

Correlation of the Rates of Solvolyses of Cinnamyl Bromide

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Solvolytic rate constants at 25 °C are reported for solvolyses of cinnamyl bromide (**1**) in binary mixtures of water with acetone, ethanol, methanol, methanol-*d*, and 2,2,2-trifluoroethanol. Product selectivities are reported for solvolyses of **1** in aqueous ethanol and methanol. Rate ratios in solvents of the same Y_{Br} value and different nucleophilicity provide measures of the minimum extent of nucleophilic solvent assistance (e.g. $[k_{40EW}/k_{97TFE}]_Y = 2.88$, EW = ethanol-water). With use of the extended Grunwald-Winstein equation, the l and m values are similar to the values of 0.43 and 0.88 obtained for the solvolyses of **1** using the equation (see below) which includes a parameter (I) for solvation of aromatic rings. The magnitude of l and m values associated with a change of solvent composition predicts the S_N1 reaction mechanism rather than an S_N2 channel.

Product selectivities (S), defined by

$$S = [\text{ether product}]/[\text{alcohol product}] \times [\text{water}]/[\text{alcohol solvent}]$$

are related to four rate constants for reactions involving one molecule of solvent as nucleophile and another molecule of solvent as general base catalyst. A linear relationship between $1/S$ and molar ratio of solvent is derived theoretically and validated experimentally for solvolyses of the above substrates from water up to 75%

$$1/S = (k_{wa}/k_{aw})([\text{alcohol solvent}]/[\text{water}]) + k_{ww}/k_{aw}$$

alcohol-water. The results are best explained by product formation from a "free" carbocation intermediate rather than from a solvent-separated ion pair.

Key Words : Solvolysis, Selectivity, General base catalyst, Solvent-separated ion pair

Introduction

Dispersion into separate lines in the correlation of the specific rates of solvolysis of a substrate in various binary mixtures was documented¹⁻⁵ in earlier treatments using the Grunwald-Winstein Eqn. (1).⁶⁻¹¹

$$\log(k/k_0) = mY + c \quad (1)$$

In general, dispersion effects in unimolecular solvolysis^{10,11} make smaller contributions to the overall linear free energy relationships (LFER) than solvent nucleophilicity effects in bimolecular solvolysis.^{6,12} It was suggested that a second term which is governed by the sensitivity l to solvent nucleophilicity N , should be added to Eqn. (1) for bimolecular solvolysis.¹² The resulting Eqn. (2) is often referred to as the extended Grunwald-Winstein equation.¹²

$$\log(k/k_0) = mY + lN + c \quad (2)$$

Kevill *et al.*, recently suggested that, since the dispersion seems to follow a consistent pattern, it should be possible to develop an aromatic ring parameter (I), which, as qualified by the appropriate sensitivity (h), can be added to Eqns. (1) or (2) to give Eqns. (3) or (4), respectively.¹³⁻¹⁶

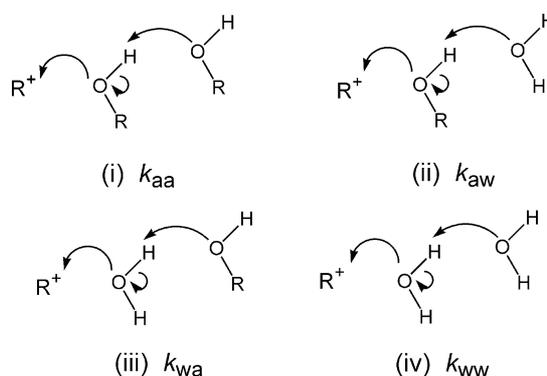
$$\log(k/k_0) = mY + hI + c \quad (3)$$

$$\log(k/k_0) = mY + lN + hI + c \quad (4)$$

Solvolytic reactions in alcohol-water mixtures lead to alcohol and ether products from which selectivities can be calculated using Eqn. (5).¹⁷⁻¹⁹

$$S = \frac{[\text{ether product}]}{[\text{alcohol product}]} \times \frac{[\text{water}]}{[\text{alcohol solvent}]} \quad (5)$$

Nucleophilic attack may occur on 'free' carbocations, solvent-separated ion pairs or possibly contact ion pairs, and the solvent-dependence of S may be due to changes in the nature of the intermediate leading to product.^{19,20} An



Scheme 1. General base catalyzed reactions for nucleophilic attack on carbocations.

alternative explanation involving nucleophilic attack solely on 'free' carbocations is given below to account for the solvent dependence of S for S_N1 reactions.¹⁹

If nucleophilic attack by one solvent molecule were assisted by a second molecule of solvent acting as general base, in alcohol-water mixtures there will be four competing product-determining steps (Scheme 1) defined by the following third order rate constants: k_{ww} in which water is both nucleophile and general base; k_{aw} , the alcohol solvent acts as nucleophile and water acts as general base; k_{wa} , water acts as nucleophile and alcohol acts as general base; k_{aa} , alcohol acts as both nucleophile and general base.¹⁹

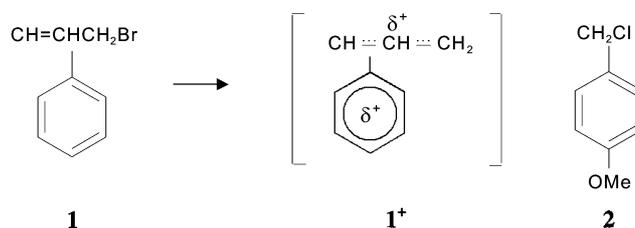
Scheme 1 is of the same kinetic form as the one we have previously derived for acyl^{21a,21b} and sulfonyl^{21c} transfer reactions which are shown to proceed by S_N2 or $S_A N$ reaction mechanisms.²¹ However, we reported the new relationship between S and solvent composition for S_N1 reactions involving nucleophilic attack on carbocations and general base catalysis.¹⁹ It should be noted that in an S_N2 processes the rate determining and product determining steps are the same, while in the S_N1 processes the rate depends on the ionizing power of the solvent but the product is determined by the attack of solvent on the cationic intermediate, R^+ , *i.e.*, the rate determining and product determining steps are not the same. We assume that the rates of product-determining reactions shown in Scheme 1 are given by third order rate constants multiplied by the appropriate molar concentrations of solvent. If, in highly aqueous media, $k_{aw}[\text{water}][\text{ROH}] \gg k_{aa}[\text{ROH}]^2$, the relationship given in Eqn. (6) can readily be derived.^{21b}

$$1/S = (k_{wa}/k_{aw})([\text{alcohol solvent}]/[\text{water}]) + k_{ww}/k_{aw} \quad (6)$$

The intercept of Eqn. (6) implies that the maximum S value in highly aqueous media is determined by the ratio of third order rate constants k_{aw}/k_{ww} ; this ratio represents the rate of nucleophilic attack by alcohol compared with the rate of nucleophilic attack by water, in water as solvent and with water acting as a general base catalyst.

It has been shown that the plot of $1/S$ versus $[\text{cosolvent}]/[\text{water}]$ (Eqn. 6) lead to a straight line of positive slope (k_{wa}/k_{aw}) in the general base catalysed (Scheme 1) solvolysis of acyl halides in which the k_{wa} (and/or k_{aw}) term is important in the product formation step.^{19b} The main purpose of the work described in this paper is to test whether there is a linear relationship between $1/S$ and the alcohol/water molar ratio, as predicted by Eqn. (6) for the product-determining step of S_N1 reaction of carbocations. Previous work has been restricted to solvolyses of *p*-methoxybenzyl and diphenylmethyl chlorides.²¹

Suitable substrates on which to test Eqn. (6) are those which react only by substitution pathways *via* 'free' carbocations.¹⁹ Also, the substrates should permit rapid mixing in highly aqueous mixtures, and should not be so reactive that their solvolyses are interfacial.⁴ We chose cinnamyl bromide (**1**) as a suitable substrate which reacts only is expected to react through an S_N1 pathway *via* 'free' carbocation in view of the stable cationic form, **1**⁺.



In this work, we determined rate constants for solvolyses of **1** in aqueous binary mixtures of acetone, ethanol, methanol, methanol-*d*, 97% TFE (w/w) and pure water at 25.0 °C, and the transition state variation is discussed by applying the Grunwald-Winstein equation, extended Grunwald-Winstein equation, aromatic ring parameter equation and kinetic solvent isotope effect. We test whether there is a linear relationship between $1/S$ and the alcohol/water molar ratio, as predicted by Eqn. (6) for the product-determining step of S_N1 reactions.

Results

Rate constants for solvolyses of **1** in aqueous binary mixtures of acetone, ethanol, methanol, methanol-*d*, water, NaClO₄, and LiBr in 90MeOH, TFE-water (%w/w) and TFE-ethanol (%v/v) at 25.0 °C are reported in Tables 1, 2 and 3, and results at various temperatures are given in Table 4 along with the corresponding Arrhenius parameters.

Selectivity data were obtained from product analyses as soon as possible after completion of 10 half-lives of reaction.¹⁹ Selectivity data for solvolyses of **1** in ethanol-water and methanol-water mixtures at 25.0 °C are given in Table 5.

Discussion

Kinetic Data. For binary solvents having the same % v/v water, rate constants (Table 1) increase in the order acetone-

Table 1. Rate constants (k , s⁻¹) for solvolyses of cinnamyl bromide (**1**) in aqueous binary mixtures at 25 °C

v/v %	MeOH	EtOH	Acetone
	$k \times 10^3$		
100	0.113 ^b	0.0156	
90	0.514	0.113	
80	1.87	0.478	0.0287
70	6.04	1.31	0.0168
60	17.6	3.40	0.741
50	41.5	12.5	3.21
40	113	39.2	15.6
30	304	269	59.7
20	393	393	206
10	445	434	382
H ₂ O	505	505	505

^aDetermined conductimetrically at least in duplicate; typical error $\pm 3\%$.
^bSolvolyses in MeOD give $k = 0.0890 (\pm 0.10) \times 10^{-3} \text{ s}^{-1}$ giving a kinetic solvent isotope effect of 1.27.

Table 2. Rate constants (k , s^{-1}) for solvolyses of cinnamyl bromide (**1**) in 2,2,2-trifluoroethanol binary mixtures at 25 °C

w/w %	TFE-H ₂ O		TFE-EtOH
	$k \times 10^3$	v/v %	
97 TFE – 3 H ₂ O	15.3	80 TFE – 20 EtOH	2.36
90 TFE – 10 H ₂ O	31.5	60 TFE – 40 EtOH	0.493
70 TFE – 30 H ₂ O	66.4	40 TFE – 60 EtOH	0.132
50 TFE – 50 H ₂ O	108	20 TFE – 0 EtOH	0.0400

^aDetermined conductimetrically at least in duplicate; typical error \pm 3%.
^b $k_{40EW}/k_{97TFE} = 2.88$ (k_{40EW} taken from Table 1).

Table 3. Rate constants (k , s^{-1}) for solvolyses of cinnamyl bromide (**1**) in salt solution in 90% methanol–10% water binary mixtures at 25 °C^a

Added salts (Concn, M)	$k \times 10^3$, sec^{-1}
0.01 M LiBr	0.577
0.003 M LiBr	0.547
0.01 M NaClO ₄	0.589
0.003 M NaClO ₄	0.523

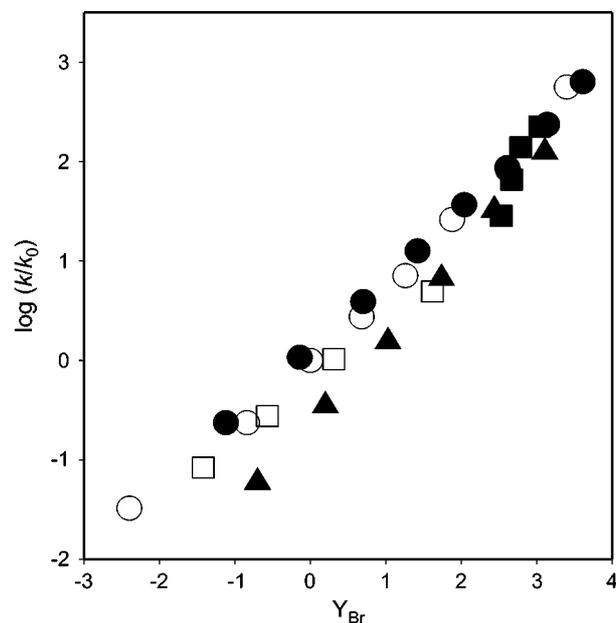
^aDetermined titrimetrically at least in duplicate; typical error \pm 5%.

Table 4. Additional rate constants (k , s^{-1}) for solvolyses of cinnamyl bromide (**1**) in aqueous alcohol mixtures and the activation parameters calculated using data at 35 and 45 °C

T/°C	k/s^{-1}	$\Delta H^\ddagger/kcal\ mol^{-1}$	$\Delta S^\ddagger/cal\ mol^{-1}K^{-1}$
70% EtOH 35 °C	4.83×10^{-3}	20.5	-0.932
70% EtOH 45 °C	1.33×10^{-2}		
70% MeOH 35 °C	1.68×10^{-2}	18.3	-5.14
70% MeOH 45 °C	4.88×10^{-2}		
90% MeOH 35 °C	1.67×10^{-3}	19.8	-5.15
90% MeOH 45 °C	4.86×10^{-2}		

H₂O < ethanol-H₂O < methanol-H₂O. The rates increase appreciably as the water content of a mixture increases, thus the rate is the highest in the solvents with high ionizing power, Y , suggesting extensive bond breaking in the transition state. First-order rate constants for solvolyses of **1** vary over 3.2×10^4 -fold in alcohol-water mixtures, whereas the observed first-order rate constants for phenyl chloroformates, which is known to react *via* an S_N2 like S_AN or carbonyl addition-elimination reaction mechanism vary only over 20-fold.^{6,22} These results indicate that the rate-determining step is not the bond formation step which was found in the reaction of phenyl chloroformates but the bond breaking step for the reaction of **1** where the transition state is sensitive to solvent ionizing power. These results are very similar to those of solvolytic reactions of *p*-methoxybenzoyl chloride,^{2,23a} cinnamyl chloride,^{23b} and thenoyl chloride,²⁵ but they are different from the results of solvolytic reactions of *p*-nitrobenzoyl chloride,^{21b,26} *p*-nitrobenzenesulfonyl chloride^{1,21} and furoyl chloride.²⁵

The close similarity of the rate-determining step for solvolyses of **1**, *p*-methoxybenzyl chloride (**2**) and cinnamyl chloride is shown by similar solvent kinetic isotope effects in methanol (1.27 for **1**, 1.22 for **2**, and 1.11 for cinnamyl

**Figure 1.** Logarithms of first-order rate constants for solvolysis of cinnamyl bromide at 25 °C; $\log(k/k_0)$ vs. Y_{Br} (solvent code: ●, methanol; ○, ethanol; ▲, acetone; ■, TFE-H₂O; □, TFE-EtOH).

chloride; see footnote in Table 1)²³ and by the very similar rate-rate profiles of solvent effects on reactivity (Fig. 3). Extensive studies have been made of solvent isotope effects in water, but relatively little has been published for methanol.^{23,27} The value of 1.27 (footnote in Table 1) is in the range expected for an S_N1 reaction mechanism.²³ Also, similar solvent isotope effects in water and methanol have previously been obtained for solvolyses of acetic anhydride.²⁷

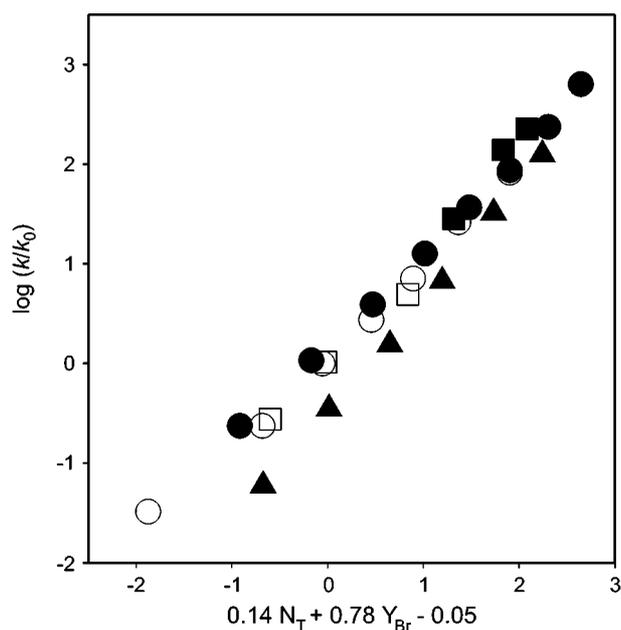
It is no common ion depression (Table 3), perhaps indicating nucleophilic attack on a contact ion pair, because we think that nucleophile attack on a solvent-separated ion pair would give a constant selectivity, and nucleophilic attack on a free cation should also show common ion rate depression.

The Grunwald-Winstein plots (Eqn. 1) of the rates in Table 1 are presented in Figure 1 using the solvent ionizing power scale Y_{Br} , based on 1-adamantyl bromide.^{3,6-9} Examination of Figure 1 shows that the plots for the three aqueous mixtures exhibit dispersions into two separate lines. The plots for all binary mixed solvents show a large m value, $m = 0.74$ ($r > 0.977$) and the data point for 97% TFE solvent mixture shows only no significant deviation from the correlation of Grunwald-Winstein plots for aqueous alcohol and acetone.

Rate ratios in two solvents which have same Y_{Br} value, *i.e.* the same degree of solvent assistance for bond cleavage but different nucleophilicity provide measures of the minimum extent of nucleophilic solvent assistance (*e.g.* $[k_{40EW}/k_{97TFE}]_Y = 2.88$, EW = ethanol-water).²⁸ Thus the large m values found (0.74) and the small value (= 2.88) of k_{40EW}/k_{97TFE} shown in Table 2 imply that the solvolyses of **1** in the binary mixtures proceed by an S_N1 pathway channel rather than by an S_N2 channel (See Table 5).

Table 5. Values of solvolysis rate ratios in 40% EtOH/H₂O compared with 97% TFE/H₂O at 25 °C^a

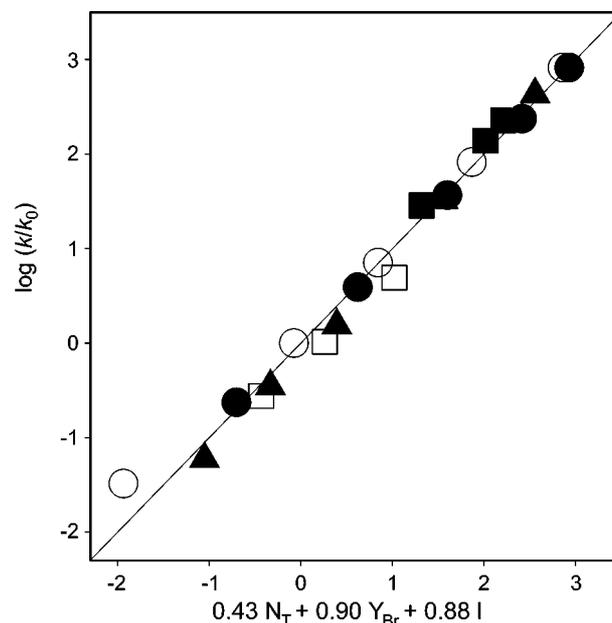
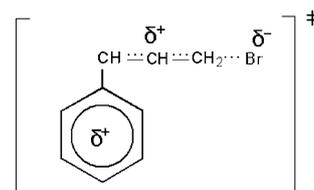
Substrates	$k_{40\text{EtOH}}/k_{97\text{TFE}}$	Substrates	$k_{40\text{EtOH}}/k_{97\text{TFE}}$
Ph ₂ CHCl	0.18	<chem>CC(C)C(C)Cl</chem>	4.6
CH ₃ OφCH ₂ Cl	0.70	<chem>CC(C)(C)C(Cl)C</chem>	7.1
<chem>CC(C)C(Cl)C</chem>	2.0	<chem>CC(C)C(Cl)C</chem>	1.9
<chem>CC(C)(C)C(Cl)C</chem>	11	<chem>c1ccc(cc1)C=CCl</chem>	1.48

^aData from refs 23 and 28.**Figure 2.** Plot of $\log(k/k_0)$ for cinnamyl bromide against $(0.14N_T + 0.78Y_{Br} - 0.05)$; $r = 0.980$ (solvent codes: ●, methanol; ○, ethanol; ▲, acetone; ■, TFE-H₂O; □, TFE-EtOH).

In order to examine the cause of this dispersion phenomenon, we correlated the rate data in Table 1 using Eqn. (2). The nucleophilicity parameter (N_T) has been shown to give a very poor correlation when an $\ln N_T$ term is added to the original Grunwald-Winstein Eqn. (1) (see Figure 2). Therefore such phenomenon can not be explained as dispersion effect caused by solvent nucleophilicity parameter.

Inclusion of the ring parameter (I) and nucleophilicity parameter (N_T) gives considerable improvement to correlations of the solvolyses of **1** (see Figure 3).

Therefore, this shows an importance of aromatic ring parameter rather than solvent nucleophilicity parameter for solvolysis of **1**. Therefore such phenomenon can be explained as dispersion effects caused by aromatic ring. The dispersions in the Grunwald-Winstein correlations in the

**Figure 3.** Plot of $\log(k/k_0)$ for cinnamyl bromide against $(0.43N_T + 0.90Y_{Br} + 0.88I)$; $r = 0.995$ (solvent codes: ●, methanol; ○, ethanol; ▲, acetone; ■, TFE-H₂O; □, TFE-EtOH).**Scheme 2.** Proposed Transition State Structure.

present studies are caused by the conjugation between the reaction center and aromatic ring through the conjugated C=C double bond leading to the medium value of sensitivity of I ($h = 0.88$) in the equation (4). The through conjugation of the ring π system with the reaction center has also been found in the solvolysis of phenyl chlorothionoformate²⁹ and pyridinolysis of phenyl chloroformates.³⁰ In this study this is interpreted to indicate the overlap of the reaction center which has strong cationic charge in the TS and aromatic ring π system through C=C double bond in the rate-limiting step. With use of the equation (4), the m and h values of 0.90 and 0.88 are obtained for the solvolyses of **1** (Fig. 3).

This study has shown that the magnitude of m and h values associated with a change of solvent composition is able to predict the S_N1 reaction mechanism rather than an S_N2 channel. In the TS, the cationic charge is delocalized to the ring, which causes to change the solvation of the ring and leads to a better correlations with the ring parameter, I .

Assuming general base catalyzed reactions as shown in Scheme 1 for nucleophilic attack on the substrate, the observed rate, k_{obs} , can be given for the reactions in acetone-water mixtures as Eqn. (7).¹⁹

$$k_{\text{obs}} = k_{\text{ww}} [\text{water}]^2 + k_{\text{wa}} [\text{water}][\text{acetone}] + k_{\text{aa}} [\text{acetone}]^2 \quad (7)$$

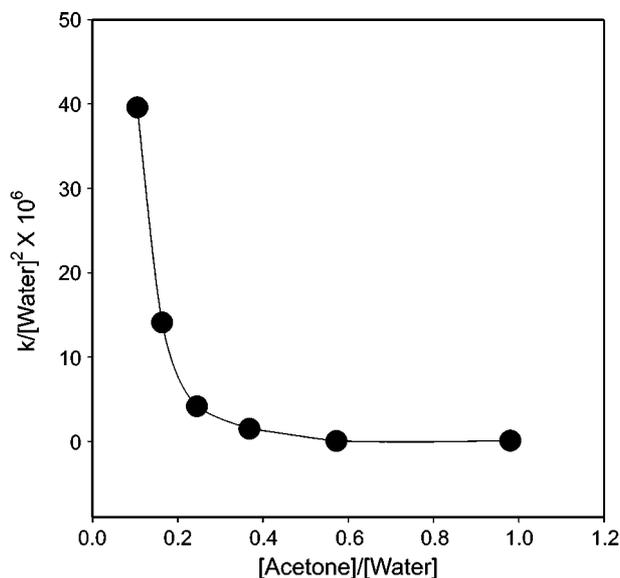


Figure 4. Plot of third-order rate constant versus [Acetone]/[Water].

The third-order rate constant k_{wa} can be obtained by the plot of $k_{obs}/[\text{Water}]^2$ versus [acetone]/[Water] (Eqn. 8).

$$k_{obs}/[\text{Water}]^2 = k_{ww} + k_{wa} [\text{acetone}]/[\text{Water}] \quad (8)$$

This type of plot showed a straight line of positive slope for the reactions proceeding with an S_N2 mechanism (e.g. solvolysis of *p*-nitrophenyl chloroformate).^{19b} The plot for the solvolysis of **1**, however, leads to an exponential decay curve, Figure (4). This means that the rate is independent of the catalysis, k_{wa} is non-existent and dependent on the ionizing power of the solvent, Y . In fact the plot of $\log k$ (third order) against volume percentage acetone shows a straight line of negative slope, which is exactly similar to that of $\log k$ versus Y plot. We therefore conclude that the rate is solely dependent on Y (as required for an S_N1 reaction) and slow step of this reaction is ionization and there is no catalysis and is not third order by solvent molecule.

The activation parameters for the reaction at 25 °C are summarized in Table 4. The small negative ΔS^\ddagger value and large positive ΔH^\ddagger values are again in good agreement with our proposed S_N1 reaction channels.

Rate-product correlations. Both the rate constants (Table 1) and S values [Eqn. (5), Table 6] increase as water is added to ethanol or to methanol. The increases in rates can be explained in various ways (e.g. by an increase in solvent ionizing power as the water content of the solvent increases), but the increase in S is more difficult to explain.³¹ Competing second-order reactions would give constant S values, in the absence of medium effects,^{31,32} but the increase in S can be explained by a third-order mechanism.¹⁹

Plot of $1/S$ versus the molar ratio of alcohol/water [Eqn. (6)] are linear from 10% to at least 70% alcohol-water mixtures and thereafter curvature is significant (Fig. 5) which is consistent with increasing contributions from the k_{aa} term (Scheme 1). Selectivities for methanol-water are significantly greater than for ethanol-water, and the slopes of $1/S$

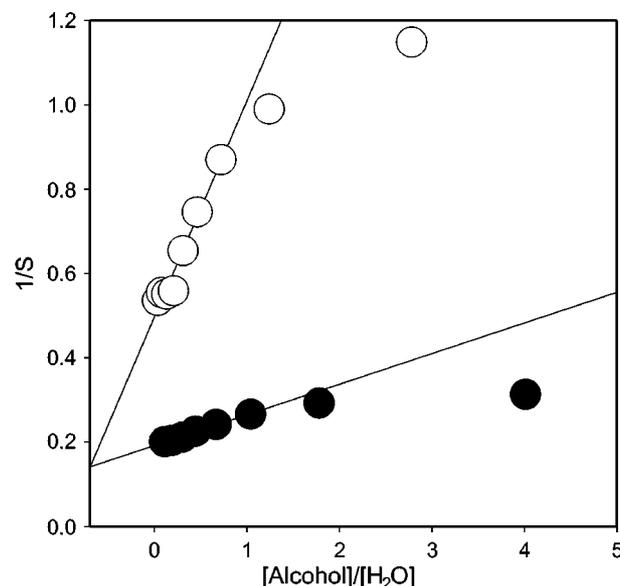


Figure 5. Plot of $1/S$ vs. molar ratio of solvents versus [alcohol]/[water], for solvolysis of cinnamyl bromide at 25 (solvent codes: ●, methanol; ○, ethanol) (ethanol: slope = 0.513, intercept = 0.497; methanol: slope = 0.00924, intercept = 0.276).

plots are lower for methanol-water than for ethanol-water (Fig. 5). These trends can be explained by the k_{aw} term, which compared with the k_{wa} term, is greater for methanol-water than ethanol-water. The intercept of Eqn. (6) is determined by the ratio of third order rate constants k_{ww}/k_{aw} , so nucleophilic attack by methanol assisted by water is 3.6 times more favourable than nucleophilic attack by water assisted by water.

The valid linear correlation obtained with eqn. (6) indicates that the product determining step, i.e. the attack of solvent on the cationic intermediate (R^+), is catalysed by the cosolvent as in the S_N2 processes where the neutral substrate is attacked by the solvent.^{19b}

Experimental Section

Material. Methanol, ethanol, and acetone were Merck GR grade (< 0.1% H_2O), and CH_3OD were from Aldrich (99.9%

Table 6. Selectivities (S) for solvolyses of cinnamyl bromide in aqueous binary mixtures at 25 °C

v/v %	Methanol/Water		Ethanol/Water	
	[Ether]/[Alcohol]	S	[Ether]/[Alcohol]	S
90	12.8	3.19	2.43	0.87
80	6.08	3.42	1.24	1.01
70	3.90	3.76	0.829	1.15
60	2.75	4.13	0.619	1.34
50	1.97	4.43	0.473	1.53
40	1.40	4.73	0.369	1.79
30	0.937	4.91	0.241	1.99
20	0.554	4.98	0.142	1.80
10	0.239	4.83	0.0639	1.81

Table 7. Correlation analyses against different Y , N and I

Parameters	n	r	m (s.d) ^a	l (s.d) ^a	h (s.d) ^a
Y_{Br}	30	0.977	0.74 (0.03)		
Y_{Br}, N_{OTs}	27	0.985	0.76 (0.03)	0.20 (0.08)	
Y_{Br}, N_{T}	16	0.980	0.78 (0.05)	0.14 (0.09)	
Y_{Br}, I	21	0.987	0.74 (0.04)		0.16 (0.21)
Y_{Br}, N_{OTs}, I	22	0.992	0.86 (0.06)	0.43 (0.11)	0.85 (0.26)
Y_{Br}, N_{T}, I	21	0.995	0.90 (0.02)	0.43 (0.05)	0.88 (0.14)

^aStandard deviation.

D). Distilled water was redistilled with Buchi Fontavapor 210 and treated using ELGA UHQ PS to obtain specific conductivity of less than 1×10^{-6} mhos/cm. Cinnamyl bromide (Aldrich GR grade, 99%) was used without further purification.

Rate measurements and product selectivities. The rates were measured conductometrically at $25 (\pm 0.03) ^\circ\text{C}$ at least in duplicate as described previously,^{1,22,33} with concentrations of substrate *ca.* 10^{-3} M.

The solvolysis products, ether and alcohol, were determined by HPLC analysis as described previously²¹ and the product-selectivities, S , were calculated from Eqn. (5). The S values were calculated from the observed peak area ratios of ether and alcohol, divided by the appropriate response factor. For response calibrations, area ratios from solvolyses of **1** in pure alcohol and in 40% acetonitrile-water mixtures were used. The eluent was 75% methanol-water mixture and the flow rate was adjusted to 1 mL min^{-1} . The HPLC system was a Hewlett-Packard 1050 Series instrument, with 250×4 mm Spherisorb ODS reversed column.

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References

- Koo, I. S.; Bentley, T. W.; Kang, D. H.; Lee, I. *J. Chem. Soc. Perkin Trans. 2* **1991**, 175-179.
- Bentley, T. W.; Koo, I. S. *J. Chem. Soc. Perkin Trans. 2* **1989**, 1385-1392.
- Bentley, T. W.; Carter, G. E. *J. Am. Chem. Soc.* **1982**, *104*, 5741-5747.
- Bentley, T. W.; Harris, H. C.; Koo, I. S. *J. Chem. Soc. Perkin Trans. 2* **1988**, 783-789.
- Bentley, T. W.; Harris, H. C. *J. Chem. Soc. Perkin Trans. 2* **1986**, 619-624.
- Koo, I. S.; An, S. K.; Yang, K.; Koh, H. J.; Choi, M. H.; Lee, I. *Bull. Korean Chem. Soc.* **2001**, *22*, 842-846.
- Bentley, T. W.; Bowen, C. T.; Morten, D. H.; Schleyer, P. V. R. *J. Am. Chem. Soc.* **1981**, *103*, 5466-5475.
- Winstein, S.; Grunwald, E. *J. Am. Chem. Soc.* **1948**, *70*, 846-854.
- Bentley, T. W.; Dau-Schmidt, J. P.; Llewellyn, G.; Mayr, H. *J. Org. Chem.* **1992**, *57*, 2387-2392.
- Winstein, S.; Fainberg, A.; Grunwald, E. *J. Am. Chem. Soc.* **1957**, *79*, 4146-4155.
- Fainberg, A. H.; Winstein, S. *J. Am. Chem. Soc.* **1957**, *79*, 1957-1608.
- Winstein, S.; Grunwald, E.; Jones, H. W. *J. Am. Chem. Soc.* **1951**, *73*, 2700-2707.
- Kevill, D. N.; Ismail, N. H. J.; D'Souza, M. J. *J. Org. Chem.* **1994**, *59*, 6303-6312.
- Kevill, D. N.; D'Souza, M. J. *J. Chem. Soc. Perkin Trans. 2* **1995**, 973-980.
- Kevill, D. N.; D'Souza, M. J. *J. Chem. Soc. Perkin Trans. 2* **1997**, 257-263.
- Kevill, D. N.; Bond, M. W.; D'Souza, M. J. *J. Org. Chem.* **1997**, *62*, 7869-7871.
- (a) Harris, J. M.; Clark, D. C.; Becker, A.; Fagan, J. F. *J. Am. Chem. Soc.* **1974**, *96*, 4478-4484; (b) Harris, J. M.; Becker, A.; Fagan, J. F.; Walden, F. A. *J. Am. Chem. Soc.* **1974**, *96*, 4484-4489.
- (a) Karton, Y.; Pross, A. *J. Chem. Soc. Perkin Trans. 2* **1977**, 1860-1863; (b) McLennan, D. J.; Martin, P. L. *J. Chem. Soc. Perkin Trans. 2* **1982**, 1099-1105.
- (a) Bentley, T. W.; Ryu, Z. H. *J. Chem. Soc. Perkin Trans. 2* **1994**, 761-767; (b) Koo, I. S.; Yang, K.; Kang, K.; Lee, I.; Bentley, T. W. *J. Chem. Soc. Perkin Trans. 2* **1998**, 1179-1183.
- Song, B. D.; Jencks, Y. P. *J. Am. Chem. Soc.* **1989**, *111*, 8470-8479.
- (a) Bentley, T. W.; Jones, R. O. *J. Chem. Soc. Perkin Trans. 2* **1993**, 2351-2357; (b) Bentley, T. W.; Jones, R. O.; Koo, I. S. *J. Chem. Soc. Perkin Trans. 2* **1994**, 753-759; (c) Jones, R. O. *M. Phil. Thesis*; University of Wales: 1991.
- Koo, I. S.; Yang, K.; Kang, K.; Lee, I. *Bull. Korean Chem. Soc.* **1998**, *19*, 968-973.
- (a) Bentley, T. W.; Koo, I. S.; Norman, S. *J. Org. Chem.* **1991**, *56*, 1604-1609; (b) Koo, I. S.; An, S. K.; Yang, K.; Lee, I.; Bentley, T. W. *J. Phys. Org. Chem.* **2002**, *15*, 758-764.
- Liu, K. T.; Duann, Y. F.; Hou, S. H. *J. Chem. Soc. Perkin Trans. 2* **1998**, 2181-2185.
- Oh, J.; Yang, K.; Koo, I. S.; Lee, I. *J. Chem. Res.* **1993**, 310-311.
- Bentley, T. W.; Harris, H. C. *J. Org. Chem.* **1988**, *53*, 724-728.
- Gold, V.; Grist, S. *J. Chem. Soc. B* **1971**, 2285-2286.
- Bentley, T. W.; Llewellyn, G.; Ryu, Z. H. *J. Org. Chem.* **1998**, *63*, 4654-4659.
- Koo, I. S.; Yang, K.; Kang, D. H.; Park, H. J.; Kang, K.; Lee, I. *Bull. Korean Chem. Soc.* **1999**, *20*, 577-580.
- Koh, H. J.; Lee, H. C.; Lee, H. W.; Lee, I. *Bull. Korean Chem. Soc.* **1996**, *17*, 712-715.
- Bentley, T. W.; Ebdon, D.; Llewellyn, G.; Abduljaber, M. H.; Miller, B.; Kevill, D. N. *J. Chem. Dalton Trans.* **1997**, 3819-3825.
- Ta-Shma, R.; Rapport, Z. *Adv. Phys. Org. Chem.* **1992**, *27*, 239.
- Yang, K.; Koo, I. S.; Lee, I.; Jo, D.-S. *Bull. Korean Chem. Soc.* **1994**, *15*, 280-284.