

Population and Interconversion of Neutral and Zwitterionic Forms of L-Alanine in Solution

Young Kee Kang,* Byung Jin Byun, Yong Hyun Kim,† Yun Ho Kim,† Dong Hwa Lee,† and Joo Yun Lee

Department of Chemistry, Chungbuk National University, Cheongju, Chungbuk 361-763, Korea

*E-mail: ykkang@chungbuk.ac.kr

†Daejeon Science High School, Daejeon 305-338, Korea

Received May 2, 2008

The conformational study on neutral and zwitterionic L-alanines (N-Ala and Z-Ala, respectively) and the transition state (TS) for their interconversion is carried out using *ab initio* HF and density functional B3LYP methods with the self-consistent reaction field method in the gas phase and in solution. At both the HF and B3LYP levels of theory, the local minimum N1 for N-Ala is found to be most preferred in the gas phase and a weak asymmetric bifurcated hydrogen bond between the amino hydrogens and the carbonyl oxygen appears to play a role in stabilizing this conformation. The local minima N2a and N2b are found to be the second preferred conformations, which seem to be stabilized by a hydrogen bond between the amino nitrogen and the carboxylic hydrogen. The relative stability of the local minimum N2b is remarkably increased in solution than that in the gas phase. The local minimum N2b becomes more stable than the local minimum N2a in most of the solution. On the whole the relative free energies of Z-Ala and TS become more lowered, as the solvent polarity increases. N-Ala prevails over Z-Ala in aprotic solutions but Z-Ala is dominantly populated in ethanol and water. In aprotic solutions, the population of Z-Ala increases somewhat with the increase of solvent polarity. The barrier to Z-Ala-to-N-Ala interconversion increases on the whole with the increase of solvent polarity, which is caused by the increase of stability for Z-Ala.

Key Words : L-Alanine, Population, Interconversion, Solvent effects

Introduction

L-Alanine exists as a zwitterion in the crystal^{1,2} and in aqueous solution,^{3,4} whereas it has a neutral form in the gas phase,⁵⁻⁷ as known for other amino acids. Because the zwitterion possesses an extremely high electric moment, it is very soluble in water but nearly insoluble in organic solvents.³ However, the neutral form has a greater affinity for organic solvents than for water, due to the relatively smaller electric moment.³ It was suggested that the concentration ratio of zwitterions to neutral molecules should decrease, as the dielectric constant (*i.e.*, the polarity) of the solvent is lowered.³

Considerable quantum chemical computations have been carried out on the conformational preferences of neutral alanine in the gas phase,⁸⁻¹⁵ and neutral and/or zwitterionic alanines in aqueous solution^{13,14,16-21} and in a KBr matrix.^{22,23} For most of the previous works in solution, solvents were considered as a continuum and the self-consistent reaction field (SCRF) methods were employed to calculate solvation free energies.^{13,14,16,22,23} A few works have adopted the models of zwitterionic¹⁷⁻²⁰ or neutral²⁰ alanine with explicit waters embedded in a SCRF.

In the gas phase, the most preferred conformation and up to 13 local minima for the neutral alanine were identified depending on the levels of theory and starting structures. In particular, the zwitterionic alanine can be located as a local minimum at the HF/6-31G(d) level of theory, but it is not a local minimum at B3LYP/6-31+G(d) and B3LYP/6-

311++G(d,p) levels of theory.¹³ It was suggested that at least two water molecules are needed to form a stable alanine zwitterion-water cluster in the gas phase.²⁰ The zwitterionic form was found to be more stable than the neutral form when the solvation shell reaches eight water molecules.²¹

In water, the zwitterionic alanine was found to be more stable by about 3-4 kcal/mol in free energy than the neutral alanine using polarizable continuum models.^{16,20} There are limited works reported on the interconversion between neutral and zwitterionic alanines in water^{20,24} and in organic solvents,²⁵ in which only representative conformations of the neutral alanine were employed.

We studied here all feasible conformations for neutral and zwitterionic L-alanines and the transition states for their interconversion at the HF and B3LYP levels of theory with the conductor-like polarizable continuum model (CPCM) in solution in order to investigate their relative populations and interconversion with the increase of solvent polarity.

Computational Methods

Chemical structure and atomic labels of L-alanine are shown in Figure 1. All *ab initio* HF and hybrid density functional B3LYP calculations were carried out using the Gaussian 03 package.²⁶ The 13 local minima of L- and D-alanines optimized at the HF/6-311++G(d,p) level of theory¹⁰ and the 9 local minima of Ac-L-Ala-NHMe optimized by the ECEPP/2 force field²⁷ were used as starting conformations for optimization of the neutral L-alanine (N-Ala) at HF/6-

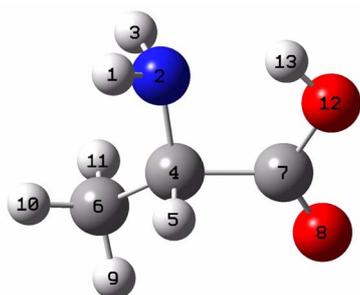


Figure 1. Chemical structure and atomic labels for L-alanine.

31+G(d) and B3LYP/6-31+G(d) levels of theory in the gas phase. The local minima at these HF and B3LYP levels of theory were edited by the GaussView²⁸ to obtain the starting structures for optimization of the zwitterionic L-alanine (Z-Ala) in the gas phase. After optimization, a single local minimum for Z-Ala was located at the HF/6-31+G(d) level of theory, but no stationary point was found at the B3LYP/6-31+G(d) level of theory, as seen previously.¹³ The transition state (TS) between N-Ala and Z-Ala was located using the Synchronous Transit and Quasi-Newton (STQN) method²⁹ with the QST3 option of the Gaussian 03³⁰ at the HF/6-31+G(d) level of theory. N-Ala with a similar configuration of amino and carboxylic groups to Z-Ala was chosen as initial structure in generating the TS structure.

We employed the conductor-like polarizable continuum model (CPCM),^{31,32} implemented in the Gaussian 03 package,²⁶ to compute solvation free energies (ΔG_{solv}) at the HF/6-31+G(d) level of theory with the united atom Kohn-Sham (UAKS) cavities, which are the united atom topological model (UATM) radii optimized at the density functional PBE0/6-31G(d) level of theory.^{33,34} The solvation free energy is the sum of the electrostatic free energy and the non-electrostatic energy terms.³⁵ The latter is composed of the cavitation, dispersion, and repulsion energy terms. For CPCM-UAKS calculations, the default average areas of 0.2 Å² for tesseræ were used. The solvents considered here are chloroform, dichloroethane, ethanol, acetonitrile, DMSO, and water, whose dielectric constants are 4.9, 10.4, 24.6, 36.6, 46.7, and 78.4 at 25 °C, respectively.³⁰ Recently, the CPCM-UAKS calculations for a number of neutral and charged organic molecules at the HF/6-31+G(d)/HF/6-31+G(d) and HF/6-31+G(d)/B3LYP/6-31+G(d) levels of theory provided hydration free energies in agreement with available experimental data.³⁶

All local minima for N-Ala optimized at HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory in the gas phase were used as starting structures for optimizations at CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory in solution, respectively. Z-Ala and TS optimized at the HF/6-31+G(d) level of theory in the gas phase were used as starting structures for optimizations at both the CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory in solution. The B3LYP/6-311++G(d,p) single-point energies were calculated for all local minima and transition states of L-alanine located at CPCM HF/6-31+G(d) and

CPCM B3LYP/6-31+G(d) levels of theory in solution.

Vibrational frequencies were calculated for all stationary points at both the HF and B3LYP levels of theory in the gas phase and in solution, which were used to compute enthalpies and Gibbs free energies with the scale factors of 0.89 and 0.98 at HF and B3LYP levels of theory, respectively, at 25 °C and 1 atm. These scale factors were chosen to reproduce experimental frequencies for the amide I band of *N*-methylacetamide in Ar and N₂ matrixes.³⁷ The zero-point energy correction and the thermal energy corrections were used to calculate the enthalpy (H) and entropy (S) of each conformation.^{30,38} The analysis uses the standard thermodynamic expressions for an ideal gas in the canonical ensemble. Each transition state was confirmed by checking whether it had one imaginary frequency after frequency calculations at both the HF and B3LYP levels of theory. The relative total free energy (ΔG) for each conformation in solution was computed by taking the sum of the relative conformational free energy (ΔE_c), the thermal contributions, and the entropic contribution. The relative conformational free energy (ΔE_c) is the sum of the conformational electronic energy ($\Delta E_{e,s}$) and the relative solvation free energy ($\Delta\Delta G_{\text{solv}}$) in solution. The relative total free energies are used here to interpret the population and interconversion between N-Ala and Z-Ala in the gas phase and in solution.

Results and Discussion

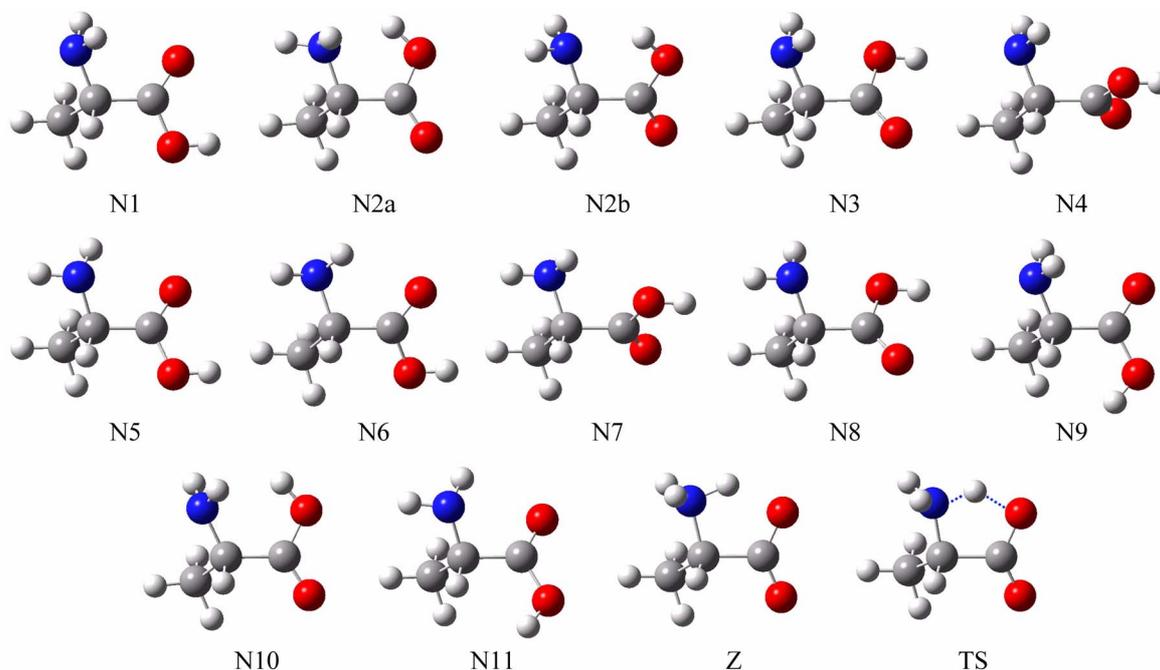
Conformational Preferences in the Gas Phase. Thermodynamic properties of N-Ala, Z-Ala, and their TS calculated at HF/6-31+G(d), B3LYP/6-311++G(d,p)/HF/6-31+G(d), B3LYP/6-31+G(d), and B3LYP/6-311++G(d,p)/B3LYP/6-31+G(d) levels of theory in the gas phase are listed in Tables 1-4, respectively. The conformations of all feasible local minima and TS found at HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory in the gas phase are shown in Figure 2. Their selected bond lengths, bond angles, and torsion angles optimized at HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory in the gas phase are listed in the Supporting Information.

At the HF/6-31+G(d) level of theory, the 12 local minima are found for N-Ala in the gas phase (Tables 1 and 2). The eight local minima of them such as N1, N2b, N4, N5, N8, N9, N10, and N11 correspond to the local minima I, IIB, IIB, IVA, VB, VI, VII, and VIIIA for L-alanine found at the HF/6-311++G(d,p) level of theory, respectively, which are listed in Tables 2 and 3 of ref 10. The local minimum N2a is also identified as a local minimum at MP2/6-311++G(d,p)⁷ and B3LYP/aug-cc-pVDZ¹¹ levels of theory. The conformations of the remaining three local minima such as N3, N6, and N7 are similar to those of the local minima N8, N1, and N4, respectively, except having the different orientations of amide hydrogens. The relative conformational stabilities of N-Ala in the gas phase are calculated to be in the order N1 > N3 > N4 > N5 > N6 > N7 > N8 > N2a ≈ N2b > N9 > N10 > N11 at the HF/6-31+G(d) level of theory by both the relative electronic energies (ΔE_c) and free energies (ΔG). At the

Table 1. Thermodynamic Properties of Neutral and Zwitterionic L-Alanines and Their Transition State at the HF/6-31+G(d) Level of Theory in the Gas Phase and in Solution^a

solvent		N1	N2a	N2b	N3	N4	N5	N6	N7	N8	N9	N10	N11	Z	TS
gas phase	ΔE_c	0.00	2.38	2.47	1.22	1.56	1.69	1.80	1.92	2.25	7.14	8.48	8.86	24.66	25.37
	ΔH	0.00	2.54	2.58	1.22	1.52	1.61	1.68	1.80	2.16	7.02	8.17	8.66	24.63	23.09
	ΔG	0.00	2.96	2.99	1.08	1.28	1.67	1.83	1.97	2.23	7.01	8.25	8.53	24.82	24.17
CHCl ₃	ΔE_c	0.00	0.59	0.52	0.99	1.41	1.30	1.23	1.47	1.60	4.69	5.92	5.91	7.76	16.94
	ΔH	0.00	0.82	0.67	1.00	1.38	1.26	1.15	1.40	1.56	4.66	5.75	5.85	8.75	14.63
	ΔG	0.00	1.24	1.06	0.85	0.96	1.31	1.22	1.44	1.51	4.71	5.79	5.70	7.86	15.83
CH ₂ ClCH ₂ Cl	ΔE_c	0.00	0.25	0.15	0.95	1.38	1.19	1.12	1.38	1.48	4.25	5.46	5.34	4.61	15.64
	ΔH	0.00	0.49	0.30	0.96	1.36	1.16	1.04	1.32	1.44	4.24	5.31	5.29	5.65	13.34
	ΔG	0.00	0.91	0.70	0.82	0.92	1.16	1.08	1.35	1.39	4.30	5.34	5.15	5.51	14.55
EtOH	ΔE_c	0.00	1.38	0.93	1.00	1.28	1.14	0.96	1.45	1.47	2.98	4.89	4.00	-3.14	15.40
	ΔH	0.00	1.77	1.01	1.08	1.20	1.16	0.81	1.32	1.50	3.03	4.92	4.06	-1.48	13.74
	ΔG	0.05	2.08	1.41	0.99	0.00	1.12	0.93	1.33	1.37	3.27	5.00	4.06	-1.18	15.06
CH ₃ CN	ΔE_c	0.10	0.09	0.00	1.01	1.48	1.21	1.14	1.43	1.50	4.02	5.21	5.01	2.59	14.91
	ΔH	0.00	0.23	0.05	0.92	1.35	1.08	0.96	1.27	1.37	3.90	4.96	4.86	3.55	12.51
	ΔG	0.00	0.55	0.42	0.76	0.90	1.08	1.00	1.27	1.30	3.96	4.99	4.76	3.47	13.72
DMSO	ΔE_c	0.15	0.12	0.00	1.07	1.50	1.26	1.15	1.45	1.55	4.09	5.27	5.08	2.39	14.82
	ΔH	0.01	0.21	0.00	0.93	1.32	1.08	0.93	1.23	1.37	3.93	4.98	4.89	3.30	12.38
	ΔG	0.00	0.57	0.38	0.75	0.79	0.99	0.96	1.22	1.28	3.98	5.01	4.74	3.21	13.58
water	ΔE_c	0.00	1.26	0.67	0.98	1.18	0.95	0.83	1.39	1.47	2.62	4.30	3.51	-4.63	14.94
	ΔH	0.00	1.66	0.73	1.09	1.10	0.93	0.68	1.28	1.52	2.65	3.71	3.51	-2.92	13.34
	ΔG	0.00	2.05	1.18	1.04	0.33	0.79	0.81	1.30	1.41	3.03	5.42	3.62	-2.60	14.67

^a ΔE_c , ΔH , and ΔG are relative electronic energies, enthalpies, and Gibbs free energies, respectively; units in kcal/mol. ΔH and ΔG are calculated at 25 °C.

**Figure 2.** Conformations of all feasible local minima for N-Ala and Z-Ala and TS found at HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory in the gas phase. Conformations N1-N11 are for N-Ala. The local minimum of Z-Ala is denoted by Z.

B3LYP/6-311++G(d,p)//HF/6-31+G(d) level of theory, the two local minima N2a and N2b, however, become more favored than the local minimum N3 but the relative stabilities of the other local minima are the same as those at the

HF/6-31+G(d) level of theory. The calculated stabilities of local minima at HF/6-31+G(d) and B3LYP/6-311++G(d,p)//HF/6-31+G(d) levels of theory are similar to those by the ΔE_c at HF/6-311++G(d,p) and B3LYP/6-311++G(d,p) levels

Table 2. Thermodynamic Properties of Neutral and Zwitterionic L-Alanines and Their Transition State at the B3LYP/6-311++G(d,p)//HF/6-31+G(d) Level of Theory in the Gas Phase and in Solution^a

solvent		N1	N2a	N2b	N3	N4	N5	N6	N7	N8	N9	N10	N11	Z	TS
gas phase	ΔE_e	0.00	0.36	0.40	1.10	1.34	1.35	1.48	2.01	2.29	5.74	6.08	7.01	17.55	14.58
	ΔH	0.00	0.52	0.51	1.11	1.30	1.27	1.35	1.90	2.20	5.62	5.78	6.80	17.52	12.29
	ΔG	0.00	0.95	0.92	0.96	1.06	1.33	1.51	2.07	2.27	5.61	5.85	6.68	17.71	13.37
CHCl ₃	ΔE_e	1.77	0.21	0.00	2.67	3.00	2.78	2.74	3.36	3.43	4.96	5.34	5.89	6.14	6.75
	ΔH	1.62	0.30	0.00	2.54	2.82	2.60	2.50	3.14	3.24	4.78	5.03	5.68	6.98	4.30
	ΔG	1.24	0.33	0.00	1.99	2.02	2.26	2.18	2.79	2.80	4.44	4.69	5.14	5.70	5.11
CH ₂ ClCH ₂ Cl	ΔE_e	2.18	0.25	0.00	3.04	3.38	3.08	3.03	3.67	3.70	4.90	5.29	5.73	3.48	5.80
	ΔH	2.03	0.34	0.00	2.91	3.21	2.90	2.80	3.46	3.51	4.74	4.99	5.53	4.37	3.36
	ΔG	1.63	0.36	0.00	2.36	2.37	2.50	2.45	3.09	3.06	4.39	4.63	4.98	3.83	4.17
EtOH	ΔE_e	1.14	0.55	0.00	2.06	2.28	1.99	1.76	2.61	2.67	2.49	3.76	3.29	-5.35	4.79
	ΔH	1.06	0.86	0.00	2.06	2.12	1.94	1.54	2.40	2.62	2.46	3.72	3.26	-3.77	3.06
	ΔG	0.71	0.77	0.00	1.57	0.52	1.50	1.26	2.01	2.09	2.30	3.40	2.87	-3.87	3.97
CH ₃ CN	ΔE_e	2.47	0.17	0.00	3.30	3.64	3.29	3.24	3.91	3.95	4.82	5.23	5.57	1.62	5.23
	ΔH	2.32	0.26	0.00	3.17	3.47	3.11	3.02	3.70	3.77	4.66	4.94	5.37	2.53	2.79
	ΔG	1.95	0.21	0.00	2.62	2.65	2.74	2.68	3.33	3.33	4.35	4.59	4.90	2.07	3.63
DMSO	ΔE_e	2.51	0.22	0.00	3.35	3.70	3.34	3.25	3.92	3.99	4.89	5.29	5.63	1.36	5.14
	ΔH	2.36	0.31	0.00	3.22	3.51	3.16	3.02	3.70	3.81	4.73	5.00	5.44	2.27	2.70
	ΔG	1.97	0.29	0.00	2.65	2.61	2.68	2.67	3.31	3.34	4.40	4.65	4.91	1.79	3.52
water	ΔE_e	1.42	0.68	0.00	2.34	2.49	2.07	1.92	2.85	2.97	2.41	3.43	3.04	-6.60	4.66
	ΔH	1.36	1.03	0.00	2.39	2.34	1.99	1.72	2.68	2.97	2.38	2.78	2.99	-4.95	3.00
	ΔG	0.91	0.96	0.00	1.88	1.12	1.40	1.38	2.25	2.39	2.31	4.03	2.64	-5.08	3.87

^aSee footnote *a* of Table 1.**Table 3.** Thermodynamic Properties of Neutral and Zwitterionic L-Alanines and Their Transition State at the B3LYP/6-31+G(d) Level of Theory in the Gas Phase and in Solution^a

solvent		N1	N2a	N2b	N3	N5	N6	N7	N8	N9	N11	Z	TS
gas phase	ΔE_e	0.00	0.01	0.06	1.09	1.44	1.66	2.13	2.33	5.98	7.37		
	ΔH	0.00	0.13	0.17	1.12	1.36	1.49	1.97	2.23	5.80	7.09		
	ΔG	0.00	0.63	0.69	0.96	1.46	1.71	2.11	2.32	5.72	7.10		
CHCl ₃	ΔE_e	1.95	0.08	0.00	2.83	3.08	3.11	3.63	3.70	5.78	6.88	4.14	7.38
	ΔH	1.82	0.06	0.00	2.74	2.91	2.85	3.38	3.52	5.60	6.65	4.46	5.11
	ΔG	1.47	0.00	0.08	2.22	2.58	2.70	2.99	3.10	5.30	6.20	4.67	5.99
CH ₂ ClCH ₂ Cl	ΔE_e	2.36	0.12	0.00	3.19	3.40	3.42	3.95	3.99	5.79	6.80	2.16	6.55
	ΔH	2.24	0.11	0.00	3.11	3.24	3.17	3.71	3.81	5.63	6.59	2.66	4.27
	ΔG	1.80	0.21	0.00	2.53	2.85	2.85	3.21	3.30	5.27	6.08	2.71	5.09
EtOH	ΔE_e	1.44	0.28	0.00	2.10	2.32	2.16	2.81	2.84	3.57	4.62	-3.37	5.98
	ΔH	1.60	0.00	0.20	2.26	2.42	1.87	2.53	2.92	3.65	4.72	-1.32	4.77
	ΔG	0.71	1.00	0.00	1.21	1.31	1.16	1.71	1.90	3.02	3.92	-2.19	5.24
CH ₃ CN	ΔE_e	2.64		0.00	3.44	3.61	3.64	4.18	4.19	5.75	6.69	0.82	6.06
	ΔH	2.52		0.00	3.36	3.47	3.40	3.95	4.03	5.61	6.50	1.43	3.78
	ΔG	2.06		0.00	2.78	3.05	3.05	3.41	3.49	5.23	6.02	1.39	4.61
DMSO	ΔE_e	2.70	0.08	0.00	3.51	3.66	3.66	4.20	4.24	5.84	6.77	0.62	5.97
	ΔH	2.59	0.04	0.00	3.43	3.53	3.42	3.97	4.08	5.70	6.58	1.23	3.69
	ΔG	2.39	0.00	0.27	3.09	3.48	3.35	3.70	3.80	5.59	6.34	1.46	4.78
water	ΔE_e	1.47		0.00	2.36	2.40	2.25	2.96	3.03	3.48	4.39	-4.52	5.83
	ΔH	1.52		0.00	2.56	2.45	1.91	2.70	3.12	3.56	4.42	-2.40	4.71
	ΔG	0.71		0.00	1.89	2.03	1.49	2.17	2.43	3.24	4.03	-3.03	5.47

^aSee footnote *a* of Table 1.

Table 4. Thermodynamic Properties of Neutral and Zwitterionic L-Alanines and Their Transition State at the B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) Level of Theory in the Gas Phase and in Solution^a

solvent	N1	N2a	N2b	N3	N5	N6	N7	N8	N9	N11	Z	TS	
gas phase	ΔE_c	0.00	0.03	0.02	1.07	1.32	1.46	1.98	2.25	5.73	6.95		
	ΔH	0.00	0.15	0.12	1.10	1.24	1.29	1.82	2.15	5.55	6.68		
	ΔG	0.00	0.66	0.64	0.94	1.33	1.51	1.95	2.23	5.47	6.69		
CHCl ₃	ΔE_c	1.96	0.13	0.00	2.83	3.01	2.95	3.51	3.66	5.56	6.59	6.18	8.24
	ΔH	1.83	0.11	0.00	2.74	2.84	2.69	3.27	3.47	5.38	6.35	6.49	5.97
	ΔG	1.42	0.00	0.02	2.17	2.45	2.49	2.81	3.01	5.02	5.85	6.65	6.80
CH ₂ ClCH ₂ Cl	ΔE_c	2.36	0.18	0.00	3.19	3.32	3.26	3.82	3.94	5.56	6.51	4.28	7.32
	ΔH	2.24	0.17	0.00	3.11	3.16	3.01	3.59	3.76	5.40	6.30	4.78	5.04
	ΔG	1.80	0.27	0.00	2.53	2.77	2.69	3.09	3.25	5.04	5.79	4.83	5.86
EtOH	ΔE_c	1.47	0.31	0.00	2.15	2.26	2.06	2.74	2.85	3.38	4.37	-0.97	6.42
	ΔH	1.60	0.00	0.17	2.28	2.33	1.73	2.43	2.89	3.43	4.44	1.05	5.18
	ΔG	0.74	1.03	0.00	1.26	1.25	1.05	1.65	1.91	2.83	3.68	0.22	5.68
CH ₃ CN	ΔE_c	2.63		0.00	3.43	3.53	3.47	4.05	4.14	5.52	6.40	2.98	6.77
	ΔH	2.51		0.00	3.36	3.39	3.23	3.82	3.98	5.38	6.21	3.59	4.49
	ΔG	2.05		0.00	2.78	2.97	2.88	3.28	3.43	5.00	5.73	3.55	5.31
DMSO	ΔE_c	2.69	0.12	0.00	3.51	3.58	3.49	4.07	4.19	5.61	6.48	2.78	6.68
	ΔH	2.58	0.08	0.00	3.43	3.44	3.25	3.84	4.02	5.47	6.30	3.40	4.40
	ΔG	2.34	0.00	0.23	3.05	3.35	3.14	3.53	3.71	5.31	6.01	3.58	5.44
water	ΔE_c	1.50		0.00	2.42	2.34	2.15	2.90	3.03	3.28	4.13	-2.11	6.24
	ΔH	1.55		0.00	2.62	2.40	1.81	2.63	3.12	3.36	4.16	0.02	5.11
	ΔG	0.74		0.00	1.95	1.97	1.39	2.10	2.43	3.05	3.78	-0.61	5.88

^aSee footnote a of Table 1.of theory, respectively.¹⁰

At the B3LYP/6-31+G(d) level of theory, the ten local minima are found for N-Ala in the gas phase (Tables 3 and 4). In particular, the two local minima N4 and N10 at the HF/6-31+G(d) level of theory do not exist any longer at the B3LYP/6-31+G(d) level of theory, as found at the B3LYP/6-311++G(d,p) level of theory.¹⁰ The relative conformational stabilities of N-Ala are calculated to be in the order N1 > N2a ≈ N2b > N3 > N5 > N6 > N7 > N8 > N9 > N11 at the B3LYP/6-31+G(d) level of theory by both the ΔE_c and ΔG . At the B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) level of theory, the relative stabilities of local minima are calculated to be similar to those at the B3LYP/6-31+G(d) level of theory, except that the local minimum N2b becomes a little more preferred over the local minimum N2a. In particular, the calculated stabilities of local minima at B3LYP/6-31+G(d) and B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) levels of theory are quite similar to those by the ΔE_c at both the B3LYP/6-311++G(d,p) and MP2/6-311++G(d,p) levels of theory.¹⁰

At all the levels of theory employed here, the local minimum N1 for N-Ala is found to be most preferred in the gas phase and a weak asymmetric bifurcated hydrogen bond between the amino hydrogens and the carbonyl oxygen appears to play a role in stabilizing this conformation (Figure 2), whose distances are computed to be 2.68 and 2.87 Å at the HF/6-31+G(d) level of theory and 2.68 and 2.94 Å at the B3LYP/6-31+G(d) level of theory. The corresponding values are 2.70 and 2.88 Å at the MP2/6-

311++G(d,p) level of theory.⁷ The local minima N2a and N2b are found to be the second preferred conformations at B3LYP/6-311++G(d,p)//HF/6-31+G(d), B3LYP/6-31+G(d), and B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) levels of theory, as found at MP2/6-311++G(d,p)⁷ and B3LYP/aug-cc-pVDZ¹¹ levels of theory. The values of ΔE_c and ΔG for the local minima N2a and N2b are computed to be 0.36 and 0.95 kcal/mol and 0.40 and 0.92 kcal/mol, respectively, at the B3LYP/6-311++G(d,p)//HF/6-31+G(d) level of theory, 0.01 and 0.63 kcal/mol and 0.06 and 0.69 kcal/mol, respectively, at the B3LYP/6-31+G(d) level of theory, 0.03 and 0.66 kcal/mol and 0.02 and 0.64 kcal/mol, respectively, at the B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) level of theory. These two local minima N2a and N2b seem to be stabilized by a hydrogen bond between the amino nitrogen and the carboxylic hydrogen (Figure 2), with the distances of 2.05 and 2.03 Å at the HF/6-31+G(d) level of theory, respectively, and 1.91 and 1.90 Å at the B3LYP/6-31+G(d) level of theory, respectively. The corresponding distance of the local minimum N2a is 1.96 Å at the MP2/6-311++G(d,p) level of theory.⁷ From electron diffraction,⁶ microwave spectroscopy,⁷ and matrix-isolation infrared¹¹ experiments, it has been observed that L-alanine adopts the two most stable conformations N1 and N2a. Our results for the relative stabilities obtained at the B3LYP/6-311++G(d,p)//HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory are reasonably consistent with these observations, although the difference in energy or free energy between two local minima N2a and N2b is not remarkable.

Although the local minima for Z-Ala and TS are obtained at the HF/6-31+G(d) level of theory, the free energy of Z-Ala is calculated to be higher than that of TS by 0.65 kcal/mol at the HF/6-31+G(d) level of theory and 4.34 kcal/mol at the B3LYP/6-311++G(d,p)//HF/6-31+G(d) level of theory.

Conformational Preferences in Solution. Thermodynamic properties of N-Ala, Z-Ala, and their TS calculated at HF/6-31+G(d), B3LYP/6-311++G(d,p)//HF/6-31+G(d), B3LYP/6-31+G(d), and B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d) levels of theory using the CPCM method in solution are listed in Tables 1-4, respectively. The selected torsion angles of all local minima and TS in solution and their selected bond lengths and bond angles in water at the CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory are listed in the Supporting Information. For local minima of N-Ala and Z-Ala and their TS in solution, there are no remarkable shifts in torsion angles, *i.e.* conformations, from those in the gas phase at both the CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory. At both the CPCM HF and B3LYP levels of theory, most of the local minima for N-Ala and Z-Ala and their TS obtained in the gas phase are retained as stationary points in solution.

At all the levels of theory employed here, the relative stability of the local minimum N2b by ΔG is remarkably increased in solution than that in the gas phase. The local minimum N2b becomes more stable than the local minimum N2a in most of the solution. In particular, the local minimum N2a does not exist as a stationary point any longer at the CPCM B3LYP/6-31+G(d) level of theory in acetonitrile and water.

Although the relative conformational stabilities of N-Ala by ΔG are different from solution to solution at the CPCM HF/6-31+G(d) level of theory, the local minimum N1 is most preferred in most of the solution, except that the local minimum N4 is a little favored than the local minimum N1 in ethanol. However, the local minimum N2b becomes the most preferred conformation by ΔG and is followed by the local minimum N2a at the B3LYP/6-311++G(d,p)//CPCM HF/6-31+G(d) level of theory in solution, except in ethanol and water. The local minima N4 and N1 are the second preferred conformations in ethanol and water, respectively, which are more stabilized by 0.25 and 0.05 kcal/mol in ΔG than the local minimum N2a, respectively. At the CPCM B3LYP/6-31+G(d) and B3LYP/6-311++G(d,p)//CPCM B3LYP/6-31+G(d) levels of theory in solution, the local minimum N2b appears to be most preferred, although its free energy is higher by 0.08 and 0.27 kcal/mol than the local minimum N2a in chloroform and DMSO, respectively. In particular, the local minimum N1 is found to be the second preferred conformation in ethanol, acetonitrile, and water.

As found in the gas phase, the local minimum N1 for N-Ala is found to have a weak asymmetric bifurcated hydrogen bond between the amino hydrogens and the carbonyl oxygen in solution, whose the shorter distance becomes longer by 0.01-0.02 Å in all the solution and the longer distance are

shortened by 0.01-0.03 Å only in ethanol, acetonitrile, and water at the CPCM HF/6-31+G(d) level of theory, compared to the values at the HF/6-31+G(d) level of theory in the gas phase. At the CPCM B3LYP/6-31+G(d) level of theory in all the solution, the former distance is a little lengthened by up to 0.03 Å and the latter distance becomes shorter by 0.01 Å, except by 0.06 Å in water.

In solution, the two local minima N2b and N2a seem to be stabilized by a hydrogen bond between the amino nitrogen and the carboxylic hydrogen, as seen in the gas phase. At the CPCM HF/6-31+G(d) level of theory, the hydrogen-bond distances of these two local minima become shorter by 0.03-0.04 Å in aprotic solutions and a little longer by 0.01 Å in ethanol and water, compared to those in the gas phase. At the CPCM B3LYP/6-31+G(d) level of theory, these hydrogen-bond distances become more shortened by 0.06-0.11 Å in solution. Thus, the shorter hydrogen-bond distances of these two local minima N2b and N2a seem to play a role in more stabilizing these conformations over the conformation N1 in solution. It should be noted that the local minimum N2a does not exist as a stationary point any longer at this CPCM B3LYP levels of theory in acetonitrile and water, as discussed above.

Z-Ala and TS exist as stationary points at both the CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory in solution. On the whole the relative free energies of Z-Ala and TS become more lowered at all the levels of theory considered here, as the solvent polarity increases. In particular, Z-Ala becomes more stable than the most preferred conformations N2b, N2a, or N1 at all the levels of theory in ethanol and water, except at the B3LYP/6-311++G(d,p)//CPCM B3LYP/6-31+G(d) level of theory in ethanol.

In water, the relative free energy of Z-Ala to the local minimum N1 is calculated to be -2.60 kcal/mol at the CPCM HF/6-31+G(d) level of theory, which was reported to be -4.48 kcal/mol at the B3LYP/6-31G(d) level of theory using the polarizable continuum model (PCM).¹⁶ The relative free energy of Z-Ala to the local minimum N2b is calculated to be -5.08 and -3.03 kcal/mol at B3LYP/6-311++G(d,p)//CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory, respectively. The corresponding relative free energy was computed to be -3.13 kcal/mol at the B3PW91/6-31++G(d,p) level of theory with the PCM.²⁰

Populations of N-Ala and Z-Ala in Solution. The populations of N-Ala and Z-Ala in solution are listed in Table 5. Each population was computed using the normalized Boltzmann weight by ΔG in solution, shown in Tables 1-4. At the CPCM HF/6-31+G(d), B3LYP/6-311++G(d,p)//CPCM HF/6-31+G(d), and CPCM B3LYP/6-31+G(d) levels of theory, N-Ala prevails over Z-Ala in aprotic solutions but Z-Ala is dominantly populated in ethanol and water. In aprotic solutions, the population of Z-Ala increases somewhat with the increase of solvent polarity, except that the population of Z-Ala in DMSO is lower than that in acetonitrile. At the B3LYP/6-311++G(d,p)//CPCM B3LYP/6-31+G(d) level of theory, N-Ala prevails over Z-Ala in both aprotic solutions and ethanol. In water, N-Ala is populated

Table 5. Populations and Ratios of Concentrations for N-Ala and Z-Ala in Solution

level of theory	pop.	CHCl ₃	CH ₂ ClCH ₂ Cl	EtOH	CH ₃ CN	DMSO	water
HF/6-31+G(d)	N% ^a	100.00	100.00	27.51	99.90	99.86	3.20
	Z% ^b	0.00	0.00	72.49	0.10	0.14	96.80
	log K _Z ^c	-6.09	-4.43	0.42	-3.01	-2.84	1.48
B3LYP/6-311++G(d,p)// HF/6-31+G(d)	N% ^a	100.00	99.91	0.34	98.34	97.22	0.04
	Z% ^b	0.00	0.09	99.66	1.66	2.78	99.96
	log K _Z ^c	-4.44	-3.03	2.47	-1.77	-1.54	3.45
B3LYP/6-31+G(d)	N% ^a	99.98	99.43	4.69	91.67	95.15	0.90
	Z% ^b	0.02	0.57	95.31	8.33	4.85	99.10
	log K _Z ^c	-3.73	-2.24	1.31	-1.04	-1.29	2.04
B3LYP/6-311++G(d,p)// B3LYP/6-31+G(d)	N% ^a	100.00	99.98	74.11	99.76	99.86	34.83
	Z% ^b	0.00	0.02	25.89	0.24	0.14	65.17
	log K _Z ^c	-5.20	-3.78	-0.46	-2.63	-2.86	0.27

^aPopulation of N-Ala. ^bPopulation of Z-Ala. ^cThe ratio of the concentrations for Z-Ala and N-Ala; $K_Z = [Z\text{-Ala}]/[N\text{-Ala}]$.

Table 6. Barriers to Interconversion ($\Delta G^\ddagger_{Z \rightarrow N}$) of Z-Ala to N-Ala in Solution^a

level of theory	CHCl ₃	CH ₂ ClCH ₂ Cl	EtOH	CH ₃ CN	DMSO	water
HF/6-31+G(d)	7.97	9.04	16.24	10.25	10.37	17.27
B3LYP/6-311++G(d,p)//HF/6-31+G(d)	-0.59	0.34	7.84	1.56	1.73	8.95
B3LYP/6-31+G(d)	1.32	2.38	7.43	3.22	3.32	8.50
B3LYP/6-311++G(d,p)//B3LYP/6-31+G(d)	0.15	1.03	5.46	1.76	1.86	6.49

^aUnits in kcal/mol.

by ~35%, although Z-Ala is dominated in water.

The ratios of the concentrations for Z-Ala and N-Ala can be computed using the equation $\log K_Z = [Z\text{-Ala}]/[N\text{-Ala}]$ and their values are listed in Table 5. The values of $\log K_Z$ are positive in ethanol and water at the first three levels of theory and only in water at the fourth level of theory. Our calculated values of 3.45 and 2.04 for $\log K_Z$ at B3LYP/6-311++G(d,p)//CPCM HF/6-31+G(d) and CPCM B3LYP/6-31+G(d) levels of theory in water, respectively, are reasonably consistent with the observed value of 5.41 for α -alanine in water.³ The calculated values of 3.45 and 2.04 for $\log K_Z$ correspond to 99.96 and 99.10% of Z-Ala in water, respectively.

Barriers to Z-Ala-to-N-Ala Interconversion in Solution.

The barriers to Z-Ala-to-N-Ala interconversion ($\Delta G^\ddagger_{Z \rightarrow N}$) in solution are shown in Table 6. The relative free energies of Z-Ala and TS to the most preferred conformation of N-Ala at the CPCM B3LYP/6-31+G(d) level of theory in each solution are plotted in Figure 3.

At all the levels of theory, the value of $\Delta G^\ddagger_{Z \rightarrow N}$ increases on the whole with the increase of solvent polarity. The barriers in ethanol are similar to those in water, although their values in ethanol are lower than those in water. As the solvent polarity increases, the increase in barrier to Z-Ala-to-N-Ala interconversion is caused by the increase of stability for Z-Ala, as described in the previous section. In particular, the barrier of -0.59 kcal/mol at the B3LYP/6-311++G(d,p)//CPCM HF/6-31+G(d) level of theory in chloroform indicates that TS lies 0.59 kcal/mol below the zwitterionic Z-Ala and the intermediate is not stable. Therefore, the CPCM B3LYP/6-31+G(d) level of theory appears to be the most appropriate method in computing the population and inter-

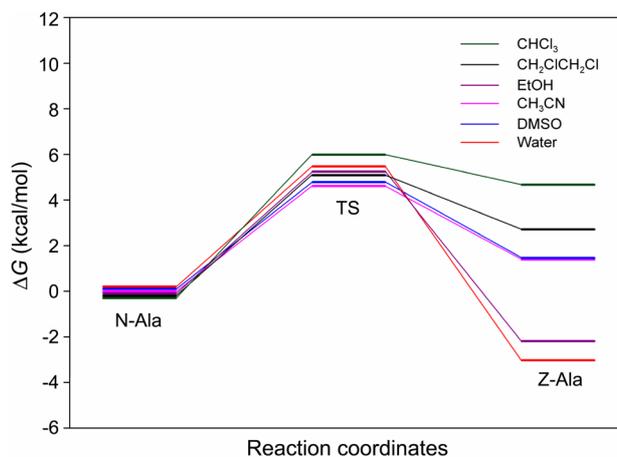


Figure 3. Relative free energies of Z-Ala and TS to the most preferred conformation of N-Ala at the CPCM B3LYP/6-31+G(d) level of theory in solution.

conversion of neutral and zwitterionic L-alanines in solution by the comparison of the ratios of the concentrations for Z-Ala and N-Ala in water and the barrier to Z-Ala-to-N-Ala interconversion in solution.

At the CPCM B3LYP/6-31+G(d) level of theory, the barriers to Z-Ala-to-N-Ala interconversion are computed to be 7.43, 3.22, and 8.50 kcal/mol in ethanol, acetonitrile, and water, respectively. These barriers were reported to be 3.38 and 3.85 kcal/mol at the B3LYP/6-31++G(d,p) level of theory in ethanol and acetonitrile, respectively,²⁵ and 5.86 kcal/mol at the MP2/6-31++G(d,p) level of theory in water,²⁴ with the SCRF method using multipolar expansions in an ellipsoidal cavity.

Conclusions

At both the HF and B3LYP levels of theory, the local minimum N1 for N-Ala is found to be most preferred in the gas phase and a weak asymmetric bifurcated hydrogen bond between the amino hydrogens and the carbonyl oxygen appears to play a role in stabilizing this conformation. The local minima N2a and N2b are found to be the second preferred conformations, which seem to be stabilized by a hydrogen bond between the amino nitrogen and the carboxylic hydrogen.

Most of the local minima for N-Ala and Z-Ala and their TS obtained in the gas phase are retained as stationary points at both the CPCM HF and B3LYP levels of theory in solution. The relative stability of the local minimum N2b is remarkably increased in solution than that in the gas phase. The local minimum N2b becomes more stable than the local minimum N2a in most of the solution.

On the whole the relative free energies of Z-Ala and TS become more lowered, as the solvent polarity increases. N-Ala prevails over Z-Ala in aprotic solutions but Z-Ala is dominantly populated in ethanol and water. In aprotic solutions, the population of Z-Ala increases somewhat with the increase of solvent polarity. The barrier to Z-Ala-to-N-Ala interconversion increases on the whole with the increase of solvent polarity, which is caused by the increase of stability for Z-Ala. The barriers in ethanol are similar to those in water, although their values in ethanol are lower than those in water.

Acknowledgements. This work is supported by a grant from Chungbuk National University in 2007.

Supporting Information Available. The selected torsion angles of all local minima and transition states in the gas phase and in solution and their selected bond lengths and bond angles in water optimized at the HF/6-31+G(d) and B3LYP/6-31+G(d) levels of theory. This material is available at <http://www.kcsnet.or.kr/bkcs> or on request from the correspondence author.

References

- Simpson, H. J., Jr.; Marsh, R. E. *Acta Cryst.* **1966**, *20*, 550.
- Lehmann, M. S.; Koetzle, T. F.; Hamilton, W. C. *J. Am. Chem. Soc.* **1972**, *94*, 2657.
- Edsall, J. T.; Blanchard, M. H. *J. Am. Chem. Soc.* **1933**, *55*, 2337.
- Ellzy, M. W.; Jensen, J. O.; Hameka, H. F.; Kay, J. G. *Spectrochim. Acta, Part A* **2003**, *59*, 2619.
- Godfrey, P. D.; Firth, S.; Hatherley, L. D.; Brown, R. D.; Pierlot, A. P. *J. Am. Chem. Soc.* **1993**, *115*, 9687.
- Iijima, K.; Nakano, M. *J. Mol. Struct.* **1999**, *485-486*, 255.
- Blanco, S.; Lesarri, A.; López, J. C.; Alonso, J. L. *J. Am. Chem. Soc.* **2004**, *126*, 11675.
- Cao, M.; Newton, S. Q.; Pranata, J.; Schäfer, L. *THEOCHEM* **1995**, *332*, 251.
- Gronert, S.; O'Hair, R. A. *J. Am. Chem. Soc.* **1995**, *117*, 2071.
- Császár, A. G. *J. Phys. Chem.* **1996**, *100*, 3541.
- Stepanian, S. G.; Reva, I. D.; Radchenko, E. D.; Adamowicz, L. *J. Phys. Chem. A* **1998**, *102*, 4623.
- Selvarangan, P.; Kolandaivel, P. *THEOCHEM* **2004**, *671*, 77.
- Kumar, S.; Rai, A. K.; Rai, S. B.; Rai, D. K.; Singh, A. N.; Singh, V. B. *J. Mol. Struct.* **2006**, *791*, 23.
- Upadhyay, D. M.; Rai, A. K.; Rai, D. K.; Singh, A. N.; Kumar, A. *Spectrochim. Acta, Part A* **2007**, *66*, 909.
- Maul, R.; Ortmann, F.; Preuss, M.; Hannewald, K.; Bechstedt, F. *J. Comput. Chem.* **2007**, *28*, 1817.
- Gontrani, L.; Mennucci, B.; Tomasi, J. *THEOCHEM* **2000**, *500*, 113.
- Frimand, K.; Bohr, H.; Jalkanen, K. J.; Suhai, S. *Chem. Phys.* **2000**, *255*, 165.
- Jalkanen, K. J.; Nieminen, R. M.; Frimand, K.; Bohr, J.; Bohr, H.; Wade, R. C.; Tajkhorshid, E.; Suhai, S. *Chem. Phys.* **2001**, *265*, 125.
- Abdali, S.; Jalkanen, K. J.; Bohr, H.; Suhai, S.; Nieminen, R. M. *Chem. Phys.* **2002**, *282*, 219.
- Ahn, D.-S.; Park, S.-W.; Jeon, I.-S.; Lee, M.-K.; Kim, N.-H.; Han, Y.-H.; Lee, S. *J. Phys. Chem. B* **2003**, *107*, 14109.
- Chuchev, K.; BelBruno, J. J. *THEOCHEM* **2008**, *850*, 111.
- Cao, X.; Fischer, G. *Spectrochim. Acta, Part A* **1999**, *55*, 2329.
- Cao, X.; Fischer, G. *Chem. Phys.* **2000**, *255*, 195.
- Sambrano, J. R.; de Sousa, A. R.; Queralt, J. J.; Andrés, J.; Longo, E. *Chem. Phys. Lett.* **1998**, *294*, 1.
- Nóbrega, G. F.; Sambrano, J. R.; de Souza, A. R.; Queralt, J. J.; Longo, E. *THEOCHEM* **2001**, *544*, 151.
- Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Montgomery, J. A., Jr.; Vreven, T.; Kudin, K. N.; Burant, J. C.; Millam, J. M.; Iyengar, S. S.; Tomasi, J.; Barone, V.; Mennucci, B.; Cossi, M.; Scalmani, G.; Rega, N.; Petersson, G. A.; Nakatsuji, H.; Hada, M.; Ehara, M.; Toyota, K.; Fukuda, R.; Hasegawa, J.; Ishida, M.; Nakajima, T.; Honda, Y.; Kitao, O.; Nakai, H.; Klene, M.; Li, X.; Knox, J. E.; Hratchian, H. P.; Cross, J. B.; Adamo, C.; Jaramillo, J.; Gomperts, R.; Stratmann, R. E.; Yazyev, O.; Austin, A. J.; Cammi, R.; Pomelli, C.; Ochterski, J. W.; Ayala, P. Y.; Morokuma, K.; Voth, G. A.; Salvador, P.; Dannenberg, J. J.; Zakrzewski, V. G.; Dapprich, S.; Daniels, A. D.; Strain, M. C.; Farkas, O.; Malick, D. K.; Rabuck, A. D.; Raghavachari, K.; Foresman, J. B.; Ortiz, J. V.; Cui, Q.; Baboul, A. G.; Clifford, S.; Cioslowski, J.; Stefanov, B. B.; Liu, G.; Liashenko, A.; Piskorz, P.; Komaromi, I.; Martin, R. L.; Fox, D. J.; Keith, T.; Al-Laham, M. A.; Peng, C. Y.; Nanayakkara, A.; Challacombe, M.; Gill, P. M. W.; Johnson, B.; Chen, W.; Wong, M. W.; Gonzalez, C.; Pople, J. A. *Gaussian 03*, revision C.02; Gaussian, Inc.: Wallingford, CT, 2004.
- Vásquez, M.; Némethy, G.; Scheraga, H. A. *Macromolecules* **1983**, *16*, 1043.
- Frisch, A.; Dennington, R. D., II; Keith, T. A. *GaussView*, version 3.0; Gaussian, Inc.: Pittsburgh, PA, 2003.
- Peng, C.; Schlegel, H. B. *Israel J. Chem.* **1993**, *33*, 449.
- Frisch, A.; Frisch, M. J.; Trucks, G. W. *Gaussian 03 User's Reference*, version 7.0; Gaussian, Inc.: Pittsburgh, PA, 2003.
- Barone, V.; Cossi, M. *J. Phys. Chem. A* **1998**, *102*, 1995.
- Cossi, M.; Rega, N.; Scalmani, G.; Barone, V. *J. Comput. Chem.* **2003**, *24*, 669.
- Barone, V.; Cossi, M.; Tomasi, J. *J. Chem. Phys.* **1997**, *107*, 3210.
- Cossi, M.; Scalmani, G.; Rega, N.; Barone, V. *J. Chem. Phys.* **2002**, *117*, 43.
- Tomasi, J.; Mennucci, B.; Cammi, R. *Chem. Rev.* **2005**, *105*, 2999.
- Takano, Y.; Houk, K. N. *J. Chem. Theory Comput.* **2005**, *1*, 70.
- Kang, Y. K. *THEOCHEM* **2001**, *546*, 183.
- Hehre, W. J.; Radom, L.; Schleyer, P. v. R.; Pople, J. A. *Ab Initio Molecular Orbital Theory*; John Wiley & Sons: New York, 1986; Chapter 6.