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Complexation Behaviour of LiCl and LiPF₆ – Model Studies in the Solid-State and in Solution Using a Bidentate Picolyl-Based Ligand

Received 00th January 20xx, Accepted 00th January 20xx Noel Angel Espinosa-Jalapa,^a Nele Berg,^b Michael Seidl,^a Ilya G. Shenderovich,^b Ruth M. Gschwind^b and Jonathan O. Bauer^{*a}

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Structural knowledge on ubiquitous lithium salts in solution and in the crystalline state is of paramount importance for our understanding of many chemical reactions and of the electrolyte behaviour in lithium ion batteries. A bulky bidentate Si-based ligand (6) was used to create simplified model systems suitable for correlating structures of LiCl and LiPF₆ complexes in the solid-state and in solution by combining various experimental, spectroscopic, and computational methods. Solution studies were performed using ¹H DOSY, multinuclear variable temperature NMR spectroscopy, and quantum chemical calculations. $[Ph_2Si(2-CH_2Py)_2 \bullet LiCl]_2$ (3) dissociates into a monomeric species (9) in THF. For $[Ph_2Si(2-CH_2Py)_2 \bullet LiPF_6]_2$ (11), low temperature NMR studies revealed an unprecedented chiral coordination mode (12) in non-coordinating solvents.

Structural knowledge on synthetically and technically relevant lithium salts like LiX (X = halogenide) and LiPF₆ in the solid-state and in solution in the presence of coordinating solvents or ligands is particularly important for a better understanding of salt effects both in chemical reactions and technical applications such as in lithium ion batteries.¹⁻⁴ LiCl is very often a by-product in chemical reactions with lithiated species involved and has long been known to be a potential key to reactivity and selectivity in reactions mediated by organometallic reagents.⁵ Combinations of inorganic salts or donor ligands with certain organometallic reagents (e.g. organolithium, -magnesium, and -zinc compounds) have gained broad applications in preparative chemistry.^{6,7} Significant progress has recently also been made in the field of LiCl catalysis.8 Understanding cation-anion coordination modes between Li⁺ and PF₆⁻ in LiPF₆ electrolytes is of great importance and has contributed significantly to the improvement of lithium ion batteries.² The development of defined [ligand • metal salt] combinations can open up entirely new paths in synthetic chemistry and provide hitherto unknown information on cation-anion interactions. However, the structural behaviour in solution often remains a black-box and

to a limited extent.⁹ Structural investigations on lithium compounds in solution using diffusion-ordered NMR spectroscopic methods have therefore become an increasingly important field, but have been reported mainly for lithium amide, mixed-metal, and organolithium reagents.¹⁰ Simplified model systems that use bulky and specially designed ligands have proven to be helpful tools for obtaining in-depth information on structure-forming principles of fundamentally important metal compounds, which are either used in preparative chemistry or found in nature.¹¹ Especially in solution, the structural space often shows great versatility and flexibility and can be efficiently restricted using bulky ligands with diagnostic molecular patterns in order to achieve information on coordination modes that are otherwise difficult to identify.¹²

correlations between solid-state and solution are only possible

This gave us the idea to introduce diphenylbis(2-picolyl)silane [Ph₂Si(2-CH₂Py)₂] (**6**) (Scheme 1) because we expected this to be a promising backbone for the design of new bulky bidentate ligands that would enable detailed model studies of widely used synthetically and technically important lithium salts by showing defined structural patterns in solution. Different to phosphorus-based bis(2-picolyl) NPN ligands,¹³ tetravalent bis(2-picolyl)-substituted group 14 elements do not have an additional free donor site capable of coordination, but sterically demanding substituents can easily be incorporated.¹⁴ This



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Scheme 1. Synthesis of $[Ph_2Si(2-CH_2Py)_2 \bullet LiCl]_2$ (3) (route A), the free ligand 6 via the borane-protected intermediates 4 and 5 (route B), and the lutidine derivative 8 (route C).

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makes them proper target molecules for use as bulky bidentate ligand systems that can limit structural flexibility in coordination modes. With this ligand in hands, we shed light on the structural behaviour of an LiCl and an LiPF₆ complex both in the solid-state and in solution by combining various NMR spectroscopic methods with quantum chemical calculations.

We first set out to investigate a straightforward synthetic route to the new bidentate ligand (6) (Scheme 1). 2-Methylpyridines (2-picolines) that are functionalized by main group elements at the methyl position have long been the subject of intensive research due to their interesting electronic properties and coordinating abilities.¹⁵ However, only little is known about the coordination abilities of multidentate, unsubstituted (2picolyl)silanes.¹⁶ Direct reaction of dichlorodiphenylsilane (1) with two equiv. of 2-picolyllithium¹⁷ (2) resulted in the formation of a dimeric LiCl complex (3) coordinated by the new bidentate ligand diphenylbis(2-picolyl)silane (6). Compound 3 could easily be isolated by extraction of the formed solids with DCM in 68% yield (Scheme 1, A) and was characterized by singlecrystal X-ray diffraction analysis. The LiCl complex 3 crystallized from acetonitrile/THF as colourless blocks in the triclinic crystal system, space group P1 (Figure 1, top). The LiCl core forms a dimeric rhombic structural unit in which each lithium atom is additionally coordinated by the two pyridyl nitrogen atoms of one ligand each (average N–Li bond length: 2.073(5) Å). An ideal tetrahedral geometry is retained around the silicon atom so that no significant structural changes of the bidentate ligand are required for an efficient coordination.

Interestingly, $Ph_2Si(2-CH_2Py)_2$ (6) is quite similar to the phosphorus analog¹⁸ PhP(2-CH_2Py)_2 concerning its coordination behaviour to LiCl.^{18c} In the molecular structure of the latter, only the two pyridyl nitrogen atoms are involved in the coordination to LiCl forming a similar dimer as reported herein.^{18c}

This prompted us to investigate the structure of $[Ph_2Si(2-CH_2Py)_2 \bullet LiCl]_2$ (3) in more detail in organic solvents. For this purpose, we developed a two-step synthesis to obtain $Ph_2Si(2-CH_2Py)_2(6)$ in a LiCl-free fashion using BH_3 as protecting agent (Scheme 1, B). In addition, we made the lutidine derivative 8 available to study the influence of a methyl group in the 6-position of the pyridyl ring on the coordination behaviour (Scheme1, C). In contrast to the picoline derivative 6, the bis(2,6-lutidinyl)silane 8 could easily be obtained without complexation of LiCl when carrying out the direct reaction of two equiv. 2-



Figure 1. Molecular structures of complexes 3 (top) and 11 (bottom) in the crystal (displacement ellipsoids set at the 50 % probability level, hydrogen atoms and included solvent molecules are omitted for clarity).

picolyllithium with dichlorosilane **1** (Scheme 1, **C**)_{iev}Bothe come pounds **6** and **8** were obtained in crystalline to the solution of the single-crystal X-ray diffraction analysis (for details, see the SI).

Solid-state ⁷Li NMR spectroscopy of dimer **3** shows a signal at δ = 2.6 ppm. In DCM, the ⁷Li NMR chemical shift of complex **3** (δ = 2.5 ppm) fits perfectly with the results of the solid-state NMR spectroscopy, which led us to the conclusion that the dimeric structure is retained in this weakly coordinating solvent. However, in a THF solution of 3, the ⁷Li NMR signal is highfieldshifted to δ = 0.8 ppm while the ¹H NMR chemical shift of the pyridinyl-H atom in 6-position ($\delta = 8.50$ ppm) still shows an interaction of the ligand with the lithium salt when compared with the ¹H NMR spectrum of the free ligand Ph₂Si(2-CH₂Py)₂ (6) (δ = 8.34 ppm). LiCl, dissolved in THF without any ligand, shows a ⁷Li NMR signal at δ = 0.5 ppm. We therefore assumed that in THF, dimer 3 breaks up into a monomeric species (9) as a result of the good coordination properties of the solvent (Scheme 2). This was further supported by quantum chemical calculations at the B3LYP/6-31+G(d) level of theory using the Polarizable Continuum Model (PCM) (solvent: THF),19 which showed that this deaggregation is indeed a slightly exothermic process in THF ($\Delta H = -5$ kJ mol⁻¹), whereas the complete dissociation into the free ligand 6 and a plausible LiCl • THF species^{1b,20} (10) is a thermodynamically unfavorable process ($\Delta H = +10$ kJ mol⁻¹ with respect to monomer 9). Compound 10 was previously identified as a crucial species in equilibrium processes of organometallic reagents in the presence of LiCl in THF solutions.²⁰ The same calculations for the lutidine derivative show a highly exothermic process ($\Delta H = -63$ kJ mol⁻¹ in total) for the dissociation of a hypothetical dimer into the free ligand 8 and [LiCl • 2 THF]₂ (10) (for details, see the SI), so that even the formation of a monomeric LiCl species can be excluded if there is a methyl group in the 6-position of the pyridyl ring. ⁷Li NMR spectroscopy of 8 with one equiv. LiCl in THF shows one signal at δ = 0.5 ppm in accordance with free LiCl in THF. DFT calculations of the ⁷Li NMR chemical shift at the B3LYP/6-311+G(2d,p)//B3LYP/6-31+G(d) level²¹ strengthen the dissociation of dimer 3 into monomer 9 in THF. For the monomeric LiCl complex **9**, a ⁷Li NMR chemical shift of δ = 1.3 ppm was calculated, which matches well with the experimentally observed ⁷Li NMR signal (δ = 0.8 ppm) of the LiCl adduct in THF. The calculated ⁷Li NMR chemical shift of dimer 3 was clearly downfield-shifted both in THF (δ = 1.9 ppm) and in the gas phase (δ = 2.3 ppm). Although the differences in the ⁷Li NMR chemical shifts are only marginal, the trend agrees well with the proposed model.

We finally performed ¹H-diffusion-ordered NMR spectroscopic (¹H DOSY)²² measurements to get additional support for the formation of a monomeric Ph₂Si(2- CH₂Py)₂ • LiCl adduct (**9**) in THF solution. The hydrodynamic radius ($r_{\rm H}$) of ligand **6** in THF was determined to be 4.70 Å. For a solution of complex **3** in THF, the ¹H DOSY experiment gave a radius of $r_{\rm H}$ = 6.12 Å. This radius corresponds much more to a monomeric LiCl complex such as **9**, as assumed in THF (see Scheme 2), than to the intact dimer **3** (for details on the measurement, see the SI). For comparison, ¹H-diffusion measurements of the lutidine derivative **8** with and



Scheme 2. Calculated reaction enthalpies (ΔH) for the proposed dissociation of dimer 3 into a monomeric LiCl complex (9) and for the unlikely dissociation into the free ligand 6 and a dissolved LiCl • THF complex (10) in THF [B3LYP/6-31+G(d)] (PCM calculations; solvent: THF)

without addition of LiCl gave hydrodynamic radii of $r_{\rm H}$ = 4.08 Å (without LiCl) and $r_{\rm H}$ = 4.88 Å (addition of one equiv. LiCl), respectively, which indicates interactions between compound 8 and LiCl in THF, but the formation of a stable complex can be excluded.

In order to get more insight into the coordinating abilities of ligand 6 towards technically important lithium salts, we performed a ligand exchange reaction on adduct 3. The LiCl complex 3 could easily be converted into the dinuclear LiPF₆ complex 11 by anion exchange with TIPF₆ (Scheme 3). 11 crystallized as colourless plates in the monoclinic crystal system, space group $P2_1/n$. The molecular structure of **11** was determined by singlecrystal X-ray crystallography (Figure 1, bottom) and is another example for the rare bridging coordination of two lithium cations by two PF₆⁻ anions, each of which providing two fluorine atoms.23 The P-F-Li angles of the dinuclear eight-membered core range from 136.10(10)° [P(1)-F(1)-Li(1)] to 157.80(12)° [P(2)-F(8)-Li(1)] and differ greatly from an almost linear arrangement, as found in the [N,N,N',N'',N''-pentamethyldiethylenetriamine • LiPF₆]₂ complex (mean P-F-Li angle: 175.5°).²³ Solid-state ⁷Li NMR spectroscopy of **11** shows a signal at δ = 0.7 ppm. In DCM at room temperature, dimer 11 can still be detected in very small amounts [δ (⁷Li) = 0.6 ppm], but is mainly converted into a new species [δ (⁷Li) = 1.7 ppm]. In view of the high structural flexibility of $Li^{+} \cdots PF_{6}^{-}$ interactions,²⁴ we carried out preliminary and for the first time variable temperature multinuclear NMR studies on a LiPF₆ complex, which gave interesting information on the structural behaviour in solution (Figure 2). In DCM, a defined major species [δ (⁷Li) = 1.2 ppm] becomes more and more evident starting from -30 to -80 °C, while at the same time formation of a minor species [δ (⁷Li) = 3.3 ppm] occurs. Remarkably, as the temperature decreases, ${}^{2}J_{HH}$ coupling (13.2 Hz) within the CH₂ groups of the main compound becomes increasingly clear in the ¹H NMR spectrum. The diastereotopicity of the methylene hydrogen atoms at -80 °C is a stereochemical probe that gives a clear hint to the formation of a defined and unprecedented complex exhibiting a chiral coordination sphere. A suggestion of a plausible structure (12) of the LiPF₆ complex in DCM in accor-dance with the solution NMR studies is also shown in Figure 2. Two lithium cations are each chelated by two fluorine atoms of one PF6⁻ anion in a chiral complexation mode.²⁵ A flipping coor-dination of the lithium centers and a rapid exchange between the coordinating and non-coordinating counteranion can be ex-pected and matches with the ¹⁹F and ³¹P NMR spectra at ambient and low tempera-



Scheme 3. Synthesis of [Ph₂Si(2-CH₂Py)₂ • LiPF₆]₂ (11) from the LiCl complex 3 via anion exchange with TIPE

ture. The formation of such a species can open new possibilities for generating polar chiral environments in solution. Furthermore, knowledge on ionic association in the presence of coordinating additives on a molecular level can provide helpful information concerning the conductivity behavi-our of LiPF₆containing nonaqueous electrolyte solutions.²⁶

In summary, we presented a combined experimental, spectroscopic, and computational model study concerning structures of fundamentally important lithium salts in the solid-state and in solution in the presence of a new bidentate silicon-based ligand (6). We have gained detailed insights into fundamental dissociation and association processes in coordinating and noncoordinating solvents. It was shown that complex 3 breaks up into a monomeric LiCl adduct (9) in THF solution, but remains a dimer in less-coordinating DCM. The chloride ions could easily be exchanged by PF₆⁻ ions, while maintaining the dinuclear structure in the solid-state. $[Ph_2Si(2-CH_2Py)_2 \bullet LiPF_6]_2$ (11) shows a rare bridging coordination of the lithium centers in the solid-state and forms a unique chiral species in solution, which was elucidated by variable temperature NMR measurements. In our ongoing studies we will address the targeted functionalization of ligand **6** at the silicon α -position. This might lead to new prototypes of multifunctional bulky bidentate C2-symmetric ligands,²⁷ which can be used as stereochemical probes to reveal unusual coordination spheres in solution.

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Figure 2. Left: Variable temperature NMR measurement (400 MHz) of complex 11 in $\mathsf{DCM-}d_2$ showing the diastereotopic splitting of the CH_2 groups in the ${}^1\mathsf{H}$ NMR spectrum and the ⁷Li NMR signals. Right: Suggestion of a plausible structure (12) of the LiPF₆ complex in solution.

Conflicts of interest

There are no conflicts to declare.

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Using a new bulky bidentate ligand and combining various structure elucidation methods, coordination modes of [ligand • LiX] (X = Cl, PF_6) complexes both in solid-state and in solution have been revealed.