



Cite this: DOI: 10.1039/c6tc04119a

Rational design and characterization of heteroleptic phosphorescent iridium(III) complexes for highly efficient deep-blue OLEDs†

Yansong Feng,^a Xuming Zhuang,^a Dongxia Zhu,^b Yu Liu,^{*a} Yue Wang^a and Martin R. Bryce^{*c}

Two new deep-blue iridium(III) complexes, (dfppy)₂IrFptz (**Ir1**) and (Medfppy)₂IrFptz (**Ir2**), comprising difluoro-bipyridyl (dfppy) derivatives as cyclometalating ligands and a chelated pyridyl-triazole (Fptz) ancillary ligand are reported. The bipyridyl ligands lead to a significantly increased HOMO–LUMO gap and a hypsochromic shift of the phosphorescence compared to phenylpyridyl analogs. Density function theory (DFT) calculations and electrochemical measurements for **Ir1** and **Ir2** support their genuine blue phosphorescent emission. The combination of ancillary and cyclometalating ligands in **Ir1** and **Ir2** significantly influences the molecular orbitals of both complexes, leading to clearly distinct electron density distributions of the HOMO and LUMO compared with other blue-emitting Ir(III) derivatives. Both complexes **Ir1** and **Ir2** show deep-blue emission with λ_{max} values in the region of 435–465 nm with high PLQYs and short excited-state lifetimes. The phosphorescent organic light emitting diodes (PhOLEDs) based on **Ir1** and **Ir2** achieve remarkably high EL performance with low efficiency roll-off at high luminance. The bluest color (CIE_{xy}, 0.14, 0.11) and the highest EL efficiency were achieved in the device based on **Ir2** (Device 2), where the peak EQE/PE of 13.0%/11.2 lm W^{−1} together with the corresponding values of 12.6%/8.8 lm W^{−1} and 10.1%/5.0 lm W^{−1} at the practical luminances of 100 and 1000 cd m^{−2} respectively, strongly compete with those of any deep-blue fluorescent and/or phosphorescent OLEDs with similar CIE coordinates previously reported.

Received 21st September 2016,
Accepted 12th October 2016

DOI: 10.1039/c6tc04119a

www.rsc.org/MaterialsC

Introduction

Since the pioneering report by Baldo *et al.*,¹ phosphorescent organic light-emitting devices (PhOLEDs) based on metal complexes have been intensively studied because internal quantum efficiencies (IQEs) approaching 100% can be achieved by harvesting both singlet and triplet excitons for light emission. The external quantum efficiencies (EQEs) of red and green PhOLEDs,² together with the sky-blue (y coordinate >0.20 for the Commission Internationale de L'Eclairage coordinates, CIE_{xy})³

and the quasi-blue ($0.15 \leq y \leq 0.20$) PhOLEDs,⁴ reach 20–25% which is *ca.* four times higher than those of conventional fluorescent OLEDs. Thus they have already been commercialized or are competing with the conventional fluorescent materials in OLEDs for displays and solid-state lighting. However, although much progress has been made in the development of new blue phosphorescent emitters, there are only a few reports of deep-blue PhOLEDs with CIE_x and y coordinates both ≤ 0.15 . Notably, devices with excellent blue chromaticity, that is, with x and y values simultaneously close to 0.15 and 0.10, respectively,⁵ showed EQE values of less than 8%, which are only slightly enhanced compared to the highest level of the conventional fluorescent blue OLEDs with similar CIE coordinates.⁶

Chou *et al.* prepared two iridium complexes using pyridyl azolates and benzyl phenylphosphine ligands,⁷ which are among the most efficient true-blue PhOLEDs documented to date with CIE_{xy} of $\sim(0.15, 0.11)$ and the peak EQEs/PEs (power efficiencies) of $\sim 11\%/8 \text{ lm W}^{-1}$, respectively. Nevertheless, they exhibited relatively high turn-on voltages of $\geq 4.4 \text{ V}$ as well as significant efficiency roll-off ($\sim 8\%$ and 3 lm W^{-1} at the practical luminance of 100 cd m^{-2}). Forrest and coworkers have very recently reported that PhOLEDs of an Ir complex of N-heterocyclic carbene ligands can attain EQE_{max} of $10.1 \pm 0.2\%$

^a State Key Laboratory of Supramolecular Structure and Materials, College of Chemistry, Jilin University, Changchun 130012, P. R. China. E-mail: yuliu@jlu.edu.cn

^b Institute of Functional Material Chemistry, Faculty of Chemistry, Northeast Normal University, Renmin Road 5268, Changchun, Jilin 130024, P. R. China

^c Department of Chemistry, Durham University, Durham, DH1 3LE, UK. E-mail: m.r.bryce@durham.ac.uk

† Electronic supplementary information (ESI) available: Synthesis and characterization details; differential scanning calorimetric (DSC) and thermogravimetric analysis (TGA) data; X-ray crystallographic data files (CIF), crystal data and structure refinement parameters; cyclic voltammograms for **Ir1** and **Ir2**. CCDC 1052634 and 1052689. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6tc04119a

at very low luminance and EQE $9.0 \pm 0.1\%$ at 1000 cd m^{-2} with CIE_{xy} 0.16, 0.09.⁸ Thus, it remains very important to develop blue phosphors for realizing very efficient deep-blue PhOLEDs that meet the requirements for practical applications.

Generally, the effective molecular design concept for deep-blue phosphors is to introduce electron-withdrawing or electron-donating groups at selective positions on the cyclometalated (C^N) ligands. For example, fluorine substitution leads to the 4',6'-difluorophenylpyridinato ligand in the well-known sky-blue iridium complex FIrpic,^{2b,3a,c,9} resulting in a larger band gap and blue shifted emission.¹⁰ Alternatively, fluorine-substituted 2,3'-bipyridine derivatives, where the additional nitrogen atom has a similar or even stronger effect of lowering the HOMO (highest occupied molecular orbital) level (as opposed to introducing the electron-withdrawing groups at the *para* position on the phenyl ring),¹¹ have shown potential as the cyclometalating ligands for efficient deep-blue iridium complexes. It is also well known that the blue color can be tuned through structural variations of the ancillary ligand.¹² For example, chelated pyridyl-triazole derivatives, which usually give very low total synthesis yields of $\leq 10\%$ for the resulting complexes, afford stable iridium complexes that show remarkably shorter emission maxima than that of FIrpic, due to their blue-shifting effect compared to picolinate and other ancillary ligands.^{4c,5b,10a,11c,12b}

The aim of the present work was to develop new phosphors and to fabricate deep-blue PhOLEDs with enhanced efficiency. To achieve this, we designed and characterized two new heteroleptic Ir(III) complexes (**Ir1** and **Ir2**) containing the unique combination of 2',6'-difluoro-2,3'-bipyridyl ligands (dfppy or Medfppy) and a 5-(2'-pyridyl)-3-trifluoromethyl-1,2,4-triazole ancillary ligand (Fptz). We demonstrate that both **Ir1** and **Ir2** display true-blue phosphorescence in their solution and film states, and they serve in the emitting layer (EML) of PhOLEDs that show deep-blue emission with very high EL efficiencies. The **Ir2**-based device (Device 2) represents the deepest blue emission with CIE_{xy} of (0.14, 0.11) near to the ideal value of (0.14, 0.08). The peak EQE/PE values are $13\%/11.2 \text{ lm W}^{-1}$, which maintain very high levels of $12.6\%/8.8 \text{ lm W}^{-1}$ and $10.1\%/5.0 \text{ lm W}^{-1}$ at luminances of 100 and 1000 cd m^{-2} , respectively. To the best of our knowledge, these EL efficiency values strongly compete with, and even exceed, those previously reported for PhOLEDs with similar CIE coordinates.^{5-7,13a} Additionally, an effective synthetic route for Fptz^{13b,c} provides high yields of **Ir1** and **Ir2** ($> 70\%$), facilitating the low cost of both complexes in the quest for deep-blue PhOLEDs in the future.

Experimental section

General information

All reagents and solvents obtained from commercial suppliers were used without further purification. Anhydrous tetrahydrofuran was distilled with sodium benzophenone ketyl under nitrogen atmosphere and degassed by the freeze-pump-thaw method. All glassware, syringes, magnetic stirring bars and needles were dried in a convection oven for at least 4 h. Reactions were monitored

with thin layer chromatography (TLC). Commercial TLC plates (silica gel 60 F254, Merck Co.) were visualized under UV light at 254 and 365 nm. Silica column chromatography used silica gel 60 G (particle size 5–40 μm , Merck Co.). ¹H NMR spectra were recorded on a 500 MHz Bruker Avance spectrometer using CDCl₃ as solvent and tetramethylsilane (TMS) as an internal standard ($\delta = 0.00 \text{ ppm}$). GC/MS data were recorded on a Thermo Fisher ITQ1100 instrument. UV-vis spectra were measured on a Shimadzu UV-2550 spectrophotometer. Fluorescence spectra were recorded using a Shimadzu RF-5301 PC spectrometer and a Maya2000Pro optical fiber spectrophotometer. Solid state PL efficiencies were measured using an integrating sphere (C-701, Labsphere Inc.), with a 365 nm Ocean Optics LLS-LED as the excitation source, and the light was introduced into the integrating sphere through an optical fiber. Single crystal X-ray diffraction intensity data were collected on a Rigaku RAXIS-PRID diffractometer using the ω -scan mode with graphite-monochromator Mo K α radiation. DSC experiments were recorded on a NETZSCH DSC 204 instrument at a scanning rate of 10 K min^{-1} . Thermogravimetric analysis (TGA) was performed on a TA Q500 thermogravimeter by measuring the weight loss while heating at a rate of $10 \text{ }^\circ\text{C min}^{-1}$ from 25 to $800 \text{ }^\circ\text{C}$ under nitrogen. Electrochemical measurements were performed with a BAS 100W Bioanalytical electrochemical work station, using Pt working electrode, platinum wire as auxiliary electrode, and a porous glass wick Ag/Ag⁺ as reference electrode. The voltammograms were referenced to the ferrocene/ferrocenium couple and the scan rate was 100 mV s^{-1} .

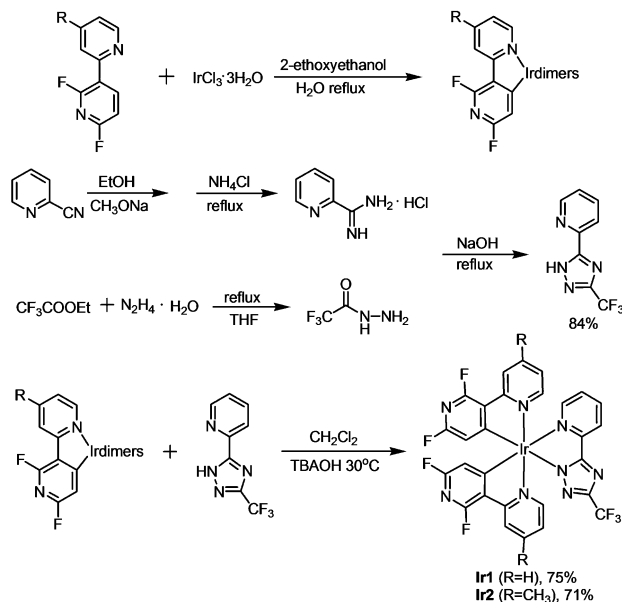
Fabrication of the OLEDs and EL measurements

The ITO (indium-tin oxide) coated glass substrates ($20 \Omega \text{ square}^{-1}$) were first cleaned in detergent solution, ethanol, acetone, isopropanol and deionized water. All organics were purified by gradient sublimation and thermally evaporated at a rate of $1.0 \text{ }^\circ\text{A s}^{-1}$ at a pressure of *ca.* $3.5 \times 10^{-4} \text{ Pa}$. A LiF layer (1 nm) was deposited at a rate of $0.2 \text{ }^\circ\text{A s}^{-1}$. The Al cathode was deposited at a rate of $10 \text{ }^\circ\text{A s}^{-1}$; the active area of the diode segments was $2 \times 2.5 \text{ mm}^2$. EL spectra and brightness-current density-voltage characteristics were measured by combining a Spectrascan PR-650 spectrophotometer with a computer-controlled direct current power supply Keithley model 2400 voltage-current source under ambient conditions at room temperature.

Results and discussion

Synthesis, single crystal structures and theoretical calculations

The chemical structures of **Ir1** and **Ir2** are shown in Scheme 1. The two cyclometalating ligands, dfppy and Medfppy, were synthesized by Suzuki coupling following the reported method.^{11a} The corresponding μ -chloro-bridged dimers [(dfppy)₂Ir(μ -Cl)]₂¹³ and [(Medfppy)₂Ir(μ -Cl)]₂^{11c} were prepared by the reaction of dfppy and Medfppy with IrCl₃·3H₂O in a mixture of 2-ethoxyethanol and water. Here, it is worth noting that overall yields of $> 70\%$ were obtained for both new iridium complexes **Ir1** and **Ir2**, which represent the most effective syntheses for deep-blue phosphors reported in literature.^{5b,11c}



Scheme 1 The synthesis and chemical structures of **Ir1** and **Ir2**.

The molecular structures of **Ir1** and **Ir2** in single crystals were determined by X-ray diffraction. As shown in Fig. 1a and Table S1 in the ESI,[†] both complexes exhibit distorted octahedral geometries around the iridium center with the classical C,C-*cis* and N,N-*trans* configurations. The bond lengths and angles are summarized in Table S2 (ESI[†]). The Ir–C bond lengths, ranging from 1.981 to 2.008 Å, and Ir–N (in dfppy) bond lengths, ranging from 2.040 to 2.055 Å, are comparable to those observed in several blue-emitting Ir(III) complexes (Ir–C = 1.997–2.005 Å for Ir(dfppy)₃; Ir–C = 1.992–2.009 Å for Ir(dfppy)₂(pic); Ir–C = 2.009–2.013 Å for Ir(dfppy)₂(fppz); Ir–N = 2.116–2.136 Å for Ir(dfppy)₂(pic); Ir–N = 2.044–2.048 Å for Ir(dfppy)₂(fppz)).¹¹

Density functional theory (DFT) calculations were performed in order to estimate the energy levels and electron density distributions of the orbitals of **Ir1** and **Ir2** together with two reported reference complexes **RIr1** and **RIr2**,^{11c,12b} which have two difluorophenylpyridinato-N,C² main ligands (dfppy and Medfppy) instead of dfppy and Medfppy, respectively (see Fig. S3 in ESI[†]). Fig. 1b and c show the optimized structures and a schematic representation of the energy levels of the four complexes. The HOMOs of **RIr1** and **RIr2** are mostly localized on both dfppy ligands together with the iridium atom, whereas the HOMOs of **Ir1** and **Ir2** are distributed over one of the two dfppy ligands and the metal d orbitals, also with a significant contribution from the ancillary ligand orbital. This clear difference results in much lower HOMO levels of **Ir1** and **Ir2** (–6.29 and –6.19 eV, respectively) compared to **RIr1** and **RIr2** (–5.73 and –5.63 eV, respectively). On the other hand, the LUMOs are distributed similarly in all four complexes over part of one of the main ligands as well as the ancillary ligand. Thus, although **Ir1** and **Ir2** show lower LUMO levels of –2.10 and –1.99 eV compared to **RIr1** and **RIr2** (–1.79 and –1.72 eV) respectively, the difference of the LUMO energy between **Ir1** and **Ir2** and the corresponding reference complexes is lower than for the HOMO. These data lead to the larger

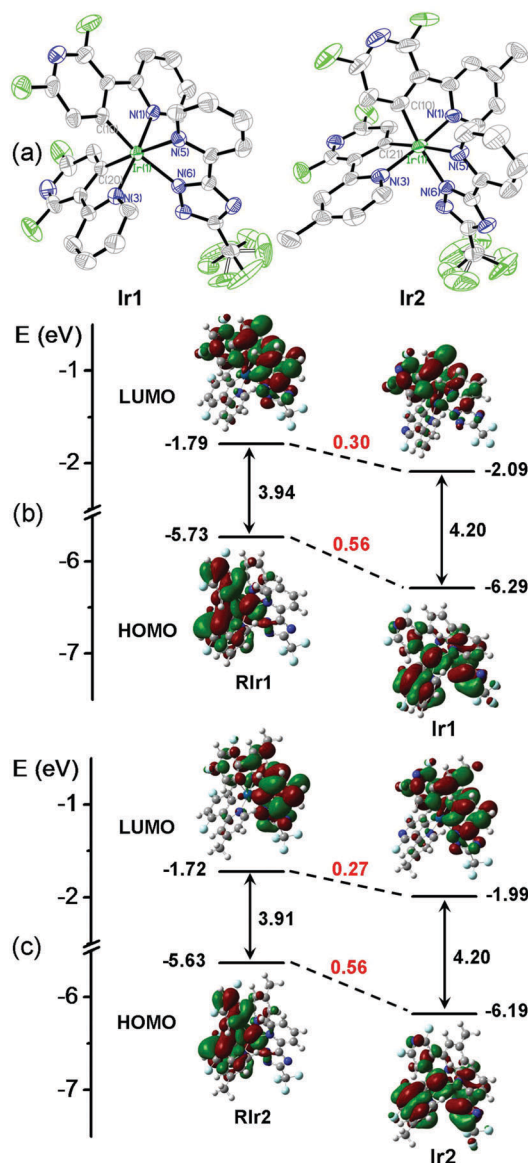


Fig. 1 (a) ORTEP plots of the X-ray molecular structures of **Ir1** and **Ir2**. Hydrogen atoms are omitted for clarity. (b and c) Molecular orbital diagrams for **Ir1**/**Ir2** and two reference complexes (**RIr1**/**RIr2**) with selected isodensity frontier molecular orbitals mainly involved in the electronic transitions. All the DFT energy values are given in electron volts (eV). The structures of **RIr1** and **RIr2** are shown in Fig. S3 (ESI[†]).

HOMO–LUMO energy gaps of ~4.2 eV for **Ir1** and **Ir2**, compared to ~3.9 eV for **RIr1** and **RIr2**. These data confirm that the bipyridyl cyclometalating ligand of **Ir1** and **Ir2** is beneficial for achieving the desired blue shift in the phosphorescent emission compared to the phenylpyridyl analogs. The improvement in the purity of blue emission from **Ir1** and **Ir2** is experimentally confirmed by the photophysical and electrochemical studies described below.

Photophysical and electrochemical properties

The UV-vis absorption and phosphorescence spectra of **Ir1** and **Ir2** are shown in Fig. 2a, measured in CH₂Cl₂ solution at room

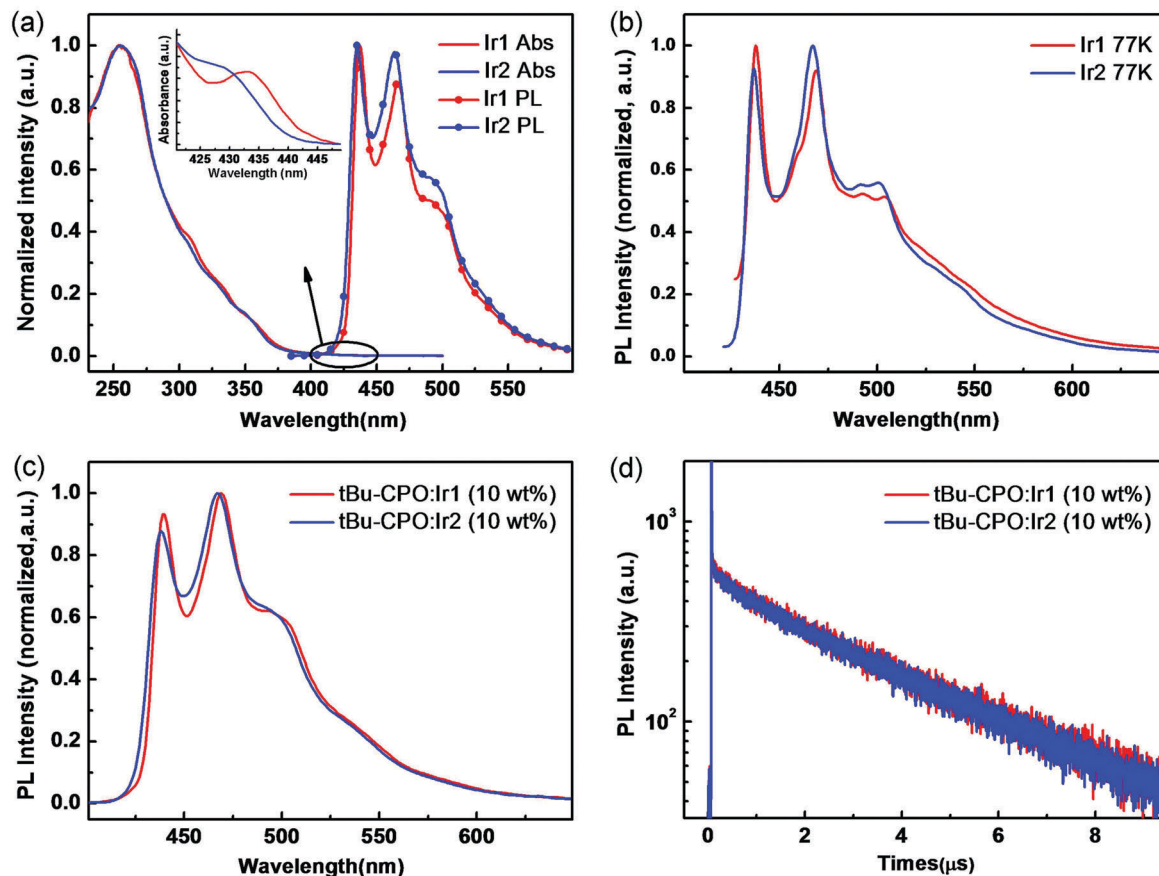


Fig. 2 UV-vis absorption and PL spectra in CH_2Cl_2 at room temperature (a) and PL spectra in THF at low temperature (77 K) (b) of **Ir1** and **Ir2**. PL spectra (c) and luminescence decay profiles (d) of tBu-CPO films doped with 10 wt% **Ir1** and **Ir2**. Emission data were obtained in degassed solutions. The structure of tBu-CPO is shown in Fig. S3 (ESI†).

temperature (RT). Both complexes display strong absorption bands at approximately 253 and 255 nm assigned to intraligand $\pi-\pi^*$ transitions.¹⁴ The lowest-energy absorption band at λ_{max} 330–355 nm can be attributed to the spin-allowed singlet metal-to-ligand charge-transfer (¹MLCT) transitions. Moreover, the spin-forbidden triplet ³MLCT or ³LC transitions (or both) appear as a weak long-tail absorption at the range of 410–430 nm (Fig. 2a inset), implying that the excitation energy for the triplet state (T_1) of both complexes is ~ 2.9 eV.¹⁴

Both **Ir1** and **Ir2** show similar intense blue phosphorescence in solution (RT) with emission at λ_{max} 437, 466 nm (RT)/438, 468 nm (77 K) for **Ir1** and 435, 464 nm (RT)/437, 467 nm (77 K) for **Ir2** (Fig. 2b), which are strongly blue shifted compared to **RIr1** and **RIr2** (~ 460 , 490 nm)^{11c,12b} in agreement with the DFT calculations above. The photoluminescence (PL) emission and the transient PL properties of thin films using t-BuCPO¹⁵ as host doped with complexes of **Ir1** and **Ir2** (10 wt%) were measured. Both films displayed strong and structured phosphorescence spectra with similar emission maxima in the deep blue 430–470 nm region (Fig. 2c) to those in solution. The PL quantum yields (PLQYs) were 0.65 ± 0.03 and 0.70 ± 0.03 , by employing an integrating sphere. The excited-state lifetimes of **Ir1** and **Ir2** are monoexponential and are 2.97 μs and 3.01 μs , respectively, consistent with their emission originating from a

triplet excited state. The PL spectra and transient PL of the doped films are shown in Fig. 2c and d.

The solution electrochemical properties of **Ir1** and **Ir2** were investigated by cyclic voltammetry (CV), and both complexes exhibit reversible reduction and oxidation behaviour as shown in Fig. S6 (ESI†). The respective HOMO and LUMO energy levels were experimentally determined from the redox curves relative to the vacuum level, and are -6.27 and -2.57 eV for **Ir1**, -6.24 and -2.56 eV for **Ir2**. These data show that **Ir1** and **Ir2** have larger HOMO–LUMO gaps than **RIr1** and **RIr2** with the HOMO/LUMO values of $\sim -5.6/-2.9$ eV.^{11c} Therefore, the electrochemical, photophysical and computational data are entirely consistent and demonstrate that the novel combination of main and ancillary ligands in **Ir1** and **Ir2**, where cyclometalating C[^]N bipyridine ligands are employed instead of phenylpyridine ligands, leads to a distinctively different HOMO/LUMO distribution compared to previous blue-emitting Ir(III) complexes^{10,11} and explains the observed deep blue emission.

Characterization of deep-blue phosphorescent OLEDs

To explore the electroluminescence (EL) characteristics of **Ir1** and **Ir2**, OLEDs based on a wide energy-gap host molecule t-BuCPO¹⁵ were fabricated by vacuum evaporation processes. The PhOLEDs have the following configuration: ITO/TAPC (30 nm)/mCP

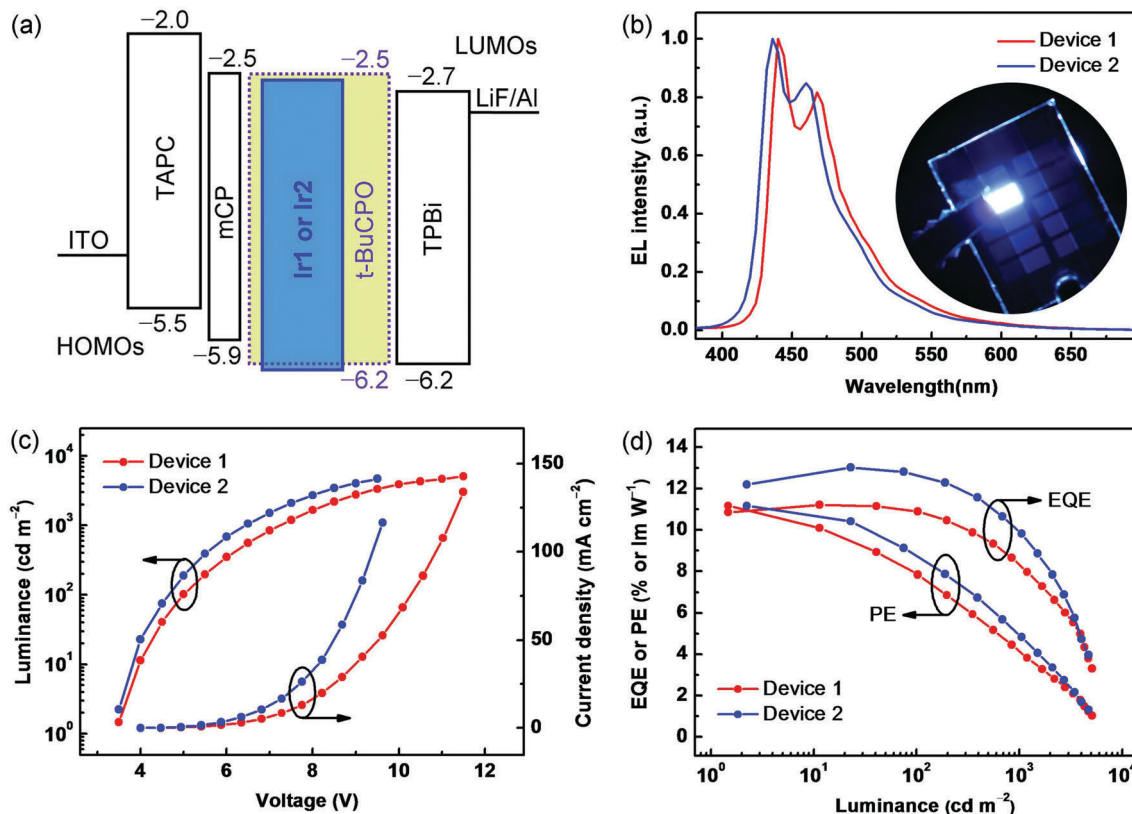


Fig. 3 (a) Energy diagram of the materials used in the OLEDs. (b) EL spectra of Devices 1 and 2. Inset: A photograph of Device 2 illuminating at 1000 cd m⁻². (c) Current density–voltage–luminance (*J*–*V*–*L*) curves of Devices 1 and 2. (d) Power efficiency (PE)–luminance (*L*)–external quantum efficiency (EQE) curves of Devices 1 and 2.

(10 nm)/EML (30 nm)/TPBi (30 nm)/LiF (1 nm)/Al (200 nm). In each device, TAPC (1,1-bis{4-[*N,N*-di(*p*-tolyl)amino]phenyl}-cyclohexane) served as a hole-transporting layer, mCP (1,3-bis-(carbazol-9-yl)benzene) as an electron-blocking layer (EBL) to prevent the formation of an exciplex between TAPC and the EMLs, TPBi (1,3,5-tris(*N*-phenylbenzimidazol-2-yl)benzene) as an electron-transporting layer (ETL), with ITO (indium-tin oxide) anode and LiF/Al cathode. The devices using doped *t*-BuCPO: Ir1 or Ir2 (10 wt%) films as the EMLs, are denoted as Device 1 and Device 2, respectively. As shown in Fig. 3a, there are slight barriers (0.2–0.3 eV) against hole injection from mCP into the EMLs based on *t*-BuCPO, and against electron injection from TPBi into the EMLs. Furthermore, *t*-BuCPO was chosen as the host due to its sufficiently high *T*₁ of ~3.2 eV to sensitize the emission of Ir1 or Ir2, together with its rather deep HOMO energy level and its favorable bipolar donor–acceptor structure.¹⁵ As a result, both holes and electrons can conduct smoothly and are confined effectively within the EMLs, ensuring that highly efficient exciton-formation/-transfer can occur even at high current densities.

The devices showed stable EL spectra (Fig. 3b) within a range of driving voltages (4–8 V) (Fig. S8, ESI†), and no host emission is observed in the EL spectra, which also exhibit good color durability after continuous open-condition operation for 2 hours. Deep-blue EL with two $\lambda_{\text{max}}^{\text{EL}}$ bands at 430–440 nm and 460–470 nm at a current density of 5 mA cm⁻² (~500 cd m⁻²)

resemble the PL spectra of the complexes. Owing to the blue-region emitting peaks of both devices, together with their full width at half maximum (FWHM) of less than 50 nm, the similar CIE coordinates of (0.15, 0.13) and (0.14, 0.11) were obtained. Fig. 3b shows a photograph of Device 2 demonstrating the observed color purity. It is interesting to note that the relative intensities at the shortest wavelength bands of Ir1 and Ir2 are higher in the EL than in the PL spectra due to optical micro-cavity effects,¹⁶ leading to the desired deep-blue EL. The current density–voltage–luminance (*J*–*V*–*L*) and EL efficiency–luminance (PE/EQE–*L*) characteristics are shown in Fig. 3c and d, respectively, and the EL data are summarized in Table 1. Devices 1 and 2 exhibit rather low turn-on voltages of ~3.5 V, indicating also that these HTL–EML–ETL systems possess the predominant factor facilitating both hole and electron injection and transport,^{8,17} which also dominated the rapidly increasing *J*–*V* and *L*–*V* curves after the onset, where the driving voltages are as low as 5.0/4.7 V and 7.3/6.4 V at the practical luminances of 100 and 1000 cd m⁻², respectively. Furthermore, the peak EQEs/PEs of 11.2%/11.1 lm W⁻¹ and 13.0%/11.2 lm W⁻¹ with remarkably low roll-off were obtained in Devices 1 and 2, respectively. Device 2 shows the bluest color and the higher EQE/PE in terms not only of the peak values above, but also by maintaining such high levels as 12.6%/8.8 lm W⁻¹ and 10.1%/5.0 lm W⁻¹ at the luminances of 100 and 1000 cd m⁻², respectively, which are needed for practical displays and/or

Table 1 Summary of the EL performance for Devices 1 and 2, using complexes **Ir1** and **Ir2**, respectively

	V_{on}^a [V]	L_{max} [cd m ⁻²]	η_{ext}^b [%]	η_p^b [lm W ⁻¹]	CIE(x,y) ^c
1	3.5	5080	11.2, 10.8, 8.3	11.1, 7.9, 4.2	0.15, 0.13
2	3.5	4710	13.0, 12.6, 10.1	11.2, 8.8, 5.0	0.14, 0.11

^a Applied voltage required to reach a luminance of 1 cd m⁻². ^b The efficiencies listed are the maximum values, and the values at 100 and 1000 cd m⁻², respectively. ^c Recorded at 1000 cd m⁻².

solid-state lighting. Here, the slight improved EL performance of **Ir2** compared to **Ir1** indicates that the methyl substituent could partly suppress the aggregation of the emitting core in the solid thin films, and thereby favour enhanced device EL efficiency.¹⁸ To the best of our knowledge, these EL efficiencies strongly compete with, and even exceed, any previously reported values for deep-blue PhOLEDs with similar CIE_{x,y} values.^{5–8}

For comparison, we note that deep-blue OLEDs using a thermally-activated delayed fluorescence (TADF) organic (metal-free) emitter achieve 20% EQE at 100 cd m⁻² with roll-off to ca. 12% at 1000 cd m⁻².¹⁹ However, the TADF-OLEDs' CIE_{x,y} coordinates (0.15, 0.13) are less blue than our Device 2, and the complicated TADF device architecture with seven separate organic layers could limit practical applications.

Conclusions

In summary, a straightforward synthetic route is reported for the high-yielding synthesis of two new heteroleptic phosphorescent complexes **Ir1** and **Ir2** consisting of cyclometalating ligands based on difluorobipyridyl derivatives, in combination with a chelated pyridyl-triazole ancillary ligand. This rational materials design has successfully achieved a very wide HOMO–LUMO gap. As a result, both phosphors show deep-blue emission with λ_{max} values in the 435–465 nm range, high PLQYs and short excited-state lifetimes. In particular, the PhOLEDs based on **Ir1** and **Ir2** realized high performance deep-blue EL. The deepest blue color (CIE_{x,y} 0.14, 0.11) and the highest EL efficiency were achieved in the device based on **Ir2** (Device 2), where the peak EQE/PE (13.0%/11.2 lm W⁻¹) together with the corresponding values at the practical luminances of 100 cd m⁻² (12.6%/8.8 lm W⁻¹) and 1000 cd m⁻² (10.1%/5.0 lm W⁻¹) strongly compete with any previously reported deep-blue fluorescent and/or phosphorescent OLEDs with similar CIE coordinates. This work represents a significant advance in the design and characterization of new transition metal complexes which display outstanding efficiency in deep-blue OLEDs.

Acknowledgements

This work was supported by National Basic Research Program of China (973 Program, 2013CB834805), Natural Science Foundation of China (91333201, 21221063, 51373062, 51473028) and the key scientific and technological project of Jilin province (20150204011GX). The work at Durham University was supported

by EPSRC (Engineering and Physical Sciences Research Council) grant EP/K039423/1.

Notes and references

- (a) M. A. Baldo, D. F. O'Brien, Y. You, A. Shoustikov, S. Sibley, M. E. Thompson and S. R. Forrest, *Nature*, 1998, **395**, 151–154; (b) M. A. Baldo, S. Lamansky, P. E. Burrows, M. E. Thompson and S. R. Forrest, *Appl. Phys. Lett.*, 1999, **75**, 4–6; (c) M. Ikai, S. Tokito, Y. Sakamoto, T. Suzuki and Y. Taga, *Appl. Phys. Lett.*, 2001, **79**, 156–158.
- (a) W. Soichi, I. Nobuhiro and K. Junji, *Jpn. J. Appl. Phys.*, 2007, **46**, 1186–1188; (b) S.-J. Su, T. Chiba, T. Takeda and J. Kido, *Adv. Mater.*, 2008, **20**, 2125–2130; (c) Y. Tao, Q. Wang, C. Yang, C. Zhong, J. Qin and D. Ma, *Adv. Funct. Mater.*, 2010, **20**, 2923–2929; (d) S. Lee, K.-H. Kim, D. Limbach, Y.-S. Park and J.-J. Kim, *Adv. Funct. Mater.*, 2013, **23**, 4105–4110; (e) G. Li, D. Zhu, T. Peng, Y. Liu, Y. Wang and M. R. Bryce, *Adv. Funct. Mater.*, 2014, **24**, 7420–7426; (f) Y. Feng, P. Li, X. Zhuang, K. Ye, T. Peng, Y. Liu and Y. Wang, *Chem. Commun.*, 2015, **51**, 12544–12547.
- (a) H. Shin, S. Lee, K.-H. Kim, C.-K. Moon, S.-J. Yoo, J.-H. Lee and J.-J. Kim, *Adv. Mater.*, 2014, **26**, 4730–4734; (b) K. Udagawa, H. Sasabe, C. Cai and J. Kido, *Adv. Mater.*, 2014, **26**, 5062–5066; (c) J.-H. Lee, S.-H. Cheng, S.-J. Yoo, H. Shin, J.-H. Chang, C.-I. Wu, K.-T. Wong and J.-J. Kim, *Adv. Funct. Mater.*, 2015, **25**, 361–366.
- (a) S. O. Jeon, S. E. Jang, H. S. Son and J. Y. Lee, *Adv. Mater.*, 2011, **23**, 1436–1441; (b) M. S. Park, D. H. Choi, B. S. Lee and J. Y. Lee, *J. Mater. Chem.*, 2012, **22**, 3099–3104; (c) J.-B. Kim, S.-H. Han, K. Yang, S.-K. Kwon, J.-J. Kim and Y.-H. Kim, *Chem. Commun.*, 2015, **51**, 58–61.
- (a) C.-H. Hsieh, F.-I. Wu, C.-H. Fan, M.-J. Huang, K.-Y. Lu, P.-Y. Chou, Y.-H. O. Yang, S.-H. Wu, I. C. Chen, S.-H. Chou, K.-T. Wong and C.-H. Cheng, *Chem. – Eur. J.*, 2011, **17**, 9180–9187; (b) S. Lee, S.-O. Kim, H. Shin, H.-J. Yun, K. Yang, S.-K. Kwon, J.-J. Kim and Y.-H. Kim, *J. Am. Chem. Soc.*, 2013, **135**, 14321–14328.
- (a) J. N. Moorthy, P. Venkatakrishnan, D.-F. Huang and T. J. Chow, *Chem. Commun.*, 2008, 2146–2148; (b) T. Peng, K. Ye, Y. Liu, L. Wang, Y. Wu and Y. Wang, *Org. Electron.*, 2011, **12**, 1914–1919; (c) W. Li, D. Liu, F. Shen, D. Ma, Z. Wang, T. Feng, Y. Xu, B. Yang and Y. Ma, *Adv. Funct. Mater.*, 2012, **22**, 2797–2803; (d) R. Kim, S. Lee, K.-H. Kim, Y.-J. Lee, S.-K. Kwon, J.-J. Kim and Y.-H. Kim, *Chem. Commun.*, 2013, **49**, 4664–4666.
- Y.-C. Chiu, J.-Y. Hung, Y. Chi, C.-C. Chen, C.-H. Chang, C.-C. Wu, Y.-M. Cheng, Y.-C. Yu, G.-H. Lee and P.-T. Chou, *Adv. Mater.*, 2009, **21**, 2221–2225.
- J. Lee, H.-F. Chen, T. Batagoda, C. Coburn, P. I. Djurovich, M. E. Thompson and S. R. Forrest, *Nat. Mater.*, 2016, **15**, 92–98.
- (a) S. Tokito, T. Iijima, Y. Suzuri, H. Kita, T. Tsuzuki and F. Sato, *Appl. Phys. Lett.*, 2003, **83**, 569–571; (b) R. J. Holmes, S. R. Forrest, Y.-J. Tung, R. C. Kwong, J. J. Brown, S. Garon

- and M. E. Thompson, *Appl. Phys. Lett.*, 2003, **82**, 2422–2424; (c) L. Xiao, S.-J. Su, Y. Agata, H. Lan and J. Kido, *Adv. Mater.*, 2009, **21**, 1271–1274; (d) E. Baranoff and B. F. E. Curchod, *Dalton Trans.*, 2015, **44**, 8318–8329.
- 10 (a) S. Takizawa, H. Echizen, J. Nishida, T. Tsuzuki, S. Tokito and Y. Yamashita, *Chem. Lett.*, 2006, **35**, 748–749; (b) S. O. Jeon, S. E. Jang, H. S. Son and J. Y. Lee, *Adv. Mater.*, 2011, **23**, 1436–1441; (c) K. S. Yook and J. Y. Lee, *Org. Electron.*, 2011, **12**, 1711–1715.
- 11 (a) S. J. Lee, K.-M. Park, K. Yang and Y. Kang, *Inorg. Chem.*, 2009, **48**, 1030–1037; (b) C.-H. Yang, M. Mauro, F. Polo, S. Watanabe, I. Muenster, R. Fröhlich and L. De Cola, *Chem. Mater.*, 2012, **24**, 3684–3695; (c) H. J. Park, J. N. Kim, H.-J. Yoo, K.-R. Wee, S. O. Kang, D. W. Cho and U. C. Yoon, *J. Org. Chem.*, 2013, **78**, 8054–8064; (d) Y. Kang, Y.-L. Chang, J.-S. Lu, S.-B. Ko, Y. Rao, M. Varlan, Z.-H. Lu and S. Wang, *J. Mater. Chem. C*, 2013, **1**, 441–450; (e) F. Kessler, Y. Watanabe, H. Sasabe, H. Katagiri, M. K. Nazeeruddin, M. Grätzel and J. Kido, *J. Mater. Chem. C*, 2013, **1**, 1070–1075.
- 12 (a) R. J. Holmes, B. W. D'Andrade, S. R. Forrest, X. Ren, J. Li and M. E. Thompson, *Appl. Phys. Lett.*, 2003, **83**, 3818–3820; (b) S. J. Yeh, M. F. Wu, C. T. Chen, Y. H. Song, Y. Chi, M. H. Ho, S. F. Hsu and C. H. Chen, *Adv. Mater.*, 2005, **17**, 285–289; (c) L. Xiao, Z. Chen, B. Qu, J. Luo, S. Kong, Q. Gong and J. Kido, *Adv. Mater.*, 2011, **23**, 926–952; (d) K. S. Yook and J. Y. Lee, *Adv. Mater.*, 2012, **24**, 3169–3190.
- 13 (a) H. Oh, K.-M. Park, H. Hwang, S. Oh, J. H. Lee, J.-S. Lu, S. Wang and Y. Kang, *Organometallics*, 2013, **32**, 6427–6436; (b) H. Xu, Y. Yue, L. Qu, Y. Hao, H. Wang, L. Chen and B. Xu, *Dyes Pigm.*, 2013, **99**, 67–73; (c) X. Zhang, J. K. Rueter, Y. Chen, M. Moorjani, M. C. Lanier, E. Lin, R. S. Gross, J. E. Tellev, J. P. Williams, S. M. Lechner, S. Markison, T. Joswig, S. Malany, M. Santos, J. C. Castro-Palomino, M. I. Crespo, M. Prat, S. Gual, J.-L. Diaz, J. Saunders and D. H. Slee, *Bioorg. Med. Chem. Lett.*, 2008, **18**, 1778–1783.
- 14 (a) S. Lamansky, P. Djurovich, D. Murphy, F. Abdel-Razzaq, H.-E. Lee, C. Adachi, P. E. Burrows, S. R. Forrest and M. E. Thompson, *J. Am. Chem. Soc.*, 2001, **123**, 4304–4312; (b) P. J. Hay, *J. Phys. Chem. A*, 2002, **106**, 1634–1641.
- 15 A. Wada, T. Yasuda, Q. Zhang, Y. S. Yang, I. Takasu, S. Enomoto and C. Adachi, *J. Mater. Chem. C*, 2013, **1**, 2404–2407.
- 16 (a) V. Bulović, V. B. Khalfin, G. Gu, P. E. Burrows, D. Z. Garbuzov and S. R. Forrest, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1998, **58**, 3730–3740; (b) Y. Zhang, J. Lee and S. R. Forrest, *Nat. Commun.*, 2014, **5**, 5008.
- 17 (a) F.-M. Hsu, C.-H. Chien, C.-F. Shu, C.-H. Lai, C.-C. Hsieh, K.-W. Wang and P.-T. Chou, *Adv. Funct. Mater.*, 2009, **19**, 2834–2843; (b) H. Sasabe, Y. Seino, M. Kimura and J. Kido, *Chem. Mater.*, 2012, **24**, 1404–1406.
- 18 K. Q. Ye, J. Wang, H. Sun, Y. Liu, Z. C. Mu, F. Li, S. M. Jiang, J. Y. Zhang, H. X. Zhang, Y. Wang and C.-M. Che, *J. Phys. Chem. B*, 2005, **109**, 8008–8016.
- 19 I. Lee and J. Y. Lee, *Org. Electron.*, 2016, **29**, 160–164.