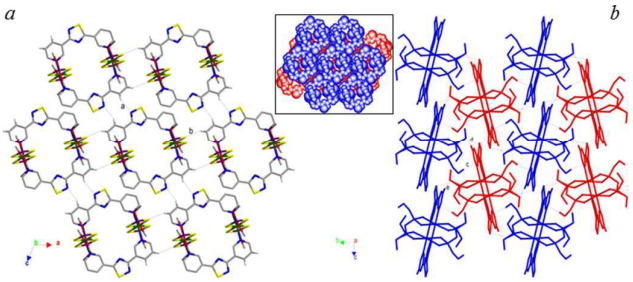

Fig. 8 Packing views showing the (***a***) layers formed by interacting $(\mathbf{1}\cdot\mathbf{L2})_2$ dimers and (***b***) packing of intercalating layers evidenced in different colours; in the inset the intercalating layers packing along the 010 direction are presented. All the hydrogen atoms with the exception of those involved in the showed interactions have been omitted. Interactions: a: C14i–H14^{i}···N3, 2.70(9), 3.331(13), 124; b: C15^{ii}–H15^{ii}···S2, 2.93(9), 3.618(8), 127; c: C1^{iii}–H1b^{iii}···N2, 2.57, 3.365(19), 138; d: C2^{iii}–H2a^{iii}···S5, 2.88, 3.542(8), 126; e: C8^{iv}–H8^{iv}···O3, 2.59 Å, 3.504(9) Å, 161°. Symmetry codes: ^{i} −1+x, y, −1+z; ^{ii} 1−x, 1−y, 1−z; ^{iii} −x, 0.5+y, 0.5−z; ^{iv} x, 1.5−y, 0.5+z.

Conclusions

The reaction of **L1** with the differently P-substituted dithiophosphonato, dithiophosphato, and dithiophosphito complexes **1–4** yielded the corresponding coordination polymers $(\mathbf{1–4}\cdot\mathbf{L1})_\infty$ featuring polymeric assemblies $[\cdots(Py\cap PyH^+\cdots Py\cap PyH^+)_n\cdots]$ built up through the moderate $NH^+\cdots N$ bonds derived by the protonation of **L1** molecules. These results indicate that **L1** can be used as a spacer for the predictable assembly of smoothly undulating chains independent of the nature of the interacting Lewis acid, since the orientation of the nitrogen atoms *para*-positioned in the outwards pyridyl rings of **L1** self-governs the geometry of the resulting supramolecular construct. On the contrary, **L2** allows for the existence of different supramolecular constructs ensuing from different ligand conformations deriving from the rotation of the pyridyl rings.

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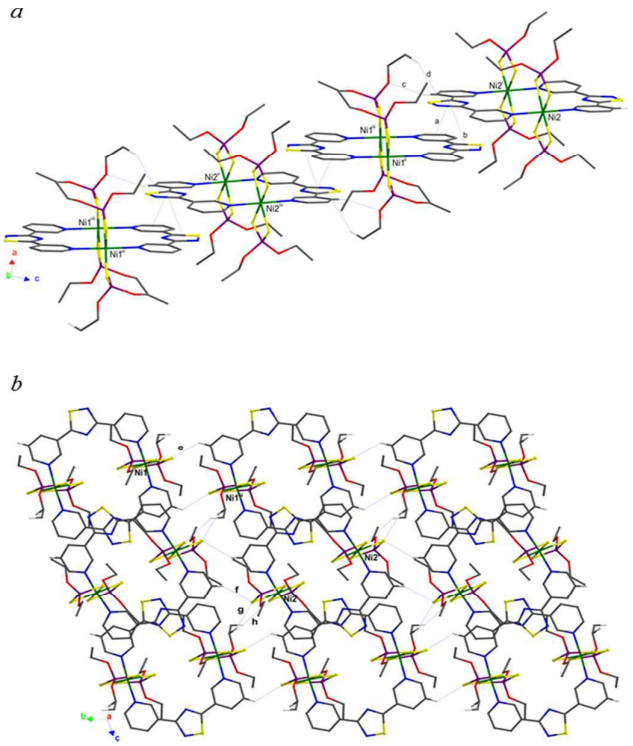

Fig. 9 Packing views showing (***a***) the ribbons formed by π-π interacting (**2**·**L2**)$_2$ dimers and (***b***) packing view along 100 direction of interacting ribbons. All the hydrogen atoms with the exception of those involved in the showed interactions have been omitted. Showed π–π interactions (a, b), weak contacts (c, d, e, g, h), and the H-bond (f): a, $Cnt_{Py(N8)}\cdots Cnt_{Ph(N1)}{}^{iii}$ 3.58, Py(N5)^Py(N4)^{iii} 1°; b, $Cnt_{Tdz(S15)}\cdots Cnt_{Tdz(S5)}{}^{ii}$ 3.44 Å, $Tdz_{(S15)}$^$Tdz_{(S5)}{}^{ii}$ 1°; c, C38–H38···O3^{iii} 2.40 Å; d, C6^{iii}–H6c^{ix}···C38, 2.78; e, C12^{ii}–H12···S2, 2.89; g, C1^{ix}–H1a^{ix}···S12, 2.99; h, C2^{ix}–H2a^{ix}···S13, 3.04 Å; f, C32^{viii}–H32^{viii}···S12, 2.77 Å, 3.596(9) Å, 146°. Symmetry codes: ^{i} 2–x, –y, 1–z; ^{ii} 1–x, 1–y, –z; ^{iii} x, –1+y, z; ^{iv} –1+x, y, –1+z; ^{v} 1–x,–y, –z; ^{vi} –x, 1–y, –1–z; ^{vii} –1+x, –1+y, –1+z; ^{viii} 2–x, 1–y, 1–z; ^{ix} 1+x, y, 1+z.

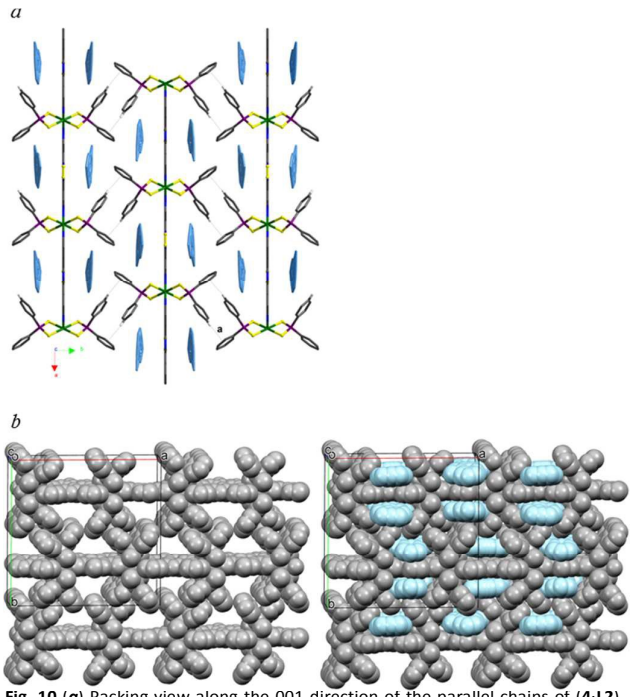

Fig. 10 (***a***) Packing view along the 001 direction of the parallel chains of (**4**·**L2**)$_\infty$ with the toluene molecules included in the crystal evidenced in light blue colour. All the hydrogen atoms have been omitted with the exception of those involved in the showed interaction a, C27–H27a···$Cnt_{(Ph(C18-C23))}{}^{i}$ 2.84 Å, ($Ph_{C24-C29}$)^ ($Ph_{C18-C23}$)^{i} 91.4°. Symmetry code: ^{i} 1.5–x, 1–y, 0.5+z. (***b***) Space-fill view of the network with (left) and without (right) the solvent molecules.

In particular, the results suggest that aromatic P-substituents capable of π-interacting with the aromatic rings of the ligand tend to favour divergent constructs. The influence of secondary interactions involving the P-substituents is confirmed by the loss of planarity of **L1** and **L2** in order to enhance inter-molecular packing interactions.

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Notes and references

‡ The 112 issue of *Chem. Rev.* (2012) and the 16 issue of *Chem. Soc. Rev.* (2014) were entirely devoted to the synthesis and applications of MOF's showing how rapidly this branch of solid state materials chemistry is being evolving.

§ Among the ligands most commonly employed as spacers the choice of using 4,4'-bipyridine and its analogues is due to their versatility. In fact, by introducing different groups between the two pyridyl rings a wide variety of either linear or bent, rigid or flexible spacers are available. See for example ref. 4.

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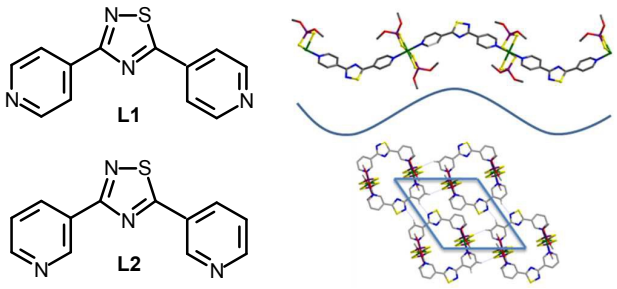
Coordination polymers and polygons using di-pyridyl-thiadiazole spacers and substituted phosphorodithioato NiII complexes: potential and limitations for inorganic crystal engineering M. C. Aragoni, M. Arca, S. J. Coles, M. Crespo, S. L. Coles (née Huth), R. P. Davies, M. B. Hursthouse, F. Isaia, R. Lai, V. Lippolis		Differently P-substituted phosphorodithioato Ni^{II} complexes were reacted with bis-pyridyl thiadiazole spacers L1 and L2 differing for the position of the donor atoms obtaining either polymeric or discrete supramolecular constructs.