

Pd-Catalyzed Triple-Fold C(sp²)-H Activation with Enaminones and Alkenes for Pyrrole Synthesis via Hydrogen Evolution

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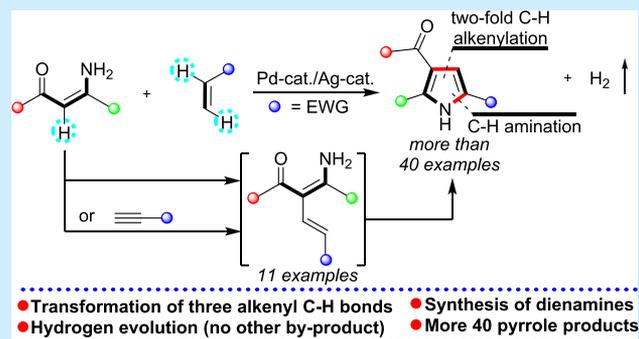


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Supporting Information

ABSTRACT: The synthesis of NH-free pyrroles via Pd-catalyzed annulation of enaminones and alkenes is reported. With the catalysis of Pd(II), the activation of triple C(sp²)-H bonds, including one internal C(sp²)-H bond in enaminone, has been activated to provide various pyrroles. The interesting evolution of hydrogen gas from the reactions has been observed by a hydrogen detector.



As one of the fundamental heteroaryl motifs showing prevalent applications in multidisciplinary areas,¹ the pyrrole ring has been an attractive target throughout the development of synthetic chemistry. The classical Paal–Knorr reaction,² Knorr reaction,³ and Hantzsch reaction,⁴ for example, have contributed enormously to the advances of pyrrole chemistry. However, following the daily increasing demand in molecular diversity-, sustainability-, and safety-oriented synthetic strategies, devising alternative and/or complementary methods for pyrrole synthesis has emerged as a significant goal in the science of synthesis.⁵ As one novel example of a newly emerged tool in pyrrole synthesis, the cascade annulation involving the key activation of one or more C–H bonds has brought an important breakthrough. For example, Stuart and co-workers reported the Rh(III)/Cu(II)-catalyzed annulation of enamides and alkynes for pyrrole synthesis via a key addition of the enamide C–H bond to the alkyne.⁶ Glorius et al. developed the pyrrole ring construction via Pd(II)-catalyzed intramolecular C–H alkenylation of imines.⁷ Ellmann and co-workers disclosed the reactions of enone-derived oxime esters and imines for pyrrole synthesis via Rh-catalyzed alkenyl C–H functionalization.⁸ In addition, reactions featured with C–H bond conversion in various other substrates have also been developed as practical protocols toward diverse pyrrole scaffolds.⁹

Among the designated methods for pyrrole synthesis via C–H activation or functionalization, however, enamines have been identified as most frequently used and reliable substrates. With the catalysis of proper transition metal reagents, stable enamines such as enaminones, enaminoesters, enaminoitriles, etc. have been used to couple alkynes,¹⁰ alkenes,¹¹ diazoesters/hydrazone,¹² or functionalization of intramolecular C–H

bonds¹³ for the synthesis of diversely functionalized pyrroles. Interestingly, although a plethora of different alkenes have been successfully used as reaction partners of stable enamines for pyrrole synthesis, no known method has hitherto been found to tolerate the transformation of acrylates for the synthesis of corresponding pyrrole products. Considering the important and broad application of acrylates in the C–H alkenylation reactions,¹⁴ it is thus highly demanding to establish new methods by activating the C–H bond of acrylates and enamines for the synthesis pyrroles with expanded product diversity.

Recently, the dehydrogenative transformations on C–H and/or heteroatom-H bonds via hydrogen evolution has emerged as a drastically sustainable tool in constructing new C–C or C–heteroatom bonds.¹⁵ Actually, the dehydrogenation of alcohols and amino alcohols has already been realized with significant success in the synthesis of pyrroles by releasing the hydrogen gas via alcohol transformation.¹⁶ However, hydrogen evolution via the activation of the stable C–H bond has not yet been reported for the construction of a pyrrole ring. In combination with the recent progress in the vinyl C–H activation and functionalization disclosed by us and others¹⁷ as well as the recent emphasis in developing novel synthetic methods via aryl ring construction using polar alkenes,¹⁸ we report herein a new method on the synthesis of pyrroles via

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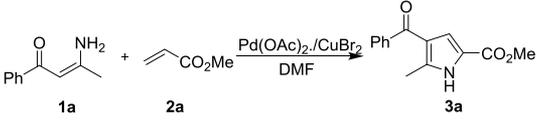
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hydrogen evolution-based coupling of three stable C–H bonds and an N–H bond by the reactions of simple enaminones and acrylates. In addition, the synthesis of conjugate dienamines has also been realized via catalytic addition of enaminones to propiolates. The efficient transformation of these dienamines into the titled pyrroles has provided strong evidence in elucidating the possible reaction mechanism.

Initially, the reaction of enaminone **1a** and methyl acrylate **2a** was conducted under different conditions for the synthesis of pyrrole product **3a**. Systematical screening on the catalyst, additive, medium, temperature, reagents loading, etc. was executed (see full data in the Supporting Information). As outlined by the typical data (Table 1), the simultaneous

Table 1. Typical Results on the Condition Optimization^a



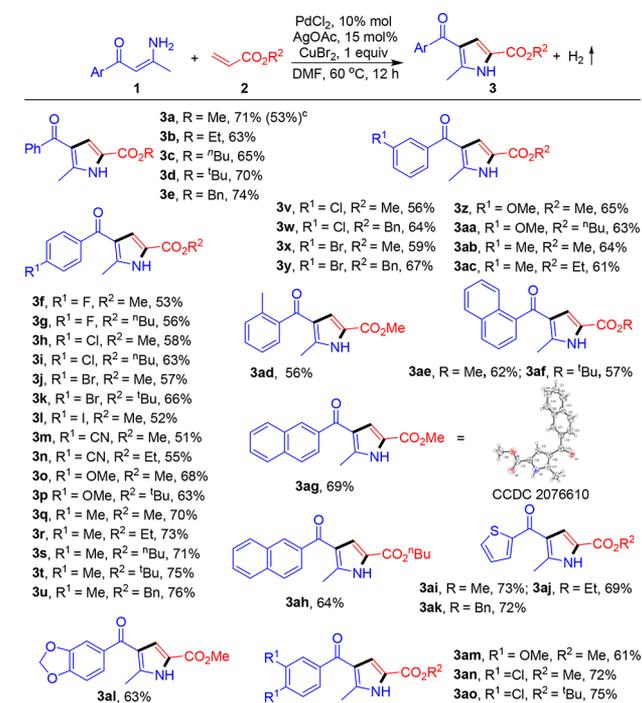
entry	variation from the general conditions	yield (%) ^b
1	no variation	43
2	Pd(OAc) ₂ (0.01 mmol) and CuBr ₂ (0.1 mmol)	24
3	Pd(OAc) ₂ (0.01 mmol), no CuBr ₂	0
4	CuBr ₂ (0.1 mmol) and no Pd(OAc) ₂	0
5	reaction at 40 °C	44
6	reaction at 60 °C	48
7	with 0.7 mmol 2a , at 60 °C	55
8 ^c	with 0.7 mmol 2a , 60 °C	63
9 ^c	with 0.7 mmol 2a , 60 °C and 0.5 mL of DMF	71

^aGeneral conditions: **1a** (0.2 mmol), **2a** (0.4 mmol), Pd(OAc)₂ (0.02 mmol), CuBr₂ (0.2 mmol) in DMF (2 mL). Stirred at 80 °C for 12 h. ^bIsolated yield. ^cWith additional AgOAc (0.03 mmol).

employment of Pd(II) and Cu(II) reagent was mandatory (entries 1–4, Table 1). In addition, conducting the reaction at a mild temperature of 60 °C was favorable to give better results (entries 5–6, Table 1). On the other hand, increasing the loading of **2a** also proved to be practical in improving the product yield (entry 7, Table 1). Notably, reducing the volume of reaction medium and employing additionally AgOAc as additive was able to further enhance the yield of **3a**, respectively (entries 7–9, Table 1).

Under the optimal conditions, the scope and limitations of this method in synthesizing pyrroles **3** were investigated.¹⁹ On one hand, when enaminone **1a** was fixed to react with different acrylic ester substrates, the products containing a methyl, ethyl, *n*-butyl, *tert*-butyl, or benzyl group (**3a–3e** and **3q–3u**, Scheme 1) were afforded in good yields. The extension of the alkene component to internal alkenyl ester, enone, vinyl nitrile, vinyl amide, styrene, and α -allylnaphthalene was also conducted, but none gave the expected pyrrole product under the present conditions (see also Scheme S2 in the Supporting Information). For the enaminones **1**, on the other hand, electron-withdrawing groups such as F, Cl, Br, I, and CN (**3f–3n** and **3v–3y**, Scheme 1) and electron-donating groups such as alkyl- and alkoxy-substituted phenyl enaminones (**3o–3u**, **3z–3ac**, Scheme 1) showed general tolerance to the synthesis. Slightly lower yields were observed in the reactions employing enaminones functionalized with electron-withdrawing groups. Moreover, enaminones featuring more than one substituent in the phenyl ring also took part in the synthesis to afford good results (**3al–3ao**, Scheme 1). As expected, the synthesis of

Scheme 1. Substrate Scope on Pyrrole Synthesis^{a,b}

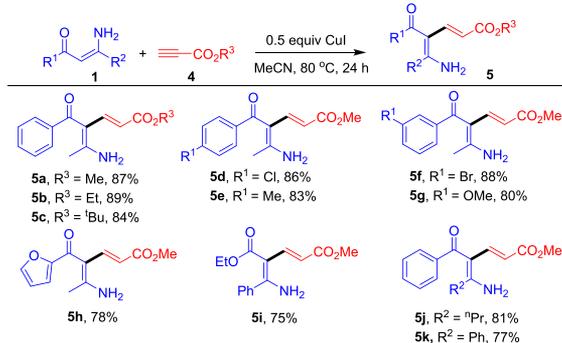


^aReaction conditions: **1** (0.2 mmol), **2** (0.7 mmol), DMF (0.5 mL), reacted under air for 12 h. ^bThe H₂ was detected by H₂ detector. ^cYield from 1 mmol scale reaction.

pyrroles bearing fused aryl (**3ag** and **3ah**, Scheme 1) and heteroaryl groups (**3ag** and **3ah**, Scheme 1) were also smoothly acquired by employing corresponding fused and heteroaryl-functionalized enaminones. Nevertheless, no expected products were observed when *N*-substituted enaminones, dimedone-derived cyclic NH₂-enaminone, and pentane-2,4-dione-derived alkyl NH₂-enaminone was independently used to react with acrylate **2a** under the standard conditions (Scheme S2 in the Supporting Information). Additionally, the reaction of aniline with **2a** was conducted with the standard catalytic conditions. However, no evident reaction was observed.

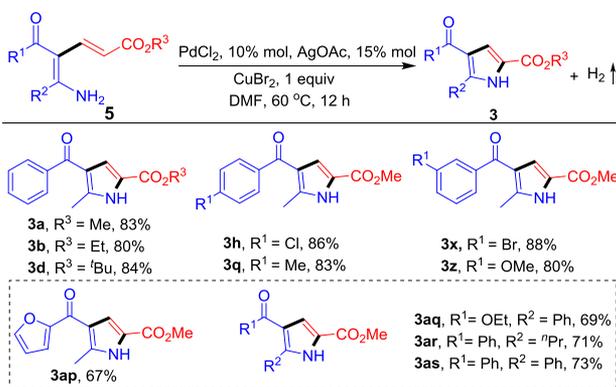
Interestingly, by means of modifying the reaction conditions, we were pleased to find that CuI could catalyze the C–H bond addition of enaminones to alkyl propiolates **4** for the selective synthesis of conjugate dienamines **5** with good to excellent yields. As depicted in Scheme 2, different alkyl acrylates (**5a–5c**, Scheme 2), as well as enaminones possessing functional groups such as halogen, methyl, methoxy, etc. in either the *meta* or *para* position (**5d–5g**, Scheme 2) provided corresponding products with satisfactory yields. Notably, this C–H addition also tolerated well the furan-functionalized enaminone and enaminoester (**5h** and **5i**, Scheme 2). Corresponding dienamines with different β -substitutions (R² = *n*-Pr or Ph) were also practically furnished (**5j** or **5k**, Scheme 2).

Having been inspired by the tunable synthesis of dienamines **5**, we thought that such compounds were the possible intermediate during the formation of the pyrrole products. Accordingly, these dienamines were found to undergo intramolecular dehydrogenative C–H amination to provide pyrroles **3** with high efficiency under the standard Pd-catalytic conditions. Besides providing products that had been

Scheme 2. Synthesis of Conjugate Dienamines^a

^aReaction conditions: **1** (0.2 mmol), **4** (0.4 mmol), MeCN (2.0 mL), stirred under air for 24 h; isolated yield was reported.

synthesized via the direct enamine–acrylate annulation (**3a**, **b**, **3d**, **3h**, **3q**, **3x**, and **3z**, Scheme 3), the intramolecular version

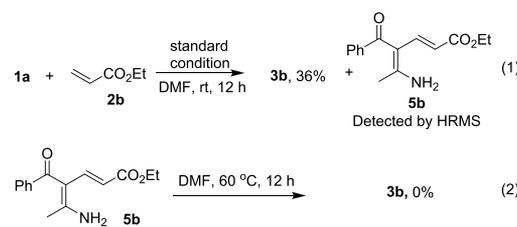
Scheme 3. Synthesis of Pyrroles from Various Dienamines^{a,b}

^aReaction conditions: **5** (0.2 mmol) in DMF (0.5 mL), stirred for 12 h under air; isolated yield was reported. ^bThe H₂ was detected by H₂ detector.

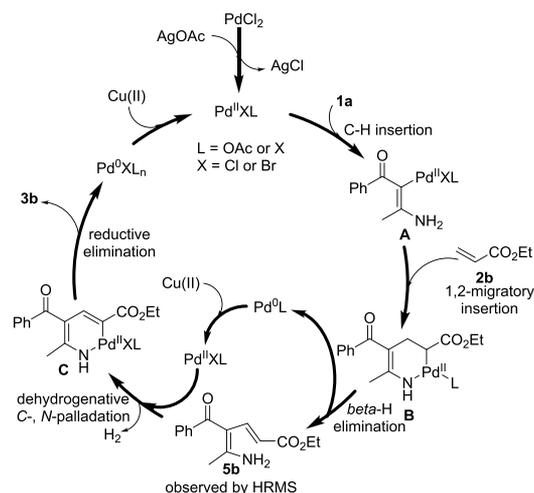
was found to be capable of providing new pyrrole products that could not be accessed by the above annulation protocol (**3ap**–**3as**, Scheme 3). Moreover, hydrogen evolution was observed from the reaction, further supporting that **5** were key reaction intermediates also in the enamine–acrylate annulation. However, directly employing enaminone **1a** and methyl propiolate (**4a**) to the standard conditions in Scheme 1 did not provide pyrrole **3a**.

For the sake of gaining more information on the reaction process, some control experiments were also designed. First, pyrrole **3b** and dienamine **5b** were obtained by performing the reaction of **1a** and **2a** at room temperature (eq 1). In addition, **3b** could not be obtained without employing the metal reagents, confirming the indispensable role of metal in the reaction (eq 2).

According to the clues provided by the control experiments, a mechanism for this NH-free pyrrole synthesis has been proposed. As outlined in Scheme 4, the reaction of AgOAc and PdCl₂ may provide active Pd^{II}XL which couples enaminone to deliver Pd(II) intermediate **A** via the C–H insertion.²⁰ The 1,2-migratory insertion of the alkene to **A** leads to the formation of palladacycle species **B**. The featured β-H elimination from **B** provides dienamine intermediate **5b**,



Scheme 4. Plausible Reaction Mechanism



which has been observed with HMRS from the reaction residue (see the Supporting Information). The dehydrogenative C₇-N-palladation of **5b** with Pd(II) gives rise to palladacycle **C**, which undergoes reductive elimination to yield the pyrrole product. The Pd(0) released thereby can be regenerated in the presence of a Cu(II) terminal oxidant.

In conclusion, we have developed a method for the synthesis of NH-free pyrroles with the dehydrogenative coupling of enaminones and alkenes via Pd-catalyzed activation of triple C(sp²)-H bonds. The reaction represents a good tolerance of a broad scope of enaminones and acrylates. In addition, we also obtained a variety of conjugate dienamines by utilizing various enaminones and propiolates, and the dienamines could be transformed to NH-free pyrroles under optimal conditions. Mechanistic studies suggested that dienamine was most likely to be an active intermediate during the whole reaction.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.orglett.1c01301>.

Experimental procedures, hydrogen gas determination, characterization data, and HRMS and NMR spectra (PDF)

Accession Codes

CCDC 2076610 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Notes

The authors declare no competing financial interest.

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