molal) was 40 times greater than that of the highest (3.5×10^{-3} molal).

A reasonable explanation lies in the assumption of polymerization reactions competitive with the association reactions already referred to. For a reaction of the type

$$y(\mathrm{NiCl_2}) \rightleftarrows (\mathrm{NiCl_2})_y$$

the expression for ϵ_A becomes

$$\epsilon_A = \frac{\epsilon_0 + \epsilon_1 k_1(\mathrm{LiCl}) + \epsilon_2 k_1 k_2(\mathrm{LiCl})^2 + y\epsilon_p k_p(\mathrm{NiCl_2})^{y-1}}{1 + k_1(\mathrm{LiCl}) + k_1 k_2(\mathrm{LiCl})^2 + y k_p(\mathrm{NiCl_2})^{y-1}} \tag{2}$$

where ϵ_p and k_p refer to the extinction coefficient and formation constant of the polymer. If the polymer is much less highly colored than the complexes so that ϵ_p is small by comparison with ϵ_1 and ϵ_2, the results found are consistent with the qualitative predictions of eq. 2. That ϵ_p is comparatively small is shown by the fact that the lowest values of ϵ_A are associated with the highest values of $\mathrm{NiCl_2}$ and the lowest values of LiCl.

The rapid decrease in ϵ_A for relatively small changes in $\mathrm{NiCl_2}$ content indicates that the tendency to form chloro complexes is small compared to the polymerization reaction since the extinction coefficients of the chloro complexes (estimated from the solutions of lowest $\mathrm{NiCl_2}$ and highest LiCl content) are probably in excess of 200. By comparison with $\mathrm{CoCl_2}$ in octanol $\mathrm{NiCl_2}$ forms the weaker chloro complexes although both salts may exist partly as polymerized species.

The results of this investigation agree qualitatively with the observations of Katzin and Gebert[14] on isopropyl and *t*-butyl alcohol solutions of $\mathrm{CoCl_2}$ in the identification of chloro complexes having 2 and 3 chlorine atoms per cobalt. The failure to find evidence of a $\mathrm{CoCl_4^=}$ complex by these authors was doubtless because of the more limited range of chloride-to-cobalt concentration ratios studied. Spectral characteristics reported for $\mathrm{CoCl_4^=}$ in acetone correspond closely to those of the highest chloro complex found in octanol.

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[14] L. I. Katzin and E. Gebert, THIS JOURNAL, **72**, 5464 (1950).

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[CONTRIBUTION FROM THE DEPARTMENT OF CHEMISTRY AND RADIATION LABORATORY, UNIVERSITY OF CALIFORNIA]

Crystal Structures of Americium Compounds[1]

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The crystal structures of several compounds of americium, element 95, have been determined by the X-ray powder diffraction method. $\mathrm{AmF_3}$ is hexagonal ($\mathrm{LaF_3}$ type) with $a = 4.067 \pm 0.001$ Å. and $c = 7.225 \pm 0.002$ Å. for the pseudo-cell which explains the powder data. $\mathrm{AmO_2}$ is cubic ($\mathrm{CaF_2}$ type) with $a = 5.383 \pm 0.001$ Å. AmOCl is tetragonal (PbFCl type) with $a = 4.00 \pm 0.01$ Å., $c = 6.78 \pm 0.01$ Å. The metal parameter is 0.18 ± 0.01. $\mathrm{Am_2O_3}$ is cubic when prepared at 600° and hexagonal when prepared at 800°. For the cubic form ($\mathrm{Mn_2O_3}$ type) $a = 11.03 \pm 0.01$ Å. The metal parameter is -0.030 ± 0.002. For the hexagonal form ($\mathrm{La_2O_3}$ type) $a = 3.817 \pm 0.005$ Å., $c = 5.971 \pm 0.010$ Å.

In coöperation with Professor B. B. Cunningham and his students, we have been observing the X-ray diffraction patterns of various compounds of americium, element 95. The isolation of americium produced by the beta decay of $\mathrm{Pu^{241}}$ has been described by Cunningham.[2] Prior to 1948 the purification was usually terminated when the impurities (chiefly lanthanum) were reduced to one or two per cent. Since that time there have been available americium stocks of much higher purity.[3,4] In this paper we report our results concerning the crystal structures of $\mathrm{AmF_3}$, $\mathrm{AmO_2}$, AmOCl and two forms of $\mathrm{Am_2O_3}$.

The powder diffraction patterns were taken with Cu Kα X-rays ($\lambda = 1.5418$ Å.) in cameras of radius 4.5 cm. The nickel filter was placed between the sample and the film to help diminish the background due to the radiations from the radioactive decay of $\mathrm{Am^{241}}$. When care was taken that the small sample was entirely in the X-ray beam, the background was not troublesome.

Americium Trifluoride.—Americium trifluoride was prepared by Eyring as described elsewhere.[5] Sample A (Table I) was pink in color, like that described by Fried.[6] Sample B was a subsequent preparation which was grey or lavender in color. Its diffraction pattern was almost identical with that of sample A. Sample C was prepared by Dr. J. C. Wallmann from the very pure americium nitrate solution used by Howland and Calvin for their magnetic investigation.[7] Some of this americium was precipitated as the hydroxide and heated with oxygen to 525°. It was then heated in hydrogen fluoride to 715°. The product was pink but after a week, part of it had turned yellow. Sample C consisted of some of this yellow material.

With each of the three samples an excellent powder diffraction pattern was obtained corresponding to the $\mathrm{LaF_3}$-type structure.[8] The unit cell dimensions listed in Table I refer to the hexagonal pseudo-cell containing two americium atoms. This cell accounts for all the diffraction lines observed in the powder patterns. Faint reflections in single crystal patterns[8,9] require an a axis which is

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[1] This research was performed under the auspices of the A.E.C.

[2] B. B. Cunningham, National Nuclear Energy Series, Plutonium Project Record, Vol. 14B, "The Transuranium Elements: Research Papers," Paper No. 19.2, McGraw-Hill Book Co., Inc., New York, N. Y., 1949.

[3] H. R. Lohr and B. B. Cunningham, THIS JOURNAL, **73**, 2025 (1951).

[4] L. Eyring, H. R. Lohr and B. B. Cunningham, *ibid.*, **74**, 1186 (1952).

[5] E. G. Westrum and L. Eyring, *ibid.*, **73**, 3396 (1951).

[6] S. Fried, *ibid.*, **73**, 416 (1951).

[7] J. J. Howland and M. Calvin, *J. Chem. Phys.*, **18**, 239 (1950).

[8] I. Oftedal, *Z. physik. Chem.*, **B5**, 272 (1929); **B13**, 190 (1931).

[9] We have observed these weak reflections for a single crystal of pure synthetic $\mathrm{CeF_3}$, showing that the effect in tysonite[8] is not due to ordering between the cerium and lanthanum atoms.

$\sqrt{3}$ times as large as those listed in Table I. Zachariasen[10,11] first observed this structure for americium trifluoride and has reported two sets of cell dimensions based on samples of uncertain purity. His results (corrected from kX. units) are included in the table. Our three samples disagree by only slightly more than the estimated limits of error. Since the material prepared by Wallmann is likely to be of the highest purity, we adopt as our best values a = 4.067 Å. and c = 7.225 Å. for the pseudo-cell. For the larger cell, a = 7.044 Å. The change from Zachariasen's values is not sufficient to affect his ionic radius for Am^{+3}, 0.99 Å.[12] We attach no structural significance to the color changes but attribute them to surface effects of some kind.

TABLE I

UNIT CELL DIMENSIONS FOR AMERICIUM TRIFLUORIDE

Sample	a, Å.	c, Å.
A[a]	4.069 ± 0.001	7.229 ± 0.002
B[a]	4.069 ± .001	7.229 ± .002
C[a]	4.067 ± .001	7.225 ± .002
Zachariasen[10,b]	4.078 ± .002	7.239 ± .004
Zachariasen[11,b]	4.081 ± .002	7.245 ± .004

[a] Based on Cu Kα, λ = 1.5418 Å. [b] Corrected from kX. units.

Americium Dioxide.—The black americium dioxide was shown by Zachariasen to have the fluorite type structure.[10,11] For samples of uncertain purity, the edge of the cubic unit cell was found to be 5.372 ± 0.005 kX. and 5.377 ± 0.003 kX. We have examined by X-ray diffraction several americium dioxide samples, three of which were especially well crystallized and gave excellent powder patterns. Sample A (Table II) was prepared by Asprey by ignition of the nitrate in air. The americium stock which was used may have contained a few per cent. lanthanum. Sample B was prepared by Eyring from a much purer americium stock by heating the oxalate in air.[4] Some of this material subsequently was heated in a small quartz bomb to 500° with oxygen at 90 atm. pressure; the product is designated as sample C.

The values observed for the unit cell dimension are compared in Table II. The results for samples B and C show that the high pressure oxidation produces no significant change in the composition of this phase. Using the purer americium stock, Asprey and Cunningham[13] measured the amount of oxygen absorbed by americium sesquioxide when heated in oxygen. The result corresponds to AmO_x for the formula of the dioxide, with x = 1.98 ± 0.02, if Am_2O_3 is the correct composition of the sesquioxide. Thus the "black dioxide" has very nearly the ideal composition, rather than an intermediate composition like the so-called Pr_6O_{11}.

The deviation of sample A from samples B and C, though small, is outside the experimental precision. The difference corresponds to an oxygen defect in sample A of 0.01 O per Am, estimated from the data of McCullough[14] for rare earth oxide systems. This defect could be the result of an impurity of about 4 atomic per cent. of a trivalent rare earth element (probably lanthanum) in sample A.

TABLE II

UNIT CELL DIMENSION OF AMERICIUM DIOXIDE

Sample	a, Å.
A	5.387 ± 0.001[a]
B	5.383 ± .001
C	5.383 ± .001
Zachariasen[10]	5.383 ± .005[a,b]
Zachariasen[11]	5.388 ± .003[a,b]

[a] Purity uncertain. [b] Corrected from kX. units.

Americium Oxychloride.—Americium oxychloride was prepared by Asprey by accidental contamination in an experiment designed to yield the sesquioxide by hydrogen reduction of the dioxide. It was identified by its powder diffraction pattern which corresponded to the PbFCl type structure,[15] with the tetragonal unit cell dimensions a = 4.00 ± 0.01 Å., c = 6.78 ± 0.01 Å. A second sample was prepared by C. W. Koch by heating Am_2O_3 at 500° in a mixture of HCl and H_2O vapor. The pattern again corresponded to the unit cell dimensions listed above.

The atoms are located in the following special positions of space group $D_{4h}^7 - P4/nmm$:

2Am in (c): $0, \frac{1}{2}, u;\ \frac{1}{2}, 0, \bar{u}$
2Cl in (c): $0, \frac{1}{2}, v;\ \frac{1}{2}, 0, \bar{v}$
2O in (a): $0, 0, 0;\ \frac{1}{2}, \frac{1}{2}, 0$

Comparison of observed with calculated intensities showed that u = 0.18 ± 0.01. The intensities are not very sensitive to the value of v. We choose v = 0.635 so that the five chlorine neighbors of each americium are at equal distances. This condition is consistent with the parameters which have been found for the isostructural compounds PuOCl,[11] LaOCl,[16] PrOCl,[11] NdOCl,[11] SmOCl[17] and HoOCl.[17]

TABLE III

POWDER DIFFRACTION DATA FOR AmOCl

	$\sin^2\theta$		Intensity		
hkl	Obsd.	Calcd.	Obsd.	Calcd.	
001	...[a]	0.0129	...[a]	16	
101	0.0501	.0501	vs+	56	50
002		.0516			6
110	.0746	.0743	s	30	
111	.0891	.0872	vs	38	3
102		.0887			35
003	.1163	.1162	vw	4	
112	.1268	.1259	m−	12	
200	.1493	.1486	m	14	
103	...	.1538	...	2	
201	.1609	.1615	vw	3	
113	.1913	.1905	m−	11	
211	.1996	.1987	m+	20	16
202		.2002			4
004	...	.2068	...	0.2	
212	.2380	.2374	m+	15	
104	.2438	.2437	w	9	

[a] Not on film.

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(10) W. H. Zachariasen, *Phys. Rev.*, **73**, 1104 (1948).

(11) W. H. Zachariasen, *Acta Cryst.*, **2**, 388 (1949).

(12) W. H. Zachariasen, American Crystallographic Association, Abstracts of Meeting, Pennsylvania State College, April 10–12, 1950.

(13) L. B. Asprey and B. B. Cunningham, to be published.

(14) J. D. McCullough, THIS JOURNAL, **72**, 1386 (1950).

(15) W. Nieuwenkamp and J. M. Bijvoet, *Z. Krist.*, **81**, 469 (1931).

(16) L. G. Sillén and A. L. Nylander, *Svensk. Kem. Tid.*, **53**, 367 (1941).

(17) D. H. Templeton and C. H. Dauben, to be published.

In Table III are listed observed and calculated intensities. No corrections for temperature or absorption have been made. In this structure each americium ion has four oxygen neighbors at 2.34 Å., and five chlorine neighbors at 3.08 Å.

Americium Sesquioxide.—Americium sesquioxide was prepared by Eyring[4] by heating the black dioxide in $^1/_3$ atmosphere of hydrogen at about 600° for half an hour. The product was a reddish brown (persimmon) color. It was first identified by its X-ray diffraction pattern, which corresponded to the cubic Mn_2O_3 type[18] structure, with the cell constant equal to 11.03 ± 0.01 Å. Subsequent experiments by Asprey and Cunningham[13] confirmed its composition.

Another americium oxide was prepared by Carniglia[19] by the ignition of the oxalate at 850°, followed by reduction with $^1/_6$ atmosphere of hydrogen at about 800°. The resulting pale tan material was identified as Am_2O_3 by its diffraction pattern which corresponded to the hexagonal La_2O_3 type structure.[20]

The dimensions of the unit cell are: $a = 3.817 \pm 0.005$ Å., $c = 5.971 \pm 0.010$ Å. A few weak lines are attributed to some unidentified impurity.

For the mineral bixbyite, $(Fe, Mn)_2O_3$, which is isostructural with the cubic Am_2O_3, Pauling and Shappell[18] give the atomic positions as:

Space group T_h^7--Ia3
32 metal atoms in $(0, 0, 0;\ ^1/_2, ^1/_2, ^1/_2) +$
8b: $(^1/_4, ^1/_4, ^1/_4;\ ^1/_4, ^3/_4, ^3/_4;\ \curvearrowleft)$
24d: $\pm(u, 0, ^1/_4;\ \curvearrowleft;\ \bar{u}, ^1/_2, ^1/_4;\ \curvearrowleft)$ with $u = -0.030 \pm 0.005$

48 oxygen atoms in $(0, 0, 0;\ ^1/_2, ^1/_2, ^1/_2) +$
48e: $\pm(x, y, z;\ \curvearrowleft;\ x, \bar{y}, ^1/_2 - z;\ \curvearrowleft;\ ^1/_2 - x, y, \bar{z};\ \curvearrowleft;\ \bar{x}, ^1/_2 - y, z; \curvearrowleft)$
with $x = 0.385 \pm 0.005$
$y = 0.145 \pm 0.005$
$z = 0.380 \pm 0.005$

The observed intensities of the cubic Am_2O_3 pattern, when compared with those calculated for various values of the parameter u, fix that parameter as -0.030 ± 0.002, in exact agreement with the result for bixbyite. This agreement is an excellent confirmation of the previous work, since the calculations for the americium compound are little influenced by the positions chosen for oxygen. The lines most useful in the determination of the parameter were those with $h^2 + k^2 + l^2$ equal to 38, 46, 48, 62 and 64. In Table IV are listed the observed intensities for the cubic Am_2O_3 and those calculated for the parameter values listed above for bixbyite. No correction was made for temperature or absorption. It should be noted that in general (*hkl*) and (*lkh*) have different structure factors. In this structure, each americium atom is surrounded by six oxygen atoms at a distance of 2.36 Å.

TABLE IV

POWDER DIFFRACTION DATA FOR CUBIC Am_2O_3

$h^2 + k^2 + l^2$	$I_{obsd.}$	$I_{calcd.}$	$h^2 + k^2 + l^2$	$I_{obsd.}$	$I_{calcd.}$
12	vs	167	44	s	58
14	vw	6	46	m−	9
16	s	79	48	m	13
18	w	9	50	vw	3
20	...	1	52	...	2
22	vw	4	54	vw	5
24	...	1	56	vw−	2
26	m−	12	58	...	0
30	vw	3	62	vw	5
32	s	62	64	w	8
34	vw	4	66	w	7
36	...	1	68	vw	3
38	w	8	70	vw	4
40	...	2	72	...	2
42	vw	6			

No intensity calculations have been carried out for the hexagonal Am_2O_3, but the similarity of the observed intensities with those for La_2O_3 indicate that the atomic coördinates are very similar in the two substances. If the atomic parameters given by Pauling and Shappell[18] for La_2O_3 are correct for Am_2O_3, then each americium atom has three oxygen neighbors at 2.35, one at 2.36 and three at 2.59 Å.

These two structures are commonplace among the sesquioxides of the rare earth elements.[21–23] Ac_2O_3 has the hexagonal structure.[11] Several studies have been made of the temperature dependence of these structures.[21,23–25] Rapid equilibrium conditions do not prevail, but the data indicate that the hexagonal form is stable at higher temperatures than the cubic form, and that the transition temperature from cubic to hexagonal increases with decreasing cationic radius. Thus, for Nd_2O_3, the cubic form was produced below 775° and the hexagonal form above 850°.[22] Our results with Am_2O_3 are analogous: the cubic form resulted at 600° and the hexagonal form at 800°. This agreement is in accord with the fact that the radii of Am^{+3} and Nd^{+3} are about the same.[10] Since the hexagonal form is unknown for oxides of rare earths smaller in radius than neodymium, Am_2O_3 is probably near the limit of stability of the hexagonal form. The oxides of elements 96 and above, which have smaller cationic radii, are more likely to occur in the cubic form.

Acknowledgment.—We are indebted to Professor Cunningham and to the several persons who prepared the americium compounds. Most of the X-ray photographs were taken by Mrs. Lee Jackson and Mrs. Helena Ruben.

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